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Comparison of the Full Catalytic Cycle of Hydroformylation Mediated by Mono- and Bis-Ligated Triphenylphosphine–Rhodium Complexes by Using DFT Calculations

Ivo Jacobs, Bas de Bruin,* and Joost N. H. Reek*^[a]

Introduction

Hydroformylation can be considered as the flagship reaction of homogeneous catalysis: it is applied commercially on a multi-million ton per year scale, but also studied in detail at the fundamental level.[1] Both terminal and internal olefins are used as feedstock to produce various products. Terminal olefins are more reactive, and selectivity issues are less challenging. Two aldehyde products can be formed from primary olefins: the linear (L) or the branched (B) aldehyde (Figure 1). There is a large market for the linear aldehyde, and many industrial processes aim for this product. In some of the processes, the aldehydes are further reacted to produce either alcohols or acids. There is an increasing demand in recent years for branched aldehydes, after linear aldehydes.[2] For example, isobutyraldehyde, which is a precursor of isobutyric acid and neopentyl glycol, is a product that is prepared by the hydroformylation of propene. Although nowadays several catalysts can produce the linear aldehyde in high selectivity, catalysts that mainly produce branched aldehydes from aliphatic alkenes are scarce, and selectivities are still rather low.

In 2001 we reported an encapsulated hydroformylation catalyst that forms by self-assembly from meta-trispyridylphosphine and three zinc tetraphenylporphyrin building blocks. The encapsulated monophosphine rhodium catalyst self-assembles upon mixing the ligands with Rh(acac)(CO)₂ (acac = acetylacetonato) under a syngas atmosphere (Figure 2).[3] This complex is an active hydroformylation catalyst that preferentially produces the branched aldehyde product from terminal alkenes. It even operates at temperatures as high as 80 °C, although under these conditions higher CO pressures are required to prevent the formation of bisphosphorus-ligated complexes.[3c] Next to terminal olefins, the catalyst can also be used to hydroformylate internal olefins with unequalled selectivity, preferably forming the most internal product.[4] In all experiments, the encapsulated catalyst showed higher activity with respect to the parent triphenylphosphine-derived catalyst.

From detailed studies it is already known for a long time that bisphosphine-ligated rhodium complexes (and trisphosphine) give a much higher selectivity for the linear product,[5] and therefore high concentrations of ligands are used in commercial applications. Although bulky monodentate phosphite ligands have been used,[6] which lead to monophosphorus-ligated rhodium complexes that are active, bulky phosphine ligands that lead to exclusive monocoordination have not been

![Figure 1. Hydroformylation yields two products: linear (L) and branched (B) aldehydes.](image)

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reported.[7] As such, the supramolecular ligand system shown in Figure 2, in which the phosphine donor is embedded, is the only system that shows such exclusive monocoordination. Matt et al. recently reported a cyclodextrin-based phosphine ligand that also leads to the exclusive formation of mono-ligated rhodium complexes under hydroformylation conditions, which resulted in the enantioselective hydroformylation of styrene.[8]

Because the properties of mono-ligated phosphine rhodium catalysts are not known experimentally, it is impossible to determine whether the high activity and unusual selectivity are the result of confinement effects or of the coordination to rhodium. To improve our understanding of how the coordination affects catalytic performance in the case of the triphenylphosphine-modified rhodium system, we performed a DFT study that compares mono- and bis-ligated rhodium species. This should facilitate the better understanding of the operational mode of supramolecular systems.

The catalytic cycle of rhodium-catalyzed hydroformylation is depicted in Figure 3. Three stages can be distinguished: In the first stage, a rhodium hydride species reacts with an olefin to form a rhodium alkyl species. In the second stage, this rhodium alkyl species reacts with CO to form a rhodium acyl species. Finally, in the third stage, the acyl species reacts with hydrogen to form the product aldehyde and the rhodium hydride is reformed. This mechanism is well established and has been extensively corroborated by using both experimental (spectroscopic) and theoretical methods.[5, 9, 10]

The kinetics of the hydroformylation reaction catalyzed by phosphine-modified rhodium are in most cases dominated by the reactions of the first stage. This results in the rate equation in a first order in olefin, a negative order in CO, and a zero order in hydrogen. From the kinetics alone it cannot be determined which individual step is rate determining, because all three possibilities give the same rate equation. Spectroscopic evidence shows that hydride species 1 is the resting state.[5b] The olefin coordination is rate determining if Xantphos is used as a ligand,[11] and several studies show that hydride migration is irreversible (for triphenylphosphine, but not for the BISBI or Xantphos ligand),[12] but there is no direct experimental evi-

dence that shows exactly which step is rate determining. However, several computational studies concluded that hydride migration is rate determining.[11,13,16a]

Some systems exhibit a different kind of kinetics, which is dominated by the third stage. This is called type II kinetics and displays first order in hydrogen, zero order in olefin, and negative order in CO. In this case, 7 is the resting state, and one of the reactions leading from it up to the release of product must be rate determining. This type of kinetics is less common, and mostly encountered only with bulky phosphite ligands.[5b] Breit’s phosphinine system may also fall under this category.[14]

The reactivities of mono- and bisphosphine complexes have been compared earlier in computational studies of the hydroformylation reaction, most recently by Hirst and Carbo.[15] The full catalytic cycle has never been investigated in any of the computed models, though. In most cases, only the first stage of the reaction was taken into account. We considered it important to perform calculations on the full catalytic cycle by using the same computational model throughout, accounting for all relevant transition states and all possible changes on the potential energy surface on migrating from the bisphosphine to the monophosphine cycles and thus allowing a proper comparison with kinetic data from experimental studies. This information is important to find out whether the changed behavior seen in the case of the encapsulated rhodium complex is only due to the enforced monocoordination or if other effects play a role.
Herein, we report DFT calculations on the entire catalytic cycle of hydroformylation mediated by mono- and bis-ligated rhodium species. These calculations show that the pathways are similar and that the rate-determining step in all cases is the hydride migration. It also became clear from these calculations that attractive van der Waals interactions between the triphenylphosphine ligands, as well as between the ligands and the substrate, are important, which is in line with a previous report by Kumar and Jackson.\(^\text{[16]}\) Because these interactions are not taken into account in (uncorrected) DFT calculations, DFT-D3 dispersion-corrected calculations were needed to get a better estimate for the energy difference between mono- and bis-triphenylphosphine complexes. Therefore, the rate-determining transition states TS-3 were recalculated with dispersion corrections to arrive at more realistic overall reaction barriers.

## Results and Discussion

### General

Several papers have appeared that report DFT calculations on (parts of) the hydroformylation catalytic cycle.\(^\text{[17]}\) In all these reports, model ligands such as PH\(_3\) or PF\(_3\) (which is the better model for PPh\(_3\)) have been used to reduce the calculation time. For our calculations, we decided to use the full PPh\(_3\) ligand because we also wanted to take into account steric hindrance in bisphosphine complexes, as it could have a strong influence on the difference between the bisphosphine pathway and the monophosphine pathway. Because we were primarily interested in activity, ethene was used as the substrate model initially. However, upon comparing the monophosphine and bisphosphine complexes, ethene was not found to be a good model, and to correct this, all propene variants of the rate-limiting transition state TS3 were also built and calculated.

For DFT calculations, we used the B3LYP functional with Stuttgart–Dresden basis sets with ECPs (ECP—effective core potential) for rhodium and phosphorus, and Popple basis sets for the rest, in Gaussian 03. This combination of functional and basis sets has previously been shown to give accurate results.\(^\text{[18]}\) Because Grimme's dispersion corrections are not implemented in Gaussian 03, the relevant structures for which we wanted to investigate the effect of dispersion corrections were reoptimized in Turbomole Version 6.5 at the b3-lyp, def2-TZVP disp3 (DFT-D3) level of theory. A small-core ECP pseudopotential was used for rhodium. The def2-TZVP basis set used was slightly larger but otherwise similar to 6-311G**.

Here, we will first discuss the catalytic cycle of the bisphosphine rhodium complexes and then compare that with the catalytic cycle of the monophosphine complex.

### Formation of the active complex

The Rh(acac)(CO), precursor can be converted to various rhodium complexes under catalytic conditions. A schematic overview of the complexes identified with high-pressure IR spectroscopy (under various conditions) and their connecting equilibria is presented in Figure 4. As known from the Rossi and Hoffmann study,\(^\text{[19]}\) the stronger sigma-donating ligands prefer to occupy the axial positions. Therefore, in all complexes the hydride ligands reside at the axial position and structures with the hydride in the equatorial plane are not even minima on the potential energy surface. Phosphine ligands are much less strong sigma donors and thus do not have a preference for the axial position as compared to CO.

The monophosphine complexes are predicted to be lower in energy than the bisphosphine complexes if regular DFT without dispersion corrections is used (Figure 4). This finding contradicts experimental observations.\(^\text{[20]}\) This is mostly due to the absence of attractive dispersive interactions in regular DFT, which leads to an underestimation of the stability of bisphosphine complexes, especially if they are situated in the cis position with respect to one another. Some of the complexes were therefore reoptimized by using Grimme's dispersion corrections (D3 version, implemented as disp3 in the Turbomole program package). The results are presented in Figure 4. These calculations show that there are significant attractive van der Waals interactions between the phosphines, and the non-dispersion-corrected calculations clearly underestimate the stability of bisphosphine complexes. In general, these errors are expected to be similar for subsequent steps in a reaction sequence starting from either mono- or bisphosphine rhodium complexes, and thus as long as no PPh\(_3\) association/dissociation occurs in the sequence the DFT-calculated reaction pathways should benefit from error cancellation. However, ee\(_1\) is 0.6 kcal mol\(^{-1}\) lower in energy than ea\(_1\), as predicted by the uncorrected calculations, which matches well with the ee\(_1)/ea\(_1\) ratio of 85:15 observed in NMR experiments by Brown and Kent.\(^\text{[10]}\) The dispersion-corrected calculations predict ea\(_1\) to be 1.7 kcal mol\(^{-1}\) lower in energy than ee\(_1\), which could be an effect of overcorrection for van der Waals interactions if gas phase calculations are compared to condensed phase experi-
ments. Ultimately approximately 2 kcal mol⁻¹ error in the energy difference between ee1 and ea1 is better than approximately 10 kcal mol⁻¹ error in the energy difference between monophosphine and bisphosphine. Because the dispersion effects are significant, we evaluated the barriers for all key transition states both with and without dispersion corrections (vide infra).[21] In the following section, we compare the catalytic cycles that start from ee1/ea1 with those that start from a1/e1.

Catalytic cycle with the bisphosphine rhodium complex as a catalyst

The (noncorrected) DFT-calculated free energy profile of the first two steps of the bisphosphine pathway is shown in Figure 5. The resting state of the triphenylphosphine-ligated catalyst is the five-coordinated rhodium hydride complex 1. This hydride complex exists as a mixture of two isomers, which are indicated by the position of phosphine ligands (e = equatorial and a = axial), and ea1 is 0.6 kcal mol⁻¹ higher in energy than ee1. The equilibrium between 1 and 2 is characterized by a low transition state (Rh−C distance is 3.1 and 3.2 Å for TS-ee1 and TS-ea1, respectively). The calculated free energy barriers for CO dissociation from ee1 and ea1 were 13.8 and 17.5 kcal mol⁻¹, respectively. The reverse reactions, CO binding to tP2 and cP2 to form ee1 and ea1, have lower free energy barriers (10.3 and 7.4 kcal mol⁻¹, respectively).

In the square planar complexes, the trans complex tP2 is 7 kcal mol⁻¹ lower in energy than the cis complex cP2, which can be due to steric hindrance between the two triphenylphosphine ligands. In cP2 the square planar complex is significantly distorted, with a P−Rh−P angle of 104°.

After coordination of ethene, the five-coordinated ethene complex is obtained. In this substrate-associated complex, both isomers ee3 and ea3 are close in energy. The rhodium–olefin interactions are underestimated in DFT calculations, and therefore the energies of these alkene intermediates are probably a few kcal mol⁻¹ lower.[22,23] The energy profile of hydride migration steps are depicted in Figure 6, and the Berry pseudorotation-like mechanism is further clarified in Figure 7 (for transition state TS-ee3-c). Migratory insertion of the olefin into the Rh−H bond of ee3 leads to the formation of the square planar cis-coordinated bisphosphine alkyl complex cP4 via low barrier transition state TS-ee3-c. Figure 5. Energy diagram of the first stage of the catalytic cycle with the bisphosphine catalyst until substrate coordination. Red line is for the ea complex, and the black line is for the ee complex.

Figure 6. Energy diagram of the hydride migration reaction with the bisphosphine complex. The way different pathways connect the intermediates is shown at the bottom.

Figure 7. Migratory insertion steps involving the bisphosphine hydrido-rhodium carbonyl-olefin complex ee3, which results in the square planar cis-bisphosphine alkyl-rhodium carbonyl complex cP4. Bond distances are given in Å.

TS-ee3-c (ΔG⁺ = 11 kcal mol⁻¹). For ea3, two pathways are possible because the olefin can rotate toward either the phosphine ligand or CO. The former has a lower barrier (ΔG⁺ = 13 kcal mol⁻¹) and results in the formation of cP4. In line with findings in the literature, the hydride migration is here predicted to have the highest overall barrier and is thus the rate-determining step. The transition state TS-ee3-c corresponds to the lowest energy pathway in this rate-determining step and lies 28.9 kcal mol⁻¹ above the corresponding resting state, ee1.

From Figure 7 it becomes clear that the distance between the two phosphine ligands changes in the hydride migration
step, and it is therefore necessary to investigate the effect of dispersion corrections for this reaction. The dispersion-corrected free energy barriers are listed in Table 1 and compared with the noncorrected values. The DFT-D3 barriers are 3–6 kcal mol\(^{-1}\) lower than the DFT barriers, but the relative order remains the same. The lower DFT-D3 barriers are partly explained by stronger ethene binding in the transition states, as close contacts between the substrate and the aryl groups of the catalyst lead to favorable interactions, which are taken into account by including dispersion corrections.

For comparison, the corresponding propene transition states were also calculated. For each of the ethene transition states, four corresponding propene transition states can be drawn, which leads to 12 transition-state structures. Six of them lead to the linear product and six to the branched product. Because dispersion is important, we evaluated these transition states both at the DFT and at the DFT-D3 level. The results of these calculations are summarized in Table 2. Dispersion corrections stabilize both the transition states and the minima. Like in the case of ethene, the barriers are lowered if dispersion corrections are applied. The barriers are lowered more in the case of propene, which can be expected because the larger propene benefits more from stronger van der Waals interactions (thus reducing unfavorable steric interactions) in the transition state.

Notably, the lowest transition state for ethene has the same structure as the lowest transition state for propene for both the linear and the branched pathway (Figure 8). The lowest barrier for the formation of the linear product is 1.2 kcal mol\(^{-1}\). The energy profile of the second stage of the catalytic cycle are depicted in Figure 9. In the first step, CO coordinates to the square planar complex to form the more stable five-coordinated rhodium species. As expected, this step is exothermic for all complexes.

The next step in the catalytic cycle involves the migratory insertion of CO into the metal–alkyl bond to form the rhodium acyl species. According to the calculated pathway, the mechanism is best described by the migration of the alkyl species to CO (Figure 10). During this reaction step, the ligands rotate in such a way that the trans-bisphosphine acyl complex TpP6 is formed from the ea–alkyl complex ea5. Similarly, the cis-bisphosphine acyl CpP is formed from the ee–alkyl complex ee5. The formation of the four-coordinate acyl complex is exothermic. The formation of a C–C bond clearly compensates for

### Table 1. Free energy barriers for the rate-determining bisphosphine transition states from ee1/ea1 with ethene as a substrate.

<table>
<thead>
<tr>
<th>Transition State</th>
<th>DFT (for ee1)</th>
<th>DFT-D3 (for ea1)</th>
</tr>
</thead>
<tbody>
<tr>
<td>TS-ea3-c</td>
<td>35.9</td>
<td>30.6</td>
</tr>
<tr>
<td>TS-ea3-t</td>
<td>31.5</td>
<td>26.8</td>
</tr>
<tr>
<td>TS-ee3-c</td>
<td>28.9</td>
<td>25.8</td>
</tr>
</tbody>
</table>

### Table 2. Free energy barriers for the rate-determining bisphosphine transition states from ee1/ea1 with propene as a substrate.

<table>
<thead>
<tr>
<th>Transition State</th>
<th>DFT (for ee1)</th>
<th>DFT-D3 (for ea1)</th>
</tr>
</thead>
<tbody>
<tr>
<td>TS-p-ea3-c-B</td>
<td>39.9</td>
<td>32.7</td>
</tr>
<tr>
<td>TS-p-ea3-c-B</td>
<td>45.4</td>
<td>36.8</td>
</tr>
<tr>
<td>TS-p-ea3-c-L</td>
<td>40.9</td>
<td>32.6</td>
</tr>
<tr>
<td>TS-p-ea3-t-B</td>
<td>38.0</td>
<td>30.8</td>
</tr>
<tr>
<td>TS-p-ea3-t-L</td>
<td>40.4</td>
<td>31.9</td>
</tr>
<tr>
<td>TS-p-ea3-t-L</td>
<td>35.7</td>
<td>28.9</td>
</tr>
<tr>
<td>TS-p-ea3-t-L</td>
<td>36.8</td>
<td>30.7</td>
</tr>
<tr>
<td>TS-p-ee3-c-B</td>
<td>34.7</td>
<td>29.4</td>
</tr>
<tr>
<td>TS-p-ee3-c-B</td>
<td>35.1</td>
<td>30.1</td>
</tr>
<tr>
<td>TS-p-ee3-c-L</td>
<td>32.3</td>
<td>28.2</td>
</tr>
<tr>
<td>TS-p-ee3-c-L</td>
<td>35.0</td>
<td>28.7</td>
</tr>
</tbody>
</table>
the formation of a relatively unfavorable unsaturated four-co-
ordinate rhodium complex. The energy barrier is low (∼14 kcal
mol⁻¹) for the ea species, and even a bit less for the ee spec-
ies. After the acyl formation step, CO is coordinated to the
complex to form the five-coordinate acyl species. This is less
energetically favorable than one would expect, mainly owing
to entropy contributions. Also, there is a significant energy bar-
rier between the four- and five-coordinated species.

Experimentally it is not clear whether the final reaction with
hydrogen occurs via an oxidative addition–reductive elimina-
tion pathway or via a concerted metathesis-like mechanism.
In our calculations, we could not find a pathway that supports
the concerted mechanism. Instead, if the rhodium acyl com-
plex is approached by hydrogen, oxidative addition occurs
relatively low energy barriers (< 20 kcal mol⁻¹).

The oxidative addition of hydrogen to the four-coordinated
rhodium acyl complexes can occur in two ways for each com-
plex: one in which both hydride atoms are situated in the cis
position of the acyl species and one with one hydrogen atom in
the cis position and one hydrogen atom in the trans
position of the acyl species. The latter geometry is expected to be
energetically less favorable because two strong sigma donors
(a hydride and the acyl species) are coordinated trans with re-
spect to one another. This is confirmed by the calculations,
showing that the trans complex is 3–8 kcal mol⁻¹ higher in
energy (Figures 11 and 12). In addition, the barriers leading to
the rhodium(III) intermediates with cis and trans hydrogen
atoms are higher. We therefore assume that the cis–trans hy-
dride complexes do not play a role here.

An analysis of the energy profiles reveals that the lowest
energy pathway for the last stage of the reaction also involves
a change in coordination geometry. The trans-bisphosphine
acyl is transformed into the cis-bisphosphine hydride after the
reductive elimination of the product aldehyde. Interestingly,
the cis-bisphosphine acyl is also preferably converted to the
cis-bisphosphine hydride during the elimination step, which
means that the catalytic cycle ends with the cis-bisphosphine
hydride. After coordination of CO, which leads to the ea rest-

 Catalytic cycle with the monophosphate rhodium complex as a catalyst

The mechanism and energy profile for the pathway of the
monophosphate rhodium complex are similar to those for the
pathway of the bisphosphate analogue. Thus, we will not dis-
cuss the pathway in detail, but instead focus the discussion on
the differences between the two pathways.

The energy diagram of the first stage is depicted in
Figure 13. The pathway starting from e1 involves the expected
low barrier CO dissociation and alkene coordination. The trans-bisphosphine
acyl is transformed into the cis-bisphosphine hydride after the
reductive elimination of the product aldehyde. Interestingly,
the cis-bisphosphine acyl is also preferably converted to the
cis-bisphosphine hydride during the elimination step, which
means that the catalytic cycle ends with the cis-bisphosphine
hydride. After coordination of CO, which leads to the ea rest-

[Figure 10. Mechanism of the CO insertion of the ea–alkyl complex, which results in the trans-acyl complex. Bond distances are given in Å.]

[Figure 11. Energy diagram of the third stage of the catalytic cycle with the bisphosphate catalyst via the cis–cis dihydride species (with respect to the acyl species). The way different pathways connect the intermediates is shown at the bottom.]

[Figure 12. Energy diagram of the third stage of the catalytic cycle with the bisphosphate catalyst via the cis–trans dihydride species (with respect to the acyl species).]
The free energy of activation for the hydride migration step (Figure 14) is similar for monophosphines and bisphosphines: 11.8 and 13.1 for mono-P compared with 11.0 and 13.0 for bis-P. The lowest energy pathway is the one in which the phosphine configuration is retained: migrating from e3 to cCO4 via TS-e3-c. Importantly, the overall energy barrier calculated from the resting state up to this transition state significantly differs for the mono- and bisphosphorus complexes. This barrier, which controls the rate of the reaction, amounts to 24.5 kcal mol$^{-1}$, which is lower than the corresponding barrier for the bisphosphine complex (28.9 kcal mol$^{-1}$). This result is in agreement with the higher reactivity observed for the monophosphine coordination complex in the supramolecular capsule. The origin of this energy difference seems to lie mostly in the substrate coordination step (2 to 3), which is 6 kcal mol$^{-1}$ more endergonic for the bisphosphine complex. This is not surprising because the bisphosphine complex is more electron-rich than the monophosphine complex.

The overall energy barriers of the non-dispersion-corrected calculations are compared with those of the dispersion-corrected calculations in Table 3. Again the barriers decrease by a few kcal mol$^{-1}$. The difference between corrected and noncorrected calculations is smaller than that in the case of bisphosphines, which is expected because there are less internal steric and van der Waals interactions.

The same trend is also seen in the case of propene as a substrate (Table 4). The DFT and DFT-D3 barriers differ slightly more than that in the case of ethene and slightly less than that in the case of bisphosphines with propene. This is all as expected taking into account the amount of internal steric interactions: The more sterically crowded bisphosphine system benefits more from the inclusion of attractive van der Waals forces.

![Figure 13. Energy diagram of the first stage of the catalytic cycle with the monophosphine catalyst until substrate coordination.](image)

![Figure 14. Energy diagram of the hydride migration reaction with the monophosphine complex. The way different pathways connect the intermediates is shown at the bottom.](image)

<table>
<thead>
<tr>
<th>Table 3. Free energy barriers for the rate-determining monophosphine transition states from a1 with ethene as a substrate.</th>
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<tr>
<td></td>
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<tr>
<td>TS-a3-c</td>
</tr>
<tr>
<td>TS-e3-c</td>
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<td>TS-e3-t</td>
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<tr>
<th>Table 4. Free energy barriers for the rate-determining monophosphine transition states from a1 with propene as a substrate.</th>
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<td></td>
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<tr>
<td>TS-p-a3-c-B</td>
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<tr>
<td>TS-p-a3-c-B</td>
</tr>
<tr>
<td>TS-p-a3-c-L</td>
</tr>
<tr>
<td>TS-p-a3-c-L</td>
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<tr>
<td>TS-p-e3-c-B</td>
</tr>
<tr>
<td>TS-p-e3-c-B</td>
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<tr>
<td>TS-p-e3-c-L</td>
</tr>
<tr>
<td>TS-p-e3-c/t-L</td>
</tr>
<tr>
<td>TS-p-e3-t-B</td>
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<tr>
<td>TS-p-e3-t-B</td>
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<tr>
<td>TS-p-e3-t-L</td>
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difference does not occur in the case of the bisphosphine pathway.

The energy diagram of the second stage of the catalytic cycle is depicted in Figure 16. CO coordinates to the unsaturated square planar complex to form the five-coordinated complex. This binding of CO to complex 4 is energetically less favorable than we found for the bisphosphine analogue, but it follows a similar pathway. Interestingly, CO insertion follows a slightly different pathway. The transition states TS-a5 and TS-e5 are similar in structure (Figure 17), which differ only in the way the phosphine ligand is rotated compared to the rest of the complex. Both these transition states subsequently lead to the same product, eCO6. The energy barriers for these transition states are significantly higher than those for the bisphosphine pathway, but because these energy barriers are after the rate-determining step this does not affect the overall rate. The binding of CO to the square planar complexes 6 to form the saturated 18-electron species is slightly more favorable than we observed for the bisphosphine, but otherwise it is similar.

For hydrogenolysis (Figures 18 and 19), the pathway is similar to that of the bisphosphine complexes; however, the free energy of activation of the oxidative addition reaction is significantly lower for the monophosphine complexes. This is surprising because oxidative addition normally becomes faster as complexes become more electron rich.[28] However, if we look at van der Waals models of the transition states leading to the six-coordinated complexes, it becomes clear that steric influences are likely important here (Figure 20). The final reductive
elimination shows similar barriers and pathways for the mono- and bis-P complexes. The cis–trans hydrogen complexes and the associated transition states are again energetically unfavorable, and therefore the reaction has to proceed via the cis–cis complex. Starting from cC06, the most favorable route is via TS-cC06-cc and TS-cCtH-cC08-t, which then regenerates tCO2.

The configuration of the complex does not change in the path from hydride migration up to CO insertion (for the lowest energy pathway), but it does in the final hydrogenolysis step. Thus, for the mono-coordinated complex, there is an overall change in the coordination mode during the catalytic cycle. The pathway that is slightly higher in energy but still relevant for the overall reaction involves a change in configuration in the hydride migration step; that is, similar to that observed for the bisphosphine pathways, the reaction ends with the higher energy resting state and a reorganization may occur before it enters the next catalytic cycle.

Our results are in line with the previously published theoretical work on hydroformylation: The hydride complex is the resting state of the reaction, and the hydride migration step is associated with the highest barrier. The former is supported by spectroscopic evidence, whereas the latter is in line with the kinetic data (see the Introduction). The absolute barriers found herein are slightly different from those published earlier, which is clearly due to the difference in method used for the calculation as well as the difference in ligands used. In one of the more recent studies, for example, by Jensen et al.[13] (one that explicitly used triphenylphosphine, although only mono-ligated), a value of 105 kJ mol\(^{-1}\) (25 kcal mol\(^{-1}\)) was found for the overall barrier for the reaction from the hydride resting state to the hydride migration rate-limiting transition state. This corresponds well with the 25.5 kJ mol\(^{-1}\) value we found without dispersion corrections for the monophosphine complex. Jensen et al. found a value of 74 kJ mol\(^{-1}\) (18 kcal mol\(^{-1}\)) for the overall barrier for the reaction from the five-coordinated acyl to the release of the product. This corresponds qualitatively with the 19.7 kJ mol\(^{-1}\) value that we found (TS-cCtH-cC08-t-e7). Hirsh and Carbo[15] recently compared the reaction barrier of ethene hydroformylation using phosphine ligands with both mono- and bis-ligated triarylphosphines and found values of 22.4 and 25.9 kJ mol\(^{-1}\), respectively, which correspond well with the values found herein.

Conclusions

We have reported here, for the first time, DFT calculations on the entire catalytic cycle of the hydroformylation reaction mediated by both mono- and bis-triphenylphosphine rhodium complexes involving the entire molecular structure. For both catalytic cycles, we found similar pathways, and according to these DFT calculations the rate-determining step is in both cases associated with the migratory insertion of the olefin into the Rh–H bond. By using these energy profiles, type I kinetics is observed in both cases, which is found experimentally for 1-octene under standard conditions.[16]

The key hydride migration step of the process was investigated in detail by comparing ethene and propene substrates, and we also investigated the effect of dispersion corrections. The relative energies of the resting states (1) are not predicted correctly without dispersion corrections. In contrast to experimental observations, the uncorrected DFT energies suggest that monophosphine complexes are lower in energy. Dispersion-corrected (DFT-D3) calculations suggest that bisphosphine complexes are lower in energy, which is in line with experimental observations. With dispersion corrections, the overall barrier for ethene hydroformylation with the bisphosphine catalyst is predicted to be 3.0 kcal mol\(^{-1}\) lower that that for ethene hydroformylation with the monophosphine catalyst. With propene as the model substrate, the overall barrier is predicted to be only 1.9 kcal mol\(^{-1}\) lower for the monophosphine pathway. In addition, the trend in selectivity that is predicted for the hydroformylation of propene mediated by the mono- and bisphosphine complexes is correct: For the bisphosphine complex, the linear reaction is predicted to be nine times faster than the branched reaction, whereas for the monophosphine complex, the linear reaction is predicted to be only three times faster than the branched reaction. This supports the calculated pathways and applied computational models.

The porphyrin-encapsulated catalysts were 10 times more active than the nencapsulated analogues. The present calculations suggest that this increase in rate is most likely a consequence of the change from bisphosphine to monophosphine coordination. The overall energy barrier for the reaction is 1.9 kcal mol\(^{-1}\) lower for the monophosphine complex, which roughly corresponds with a rate enhancement by a factor of 30.[23] The selectivity of these monophosphine complexes as predicted by these calculations, however, favors the linear product, which means that the steric hindrance imposed by the capsule play an important role in directing the selectivity toward the branched aldehyde. It is difficult to obtain precise experimental (kinetic) data on the use of monophosphine-ligated rhodium complexes in hydroformylation because these complexes so far have been obtained in a mixture of complexes. The present computational data nicely demonstrate that the high reaction rates observed for the trispyridylphosphine–porphyrin-based catalysts are mostly due to the coordi-
nation, whereas the selectivity of these catalysts is more con-
trolled by the confined space created around the active site.

Experimental Section

The uncorrected DFT calculations were performed with Gaussian 03,[29] Minima on the potential energy surface were characterized by being stationary points without negative eigenvalues. Transition states were characterized by having one negative eigenvalue. Ei-
generator following (called IRC in Gaussian), was used to confirm
which minima were connected by the transition states. The B3LYP
gegen following (called IRC in Gaussian), was used to confirm
that range. This error was probably well within the accuracy of the
hydrogen, and everything originating from these), the 6-311G**
basis set was used, and for the phenyl groups on the triphenyl-
phosphine ligands, 6-31G* was used.

The DFT-D3 dispersion-corrected geometry and transition-state op-

calculated accordingly by using standard thermodynamic relations.

A simple conformational study was performed at the PM3 level for
ree1 and ea1, which assured us that any error stemming from con-
formational freedom of the triphenylphosphine ligands would be
smaller than 2 kcal mol$^{-1}$ because almost all minima were within
that range. This error was probably well within the accuracy of the
calculations (for details, see the Supporting Information). Confor-
national freedom of the alkyl and acyl substituents was expected to
give even lower errors.

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[27] Because the back reaction from CO2 to a1 is without barrier on the (self-consistent field) potential energy surface, the free energy barrier is completely due to entropy. The enthalpy of TS-a1 should be close to the enthalpy of CO2, whereas the entropy of TS-a1 should be close to the entropy of a1: \( \Delta G_{\text{TS-a1}} \approx \Delta H_{\text{CO2}} - T \Delta S_{\text{a1}} \).


