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C–H Activation of Benzene by a Photoactivated NiII(azide): Formation of a Transient Nickel Nitrido Complex**

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In memory of Greg Hillhouse

Abstract: Photochemical activation of nickel-azido complex 2 [Ni(N₃)₃(PNP)] (PNP=2,6-di(isopropylphosphino)-4,4'-di-tolylamine) in neat benzene produces diamagnetic complex 3 [Ni(Ph)(PNP₃)]₄, which is crystallographically characterized. DFT calculations support photoinitiated N₃-loss of the azido complex to generate a rare, transient NiII nitrido species, which bears significant nitridyl radical character. Subsequent trapping of this nitrido through insertion into the Ni–P bond generates a coordinatively unsaturated NiII imidophosphorane P=N donor. This species shows unprecedented reactivity toward 1,2-addition of a C–H bond of benzene to form 3. The structurally characterized chlorido complex 4 [Ni(Cl)-(PNP₃)] is generated by reaction of 3 with HCl or by direct photolysis of 2 in chlorobenzene. This is the first report of aromatic C–H bond activation by a trapped transient nitrido species of a transition metal.

Transition metal complexes with terminal multiple bonds to main-group elements (e.g. C, O, N) play critical roles in various chemical and biological processes, including but not limited to nitrogen fixation, water splitting, and metabolism.[1] However, ligands based on multiple-bonded heteroatoms (e.g. imido or nitrido) typically exhibit strong π-donor character. For stable multiple bonding to occur between a transition metal and a strong π-donor main-group element, one or more empty metal d-orbitals of suitable symmetry and matching energy levels should be available on the metal center to accommodate the additional π-electron density.[2] Hence, complexes with Group 9–11 metals featuring multiple-bonded E-fragments are either rare or elusive, a situation sometimes referred to as “the oxo-wall dilemma”. [3] Notwithstanding the challenges associated with the preparation of these species, their potential application in for example, C–N bond formation[4] or as intermediates in alternative routes for N₂ fixation warrant continued investigation, even when using “abiological” non-Group 8 metals.

Metal azides offer access to the interesting chemistry of nitrido complexes, as thermal, photolytic, or oxidative activation may result in the expulsion of N₂. Progress in the generation of Group 9 nitrides and open-shell nitridyl radical analogues demonstrate the feasibility of this approach.[5] However, these studies are dominated by the use of the second- and third-row transition metals. Only very recently the first low-temperature characterization of a cobalt nitrido intermediate was reported, which highlights the potential for late first-row transition metal nitrido chemistry.[6]

Studies on the activation of terminal metal azides to generate M≡N species with Group 10 metals are very scarce, particularly for nickel[7–9] and no bona fide (transient) nickel nitrido has been reported to date. Photo-, thermal, or redox-induced release of N₂ would convert the monoanionic azido ligand into a trivalent nitrido (N^3⁻) ligand, which raises the formal oxidation state of the metal platform by two (Figure 1). Hence, stabilization of such (transient) species could potentially benefit from electron-donating or redox-active ligands. Terminal nitrido ligands on Group 9 metals are often too reactive to be isolated, and “trapping” these sought-after transient species by insertion into ligand C–H[10] or

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**Figure 1.** Strategies toward and reactivity with “trapped” transient nitrido species of Group 9 (previous work) and Group 10 metals (Ni, this work; arene C–H).
metal–ligand bonds is an attractive strategy. However, to the best of our knowledge, well-defined reactivity of a “trapped” nitrido species toward intermolecular arene C–H bond activation has not been reported to date.

As part of our ongoing program to utilize reactive ligands with (first-row) late transition metals for bond activation, we became interested in the preparation of nitrido species for bond activation, but this has not been reported to date. Inspired by the recent examples with Group 9 metals, we set out to explore this chemistry with Group 10 congeners. Herein, we report on the photolytic activation of a Ni(Ni)(PNP) complex, which results in a transient nitrido species that is trapped as parent imidophosphorane by an intramolecular insertion reaction. This species facilitates unexpected follow-up aromatic C–H bond activation of benzene. DFT calculations on the mechanism and reactivity of the photolysis product are presented. The combined data provide insight into the chemical reactivity of a transient Ni-bound “nitrido” ligand.

Our starting point was the previously reported chlorido complex Ni(Cl)(PNP) (1; PNP = 2,2’-di(isopropylphosphino)-4,4’-ditolylamide). This species cleanly reacted through salt metathesis with excess NaN₃ in THF to generate azido complex Ni(N₃)(PNP) (2, Scheme 1) in high yield, as confirmed by multinuclear NMR spectroscopy. IR spectroscopy suggested the quantitative formation of one well-defined diamagnetic species. The ¹H NMR spectrum (Figure 2) showed complete preservation of all hydrogen atoms within the ligand scaffold, hence excluding ligand-based intramolecular C–H activation as a potential reaction pathway. Furthermore, complete loss of the C₂ symmetry is observed, with two signals for the -CH₃ groups of the ditolyamine backbone and eight non-equivalent -CH₃ signals from the PrPr₃ groups present. Notably, the ¹³P NMR spectrum contains two doublets at δ 45.9 and 48.5 with a small coupling constant J_P of 6.4 Hz, excluding a direct trans-coordination of the P moieties. Irradiation of the corresponding ¹⁵N-labeled isotopomer of 2 resulted in two new doublets-of-doublets in the ³¹P NMR spectrum. Remarkably, both J_P coupling constants are almost identical at 29.2 and 30.5 Hz, whereas the J_P coupling is 6.5 Hz. This confirmed the presence of the azide N₃ atom and suggested that the two P-atoms are chemically inequivalent, but both couple to the ¹³N nucleus. Further analysis of the species by cold-spray ionization (CSI) MS spectrometry revealed the presence of a perdeuterophenyl fragment in the generated complex.

![Scheme 1](image1.png)

Scheme 1. Synthesis and reactivity of complex 2, (PNP)Ni(N₃), including reaction with CO to generate A, photolytic activation to species 3, and subsequent reaction with HCl to product 4.

![Figure 2](image2.png)

Figure 2. ¹H NMR spectra before (top), after irradiation of 2 in CD₈ (middle), and after irradiation in CD₃H (bottom). The insert is showing an additional (-NH) doublet at 0.22 ppm for the irradiation in CD₃H.
Based on the latter observation, we decided to switch to C6H6 as the solvent for the photolysis reaction. In line with the MS data, this resulted in the appearance of additional aromatic signals in the 1H NMR spectrum and, markedly, also a doublet at δ 0.22, coupled to a 13P nucleus (confirmed by selective 31P decoupling of the 1H NMR spectrum). The 15N-enriched material gave an additional splitting of the doublet with a large coupling constant (JN1 73.2 Hz), which suggests a direct bond between a hydrogen of the activated benzene and a nitrogen originating from the decomposed azide group. The existence of an N–H fragment was additionally confirmed by the appearance of a weak N–H vibration (ν 3365 cm⁻¹) in the IR spectrum. Close inspection of the IR spectrum of 3⁰ (product from photolysis in C6D6) revealed the presence of an N–D vibration at ν 2493 cm⁻¹. In the corresponding 15N HSQC NMR spectrum a triplet was observed at δ ~365.

Yellow-colored single crystals, suitable for X-ray diffraction, were obtained by slow evaporation of a dichloromethane/cyclopentane mixture. The X-ray structure of square-planar Ni⁶-complex 3 (Figure 3) confirms the formal insertion of an –NH fragment into the Ni₃–P bond, thus creating a parent iminophosphorane, with a P–N₃ bond length of 1.5970(20) Å, which is typical for a P=N double bond.¹⁷ The angles ∠N₃–Ni–N₁, (93.98(8)°) vs. (83.83(4)°) for ∠N₃–Ni–P₁ in 2 and ∠N₃–Ni–P₁–Ni₁ (177.48(7)°) vs. 167.38(1)° for ∠P₁–Ni₁–P₁ in 2 illustrate the expanded coordination sphere around Ni upon transformation of the PNP-donor into a 5.6ʻ-chelating PNP¹¹ donor in 3. Furthermore, C–H bond activation of benzene has resulted in the introduction of a phenyl ligand, with hydrogen transfer to the nitrogen originating from the azide precursor.¹⁸ Variable temperature NMR studies suggest that the phenyl ligand exhibits hindered rotation at room temperature, causing the broadening of the ortho- and meta-H signals in the 1H NMR spectrum (Figure 2). Decreasing the temperature of a sample in [D₈]toluene to −20°C gave rise to splitting of these broad signals into four well-defined peaks. Species 3 still displays well-behaved redox chemistry, as illustrated by cyclic voltammetry, with a reversible one-electron oxidation event occurring at E₁/₂ = −0.34 V, which is significantly more negative (ΔE = −0.41 V) than that for azide 2.

To gain more insight into the mechanism of the formation of complex 3, DFT calculations were performed. For computational ease, a simplified model of complex 2 was used, with phenyl instead of tolyl rings for the backbone and Me-substituents at the phosphorus atoms. We investigated routes at different (excited-state) spin states, both at the hybrid b3-lyp level (Scheme 2) and the pure GGA BP86 DFT functional (Scheme S3, SI). Open-shell singlet (OSS) configurations for all investigated stationary points could only be located at the hybrid b3-lyp level. At the BP86 level only closed-shell singlet (CSS) configurations were found. As a starting point the activation of the azide was investigated. The release of N₂ proceeds via a high-energy transition state TS₁ on all calculated surfaces and is therefore considered as a nonthermal, photochemical step. The reaction pathway after N₂ loss from 2 can either follow photochemical or thermal steps or a combination thereof. The meta-stable “N₃”-nitrado” intermediate B is most stable in its triplet configuration (³B), both at the BP86 and at the b3-lyp level.
The spin density predominantly resides on the N-atom (b3-lyp: 89 %, BP86: 81 %), resulting in significant nitridyl radical character (Figure 4), similar to previously reported systems.\(^{[5b,6]}\) The unusual electronic structure of \(\text{B}^1\) is perhaps best described with contributions from the following resonance structures: 

\[
\text{Ni}^{I\{\text{PNP}\}}(\text{N}^\text{I})^\text{II} \rightarrow \text{Ni}^{I\{\text{PNP}\}}(\text{N}^\text{I})^\text{II} \rightarrow \text{Ni}^{I\{\text{PNP}\}}(\text{N}^\text{I})^\text{II}.
\]

The CSS and the OSS configurations of singlet \(\text{B}^1\) collapse via a virtually barrierless transition state \(\text{TS2}\) to the coordinatively unsaturated \(\text{Ni}^{I\{\text{PNP}\}}(\text{PNP})\) intermediate \(\text{C}\), featuring two anionic N-donor atoms (i.e. the PNP-nitrogen and the “nitrido” nitrogen). In fact, the barrier from \(\text{B}^1\) to \(\text{C}\) at the closed-shell singlet BP86 configuration is so low that \(\text{B}^1\) converges spontaneously to \(\text{C}\). Conversion of \(\text{B}^1\) to \(\text{C}\) on the triplet surface also has quite a low barrier (both at the BP86 and the b3-lyp level) and thus might well proceed thermally (although photoexcitation to the singlet surface cannot be excluded).

The relative spin-state energies of intermediate \(\text{C}\) at the CSS, OSS, and triplet surface are quite similar. While the triplet configuration of \(\text{C}\) is most stable at the b3-lyp level, the CSS configuration of \(\text{C}\) is most stable at the BP86 level. The reaction from \(\text{C}\) via \(\text{TS3}\) to \(\text{D}\) may follow a thermal or photochemical pathway. Thermal conversion at the triplet surface has a prohibitively high barrier at room temperature, but spin cross-over from the triplet to the CSS surface might facilitate the overall process. A pathway involving H-atom abstraction from benzene followed by a radical rebound reaction was found to be less likely, due to a higher energy barrier. At the CSS and OSS surfaces, direct thermal conversion from \(\text{C}\) to \(\text{D}\) is energetically possible at room temperature. The alternative scenarios of outer-sphere C–H activation prior to the N-insertion step, with formation of a five-coordinate \(\text{Ni}^{I\{\text{PNP}\}}(\text{NH})(\text{Ph})\) imido-species intermediate, were also investigated but found to be less favorable.

Based on these computational data, several attempts to trap the postulated intermediate \(\text{B}\) were undertaken. Photoysis of \(\text{2}\) in benzene in the presence of Me\(_2\)NO, 50 equiv PMe\(_2\), or in pure 1,4-cyclohexadiene did not generate \(\text{Ni}^{I\{\text{NO}\}}(\text{PNP}), \text{Ni}^{I\{\text{PNP}\}}(\text{PNP}), \text{or Ni}^{I\{\text{NH}_2\}}(\text{PNP}),\) respectively, based on IR and NMR spectroscopic analysis. The use of \(\text{H}_2\) as trapping agent did not lead to the parent Ni-imido species, but instead the known \(\text{Ni}^{I\{\text{H}\}}(\text{PNP})\) complex\(^{[13]}\) was observed. \(\text{NH}_3\) was detected by \(\text{^15N}\) NMR spectroscopy and spectrophotometric assay analysis (SI). Use of \(\text{D}_2\) resulted in \(\text{ND}_2\). The \(\text{NH}_2\) and the Ni-hydride may originate from reaction of \(\text{H}_2\) with the transient nitrido, but other pathways for their formation cannot be excluded at this point.\(^{[13]}\) Initial experiments to broaden the reactivity “scope” of the trapped nitrido species \(\text{C}\) were also performed. Photoysis of \(\text{2}\) in toluene led to four sets of doublets by in situ \(\text{^13C}\) NMR spectroscopy, suggesting that \(\text{C}_\text{N}^{\cdot\text{H}}\) and \(\text{C}_\text{N}^{\cdot\text{H}}\) bond activation may be competitive processes. Photoysis of \(\text{2}\) in neat cyclohexane or heptane resulted in a myriad of unidentified species in the \(\text{^13C}\) NMR spectrum. Photoysis in 1,3,5-trideuterobenzene was thwarted by trace amounts of chlorobenzene (detected by MS), culminating in the presence of two doublets at \(\delta\) 46.8 and 45.3 in the \(\text{^13C}\) NMR spectrum, attributed to be the chlorido derivative of \(\text{3}\). \(\text{NiC}(\text{PNP})(\text{N}^\text{I})\) \(\text{(4)}\). Complex \(\text{4}\) was also prepared independently by reaction of \(\text{3}\) with HCl (Scheme 1) and fully characterized by NMR spectroscopy, mass spectrometry, and X-ray crystallography (see Figure 3 for X-ray structure). The structural parameters for \(\text{4}\) are very similar to those observed for \(\text{3}\) (Figure 3). Deprotonation of \(\text{4}\) with lithium disopropylamide (LDA) in the presence of benzene did not result in aren C–H activation. We are currently exploring the reactivity of \(\text{3}\) and \(\text{4}\), both to understand the unique structural features of these species and to deduce whether the parent iminophosphorane P=NH unit shows promising ligand reactivity.

In summary, we herein report the first example of a trapped transient nickel nitrido species, making use of a well-defined Ni\(^{I\{\text{PNP}\}}\) platform. Irradiation of the nickel-azido precursor generates a nickel nitrido, which has significant nitridyl radical character, that efficiently trapped by insertion of the nitrogen into a Ni–P bond. This trapped intermediate undergoes follow-up intermolecular C–H bond activation of aromatic solvent, with the activated exogenous phenyl fragment incorporated in the final complex \(\text{3}\), which bears a unique tridente phosphino-amido-iminophosphorane PNP\(^{I\{\text{N}\}}\) ligand in the coordination sphere of Ni\(^{I\{\text{N}\}}\). DFT calculations support the intermediacy of a rare transient formal Ni\(^{I\{\text{N}\}}\) nitrido species and subsequent 1,2-addition of the C–H bond. Substitution of the Ph-ligand for Cl is facile, generating \(\text{4}\), which is likely amenable to follow-up chemistry.

**Keywords:** azides · C–H activation · iminophosphoranes · nickel · nitridyl radical

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[8] Irradiation of a tetrahedral Ni
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(P
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)complex resulted in the formation of Ni(N
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)3 and Ni(“P”)
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[16] CCDC 1043360 (2), 1043865 (3), and 1048448 (4) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.


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