A salt-bridge structure in solution revealed by 2D-IR spectroscopy

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A salt-bridge structure in solution revealed by 2D-IR spectroscopy†

Adriana Huerta-Viga, Sérgio R. Domingos, Saeed Amirjalayer and Sander Woutersen*

Salt bridges are important interactions for the stability of protein conformations, but up to now it has been difficult to determine salt-bridge geometries in solution. Here we characterize the spatial structure of a salt bridge between guanidinium (Gdm⁺) and acetate (Ac⁻) using two-dimensional vibrational (2D-IR) spectroscopy. We find that as a result of salt bridge formation there is a significant change in the infrared response of Gdm⁺ and Ac⁻, and cross peaks between them appear in the 2D-IR spectrum. From the 2D-IR spectrum we determine the relative orientation of the transition-dipole moments of the vibrational modes of Gdm⁺ and Ac⁻, as well as the coupling between them.

Salt bridges are hydrogen-bonded ion pairs that are important for the stabilization of molecular conformations. Intermolecular salt bridges are used to create new supramolecular systems, while biological systems such as proteins often contain salt bridges that determine their structure and function. To understand the chemical functionality of salt bridges, it is essential to detect salt-bridged ion pairs and to characterize their solvated structure, which is unfortunately difficult with conventional methods such as NMR. Here we report the study of a biologically-relevant ion pair formed by guanidinium (Gdm⁺) and acetate (Ac⁻), using two-dimensional infrared (2D-IR) spectroscopy. This technique is used to measure the vibrational coupling between transition-dipole moments of individual vibrations, and through a quantitative analysis of the 2D-IR spectrum, structural information can be obtained.

The Gdm⁺···Ac⁻ ion pair is a model for salt bridges between the Gdm⁺ side group of arginine and the carboxylate side group of glutamate or aspartate, which occur commonly in proteins. The molecular structure of this ion pair is shown in Fig. 1A. Isolated Gdm⁺ has D₃ symmetry, and has a degenerate mode due to a combined CN₃ antisymmetric stretch and NH₂ scissors motion at a frequency of 1600 cm⁻¹. This degeneracy is observed in aqueous solution, but it is broken in viscous solvents. When dissolving deuterated Gdm⁺ (guanidine-DCl, >98% purity, 400 mM) in deuterated dimethylsulfoxide (DMSO), we observe a splitting between the frequencies of the two CN₃D₆ modes, as can be seen in Fig. 1B. In the following, we refer to the high- and low-frequency CN₃D₆ modes of Gdm⁺ as GdmHF and GdmLF, respectively. Interestingly, when an equimolar amount of Ac⁻ ions is added to the solution (guanidine acetate salt, >98% purity, 400 mM), the splitting between GdmHF and GdmLF becomes larger, as seen in Fig. 1B. Gdm⁺ and Ac⁻ have a strong binding affinity in DMSO, forming more than 98% dimers at the concentration used in our experiments, which suggests that the larger splitting between the Gdm⁺ modes is due to an interaction with the Ac⁻ ion. Moreover, Ac⁻ (tetrabutylammonium acetate, >97% purity) has an absorption band at 1580 cm⁻¹ in DMSO (shown in Fig. 1B) due to the COO⁻-antisymmetric-stretch mode. This mode red-shifts after dimer formation with Gdm⁺. The change in the infrared response of both the Gdm⁺ and the Ac⁻ ions upon dimer formation strongly suggests that there is a coupling between the vibrational modes of these two molecules.

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Fig. 1 (A) Molecular structure of the Gdm⁺···Ac⁻ dimer obtained using ab initio methods. The corresponding transition-dipole moments of the COO⁻-strecth mode (1) and of the CN₃D₆ low- and high- frequency modes (2 and 3, respectively) are indicated by arrows. (B) Infrared absorption spectrum of Gdm⁺···Ac⁻, Gdm⁺ and Ac⁻ in DMSO (solvent subtracted). Shifts of the COO⁻ and high-frequency CN₃D₆ bands are indicated by arrows.
The 2D-IR spectra of the Gdm$^+\cdot\text{Ac}^-$ dimer confirm unambiguously that the GdmHF$^+$ and GdmLF$^+$ modes are both coupled to the COO$^-$ stretch mode of Ac$^-$. To measure the 2D-IR spectra, we use a femtosecond pump–probe setup that has been described elsewhere. The 2D-IR spectra are shown in Fig. 2B and C for parallel and perpendicular polarization of the pump and probe pulses, and weighted difference $3\Delta\omega_1 - \Delta\omega_2$, respectively. The non-zero off-diagonal response (cross peaks) in the 2D-IR spectra indicates that there is a coupling between the COO$^-$ stretch mode and each of the Gdm$^+$ modes. These cross peaks can be seen better in slices along both the pump and probe axes of the 2D-IR spectra. Fig. 3A shows cross sections of $\nu_{\text{pump}} = \nu_{\text{COO}^-}$ for parallel and perpendicular polarization of the pump and probe pulses. In the cross sections, the cross peaks between the COO$^-$ stretch mode and each of the Gdm$^+$ modes are clearly visible. Note that the diagonal signal of the COO$^-$ stretch mode has a smaller magnitude than these cross peaks because of its smaller absorption cross section (see Fig. 1B). Fig. 3B shows a cross section along the probe axis for $\nu_{\text{probe}} = \nu_{\text{COO}^-}$ for parallel polarization of pump and probe pulses. The negative part at 1550 cm$^{-1}$ is due to the bleaching and $\nu = 1 \rightarrow 0$ stimulated emission of the COO$^-$ stretch mode on the diagonal of the 2D spectrum. The positive region centered at 1580 cm$^{-1}$ is the low-probe-frequency tail of the excited-state absorption of the GdmHF$^+$ mode on the diagonal. The negative region at 1610 cm$^{-1}$ is the negative part of the cross peak between the COO$^-$ stretch mode and the GdmHF$^+$ mode. Note that the cross peaks between GdmHF$^+$ and Ac$^-$ are more clearly visible in the parallel-polarization 2D-IR spectrum (Fig. 2B), and almost vanish in the $3\Delta\omega_1 - \Delta\omega_2$ difference spectrum (Fig. 2D). This behaviour indicates that the transition-dipole moments of these modes are approximately parallel. In contrast, the cross peaks between GdmLF$^+$ and Ac$^-$ appear clearly in the $3\Delta\omega_1 - \Delta\omega_2$ difference spectrum (Fig. 2D), which suggests that their transition-dipole moments are at a nonzero angle.

In order to obtain structural information from the 2D-IR spectra, we use a vibrational-exciton model (see ESI† for a description of the model). It has been shown before that in addition to coupling between the two CN$_3$D$_6^+$ modes, there is also energy transfer between them, but in DMSO this process is slow enough to be neglected at the delay at which we measured the 2D-IR spectra used for the structural analysis (1.5 ps, see ESI†). We determine the spectral and structural parameters from a global fit to the parallel and perpendicular 2D-IR spectra. Fig. 2F and G show the fitted 2D-IR spectra, which are in very good agreement with the measured ones. Note that the difference spectrum, shown in Fig. 2D is not fitted but the corresponding calculated spectrum, shown in Fig. 2H, is also in very good agreement with the measurement. We find that in order to obtain a good fit, an inhomogeneity in the coupling between GdmLF$^+$ and GdmHF$^+$ modes has to be taken into account in the model.
Table 1  Coupling β and angle θ between the transition-dipole moments of the COO−stretch mode of Ac− and the two CN3H6+ modes of Gdm+ as obtained from the fit (γ=0.4). See ESI for estimates of the confidence limits.

<table>
<thead>
<tr>
<th>Modes</th>
<th>β (cm⁻¹)</th>
<th>θ (°)</th>
</tr>
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<tbody>
<tr>
<td>GdmHF+–GdmLF+</td>
<td>−8</td>
<td>105</td>
</tr>
<tr>
<td>GdmHF+–Ac−</td>
<td>−11</td>
<td>180</td>
</tr>
<tr>
<td>GdmHF+–Ac−</td>
<td>8</td>
<td>75</td>
</tr>
</tbody>
</table>

The couplings between the three modes involved in the salt bridge (GdmHF+, GdmLF+ and Ac−), and the angles between their transition-dipole moments were parameters of the fit and are listed in Table 1 (see ESI† for a complete list of fit parameters). The coupling between GdmHF+ and GdmLF+ modes, which involve oscillations of the same CN and NH bonds, is ≈8 cm⁻¹. Remarkably, the couplings between the COO−stretch mode and the GdmHF+,LF+ modes are of a similar magnitude, (−11 and 8 cm⁻¹, respectively) even though these vibrational modes share neither atoms nor bond: these couplings originate purely from the salt bridge between the two molecules. The large couplings between the COO− and the GdmHF+,LF+ modes show that the frequency shifts observed in the IR spectrum upon salt-bridge formation (Fig. 1A) are mostly due to splitting of the coupled modes rather than due to a change in the local-mode frequencies. This is confirmed by the fact that the COO−peak shifts much less when acetate forms a salt bridge with protonated Gdm+, in which the two CN3H6+ modes are at a higher frequency (see ESI†). The angles between the transition-dipole moments of the salt-bridged vibrational modes are in good agreement with the planar geometry shown in Fig. 1A. The planarity of a salt bridge is often taken as a metric of its quality in X-ray studies,13 the planar geometry shown in Fig. 1A. The planarity of a salt

References

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