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Rising CO₂ Levels Will Intensify Phytoplankton Blooms in Eutrophic and Hypertrophic Lakes

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Abstract

Harmful algal blooms threaten the water quality of many eutrophic and hypertrophic lakes and cause severe ecological and economic damage worldwide. Dense blooms often deplete the dissolved CO₂ concentration and raise pH. Yet, quantitative prediction of the feedbacks between phytoplankton growth, CO₂ drawdown and the inorganic carbon chemistry of aquatic ecosystems has received surprisingly little attention. Here, we develop a mathematical model to predict dynamic changes in dissolved inorganic carbon (DIC), pH and alkalinity during phytoplankton bloom development. We tested the model in chemostat experiments with the freshwater cyanobacterium Microcystis aeruginosa at different CO₂ levels. The experiments showed that dense blooms sequestered large amounts of atmospheric CO₂, not only by their own biomass production but also by inducing a high pH and alkalinity that enhanced the capacity for DIC storage in the system. We used the model to explore how phytoplankton blooms of eutrophic waters will respond to rising CO₂ levels. The model predicts that (1) dense phytoplankton blooms in low- and moderately alkaline waters can deplete the dissolved CO₂ concentration to limiting levels and raise the pH over a relatively wide range of atmospheric CO₂ conditions, (2) rising atmospheric CO₂ levels will enhance phytoplankton blooms in low- and moderately alkaline waters with high nutrient loads, and (3) above some threshold, rising atmospheric CO₂ will alleviate phytoplankton blooms from carbon limitation, resulting in less intense CO₂ depletion and a lesser increase in pH. Sensitivity analysis indicated that the model predictions were qualitatively robust. Quantitatively, the predictions were sensitive to variation in lake depth, DIC input and CO₂ gas transfer across the air-water interface, but relatively robust to variation in the carbon uptake mechanisms of phytoplankton. In total, these findings warn that rising CO₂ levels may result in a marked intensification of phytoplankton blooms in eutrophic and hypertrophic waters.

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Introduction

Since the start of the industrial revolution, atmospheric CO₂ concentrations have increased from 275 to 400 ppm CO₂, and climate change scenarios predict that atmospheric CO₂ will further increase [1]. Enhanced dissolution of CO₂ will lower the pH of aquatic ecosystems [2,3]. However, CO₂ in freshwater ecosystems does not only originate from dissolution of atmospheric CO₂, but also from mineralization of organic carbon obtained from terrestrial sources in the surrounding watershed [4]. Mineralization of organic carbon causes CO₂ supersaturation in many lakes, in some cases even reaching CO₂ levels exceeding 10,000 ppm [5–7].

Phytoplankton fix CO₂ for photosynthesis, and many species can also utilize bicarbonate as a carbon source [8–10]. Assimilation of inorganic carbon by dense phytoplankton blooms can deplete the dissolved CO₂ concentration [11–15], sometimes down to levels below 1 ppm [7,15], so that these waters become severely CO₂-undersaturated. CO₂ depletion will cause an increase in pH [11,16,17]. Indeed, in eutrophic lakes with dense phytoplankton blooms, pH easily exceeds values of 9 [7,15], and can reach values as high as 11 in shallow hypertrophic lakes [19].

The combination of high pH values and CO₂ depletion in freshwaters is often associated with cyanobacterial blooms [19,20]. Several of the cyanobacterial species that commonly dominate these blooms are capable of producing toxic substances [21,22]. Consequently, cyanobacterial blooms threaten the water quality of many freshwater lakes and brackish waters around the world, including Lake Erie in USA-Canada [23], Lake Taihu in China [24,25], Lake Biwa in Japan [26], Lake Victoria in Africa [27,28], the Baltic Sea in Northern Europe [29,30], and many other ecologically and economically important lakes, rivers and estuaries [21,22,31,32]. Cyanobacterial blooms are expected to benefit from global warming [32–35]. The response of cyanobacteria to rising CO₂ concentrations, however, is less well understood, although it is clear that there is a strong interaction between cyanobacterial bloom development and CO₂ availability.

As an illustration, Fig. 1 provides data from Lake Volkerak, a large eutrophic lake in The Netherlands [31,36]. In winter and spring, CO₂ concentrations in Lake Volkerak largely exceed the
CO₂ concentrations that would be predicted from equilibrium with the atmosphere, and hence the lake is supersaturated with CO₂. In summer and early fall, however, Lake Volkerak is covered by dense blooms of the harmful cyanobacterium *Microcystis* [31,36]. The photosynthetic activity of these blooms depletes the CO₂ concentration, such that the lake becomes undersaturated with CO₂ in summer while the pH rises to values above 9 for several months (Fig. 1, Text S1).

Hence, there is a strong and complex coupling between phytoplankton growth and the inorganic carbon chemistry of aquatic ecosystems that may lead to CO₂ depletion during dense blooms, even in lakes that would otherwise be supersaturated with CO₂. This biological-chemical coupling is further complicated by several additional feedbacks. For instance, dense phytoplankton blooms not only deplete CO₂ and enhance pH but also increase the turbidity of the water column as a result of self-shading, thereby reducing light available for carbon fixation by photosynthesis [31,37]. Moreover, nutrient uptake by dense blooms also affects alkalinity [38–40], which in turn feeds back upon pH and the speciation of dissolved inorganic carbon (DIC). Given the pH and total DIC concentration, it is straightforward to calculate the CO₂, bicarbonate and carbonate concentrations [41–43]. However, we still lack an integrative understanding that incorporates the different feedback loops to enable quantitative prediction of the changes in DIC concentration and pH during phytoplankton bloom development. Yet, such an integrative approach will be required to assess how rising CO₂ concentrations will affect phytoplankton blooms and carbon sequestration in aquatic systems.

In this study, we investigate the dynamic feedbacks between phytoplankton growth, DIC, alkalinity, pH and light during phytoplankton bloom development. Our study specifically focuses on eutrophic and hypertrophic waters, where an excess of mineral nutrients provides ideal conditions for phytoplankton blooms. We incorporate standard inorganic carbon chemistry into a mathematical model of phytoplankton growth with CO₂, bicarbonate and light as limiting resources. We test the model in controlled laboratory experiments at different pCO₂ levels and alkalinities using the harmful cyanobacterium *Microcystis aeruginosa*, a cosmopolitan and often toxic species that develops dense blooms in Lake Volkerak and many other eutrophic lakes worldwide [23–27,31,33]. Our model fits were in good agreement with the experimental results, and show that the coupling between phytoplankton growth and inorganic carbon chemistry is strongly affected by the CO₂ level. Subsequently, we use the experimentally validated model to explore how phytoplankton blooms in eutrophic lakes may respond to rising CO₂ availability.

![Figure 1. Seasonal dynamics of phytoplankton blooms in Lake Volkerak.](image-url)

(A) Changes in phytoplankton population density (strongly dominated by the cyanobacterium *Microcystis*) and measured dissolved CO₂ concentration ([CO₂]) during two consecutive years. The dashed line is the expected dissolved CO₂ concentration ([CO₂*]) when assuming equilibrium with atmospheric pCO₂. Dark shading indicates that the lake is supersaturated with CO₂, while light shading indicates undersaturation. (B) Changes in pH, bicarbonate and total DIC concentration. Sampling details are described in Text S1.

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The Model

General outline

Our model builds upon a long tradition of model studies in phytoplankton ecology [44–49], extending these earlier studies by the incorporation of dynamic changes in inorganic carbon availability, alkalinity and pH induced by phytoplankton blooms. The model considers a well-mixed water column, illuminated from above, with a growing phytoplankton population that is homogeneously distributed over depth. Here we introduce the key assumptions, while the model is described in full detail in Text S2 (for chemostats) and Text S3 (for lakes).

Phytoplankton population dynamics

In this study, we focus on eutrophic and hypertrophic ecosystems where all nutrients are in excess. Hence, the specific growth rate of phytoplankton does not become limited by nutrients but depends only on its cellular carbon content. The cellular carbon content is a dynamic variable, which increases by the photosynthetically-driven uptake of CO2 and bicarbonate, while it decreases by respiration and by dilution of the cellular carbon content due to population growth. More precisely, let X denote the population density of the phytoplankton, and let Q denote its cellular carbon content. Changes in phytoplankton population density and its carbon content can then be described by:

\[
\frac{dX}{dt} = \mu(Q)X - mX
\]

(1)

\[
\frac{dQ}{dt} = u_{CO2} + u_{HCO3} - r - \mu(Q)Q
\]

(2)

where \(\mu(Q)\) is the specific growth rate of the phytoplankton as function of its cellular carbon content, \(m\) is the specific loss rate (e.g., by background mortality, grazing, sedimentation), \(u_{CO2}\) and \(u_{HCO3}\) are the uptake rates of CO2 and bicarbonate, respectively, and \(r\) is the respiration rate.

We assume that the specific growth rate increases with the cellular carbon content of the phytoplankton, which require a minimum cellular carbon content in order to function (i.e., let \(\mu(Q_{MIN}) = 0\)) and reach their maximum specific growth rate when saturated with carbon (i.e., let \(\mu(Q_{MAX}) = \mu_{MAX}\)). Uptake rates of CO2 \(u_{CO2}\) and bicarbonate \(u_{HCO3}\) are increasing but saturating functions of the ambient CO2 and bicarbonate concentration according to Michaelis-Menten kinetics, and are suppressed when cells become saturated with carbon [50]. The energy for carbon assimilation comes from photosynthesis, and therefore depends on light availability. The underwater light environment is described by Lambert-Beer’s law, taking into account that a growing phytoplankton population gradually increases the turbidity of the water column through self-shading and thereby reduces the light available for further photosynthesis [31,51]. We assume that the respiration rate \(r\) increases with the cellular carbon content, approaching maximum values when cells become saturated with carbon [52]. The mathematical equations describing these relationships are presented in Text S2.

To assess to what extent phytoplankton growth is limited by carbon, we introduce a simple relative measure of the inorganic carbon availability \(f_C\):

\[
f_C = \frac{1}{u_{MAX,CO2} + u_{MAX,HCO3}}
\]

(3)

where \(u_{MAX,CO2}\) and \(u_{MAX,HCO3}\) are the maximum uptake rates of carbon dioxide and bicarbonate, respectively, and \(H_{CO2}\) and \(H_{HCO3}\) are their half-saturation constants. We note that \(0 \leq f_C \leq 1\). The level of carbon limitation \(L_C\) can then be defined as the reduction in carbon uptake due to low carbon availability: \(L_C = (1 - f_C) \times 100\%\). Accordingly, if CO2 and bicarbonate are both available in saturating concentrations, \(L_C\) will be close to 0%. Conversely, if CO2 and bicarbonate are available only in trace amounts, \(L_C\) approaches 100%.

Dissolved inorganic carbon, alkalinity and pH

On the timescales used in our model (ranging from minutes to days) the speculation of dissolved inorganic carbon is essentially in equilibrium with alkalinity and pH. Therefore, let [DIC] denote the total concentration of dissolved inorganic carbon. Changes in [DIC] can be described by:

\[
\frac{d[DIC]}{dt} = D([DIC]_{IN} - [DIC]) + \frac{g_{CO2}}{z_{MAX}}(r - u_{CO2} - u_{HCO3})X
\]

(4)

The first term on the right-hand side of Eqn (4) describes changes in the DIC concentration due to the influx \([DIC]_{IN}\) and efflux of water containing DIC, where \(D\) is the dilution rate. The second term describes exchange of CO2 gas with the atmosphere, where \(g_{CO2}\) is the CO2 flux across the air-water interface (also known as the carbon sequestration rate) and division by \(z_{MAX}\) converts the CO2 flux per unit surface into a volumetric CO2 change. The third term describes how the DIC concentration increases through respiration \((r)\) and decreases through uptake of CO2 \((u_{CO2})\) and bicarbonate \((u_{HCO3})\) by phytoplankton.

The CO2 flux across the air-water interface is proportional to the difference between the dissolved CO2 concentration in the water column and in the atmosphere \([CO2^*]\) and the actual dissolved CO2 concentration [53,54]:

\[
g_{CO2} = v([CO2^*] - [CO2])
\]

(5)

where \(v\) is the gas transfer velocity. The equilibrium value \([CO2^*]\) is calculated from Henry’s law, i.e., \([CO2^*] = K_h pCO2\), where \(pCO2\) is the partial pressure of CO2 in air and \(K_h\) is the solubility constant of CO2 gas in water. In our experiments, gas exchange will increase with the gas flow rate \((a)\). Hence, we assume \(v = b a\), where \(b\) is a constant of proportionality reflecting the efficiency of gas exchange.

Changes in pH depend on alkalinity, which is a measure of the acid-neutralizing capacity of water. In our experiments, alkalinity is dominated by dissolved inorganic carbon and inorganic phosphates [40]:

\[
\frac{df}{dt} = \frac{d[DIC]}{dt} - \frac{g}{z_{MAX}}[DIC]
\]
Figure 2. Changes in inorganic carbon chemistry during phytoplankton growth in two chemostat experiments. Left panels: Chemostat experiment with low pCO$_2$ of 200 ppm in the gas flow and 500 µmol L$^{-1}$ bicarbonate in the mineral medium. Right panels: Chemostat experiment with high pCO$_2$ of 1,200 ppm in the gas flow and 2,000 µmol L$^{-1}$ bicarbonate in the mineral medium. Both chemostats were inoculated with...
CO2, bicarbonate and carbonate concentrations, (E, F) total DIC concentration and pH, and (G, H) alkalinity (ALK) and concentrations of dissolved inorganic nitrogen (DIN) and phosphorus (DIP). Symbols represent measurements, lines show the model fits. The model and its parameter values are detailed in Text S2. doi:10.1371/journal.pone.0104325.g002

We note from Eqn (6) that changes in the concentration of dissolved CO2 do not change alkalinity. Furthermore, uptake of bicarbonate for photosynthesis is accompanied by the release of a hydroxide ion or uptake of a proton, and therefore does not change alkalinity either. Hence, carbon assimilation by phytoplankton does not affect alkalinity [40]. However, nitrate, phosphate and sulfate assimilation are accompanied by proton consumption to maintain charge balance, and thus increase alkalinity [38-40]. More specifically, both nitrate and phosphate uptake increase alkalinity by 1 mole equivalent, whereas sulfate uptake increases alkalinity by 2 mole equivalents [40]. Hence, changes in alkalinity can be described as:

$$\frac{dALK}{dt} = D(ALK_{IN} - ALK) + (u_N + u_P + 2u_S) X$$

(7)

where ALKIN is the alkalinity of the water influx, and $u_N$, $u_P$, and $u_S$ are the uptake rates of nitrate, phosphate and sulfate by the growing phytoplankton population. The model keeps track of the nitrate, phosphate and sulfate concentration.

At each time step, the dissolved CO2, bicarbonate and carbonate concentration and pH are calculated from [DIC] and alkalinity (Text S2).

Materials and Methods

Experiments

Experimental system. We tested the model using two strains of the freshwater cyanobacterium *Microcystis aeruginosa*. Strain *Microcystis* CYA140 was obtained from the Norwegian Institute for Water Research (NIVA). Strain *Microcystis* HUB5-2-4 was obtained from the Humboldt University of Berlin, Germany. Both *Microcystis* strains grow as single cell populations. Although all culture equipment was autoclaved prior to the experiments, we were not able to sustain axenic conditions. However, regular microscopic inspection confirmed that abundances of heterotrophic bacteria remained low (<0.1% of the total biomass) for the entire duration of the experiments.

The experiments were carried out in laboratory-built chemostats specifically designed for phytoplankton studies [49,55,56]. Each chemostat consisted of a flat culture vessel illuminated from one side with a constant incident light intensity of $I_{IN} = 50 \pm 1 \mu mol\text{ photons m}^{-2} \text{s}^{-1}$ provided by white fluorescent tubes (Philips PL-L 24W/840/4P, Philips Lighting, Eindhoven, The Netherlands). The chemostats had an optical path length (“mixing depth”) of $z_{MAX} = 5$ cm, and an effective working volume of 1.7 L. The chemostats were supplied with a nutrient-rich mineral medium (DIN) to prevent nutrient limitation during the experiments. Under conditions of nutrient excess, phytoplankton population densities tend to become much higher in laboratory chemostats where phytoplankton is concentrated within only 5 cm depth than in lakes where the phytoplankton population is dispersed over several meters depth [51,58]. This scaling rule implies that nutrient concentrations have to be much higher in mineral media of small-scale laboratory chemostats than in eutrophic lakes to sustain these high population densities. The chemostats were maintained at a constant temperature using a metal cooling finger connected to a Colora thermocryostat, and were aerated with sterilized (0.2 μm Millex-FG Vent Filter, Millipore, Billerica, MA, USA) N2 gas enriched with different CO2 concentrations using Brooks Mass Flow Controllers (Brooks Instrument, Hatfield, PA, USA). The gas mixture was dispersed from the bottom of the chemostat vessel in fine bubbles at a constant gas flow rate ($a$) of 25 L h$^{-1}$.

Treatments. First, we studied dynamic changes in inorganic carbon chemistry and pH in six chemostats without any phytoplankton, to assess whether the model adequately describes the dissolution of CO2 and subsequent dynamic changes in inorganic carbon chemistry. These auxiliary experiments are described in Text S4.

Subsequently, we ran two chemostat experiments with *Microcystis* CYA140 to investigate dynamic changes in phytoplankton growth, inorganic carbon chemistry, alkalinity and pH. The first chemostat was provided with a low pCO2 of 200 ppm in the gas flow and 0.5 mmol L$^{-1}$ NaHCO3 in the mineral medium. The second chemostat was provided with a high pCO2 of 1,200 ppm in the gas flow and 2.0 mmol L$^{-1}$ NaHCO3 in the mineral medium. Both chemostats had a dilution rate of $D = 0.011$ h$^{-1}$. The chemostats were sampled every other day, from the inoculation of a small number of *Microcystis* CYA140 cells to steady state with high population densities.

Next, we studied the steady states of six chemostats of *Microcystis* HUB5-2-4 along a gradient from carbon-limited to light-limited conditions. The chemostats had a dilution rate of $D = 0.00625$ h$^{-1}$, and were provided with different pCO2 concentrations in the gas flow (0.5, 50, 100, 400 or 2,800 ppm CO2) and two different NaHCO3 concentrations in the mineral medium (0.5 or 2.0 mmol L$^{-1}$). The steady states were monitored for at least ten days.

Measurements. The incident light intensity ($I_{IN}$) and the light intensity transmitted through the chemostat vessel ($I_{OUT}$) were measured with a LI-COR LI-250 quantum photometer (LI-COR Biosciences, Lincoln, NE, USA) at 10 randomly chosen positions on the front and back surface of the chemostat vessel, respectively. Background turbidity ($K_b$), was calculated from the light transmission through chemostat vessels without phytoplankton using Lambert-Beer’s law, as $K_b = \ln(I_{IN}/I_{OUT})/z_{MAX}$. DIC concentrations were determined by sampling 15 mL of culture suspension, which was immediately filtered over 0.45 μm membrane filters (Whatman, Maidstone, UK). DIC was subsequently analyzed by phosphoric acid addition on a Model 700 TOC Analyzer (OI Corporation, College Station, TX, USA), with a detection limit of 0.15 ppm. Temperature and pH were measured with a SCHOTTM pH meter (SCHOTT AG, Mainz, Germany). Concentrations of dissolved CO2, bicarbonate and carbonate were calculated from DIC and pH [30], based on the dissociation constants of inorganic carbon corrected for temperature and salinity (Table S2.1 in Text S2). Alkalinity was determined in a 50 mL sample that was titrated in 0.1 to 1 mL.
Figure 3. Trajectories of dissolved CO₂ and population density. Trajectories predicted by the model for chemostats with (A) low pCO₂ of 200 ppm in the gas flow and 500 μmol L⁻¹ bicarbonate in the mineral medium, and (B) high pCO₂ of 1,200 ppm in the gas flow and 2,000 μmol L⁻¹ bicarbonate in the mineral medium. The trajectories start from a series of different initial conditions, and all converge to the same equilibrium point. Arrows indicate the direction of the trajectories. The model assumes species parameters specific for *Microcystis CYA140*, and is detailed in Text S2. doi:10.1371/journal.pone.0104325.g003
steps with 10 mmol L$^{-1}$ HCl to a pH of 3.0. The alkalinity was subsequently calculated using Gran plots [30].

Residual nitrate and phosphate concentrations in the chemostats were determined in triplicate by sampling 15 mL of culture suspension, which was immediately filtered over 0.45 μm membrane filters (Whatman, Maidstone, UK) and the filtrate was stored at −20°C. Nitrate concentrations were analyzed using a Skalar SA 400 autoanalyzer (Skalar Analytical B.V., Breda, The Netherlands), and phosphate concentrations were analyzed spectrophotometrically [59].

Phytoplankton population density, both as cell numbers and total biovolume, was determined in triplicate using a Casy 1 TTC cell counter with a 60 μm capillary (Schaere System GmbH, Reutlingen, Germany). Cell size varied considerably during the experiments, ranging from 31–66 μm$^3$ cell$^{-1}$ in Microcystis CYA140 and from 25–50 μm$^3$ cell$^{-1}$ in Microcystis HUB5-2-4. We therefore used the total biovolume (i.e. the summated volume of all cells per litre of water) as a measure of phytoplankton population density.

Samples for cellular carbon, nitrogen, phosphorus and sulfur content were pressurized at 10 bar to collapse the gas vesicles of Microcystis and subsequently centrifuged for 15 min at 2,000 g. After discarding the supernatant, the pelleted was resuspended in demineralised water, and centrifuged for 5 min at 15,000 g. The supernatant was discarded, pellets were stored at −20°C and subsequently freeze-dried and weighted to determine dry weight. The carbon, nitrogen and sulfur content of homogenised freeze-dried cell powder were analysed using a Vario EL Elemental Analyzer (Elementar Analysensysteme GmbH, Hanau, Germany). To determine the phosphorus content, cells were oxidized with potassium persulfate for 1 h at 100°C [60], and phosphate concentrations were subsequently analyzed spectrophotometrically [59].

To calculate the carbon sequestration rate of the experiments at steady state, we solved Eqns (1), (2) and (4) for zero. This yields:

$$r_{CO2} = z_{MAX}D{(\frac{[DIC]}{[DIC]_{IN}})} - z_{MAX}DQY$$

(8)

where we assumed that the specific loss rate of the phytoplankton was governed by the dilution rate of the chemostat (i.e., $m = D$). This equation shows that, at steady state, the carbon sequestration rate equals the net enhancement of the DIC concentration plus the carbon fixation rate of the phytoplankton population.

Parameter estimation

System parameters such as incident light intensity, mixing depth of the chemostats, composition of the mineral medium, dilution rate and CO2 concentration in the gas flow were measured prior to and/or during the experiments. Some phytoplankton parameters were measured experimentally, while others were estimated from fits of the model predictions to time courses of the experimental variables following the same procedures as in earlier studies [49,55]. An overview of all parameter estimates is given in Text S2.

Extrapolation to lakes

Chemostats provide ideal systems to test models under highly controlled conditions. They operate at the laboratory scale, with parameter settings tuned to the small size of the chemostat vessel. To extrapolate the model predictions to natural waters, we therefore adapted several model assumptions. Phytoplankton parameters were still based on our laboratory experiments with Microcystis HUB5-2-4. However, we used physical and chemical parameter settings typical for the summer situation in eutrophic lakes based on our data from Lake Volkerak, The Netherlands [31]. For instance, the mixing depth was increased from a chemostat of only 5 cm deep to a lake of 5 m deep. The very high phosphate and nitrate concentrations in the mineral medium of the chemostat were reduced to a lower (but still fairly high) phosphate concentration of 15 μmol L$^{-1}$ and nitrate concentration of 150 μmol L$^{-1}$, representative for hypertrophic lakes dominated by cyanobacterial blooms [15,25,31–33]. The high influx of CO2 gas into the chemostat vessel was replaced by a low gas transfer velocity across the air-water interface of lakes [54,61,62]. Full implementation of the lake model is described in Text S2 and Text S3.

Sensitivity analysis

We performed a sensitivity analysis to assess how variation in the model parameters would affect the model predictions. In this analysis, we focus on low-alcaline lakes (ALK$_{IN}=0.5$ mEq L$^{-1}$), since they are more sensitive to rising atmospheric CO2 concentrations than high-alcaline lakes. The sensitivity analysis investigates how the model predictions were affected by variation in two input parameters: (i) the atmospheric CO2 level and (ii) a second model parameter of choice. In contrast to traditional one-factor-at-a-time (OAT) sensitivity analysis, this two-dimensional approach may reveal possible interactions between the two model parameters [63]. For instance, model predictions might be more sensitive to parameter changes at low than at high atmospheric CO2 levels.

In addition, we calculated the normalized sensitivity coefficient ($SC$), which is a local sensitivity index that quantifies the relative change in model output $Y$ with respect to a relative change in input parameter $Z$ [64]:

$$SC = \frac{(ΔY/Y)}{(ΔZ/Z)}$$

(9)

The normalized sensitivity coefficient is dimensionless, and allows comparison between input and output parameters independent of their units of measurement. $|SC| >>1$ implies that the model prediction is very sensitive to a change in the input parameter, whereas $|SC| <<1$ implies that the model prediction is rather insensitive to a change in the input parameter. We based the calculation of $SC$ on a 1% increment of the input parameter. The sensitivity coefficient was calculated at two atmospheric CO2 levels, the present-day level of 400 ppm and an elevated level of...
750 ppm predicted for the year 2150 by the RCP6 scenario of the Fifth Assessment Report of the IPCC [1].

**Results**

Dynamic changes during phytoplankton growth

We studied dynamic changes in inorganic carbon chemistry during the growth of *Microcystis CYA140* in two chemostats that differed with respect to the pCO$_2$ level in the gas flow and the bicarbonate concentration in the medium (Fig. 2). In both chemostats, the population density increased after inoculation, while light penetration ($I_{OUT}$) decreased due to shading by the growing *Microcystis* populations, until steady state was reached after ~30 days (Figs. 2A and 2B). At high pCO$_2$ the population

Figure 5. Steady-state patterns predicted for phytoplankton blooms in low-alkaline lakes. Steady-state predictions of the model evaluated across a wide range of atmospheric pCO$_2$ levels. (A) Phytoplankton population density (expressed as biovolume), (B) light intensity reaching the lake sediment ($I_{OUT}$), (C) dissolved CO$_2$ concentration, (D) bicarbonate concentration, (E) pH, (F) alkalinity, (G) DIC concentration, and (H) carbon sequestration rate. Shading indicates the level of carbon limitation ($L_C$). For comparison, dashed lines show steady-state patterns predicted for low-alkaline waters without phytoplankton. The model parameters are representative for eutrophic low-alkaline lakes (ALK$_{in}$ = 0.5 mEq L$^{-1}$) dominated by the cyanobacterium *Microcystis* HUB5-2-4. The model and its parameter values are detailed in Text S2 and Text S3.

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Figure 6. Contour plots of phytoplankton blooms predicted for different pCO$_2$ levels and alkalinities. Model predictions of (A) the level of carbon limitation, and (B) phytoplankton population density (expressed as biovolume, in mm$^3$ L$^{-1}$). The vertical solid line represents the present-day atmospheric CO$_2$ level of ~400 ppm, while the vertical dashed line shows the atmospheric CO$_2$ level of 750 ppm predicted for the year 2150 by the RCP6 scenario of the Fifth Assessment Report of the IPCC. The model predictions are based on steady-state solutions across a grid of $40 \times 50 = 2,000$ simulations, using the model and parameter values detailed in Text S2 and Text S3.

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Rising CO$_2$ and Phytoplankton Blooms
Figure 7. Sensitivity of the model predictions to variation in phytoplankton traits. Contour plots of the level of carbon limitation ($L_C$) and steady-state phytoplankton population density (right panels, expressed as biovolume, in mm$^3$ L$^{-1}$) predicted for different atmospheric pCO$_2$ levels and phytoplankton traits. The phytoplankton traits are (A, B) the half-saturation constant for CO$_2$ uptake ($H_{\text{CO}_2}$), (C, D) the half-saturation constant for HCO$_3$ ($H_{\text{HCO}_3}$), (E) the maximum CO$_2$ uptake rate ($u_{\text{MAXCO}_2}$), and (F, G) the half-saturation constant for $N$ ($C_N$).
dissolved CO₂ concentration over almost two orders of magnitude, from 10⁻¹⁰⁻¹⁵ mol L⁻¹. The increased alkalinity enhances the storage capacity for bicarbonate as an inorganic carbon source. This unexpected result is counterintuitive, at pCO₂>100 ppm, the bicarbonate concentration became higher in the presence than in the absence of phytoplankton concentration.

Separation of time scales

Because of the relatively high dimensionality of our model, formal mathematical analysis of the existence, uniqueness and stability of the equilibrium point is not straightforward. Therefore, we explored the full phase space of the model by extensive numerical simulations. This did not reveal any indications for alternative stable states or non-equilibrium dynamics. Instead, we always found at most one unique positive equilibrium point that was locally and globally stable whenever it existed. Two examples are given in Fig. 3, where we used the calibrated model to investigate trajectories of dissolved CO₂ and population density from a range of different initial conditions. Interestingly, the trajectories show that the dynamics operated at two distinct time scales: fast chemical dynamics and slow biological dynamics. The inorganic carbon chemistry equilibrated with the standing population density within a few hours, as indicated by the horizontal parts of the trajectories in Fig. 3. These rapid dynamics are consistent with the inorganic carbon chemistry in our chemostat experiments without phytoplankton, which also equilibrated within 1–4 hours (Fig. S4.1 in Text S4). Subsequently, the population density slowly converged to equilibrium within a time span of several weeks. These slow dynamics are indicated in Fig. 3 by the thick curved parts of the trajectories, which ultimately lead to the equilibrium point. Hence, the inorganic carbon chemistry rapidly adjusted to the standing population, and subsequently tracked the slower changes in population density.

Steady-state patterns

We investigated steady-state patterns of phytoplankton abundance and inorganic carbon chemistry using six chemostats of Microcystis HUB3-2-4 (Fig. 4). The steady-state population density increased with pCO₂, demonstrating that it was limited by the supply of inorganic carbon. The population density leveled off when carbon limitation was alleviated at pCO₂>200 ppm (Fig. 4A). At pCO₂ levels <1 ppm, a low DIC concentration of 0.5 mmol L⁻¹ in the mineral medium provided insufficient inorganic carbon, whereas a higher DIC concentration of 2.0 mmol L⁻¹ was sufficient to sustain a steady-state population density. At pCO₂ levels >100 ppm, the influx of DIC supplied by the mineral medium was small compared to the influx of CO₂ supplied by the high gas flow rate, such that the four-fold difference in DIC concentration in the mineral medium had little effect on the steady-state population density.

The increase in population density with rising pCO₂ reduced light penetration through the chemostats (Fig. 4B), which shifted the growth conditions from carbon limitation at low pCO₂ to light limitation at high pCO₂. At pCO₂ levels <100 ppm, phytoplankton strongly depleted the dissolved CO₂ concentration to a stable level of ~0.1 µmol L⁻¹ (Fig. 4C), while pH was maintained at values around 10 (Fig. 4E). At pCO₂ levels >100 ppm, the dissolved CO₂ concentration increased and pH decreased with increasing pCO₂ (Fig. 4C, E). The pH remained consistently higher in the presence than in the absence of phytoplankton.

Counterintuitively, at pCO₂>100 ppm, the bicarbonate concentration became higher in the presence than in the absence of phytoplankton (Fig. 4D), even though phytoplankton consume bicarbonate as an inorganic carbon source. This unexpected result is caused by the shift in pH in combination with an increase in alkalinity associated with uptake of nitrate, phosphate and sulfate by the phytoplankton population (Fig. 4F; see also Eqn (7)). An increased alkalinity enhances the storage capacity for bicarbonate and carbonate in the system. The alkalinity, bicarbonate concentration and total DIC concentration all showed a similar increase with rising pCO₂ as the phytoplankton population density (compare Figs. 4D, F, G with Fig. 4A). At pCO₂>200 ppm, 70–80% of the total amount of carbon in the system was in phytoplankton biomass while 20–30% of the total carbon was DIC.

The carbon sequestration rate also showed a similar increase with rising pCO₂ as the DIC concentration and phytoplankton
Figure 8. Sensitivity of the model predictions to variation in lake properties. Contour plots of the level of carbon limitation ($L_C$) and steady-state phytoplankton population density ($X$) predicted for different atmospheric pCO$_2$ levels and lake properties. The lake properties are (A, B) lake depth ($z_{\text{MAX}}$), (C, D) CO$_2$ gas transfer velocity ($v$), (E, F) DIC concentration of the influx ([DIC]$_{\text{IN}}$), (G) salinity (g L$^{-1}$).
population density, and leveled off when the population approached its maximum productivity at >200 ppm (Fig. 4H). The model fits were in good agreement with the observed steady-state patterns in phytoplankton population density, inorganic carbon availability, alkalinity and pH along the entire CO2 gradient.

**Extrapolation to lakes**

The model was adapted to natural waters to explore the impact of rising atmospheric CO2 levels on phytoplankton blooms in lakes. Although it is difficult to capture the complex dynamics of natural systems, such a modelling exercise may help in understanding the coupling between phytoplankton blooms and inorganic carbon chemistry. As a first step, we investigated steady-state patterns of phytoplankton abundance in low-alkaline lakes, where bicarbonate concentrations are low and phytoplankton growth therefore largely depends on dissolved CO2 as a carbon source. This is a similar situation as in our chemostat experiments, and the model predictions for low-alkaline lakes are therefore qualitatively similar to the results obtained in our chemostats (compare Fig. 4 and Fig. 5). The phytoplankton population can be sustained at pCO2 levels above 0.17 ppm, and is predicted to increase strongly with pCO2 (Fig. 5A). Above 2,000 ppm, a further rise of the pCO2 level no longer enhances the population density, because the high CO2 supply in combination with self-shading in dense phytoplankton blooms has shifted phytoplankton growth from carbon-limited to light-limited conditions (Fig 5B).

Over a wide range of pCO2 levels, from 0.17 to 1,000 ppm, phytoplankton blooms exert strong control over the dissolved CO2 concentration and pH, depleting the dissolved CO2 concentration below 0.1 μmol L\(^{-1}\) and raising pH to 10 (Fig. 5C, E). The bicarbonate and total DIC concentration are reduced by the phytoplankton population for pCO2 levels ranging from 0.17 to 1,400 ppm CO2 (Fig. 5D, G). The bicarbonate concentration, total DIC concentration, alkalinity and carbon sequestration rate all increase with rising pCO2, and level off when the phytoplankton population approaches maximum densities (Fig. 5D, F-H). Above 1,000 ppm, phytoplankton blooms exert less control over CO2 availability and pH, and the dissolved CO2 concentration increases while pH decreases with a further rise in pCO2 (Fig. 5C, E).

Figure 6 summarizes the level of carbon limitation and the population density predicted for dense phytoplankton blooms in different eutrophic waters spanning a wide range of alkalinitities and pCO2 levels. In line with expectation, the model predicts that carbon limitation of dense phytoplankton blooms will be most pronounced in low-alkaline waters, where CO2 provides the main inorganic carbon source (Fig. 6A). Rising atmospheric pCO2 levels are expected to lead to a strong increase in phytoplankton population density in these low-alkaline waters (Fig. 6B). In lakes with a moderate alkalinity, where bicarbonate can partially supplement growth when CO2 is depleted, carbon limitation is predicted to be less intense but may still play a substantial role (i.e., LC = 10-50%; Fig. 6A). In high-alkaline waters and soda lakes, however, carbon will rarely be limiting at ambient atmospheric pCO2 levels (Fig. 6A). Their large DIC pools provide a sufficient supply of CO2 and bicarbonate to produce high phytoplankton population densities at ambient pCO2 levels (Fig. 6B).

**Sensitivity analysis**

**Phytoplankton traits.** As a first step, we investigated the sensitivity of the model predictions to variation in the half-saturation constant for CO2 uptake (Fig. 7A, B). Note that an increase of the half-saturation constant implies a reduced affinity. All else being equal, an increase in the half-saturation constant for CO2 therefore leads to stronger carbon limitation and lower phytoplankton population densities (Fig. 7A, B). The normalized sensitivity coefficients were small, both at 400 and at 750 ppm (Table 1). A value of SC = 0.10 implies that for a 1% increase in the half-saturation constant, the model predicts only a 0.1% increase in the level of carbon limitation. Hence, the sensitivity of the model predictions to variation in the half-saturation constant for CO2 uptake is relatively low.

The half-saturation constant for bicarbonate shows a similar pattern (Fig. 7C, D).

An increase in the maximum uptake rate of CO2 causes stronger CO2 depletion during phytoplankton blooms, which results in stronger carbon limitation and higher population densities (Fig. 7E, F). Interestingly, comparison of the sensitivity coefficients indicates that changes in the maximum uptake rate of CO2 have a larger effect on the level of carbon limitation than on the phytoplankton population density (Table 1).

Changes in the C:N stoichiometry of phytoplankton cells do not directly affect the growth rates in our model, because we assumed that all nutrients are available at saturating levels. Changes in cellular C:N stoichiometry may have a small indirect effect, however, because nitrate uptake affects alkalinity, and thereby inorganic carbon availability. Hence, as expected, the model predictions are rather insensitive to changes in cellular C:N stoichiometry (Fig. 7G, H; Table 1).

**Lake properties.** Lake depth has strong effects on the model predictions. In deep lakes, the phytoplankton population is spread out over a large water volume, and will be light-limited in deeper parts of the water column. Hence, all else being equal, CO2 depletion in deep lakes will be less intense, resulting in lower levels of carbon limitation than in shallow lakes (Fig. 8A). Phytoplankton population densities are therefore predicted to respond more strongly to rising pCO2 levels in shallow than in deep lakes (Fig. 8B).

The CO2 gas transfer velocity across the air-water interface varies with wind speed and precipitation events [54,61,62]. An increase in CO2 gas transfer velocity strongly reduces the level of carbon limitation and increases the phytoplankton population density (Fig. 8C, D). Interestingly, the sensitivity coefficients point at an interactive effect with the atmospheric CO2 level. The model predictions become more sensitive to changes in CO2 gas transfer velocity at higher atmospheric CO2 levels (Table 1).

Enhanced mineralization of organic carbon in the sediment or additional CO2 input from the surrounding watershed may cause an enhanced CO2 influx into the lake. In our model this would be represented by an increase in DIC influx without a change in alkalinity. Such an enhanced CO2 influx reduces the level of carbon limitation, thereby raising phytoplankton population density (Fig. 8E, F). The sensitivity coefficients indicate that the
drawdown of the dissolved CO\textsubscript{2} concentration by dense phyto-
aquatic ecosystems. Biological enhancement is due to the
influx of CO\textsubscript{2} chemically reacts with water, and is transferred
to growth \cite{15,67}. Chemical enhancement occurs because part of
the resultant influx of atmospheric CO\textsubscript{2} can further fuel phytoplank-
ton blooms, which enlarges the CO\textsubscript{2} concentration gradient
during bloom development \cite{38–40}. Increases in pH and alkalinity
in many freshwater lakes are sufficiently high to cover the carbon
demands of phytoplankton populations. Indeed, in addition to
bicarbonate and carbonate \cite{68}. This chemical enhancement is
promoted by the high pH and alkalinity induced by phytoplankton
blooms, which enlarge the DIC storage capacity of aquatic
ecosystems.

Interestingly, our laboratory experiments show that the
enhanced CO\textsubscript{2} influx induced by dense phytoplankton
populations can even raise the bicarbonate and total DIC
centration \cite{2E, 2F, 4G}. This may seem counterintuitive, because
phytoplankton populations consume inorganic carbon. However,
the high pH and alkalinity in phytoplankton blooms favors the
formation of bicarbonate and carbonate. Depending on the
interplay between CO\textsubscript{2} gas transfer, inorganic carbon uptake,
alkalinity and pH, this can result in either a decrease or increase in
total DIC concentration. The lake model predicts that dense
phytoplankton blooms may increase the bicarbonate and DIC
centration in lakes, but only at very high pCO\textsubscript{2} levels. At pCO\textsubscript{2}
levels below 1,400 ppm, the lake model predicts a reduced
bicarbonate and DIC concentration during phytoplankton blooms
\cite{5D, G}, which is supported by our observations from Lake
Volkerek (Fig. 1B).

### Discussion

**Coupling between phytoplankton blooms and inorganic
carbon chemistry**

Our theoretical and experimental results demonstrate that the
development of dense algal blooms can dramatically change the
dissolved CO\textsubscript{2} concentration, alkalinity and pH of aquatic ecosystems. In our experiments, phytoplankton growth induced a strong CO\textsubscript{2} drawdown, especially when provided with a low
pCO\textsubscript{2} level in the gas flow. Assimilation of CO\textsubscript{2} and nutrients such
as nitrate, phosphate and sulfate increased alkalinity and pH
during bloom development \cite{38–40}. Increases in pH and alkalinity
shifted the inorganic carbon composition towards bicarbonate and
carbonate. These findings are in good agreement with field
observations, as similar changes in DIC speciation, pH, and
alkalinity have also been documented in studies of dense
phytoplankton blooms in natural waters (Fig. 1) \cite{11,12,15}.

Dense phytoplankton blooms contribute to both ‘biological
enhancement’ and ‘chemical enhancement’ of the CO\textsubscript{2} influx into
aquatic ecosystems. Biological enhancement is due to the
drawdown of the dissolved CO\textsubscript{2} concentration by dense phyto-
plankton blooms, which enranges the CO\textsubscript{2} concentration gradient
across the air-water interface. Hence, dense phytoplankton blooms
can turn aquatic ecosystems into net carbon sinks, and the
resultant influx of atmospheric CO\textsubscript{2} can further fuel phytoplank-
ton growth \cite{13,67}. Chemical enhancement occurs because part of
the influx of CO\textsubscript{2} chemically reacts with water, and is transferred
to bicarbonate and carbonate \cite{68}. This chemical enhancement is
promoted by the high pH and alkalinity induced by phytoplankton
blooms, which enlarge the DIC storage capacity of aquatic
ecosystems.

### Table 1. Normalized sensitivity coefficients of selected model parameters at atmospheric CO\textsubscript{2} levels of 400 ppm (SC\textsubscript{400}) and 750 ppm (SC\textsubscript{750}).

<table>
<thead>
<tr>
<th>Parameter</th>
<th>Description</th>
<th>Level of carbon limitation</th>
<th>Population density</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>SC\textsubscript{400}</td>
<td>SC\textsubscript{750}</td>
</tr>
<tr>
<td>Species traits</td>
<td></td>
<td></td>
<td></td>
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<tr>
<td>$K_{CO2}$</td>
<td>Half-saturation constant for CO\textsubscript{2} uptake</td>
<td>0.07</td>
<td>0.10</td>
</tr>
<tr>
<td>$K_{HCO3}$</td>
<td>Half-saturation constant for bicarbonate uptake</td>
<td>0.03</td>
<td>0.03</td>
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<tr>
<td>$u_{MAX, CO2}$</td>
<td>Maximum uptake rate of CO\textsubscript{2}</td>
<td>0.56</td>
<td>0.88</td>
</tr>
<tr>
<td>$cH$</td>
<td>Cellular N:C ratio</td>
<td>0.04</td>
<td>0.08</td>
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<td>Lake properties</td>
<td></td>
<td></td>
<td></td>
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<td>$z_{MAX}$</td>
<td>Lake depth</td>
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<td>-0.65</td>
</tr>
<tr>
<td>$v$</td>
<td>Gas transfer velocity of CO\textsubscript{2}</td>
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<td>-0.63</td>
</tr>
<tr>
<td>$[DIC]_{IN}$</td>
<td>Concentration of DIC at influx</td>
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<td>-0.84</td>
</tr>
<tr>
<td>$Sal$</td>
<td>Salinity</td>
<td>0.00</td>
<td>-0.01</td>
</tr>
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The normalized sensitivity coefficient expresses the relative change in model output with respect to a relative change in input parameter. We used several species traits and lake properties as input parameters, and the level of carbon limitation and phytoplankton population density as model output.

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carbon source [8–10]. However, utilization of bicarbonate requires additional investments in, e.g., sodium-dependent and ATP-dependent bicarbonate uptake systems and carbonic anhydrases [9,10]. The costs of bicarbonate utilization may therefore have repercussions for the growth rates that can be achieved. *Synechococcus leopoliensis*, for instance, grows at ~80% of its maximum growth rate when provided with bicarbonate as its main carbon source [70]. Our parameter estimates indicate that *Microcystis* CYA 140 grows at ~50% while *Microcystis* HUB5-2-4 can only grow at 35% of its maximum growth rate on bicarbonate alone (Table S2.3 in Text S2). This is supported by the chemostat experiments. For instance, *Microcystis* HUB5-2-4 could just barely sustain a low population density when CO₂ was largely removed from the gas flow, even though bicarbonate was provided at a saturating concentration of 2,000 μmol L⁻¹ in the mineral medium (see the datapoint at 0.5 ppm pCO₂ in Fig. 4A). For both strains, an increase in pCO₂ level led to a clear increase in population density (Figs. 2 and 4). Hence, our experiments demonstrate that, even for cyanobacteria with their sophisticated carbon-concentrating mechanisms, increasing pCO₂ levels in bicarbonate-rich waters can cause an increase in phytoplankton population density.

In line with expectation, our model predicts that the potential for carbon limitation strongly depends on alkalinity (Fig. 6). This is consistent with studies in natural waters. Carbon limitation is often observed during algal blooms in eutrophic low-alkaline lakes, where CO₂ is the main inorganic carbon source [11,71]. Carbon limitation has also been reported for moderately alkaline lakes (Fig. 1) [12,13,72], where bicarbonate partially supplements phytoplankton growth when CO₂ is depleted. The model predicts that carbon limitation will be almost absent in high-alkaline waters and soda lakes, owing to their high inorganic carbon availability (Fig. 6). Indeed, tropical soda lakes are widely recognized to be among the world’s most productive ecosystems, and can sustain extremely dense populations of cyanobacteria [73,74].

Only high nutrient loads can sustain phytoplankton blooms dense enough to deplete the dissolved CO₂ concentration and induce carbon limitation [75]. In an analysis of 131 eutrophic lakes in the Midwestern USA, Balmer and Downing [15] showed that dissolved CO₂ decreased below atmospheric equilibrium when total phosphorus (TP) concentrations exceeded 1–2 μmol L⁻¹ and chlorophyll a levels exceeded 10–20 μg L⁻¹. Severe CO₂ depletion occurred at chlorophyll concentrations exceeding 80–100 μg L⁻¹. This matches our data from Lake Volkerak, which has a summer TP concentration of ~3 μmol L⁻¹ [31], and where the dissolved CO₂ concentration became undersaturated at chlorophyll concentrations exceeding 20 μg L⁻¹ and was severely depleted during the height of the blooms (Fig. 1A). Such conditions also seem to be representative of several other eutrophic and hypertrophic lakes with dense phytoplankton blooms. For example, TP concentrations exceeding 2 μmol L⁻¹ are also found in Lake Taihu in China [25], Lake Victoria in East Africa [76], the western part of Lake Erie, USA [23,77], the southern part of Lake Peipsi on the border of Estonia and Russia [78,79], and several smaller lakes and reservoirs [33,71,80], all of which have suffered from dense cyanobacterial blooms in summer. This indicates that the nutrient availability in these eutrophic and hypertrophic lakes is, at least in potential, high enough for dense phytoplankton blooms to induce carbon-limited conditions.

**Model limitations**

Combining models and experiments has several advantages. It allows quantitative analysis of the different processes under controlled conditions. Furthermore, it ensures that model predic-
Effects of rising CO₂

Our results support earlier reports that rising atmospheric pCO₂ levels are likely to promote phytoplankton productivity in eutrophic waters [89,90]. More specifically, both our model predictions and laboratory experiments indicate that elevation of the atmospheric pCO₂ will enhance the CO₂ influx across the air-water interface during dense phytoplankton blooms, which in turn will tend to further increase the population density of these blooms.

Our study shows that dense phytoplankton blooms are capable to deplete the dissolved CO₂ concentration and increase the pH over a relatively wide range of atmospheric CO₂ levels (Fig. 4C, E, Fig. 5C, E). Yet, at some point, rising atmospheric CO₂ levels will alleviate phytoplankton blooms from carbon limitation, such that they will exert much less control over the dissolved CO₂ concentration and pH. Beyond this point, phytoplankton blooms become carbon-saturated and a further rise in atmospheric CO₂ levels will not enhance population densities but will lead to less intense CO₂ depletion and a lower pH during the blooms. Our results indicate that the atmospheric CO₂ level at which blooms become carbon-saturated is likely to vary among phytoplankton species depending on e.g. their carbon concentrating mechanisms [8–10], and among ecosystems depending on e.g. alkalinity, lake depth and CO₂ input from the mineralization of dissolved organic carbon (Figs. 6–8). For instance, in our chemostat experiments the population density increased for atmospheric pCO₂ levels from 10 to 200 ppm, while the Microcystis population became carbon-satured at pCO₂ levels beyond 200 ppm (Fig. 4A). In contrast, for low-alkaline lakes, our baseline model predicts a strong increase in phytoplankton population density from 100 to 1,500 ppm, while the transition to carbon-satured Microcystis blooms is predicted to occur only when the pCO₂ level exceeds 1,500 ppm (Fig. 5A).

Intensification of phytoplankton blooms by rising CO₂ levels may further aggravate the problems associated with these blooms, such as anoxia and fish kills [82,91] and the loss of submerged vegetation due to an increased turbidity [83,84]. In particular, enhanced bloom formation by toxic phytoplankton species, like the cyanobacterium Microcystis of this study, can pose health risks for humans and animals, and may threaten the use of eutrophic waters for recreation, fisheries, drinking water and agricultural irrigation [21–25].

Conclusions

The interplay between data-driven models and model-driven experimentation explored by our study may offer an important step towards an improved quantitative understanding and prediction of the impacts of rising CO₂ on phytoplankton blooms. Our results demonstrate that, under controlled laboratory conditions, the coupling between phytoplankton growth, CO₂ drawdown and the inorganic carbon chemistry of aquatic systems can be captured by a relatively simple model. Extrapolation of this experimentally validated model to lakes warns that rising CO₂ levels are likely to intensify phytoplankton blooms, especially in low to moderately alkaline waters with high nutrient loads.

Supporting Information

Text S1 Sampling of Lake Volkerak. Description of the lake, sampling method and analysis of the data displayed in Fig. 1. (PDF)

Text S2 Model description and parameter estimation. Full description of the model used to predict the population dynamics, inorganic carbon chemistry, pH and alkalinity in the chemostat. Solubility and dissociation constants of dissolved inorganic carbon and phosphate are given in Table S2.1 in this text. System parameters and phytoplankton parameters are listed in Tables S2.2 and S2.3, respectively. (PDF)

Text S3 Adaptation of the model to lakes. Detailed description of the extended model formulation to apply the model to lakes. (PDF)

Text S4 Dynamic changes in inorganic carbon chemistry and pH without phytoplankton. Auxiliary experiments investigating dynamic changes in inorganic carbon chemistry and pH in six chemostats without phytoplankton. The experimental data and model fits are displayed in Fig. S4.1, and the estimated parameter values are given in Table S4.1 in this text. (PDF)

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Author Contributions

Conceived and designed the experiments: JMHV DBVDW PMV EVD JH. Performed the experiments: JMHV DBVDW JFF. Analyzed the data: JMHV DBVDW JH. Contributed reagents/materials/analysis tools: EVD JH. Wrote the paper: JMHV DBVDW JH. Designed the model: JMHV JH.

References


