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Low-Valent Iron(I) Amido Olefin Complexes as Promoters for Dehydrogenation Reactions**

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Abstract: Fe° compounds including hydrogenases show remarkable properties and reactivities. Several iron(I) complexes have been established in stoichiometric reactions as model compounds for N₂ or CO₂ activation. The development of well-defined iron(I) complexes for catalytic transformations remains a challenge. The few examples include cross-coupling reactions, hydrogations of terminal olefins, and azide functionalizations. Here we report the syntheses and properties of bimetallic complexes [MFe°-(tropdae)(solv)] with a d° Fe low-spin valence-electron configuration (M = Na, solv = 3thf; M = Li, solv = 2EtO; trop = 5H-dibenzo[a,d]cyclohepten-5-yl, dae = (N-CH₂CH₂N) with a d° Fe low-spin valence-electron configuration are reported. Both compounds promote the dehydrogenation of N,N-dimethylaminoborane, and the former is a single-molecule magnet with a high spin-reversal barrier.

Fe° compounds including hydrogenases show remarkable properties and reactivities. The anion [Fe(C(SiMe₃)₂)₃]⁻ is a single-molecule magnet with a high spin-reversal barrier. In stoichiometric reactions, iron(I) diketiminate, iron(I) tris(phosphino)borates, and related species have been established as model compounds for N₂ activation and for the cleavage and coupling of CO₂. The development of well-defined iron(I) complexes for catalytic transformations, however, remains a challenge. The few examples include cross-coupling reactions, hydrogations of terminal olefins, and azide functionalizations. Here we report the syntheses and properties of bimetallic complexes [MFe°-(tropdae)(solv)] with a d° Fe low-spin valence-electron configuration (M = Na, solv = 3thf; M = Li, solv = 2EtO; trop = 5H-dibenzo[a,d]cyclohepten-5-yl, dae = (N-CH₂CH₂N)) with a d° Fe low-spin valence-electron configuration are reported. Both compounds promote the dehydrogenation of N,N-dimethylaminoborane, and the former is a single-molecule magnet with a high spin-reversal barrier.

The reaction of the amine H₄tropdae with one equiv of [FeCl₂(thf)]₃ and three equiv of [Na(CH₂SiMe₃)] as base and reducing agent in THF at −30°C gave the bis(amido)-diolefin complex [NaFe(tropdae)(thf)] (Scheme 1). Compound 1 was isolated as a deep red, single-crystalline material, which is soluble in polar solvents such as THF but also in aromatic hydrocarbons. To study the influence of the counterion, [LiFe(tropdae)(Et₂O)] (2) was synthesized in a similar manner by using a slight excess of [Li(CH₂SiMe₃)]. Compound 2 was isolated as deep-red single crystals and shows a solubility similar to that of 1. The reactions leading to the formation of complexes 1 and 2 are complex and involve at some stage ligand coordination, deprotonation, salt meta-
amid functionalities and the two olefinic groups of the \( \text{(trop,dae)}^- \) ligand. Taking the centroids of the C=C$_{\text{olefin}}$ units as coordination points, this results in a distorted square-planar coordination sphere around Fe1 (2$^\circ$ = 369$^\circ$). Planar coordination spheres have occasionally been reported for tetracoordinate iron(II) compounds, but are rare for iron(III) compounds.

The five-membered ring (Fe1-N1-C$_{\text{olefin}}$) adopts a twisted configuration with the carbon atoms located above and below the N1-Fe1-N2 plane. As a consequence of the additional interaction with the sodium cation Na1, the Fe1-N1 distance (1.900(4) Å) is slightly longer than Fe1-N2 (1.860(3) Å), and N1 resides in a slightly more pyramidal coordination sphere ($\Sigma$(C/Fe-N1-C) = 343$^\circ$; $\Sigma$(C/Fe-N2-C) = 350$^\circ$). The Fe1-N1/2 distances are 0.02–0.05 Å shorter than those in pseudotetrahedral nitrosyl amide complexes \( \text{Fe-NiBuArF}_2\text{NO}^{-}\text{L} \) ($\text{ArF}_2 = 2.5-\text{C}_5\text{H}_4\text{FeMe}$, L = NC$_6$H$_4$, PEn$_2$) and about 0.09 Å shorter than those in square-planar complexes of type \( \text{Fe}[(\beta-\text{ diketiminate})\text{L}]^- \) (L = CO, CNBu$_2$)\[^{8b,b,d} \] The C=C$_{\text{olefin}}$ bonds, C4=C5 and C19=C20, are rather long (1.415–1.438 Å) and indicate significant electron back-donation from the iron center into the π$^*$ orbital of the olefinic binding site (for shorter coordinated C=C bonds see Refs. [2b,5c]). Compound 2 also shows a distorted planar coordination geometry around Fe1, but both nitrogen atoms of the \( \text{(trop,dae)}^- \) ligand interact with the Li ion (Figure 1b).\[^{13} \] In comparison to 1, this results in slightly elongated Fe1–N1/2 bonds (0.04 Å in average), shortened C=C$_{\text{olefin}}$ bonds (0.02 Å in average), a smaller angle sum C/Fe-N2-C around N2 ($\Delta = -15.1(3)^\circ$), and in a decreased N1-Fe1-N2 angle ($\Delta = -5.40(2)^\circ$). The five-membered ring (Fe1-N1-C31-C32-N2) in 2 shows an envelope conformation with Fe1 displaced by 1.09 Å from the N1-C31-C32-N2 plane.

The subtle structural differences between 1 and 2 and their resulting electronic structures in the solid state were further investigated by $^{57}$Fe Mössbauer spectroscopy at 77 K (see the Supporting Information). The isomer shift of compound 1 ($\delta = 0.20(1)$ mm s$^{-1}$) is significantly smaller than that of 2 ($\delta = 0.28(1)$ mm s$^{-1}$), which indicates a higher electron density in the Fe$^\text{III}$-bonds of 1 and in agreement with the smaller average Fe–L bond lengths in 1\[^{14} \] The lower symmetry of the ligand environment in 1 leads to a larger quadrupole splitting ($|\Delta E_0| = 2.87(1)$ mm s$^{-1}$) compared to 2 ($|\Delta E_0| = 2.01(1)$ mm s$^{-1}$). These quadrupole splittings fall in the usual range for literature-known Fe$^\text{III}$ complexes ($|\Delta E_0| = 0.89–3.48$ mm s$^{-1}$).\[^{16,6,c} \] However, other Fe$^\text{III}$ complexes generally show larger isomer shifts (from $\delta = 0.28$ mm s$^{-1}$ to 1.09 mm s$^{-1}$, which indicates lower electron density in Fe$^\text{III}$ orbitals)\[^{10,b,2,a,15} \] Remarkably, Fe$^\text{III}$ and Fe$^\text{II}$ compounds with a fourfold planar coordination geometry as in 1 and 2 show similar isomer shifts \( \text{Fe}^\text{II}, 0.14–0.59; \text{Fe}^\text{III}, 0.23–0.37 \) mm s$^{-1}$ but associated with much larger \( |\Delta E_0| : \text{Fe}^\text{II}, 1.16–4.63; \text{Fe}^\text{III}, 3.02–5.16 \) mm s$^{-1}$\[^{10,a,16,17} \] The low Fe$^\text{II}/\text{Fe}^\text{III}$-like isomer shifts for 1 and 2 are possibly a consequence of metal-to-ligand electron back-donation. DFT calculations on 1 and 2 gave optimized structures in good agreement with the single-crystal X-ray data (see the Supporting Information). Derivatives \[ \text{MFe(trop,dae)(L)}_2 \] (M = Li, Na; L = neutral ligand bound to M; n = 0–3) and counterion-free \( \text{Fe(trop,dae)}^- \) were also investigated. All species are mainly metal-centered radials with significant spin polarization to the nitrogen donors and olefinic carbon atoms of the \( \text{(trop,dae)}^- \) ligand (Figure 2a). Hence, they are best described as low-spin d$^5$ Fe species hosted by dianionic \( \text{(trop,dae)}^- \) ligands. Significant differences between the electronic structures of 1 and 2 arise from the larger g anisotropy, higher Fe spin density, and higher NBO charge at Fe for 2 (see Table S2 in the Supporting Information). These differences are mainly due to the coordination mode of the alkali metal (terminal as in...
1 versus bridging as in 2), which can be controlled by choice of the neutral ligand L (see derivatives 2b–d in the Supporting Information). The nature of the alkali metal itself has a smaller effect.

The electronic structures of 1 and 2 were further investigated by EPR spectroscopy at 20 K. Undiluted powdered solids of 1 and 2 gave X-band EPR spectra of reasonable quality which are clearly distinct from each other (see Figure S10 in the Supporting Information). The X-band EPR spectra of 1 and 2 in frozen THF containing 0.1 M [N(nBu)]4[PF6] are highly similar, thus indicating the formation of the free anion [Fe(trop-dae)]− in both cases. Rhombic signals without any (resolved) hyperfine couplings were observed which are characteristic for low-spin d5 Fe species (Figure 2b).

The experimental g values (g∥ = 2.009, g⊥ = 2.060, gμ = 2.160) are in experimental agreement with those predicted by DFT calculations (see Table S2 in the Supporting Information). Unexpectedly, EPR spectroscopic analysis of 1 and 2 in toluene glasses at 20 K revealed broad and complex signals which are ascribed to aggregation phenomena (see Figure S13 in the Supporting Information). The spectra also show weak “half-field signals” which can be observed when S = 1/2 systems weakly interact in the matrix (see Figure S14 in the Supporting Information). The spectra of 1 and 2 in toluene are clearly different, thus indicating distinct aggregation behavior.

The low-spin d5 electron configuration of 1 was further confirmed by SQUID magnetization measurements. In an applied field of 1 T, an effective magnetic moment of μeff = 1.96 μB at 300 K was determined, which is almost invariant over a temperature range of 10–300 K (see the Supporting Information). This value is close to the spin-only value of 1.80 μB for one unpaired electron (for g = 2.076). In good agreement with this result, determination of the effective magnetic moment in benzene solution by the Evans method gave μeff = 2.0(1) μB for compounds 1 and 2. A cyclic voltammogram of 1 in THF at 23°C as a scan rate of 0.2 V s−1 shows a quasireversible redox wave with E1/2 = −2.24 V versus Fe/Fc− (see the Supporting Information), probably for an Fei to Feµ conversion, while irreversible oxidation half-waves were recorded at −0.56 V and at 0.05 V versus Fe/Fc−. The large splitting of about 1.7 V between the oxidation and reduction waves indicates a high stability of the Feµ complexes 1 and 2 with respect to disproportionation.

No reaction of 1 with H2 (1.5 bar, T = 25°C) was observed in nonpolar solvents. However, substrates with “krypto”-hydrogen, that is hydrogen in the form of H3+ or Hµ,[21] such as N,N-dimethylaminoborane (DMAB; Me₂HN-BH₃), are efficiently dehydrogenated.[22] Few iron-based catalysts have been reported for the dehydrogenation of DMAB.[23] Recent studies by Manners and co-workers show that in situ generated Feµ nanoparticles are frequently catalytically active and only [Fe(C₅H₄)(CO)₂] acts as a homogeneous catalyst.[23a,b,24] In an open system, 5 mol% of 1 and 2 were used as catalysts for the dehydrogenation of DMAB in toluene at room temperature (Table 1 and see the Supporting Information). Under these conditions, the LiFei species 2 showed only moderate activity (4 h, 35%, entry 1) and [LiFe(trop-dae)−(thf)]2 (2d) was almost inactive (entry 2). In contrast, the NaFei compound 1 led to full conversion after 4 h (entry 3), giving the 1,3-diazao-2,4-diborolane 4 as sole product (see Table 1). The rate constant kobs at early stages of the reaction increased by a factor of 2.1(2) upon increasing the overall concentrations by a factor of 2.0, and full conversion was reached after 1.3 h (entry 4, see also the Supporting Information). In contrast, additives such as 15-crown-5 or [nBu]4[NBr increase the reaction rate, which suggests that a solvent-separated ion pair with the [Fe(trop-dae)]− anion is not the most active species (entries 5 and 6). The use of deuterated Me₂NDBH₃ as a substrate and 1 as a catalyst revealed a kinetic isotope effect of 2.0(2), thus indicating that deprotonation of the amino group is one of the rate-limiting steps (entry 7, and see the Supporting Information). Reactions of 1 with stoichiometric amounts of Me₂NBH₃ showed no H abstraction; however, an alteration of the B–H stretching frequencies in the solid-state IR spectrum of the solid obtained after workup indicate weak interactions of the BHµ group with 1. According to 11B NMR studies, such interactions might be present in toluene solutions of 1, but not of 2 (see the Supporting Information). Catalyst 1 remains active after dehydrogenation of 20 equiv of DMAB, which was shown by consecutive addition of fresh substrate. Overall, at least 3 × 20 equiv of DMAB are dehydrogenated without any loss of activity (entry 8). In a closed system, the linear species 5 (see Table 1) was detected as the major intermediate by 11B NMR spectroscopy.[25–27] Only traces of monomeric intermediate Me₂N=BH₂ were detected.

When the reaction is carried out in THF, the conversion rate drops significantly (entry 9). In selective-poisoning experiments (with 0.2 equiv of PPh₃ or 0.1 equiv of P(OMe)₃,
per Fe$^{[234]}$ with 1 as catalyst, the reactions proceed to completion albeit at decreased rates (entries 10 and 11). Time-conversion plots for DMAB dehydrogenation with 1 as a catalyst do not show an induction period. Small aliquots of the reaction solutions were analyzed by scanning electron microscopy (SEM) and gave no indications for the formation of Fe nanoparticles. Although the reaction mechanism remains obscure, these results indicate that 1 acts as a homogeneous catalyst in the dehydrogenation of DMAB. Catalyst 1 has a very high activity compared to [Fe-(C$_{6}$H$_{4}$)(CO)$_{3}$]$_{n}$, which requires 9 h and continuous irradiation with UV light for full conversion.$^{[234]}$ The counterion effect, Na$^{+}$ > Li$^{+}$, may indicate that catalytically more-active aggregated species are formed with compound 1$^{[28,29]}$ and/or more effective substrate coordination by the [Na(thf)$_{3}$]$^{+}$-containing species.

Complex 1 also efficiently catalyzes the reaction of silanes with alcohols as an intermolecular variant of the release of “krypto”-hydrogen.$^{[30]}$ We are especially interested in the dehydrogenative alcoholysis of silanes with diols to yield oligo- or poly(alkyl silyl ethers) which has been scarcely exploited to date.$^{[31–33]}$ The use of iron-based catalysts for this type of reaction is unprecedented. The simple methanolation of PhSiH$_{3}$$^{[34]}$ is efficiently catalyzed with 3 mol% of 1 (1 mol% per Si–H bond), and complete conversion is reached after 5 min in toluene (T = 25°C). The reaction solution remains homogeneous and at least three consecutive catalytic runs can be performed without apparent loss in activity (see the Supporting Information). In reactions between 1,4-benzene-dimethanol as the diol and phenylsilane or diphenylsilane, three or two equivalents of H$_{2}$ are released and full conversion is reached after 15 min or 54 min, respectively (Scheme 2). Products 6 and 7 were isolated as off-white solids with at least nine (in the case of 6) or nineteen (in the case of 7) repeating units with respect to Si, as indicated by mass spectrometric analysis.

In summary, the trop-amine-type ligand (trop.dae)$^{2-}$ strongly stabilizes low-valent iron species, thereby allowing the synthesis of rare examples of heterobimetallic d$^{6}$ iron(I) amide complexes and their application as homogeneous catalysts. The structures, electronic properties, aggregation behavior in solution, and especially the reactivities depend sensitively on the counterion [Na(thf)$_{3}$]$^{+}$ versus [Li(Et$_{3}$O)$_{4}$]$^{+}$. Although the benchmark performance of [Ni(O$_{2}$CCF$_{3}$)-(NHtrop$_{3}$)]$^{[35]}$—another metalatrolradical first row transition-metal complex, but with a low-valent d$^{6}$ nickel(I) center—is not reached, [NaFe(trop$_{3}$.dae)(thf)$_{2}$] (1) is a remarkable dehydrogenation catalyst especially for the syntheses of oligo/poly(silyl ethers) from polyols and silanes. We believe that this reaction has the potential of becoming an atom-economic method for the synthesis of oligo- and polymeric alkyl silyl ethers under mild conditions, generating no waste, only hydrogen as a valuable by-product.

**Keywords:** aminoboranes · condensation reactions · dehydrogenation · heterometallic complexes · low-valent iron.


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Na$_1$ is found in an distorted trigonal bipyramidal coordination geometry with two thf molecules and N$_1$ in the equatorial positions. 


These compounds show high-spin electron configurations, which also contributes to their higher isomer shifts.


[14] The experimentally determined isomer shifts are well reproduced by DFT calculations. The experimentally determined quadrupole splittings deviate from calculated values (see the Supporting Information).

These compounds show high-spin electron configurations, which also contributes to their higher isomer shifts.


[15] These compounds show high-spin electron configurations, which also contributes to their higher isomer shifts.

exception is $[B(C_2F_5)_3]$ for the synthesis of poly(aryl silyl ethers) (Ref. [32d]).


[35] CCDC 999856 (1), 999857 (2), 999858 (3), 1033680 (2b), 1033681 (2c), 1033682 (2d), contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.