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Distinct Effects of Sorbic Acid and Acetic Acid on the Electrophysiology and Metabolism of *Bacillus subtilis*

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Sorbic acid and acetic acid are among the weak organic acid preservatives most commonly used to improve the microbiological stability of foods. They have similar pKₐ values, but sorbic acid is a far more potent preservative. Weak organic acids are most effective at low pH values. Under these circumstances, they are assumed to diffuse across the membrane as neutral undissociated acids. We show here that the level of initial intracellular acidification depends on the concentration of undissociated acid and less on the nature of the acid. Recovery of the internal pH depends on the presence of an energy source, but acidification of the cytosol causes a decrease in glucose flux. Furthermore, sorbic acid is a more potent uncoupler of the membrane potential than acetic acid. Together these effects may also slow the rate of ATP synthesis significantly and thus (partially) explain sorbic acid’s effectiveness.

Various small weak organic acids (WOAs) have been used as food preservatives for a very long time. These weak acids slow the growth of various spoilage bacteria, yeasts, and molds without having overt undesired effects on taste or being toxic to the consumer. The undissociated states of the WOA preservatives are more effective in slowing growth than the dissociated form, although the latter may have some level of toxicity. As such, WOAs are most effective when applied at low pH values below their pKₐ value (1, 2). Under these conditions, the neutral acid is assumed to diffuse across the plasma membrane and dissociate in the cytosol, which generally has a higher pH. In this way, the proton gradient over the membrane is depleted and the anion may accumulate to potentially toxic levels inside the cell. This is known as the classical weak-acid preservative theory (3). Commonly used WOA preservatives include sorbic acid and acetic acid, which have a similar pKₐ value of 4.76 but dissimilar octanol/water partition coefficients (log Kᵣₑₐ values) of 1.33 and −0.17, respectively (International Programme on Chemical Safety [http://www.inchem.org/]). This means that at a particular pH and the same total concentration, the concentrations of both undissociated acids are the same, but sorbic acid has a higher affinity for a hydrophobic (membrane) environment. Sorbic acid is clearly the more potent preservative of the two, but the exact reason why is still not fully clear (4, 5).

It is important to distinguish the different modes of action that WOAs may have on cell physiology. The classical weak-acid preservative theory (3) assumes entry only of the undissociated acid, dissociation in the cytosol, and cytosolic acidification. While this is sometimes described as uncoupling, we use the latter phrase for compounds that shuttle protons across the membrane and are thus protonophoric uncouplers. Others have also pointed to the potential toxicity of accumulated anions (1, 6, 7). If we assume that only the undissociated acid diffuses across the membrane, it follows from the Henderson-Hasselbalch equation (8) that the pH difference over the membrane (ΔpH) is equal to log([A⁻]ₜₒᵤₖ/[A⁻]ₜᵢₙₙ), where [A⁻]ₜₒᵤₖ is the concentration of the anion outside the cell and [A⁻]ₜᵢₙₙ is the concentration of the anion inside the cell, and thus, the internal concentration of the anion may become very high when a high ΔpH remains present. Also, specific binding of sorbate to cysteine (9) has been shown and is proposed to be an explanation for the higher toxicity of sorbic acid.

Several mechanisms of resistance against WOAs have been reported for yeasts, like *Saccharomyces cerevisiae*, *Zygosaccharomyces bailii*, and *Zygosaccharomyces rouxii*. These organisms induce the expression of H⁺ ATPases to regulate their cytosolic pH. *S. cerevisiae* uses a dedicated ATP binding cassette (ABC) transporter (Pdr12) to prevent accumulation of the anion (5). Also, *S. cerevisiae* plasma membrane components likely play an important role in the modulation of the influx of lipophilic weak organic acids (10, 11). Furthermore, *Z. rouxii* and *Aspergillus niger* have been shown to degrade sorbic acid to 1,3-pentadiene (12, 13).

The responses and potential resistance mechanisms against weak acids have not been as well described for bacteria as they have been for yeasts (14, 15). The level of growth reduction was modeled some 30 years ago (1), and the effect of sorbic acid on the membrane potential in *Escherichia coli* membrane vesicles has been described (16). More recently, Ter Beek et al. (2) performed a microarray study on *Bacillus subtilis* cells exposed to sorbic acid. Their results showed a broad transcriptomic response resembling a pattern typical of that for cells responding to nutrient limitation. The authors observed an upregulation of genes encoding potential efflux pumps as well as genes involved in remodeling of the plasma membrane.

*Bacillus subtilis* has been the model Gram-positive organism for decades, because it is generally recognized as safe (GRAS), it is genetically accessible, and it has a fully sequenced genome. It also forms heat-resistant spores and as such is a recognized spoilage

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organism (17). Spores of several related bacterial species are of great concern to the food industry because they are highly resistant to most preservation techniques and, once germinated, can cause food spoilage through the growth of vegetative cells that may produce toxins (15).

To comprehensively elucidate the physiological effects of WOAs on B. subtilis, we measured the effects of sorbic acid and acetic acid on the chemical ($\Delta$pH) and electrical potential ($\Delta$\psi) components of the proton motive force (PMF). With a depleted PMF, we speculated that the cell might alter its energy needs in terms of glucose consumption and the availability of its terminal electron acceptor (i.e., O2), which were hence determined. We assessed the rate and extent of the change in the internal pH (pHin) caused by these two weak acids using B. subtilis cells that either were directly exposed to both WOAs or had been preexposed to sorbic acid and reexposed to WOAs. The latter experiment was done because we inferred from the previously collected microarray data that cells elicit an adaptive response to these compounds (2).

**MATERIALS AND METHODS**

**General growth conditions and strains.** For general purposes, B. subtilis PB2 strains were grown in tryptone broth (LB). For weak-acid stress experiments and fluorescence measurements, B. subtilis strains were grown in defined liquid medium (M3G) (18) set at pH 5.5, 6.4 (buffered with 80 mM MES [morpholinopropanesulfonic acid]), as well as 7.0 and 7.4 (buffered with 80 mM MOPS [morpholinepropanesulfonic acid]). The medium contained 5 mM glucose, 10 mM glutamate, and 10 mM NH4Cl as the carbon and nitrogen sources. All cultures were grown at 37°C under constant agitation at 200 rpm. The wild-type strain used (B. subtilis PB2) was obtained from C. W. Price and A. Ter Beek (2). The strain B. subtilis IpHluorin (pHin) caused by these two weak acids using

$$
\Delta\psi = Z \cdot \frac{1}{2} \ln \left( \frac{C_w(C_e - 1)^{-1} + \left(1 - \frac{f_{\text{cm}}K_{\text{cm}}}{C_w} \right)^{-1}}{1 + \frac{f_{\text{cm}}K_{\text{cm}}}{C_w}} \right)
$$

where $\Delta\psi$ is the transmembrane electrical potential (in mV; which is equal to $\psi_{\text{cm}} - \psi_{\text{int}}$), Z (a conversion factor) is equal to 2.303 · (RT/\text{mol}F) (where R is the universal gas constant [83.1 J K$^{-1}$ mol$^{-1}$], T is the absolute temperature [in K], n is the charge of the translocated ion, and F is Faraday’s constant [96.6 J mV$^{-1}$ mol$^{-1}$]), $C_w$ is the probe concentration in the medium without the addition of cells, $C_e$ is the extracellular probe concentration, $f_{\text{cm}}$ is the ratio of the fractional cytoplasmic membrane and intracellular volume, $K_{\text{cm}}$ is the cytoplasmic membrane partition coefficient, and $x$ is the fractional internal volume.

The factor $f_{\text{cm}}K_{\text{cm}}$ was determined by a probe binding assay (20) to be 14. In this simplified approach, one assumes that the amount of TPP$^+$ bound to extracytoplasmic components of the cell (like parts of the cell wall) is equal to the amount of probe that binds to intracellular components, like nucleic acids. At the end of the experiment, 1.5-ml samples were collected for protein quantification as described above. Three biological replicates were measured for each condition.

**Oxygen consumption measurements.** Oxygen consumption rates were measured simultaneously with measurement of the $\Delta$\psi in the same measuring cell using a Neofox fiber optic oxygen sensor (Ocean Optics Inc.). Oxygen concentrations were measured every second from the maximally aerated to the fully oxygen depleted state. The slope of the straight line’s constant at plot (at least 15 s) was used to derive the oxygen consumption rate. At the end of the experiment, 1.5-ml samples were collected for protein quantification as described above. Rates were normalized to the protein content.

**Glucose consumption and metabolite measurements.** B. subtilis PB2 and B. subtilis IpHluorin were grown as described above in M3G, pH 6.4, to an OD$_{600}$ of 0.8. The cultures were split and exposed to different stress conditions (3 and 11 mM K-S or 25 and 80 mM K-Ac), and 10 mM glucose was added to each culture. A high OD$_{600}$ was required to observe a significant decrease in glucose concentrations within the time frame of the experiment.

Samples, taken every 30 min, were snap-frozen in liquid nitrogen. The protein concentration of each sample was measured as described above. Samples were further processed for high-pressure liquid chromatography (HPLC) analysis; a 1-ml sample was mixed with 100 $\mu$g 35% perchloric acid, and subsequently, 55 $\mu$1 M KOH was added. Filtered supernatants were analyzed for the levels of glucose consumption by the cells as well as the presence of fermentation end products. Glucose, succinate, lactate, acetate, 2,3-butanediol, and ethanol contents were determined by HPLC (LKB) with a Zezox organic acid analysis column (Phenomenex) at 45°C with 7.2 mM H$_2$SO$_4$ as the eluent. An RI 1530 refractive index detector (Jasco) was used, and AZUR chromatography software was used for data integration. All measurements were performed with three biological replicates. From the obtained metabolite fluxes, a carbon balance was calculated with the assumption that similar molar amounts of CO$_2$ were produced per amount of O$_2$ consumed, that similar molar amounts of CO$_2$ were produced per mole of acetate pro-
6.4. In M3G, the growth rate is the highest, but a continuous decrease in pHin monitored for 6 h under various stress conditions. The medium pH was experiment. With 80 mM K-Ac, this was not seen, and pHin re-
cultures continued to drop from time zero until the end of the 7.5 to 7.2 during this time. The internal pH of sorbic acid-stressed
weak-acid exposure occurred within the first minutes after addition. To investigate the rate of influx of weak organic acid preservatives and their dependence on the metabolic activity of the cell, we investigated the short-term effects of WOA injections into the medium.

Before acid injection, the pHin of starved wild-type B. subtilis cells was 7.30 ± 0.05. Injection of either KCl or NaCl at concentrations identical to those of the weak acids had no effect on pHin within 1 min after injection (not shown). Upon glucose injection, the pHin rapidly increased to 7.5 within 1 min (not shown). When WOAs were injected, the pH dropped to its lowest point within 1 to 4 s. In starved cells, this pH remained stable, but in the presence of glucose, pHin recovered quickly to a new equilibrium (Fig. 2A to D). The acidification of starved cells was fitted with a first-order kinetic equation,

\[
pH_{in}(t) = pH_{in}(0) - \Delta pH(e^{kt} - 1)
\]

where \( pH_{in}(t) \) is the pHin at time \( t \), \( pH_{in}(0) \) is the offset (pHin at time zero, before injection), \( k \) is the rate constant (s\(^{-1}\)), \( B \) is the amplitude factor (\( \Delta pH/\text{mM} \)) indicating the intracellular buffering capacity, and \([HA]\) is the concentration of undissociated weak acid (in mM). Using this equation, values for \( B \) and \( k \) were determined for sorbic acid and acetic acid. The value for \( \Delta pH_{in} \) or the amplitude of acidification has a linear relation with the concentration of undissociated acid in the medium (Fig. 3). These data (with starved cells) show that sorbic acid causes a change in pHin similar to that caused by acetic acid. The rate of acidification is high and similar for both acids, 1.29 ± 0.09 s\(^{-1}\) for sorbic acid and 1.27 ± 0.33 s\(^{-1}\) for acetic acid. However, with a measuring frequency of once per second, this rate is most likely set by our detection system and the actual rate is likely higher. Cells preexposed to sorbic acid had \( B \) and \( k \) values similar to those for nonstressed cells for either WOA, showing no sign of adaptation at this level (see Table S1 in the supplemental material; our unpublished data).

With glucose added to the cultures, recovery of the pHin started immediately after the injection of weak acid. This curve, too, was fitted, but with an additional factor describing recovery:

\[
pH_{in}(t) = pH_{in}(0) - B[H\text{A}](e^{-k_{r}t} - 1) + B[H\text{A}](e^{-k_{r}t} - 1)
\]

where \( B \) is the amplitude of recovery and \( k_{r} \) is the rate constant of the pH recovery. Recovery kinetics did not seem to change between unexposed and preexposed cultures. There was, however, a clear difference between sorbic acid and acetic acid. The pHin of acetic acid-stressed cultures recovered to a higher value than that of sorbic acid-stressed cultures (Fig. 4; see Table S1 in the supplemental material). The amplitude of recovery also appeared to have a linear relation with the acid concentration. The acquired constants allowed predictions for both acidification and subsequent recovery (see Fig. S1 in the supplemental material).

**Sorbic acid affects both \( \Delta pH \) and \( \Delta \psi \).** Proton translocation by the electron transport chain through the cell membrane results in..
both a gradient in the chemical potential of protons (i.e., $\Delta \mu_p$) and an electrical potential ($\Delta \psi$). This electrochemical proton gradient exerts an inward-directed proton motive force (PMF) (26, 27). The PMF can drive protons back into the cell via the $F_1F_0$ ATPase for ATP synthesis and is also required for various membrane transport processes and for flagellar rotation. The PMF is defined as follows (e.g., see reference 28):

$$\Delta \rho = \mu_{H^+} \cdot F^{-1} = \Delta \psi - Z \cdot \Delta p \text{H}$$

where $\Delta \rho$ is the PMF (in mV), $\Delta \mu_{H^+}$ is the transmembrane electrochemical proton potential (in J/mol), $F$ is Faraday’s constant [96.6 J mV$^{-1}$ mol$^{-1}$], $\Delta \psi$ is the transmembrane electrical potential difference (in mV; $\psi_{in} - \psi_{out}$), $Z$ (a conversion factor) is equal to $2.303 \cdot (RT/nF)$ (where $R$ is the universal gas constant [8.31 J K$^{-1}$ mol$^{-1}$], $T$ is the absolute temperature [in K], and $n$ is the charge of the translocated ion), and $\Delta \text{pH}$ is the transmembrane pH gradient ($\text{pH}_{in} - \text{pH}_{out}$, where $\text{pH}_{in}$ is the pH inside the cell and $\text{pH}_{out}$ is the pH outside the cell).

In our experiments, which were performed at 37°C, $Z$ was equal to 61.4 mV. Our measurements with TPP$^+$ show that under nonstressed conditions with an extracellular pH of 6.4, $\Delta \psi$ is equal to 103 ± 9 mV and $\text{pH}_{in}$ is equal to 7.78 ± 0.05. This thus gives a $Z \cdot \Delta \text{pH}$ value of $84 \pm 3$ mV and a total PMF of 188 ± 9 mV (Fig. 5). This value for $\Delta \psi$ is similar to values reported earlier (22, 29) and the pH$_{in}$ value is slightly higher (29). Sorbic acid has a severe effect on $\Delta \psi$, reducing it by 64% even at 3 mM K-S. Acetic acid, however, does not deplete $\Delta \psi$ as strongly. With 25 mM K-Ac, the $\Delta \psi$ is 17% lower, and with 80 mM K-Ac only a 45% reduction is

**FIG 2** Acidification and recovery upon weak organic acid addition to *B. subtilis* PB2 $P_{ptsG}$-IpHluorin. Sorbic acid (A and B) or acetic acid (C and D) was injected at time zero. The medium pH was 6.4. The internal pH drops to its lowest point within seconds (A and C) and recovers fast when glucose is available (B and D). The pH recovers to higher values with acetic acid stress. Data from typical examples are shown.

**FIG 3** Amplitude of acidification by addition of sorbic or acetic acid. A linear trend can be observed between $\Delta p \text{H}$ and $-\log [HA]$ ([HA] is in M). Data from experiments at extracellular pHs of 5.5 and 6.4 are combined. Error bars indicate standard deviations.
seen. It appears that sorbic acid acts as an uncoupler more than acetic acid does.

Together, $\Delta \psi - (Z \cdot \Delta \rho H)$ comprises the proton motive force. In total, WOA stress causes a dissipation from 188 ± 9 mV in nonstressed cells to 83 ± 5 mV with 11 mM K-S and 108 ± 7 mV with 80 mM K-Ac. This, too, shows that sorbic acid has a stronger effect on the electrochemical gradient for protons than acetic acid does at concentrations that lead to similar levels of growth inhibition (see Table S2 in the supplemental material).

Oxygen consumption is reduced upon addition of high WOA concentrations. Because weak organic acids partially dissipate the PMF, we speculated that *B. subtilis* might alter its metabolism. Glucose consumption was the highest for nonstressed cells, reaching 7.8 mmol s$^{-1}$ (mg protein)$^{-1}$. Addition of weak acids lowered glucose consumption rates, with 80 mM K-Ac resulting in 1.45 mmol s$^{-1}$ (mg protein)$^{-1}$ (Fig. 7; see also Table S3 in the supplemental material).
Acetate was produced under all conditions tested and was already present in small amounts at the start of the experiment. Nonstressed cultures produced acetate at the highest rate, together with the K-S-exposed cells. Addition of K-S lowered the rate of acetate synthesis. Acetic acid-stressed cells had the lowest rate of acetate production. These cultures redirected fermentation routes by switching to 2,3-butanediol fermentation. Some 2,3-butanediol was also found in the final sample of nonstressed cells but not with sorbic acid-stressed cells.

**Carbon flux.** When confronted with a lack of an electron acceptor, *B. subtilis* employs a mixed acid fermentation and has been shown to produce lactate, acetate, acetoin, ethanol, 2,3-butanediol, and succinate as fermentation products when grown on glucose (30), but it also produces a lot of acetate as a result of overflow metabolism (31). In our experiments, only acetate and 2,3-butanediol were found.

Although batch cultures are not ideally suited for accurate carbon flux determinations, our data do seem to be in line with reported values for wild-type *B. subtilis* (32). With the assumptions made previously regarding O₂ utilization, our data (see Table S3 in the supplemental material) for WOA-stressed cells show reasonably close C balances for all stresses, apart from 25 mM K-Ac, with which cells appeared to use about 3.5 times as much O₂ as the amount estimated on the basis of carbon fluxes based solely on glycolysis and fermentation.

The calculated $q_{\text{ATP}}$ based on $q_{\text{acetate}}$ and $q_{\text{2,3-butanediol}}$ ($q_{\text{ATP}} = (2 \cdot q_{\text{acetate}}) + (2 \cdot q_{\text{2,3-butanediol}})$) through glycolysis and fermentation, displayed a similar WOA concentration-dependent behavior as $q_{\text{glucose}}$. That is, increasing concentrations of weak acid caused a reduction in fluxes. This behavior seems to be analogous to the effect on pHₜₐₓ which depends more on the concentration than the nature of the acid (as shown above). Indeed, there appears to be a linear relationship between pHₜₐₓ and $q_{\text{glucose}}$ (Fig. 8), as well as between pHₜₐₓ and the calculated amounts of ATP generated via production of $q_{\text{acetate}}$ and $q_{\text{2,3-butanediol}}$ (our unpublished calculations). When looking at correlations between $\Delta W$ and $q_{\text{glucose}}$, such a correlation does not exist for either WOA. Only for sorbic acid might a linear correlation be observed, but more experiments will be needed to confirm this relationship. The relation between $q_{\text{O₂}}$ and the decrease in pHₜₐₓ seems to be rather constant, while a decrease in $q_{\text{O₂}}$ seems to correlate with a decrease in $\Delta W$. However, more experiments will be needed for an accurate description.

**DISCUSSION**

To reduce the growth rate to a similar level, more acetic acid than sorbic acid at an identical external pH value is required. Both sorbic acid and acetic acid lower pHₜₐₓ and reduce the growth rate of *B. subtilis*. When cells are growing with WOA stress, the internal pH is lower than that of cells not exposed to WOAs during the early exponential phase. Even though in the first minute after WOA exposure some recovery takes place, the internal pH does not seem to recover fully from this acidification, even after 6 h of growth. Given that the PKₐ for both weak acids is almost equal, this makes sorbic acid a more effective preservative.

Homeostasis of the pHₜₐₓ is crucial for cell physiology. Optimal proton extrusion clearly depends on the presence of an energy source (i.e., glucose) in the medium, because recovery of the pHₜₐₓ does not reach the original values. It appears that a new $\Delta \text{pH}$ equilibrium is established within a few minutes, i.e., before most transcriptomic changes take effect. The level of cytosolic acidification after injection of weak acid into the medium is very similar for both acids without glucose present. This suggests that in starved cells, an equilibrium for acid dissociation is established. Since the only factor measured in this regard is pH or [H⁻], the anion accumulation level is unknown. It can, however, be estimated if it is assumed that $\Delta \text{pH}$ is equal to $\log ([A^-]_{\text{in}}/ [A^-]_{\text{out}})$. The cell has to establish a new equilibrium with a $\Delta \text{pH}$ low enough to reduce anion concentrations but high enough to allow metabolism. As is clear from the data, the ability to restore pHₜₐₓ differs between sorbic acid-treated cells and acetic acid-treated cells.

For sorbic acid, the inability to recover pHₜₐₓ during acid stress reflects the ability of sorbate to act as a classic uncoupler, shutting protons over the membrane, whereas the less lipophilic acetate is believed to do this to a lesser extent. The latter is corroborated by the greater effect that sorbic acid has on the membrane potential, while acetic acid carries only bulk volume protons across the membrane until a steady state is reached, allowing $\Delta \psi$ to remain relatively unaffected. Even though some authors have pointed to the similar effects of the uncoupler 2,4-dinitrophenol and sorbic acid on growth and metabolism (9, 33, 34), to our knowledge no quantitative measurements of the effect that sorbic acid has on $\Delta \psi$ have been reported thus far. While relatively little research on the effects that WOA preservatives have on the membrane potential of microorganisms has been published (6, 16, 29, 35, 36), the depletion of $\Delta \psi$ may have a plethora of effects. The consequences may range from reduced transporter activity (37) to destabilization of the bacterial cytoskeleton (38), although we have not seen evidence for the latter effect (our unpublished results). Anion efflux from the cytosol would be driven by the remaining $\Delta \psi$. This effect was clearly shown for the permeant ion picrate, which mainly acts as an uncoupler in erytred cell membranes when a high $\Delta \psi$ is established (39). It has also been observed that lipophilic compounds such as ethanol can stimulate the leakage of protons over the membrane of *S. cerevisiae* (40, 41). However, recent results show that this is not the case for sorbic acid or acetic acid (5).

Other commonly used preservatives, such as lactic acid, benzoic acid, and propionic acid, are also expected to diffuse and dissociate, in agreement with the Henderson-Hasselbalch equation (9). Future experiments should establish to what extent these and other preservatives have uncoupling properties and what the
quantitative effect of Δψ depletion on the growth rate is. It is likely that their ability to cross the membrane barrier (log K_{ac}) into the aqueous phase will play a role, as seems to be the case with sorbic acid and acetic acid (42, 43).

With a low pH outside the cell (pH_{out}), the balance of the PMF is shifted toward ΔpH, so the effectiveness of weak organic acids at low pH is 2-fold: the concentration of undissociated acid is higher, and the WOAs act on the dominant component of the PMF. In the search for food preservatives that are also active near neutral pH, the food industry should preferably consider uncoupler-like molecules similar to or better than sorbate. At an even higher pH, permeant cations might be useful to deplete the membrane potential. In food with low glucose concentrations, the recovery of pH_{in} might be less; hence, less growth is observed with lower preservative concentrations.

We observed a decrease in the O_{2} consumption rate with WOA-stressed cells. This suggests that there are fewer electrons to feed the electron transport chain and that metabolism may be affected by acidification of the cytosol. Because weak organic acids partially dissipate the PMF, the amount of ATP that the cells can generate through the F_{F_{0}}F_{1} ATPase is considerably smaller. So, to produce a sufficient amount of ATP to restore pH homeostasis and proliferate, the cell might alter its glucose metabolism to generate ATP through substrate-level phosphorylation. Because increased energy needs have been observed in WOA-stressed yeast (44, 45) and a starvation-like response was reported for sorbic acid-stressed B. subtilis (2), we measured the glucose consumption rate as well as the production of fermentation products.

Glucose metabolism slows in the presence of WOAs, depending on the concentration of undissociated acid, as has also been described for E. coli and Saccharomyces cerevisiae (24, 46, 47). This results in a linear relation between pH_{in} and \( q_{\text{glucose}} \). The observation that acetic acid continues to be produced when cells are exposed to sorbic acid [2 mmol h^{-1} (mg protein)^{-1} with 11 mM K-S] indicates that acetic acid may be an extra stress factor for cells growing under these conditions and may be part of the explanation for the observed continued decrease in pH_{in} during sorbic acid stress. Addition of K-S lowered the rate of acetate synthesis, possibly due to growth reduction or a decrease in glycolytic activity.

Glycolytic flux can be controlled, among other factors, by pH (48), affecting phosphofructokinase (PFk) activity (49) or, specifically in the case of Bacillus species, phosphoglycerate mutase activity (50), which is also very sensitive to pH changes. Glucose is generally assumed to be the preferred carbon and energy source for B. subtilis, which is taken up via the phosphoenolpyruvate-sugar phosphotransferase system (PTS). It is subsequently converted into fructose 1,6-bisphosphate (FBP) by PFk. FBP is required for the phosphorylation of CcpA, one of the main regulators in carbon catabolite repression (CCR) in B. subtilis. In this way, FBP availability causes CCR (51). Also, CCR inhibits expression of citric acid cycle enzymes when glucose concentrations are high. When a decrease of the internal pH reduces the activity of PFk, the concentration of FBP is reduced, CcpA (52) activity is reduced and CCR is released. This may explain how in earlier studies (2) a starvation-like response could be observed in B. subtilis upon sorbic acid stress.

With high concentrations of acetic acid, it is likely that fermentation routes toward acetate are diverted to 2,3-butanediol. Synthesis of 2,3-butanediol depends on bdhA, a gene that encodes acetoin reductase (30) and that has SigB-controlled expression. Acetic acid stress can trigger stressosome activity and, indeed, results in the upregulation of acetoin reductase (53, 54). Therefore, when acetate is present as a preservative, putative organoleptic properties of 2,3-butanediol should be considered when B. subtilis or related species may be present. Weak organic acid stress also causes an upregulation of citric acid cycle enzymes (2, 53). This route may also consume added acetic acid, thereby providing an extra carbon source and eliminating this weak organic acid. The high oxygen consumption rate that we observed in the presence of 25 mM and 80 mM K-Ac may be explained by the utilization of acetate. While growth rates are reduced by 80% with both 11 mM K-S and 80 mM K-Ac, the glucose consumption rate with 80 mM K-Ac is almost half of that with 11 mM K-S. Future experiments should be conducted in a turbidostat setup and include CO_{2} measurements to allow a direct measurement of the carbon balance under these conditions. Therefore, knowledge of the pH_{in} can be used to predict glycolysis, and this knowledge might be useful for modeling purposes in food fermentation (48, 55).

In summary, oxidative phosphorylation is an important source of ATP for aerobic, nonstressed B. subtilis cells. The O_{2} consumption rates follow a trend similar to that for \( q_{\text{glucose}} \), \( q_{\text{acetate}} \), and \( q_{2,3\text{-butanediol}} \) apart from the 25 mM K-Ac stress. The level of WOA stress may leave too little PMF to allow F_{F_{0}}F_{1} ATPase to produce ATP. Sorbic acid has an effect similar to that of acetic acid on pH_{in}, but it is more effective in depleting Δψ. This may partially explain the observation that sorbic acid is a more potent preservative than acetic acid. In the quest for preservatives active at nearly neutral pH, industry should best focus on uncoupler-like molecules that act similarly to or better than sorbic acid.

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REFERENCES


21. Vojinovic V, von Stockar U. 1995. Transport processes of various lipid components of Zygosaccharomyces rouxii as influenced by the quantitative determination of the proton-motive force in bac-


39. Fujita Y. 2009. Carbon catabolite control of the metabolic network in

