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DOI
10.1002/cssc.201500950

Publication date
2015

Document Version
Final published version

Published in
ChemSusChem

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Citation for published version (APA):

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Highly Soluble Benzo[ghij]perylenetriimide Derivatives: Stable and Air-Insensitive Electron Acceptors for Artificial Photosynthesis

Hung-Cheng Chen,[a] Chao-Ping Hsu,[b] Joost N. H. Reek,[a] René M. Williams,[a] and Albert M. Brouwer*[a]

A series of new benzo[ghij]perylenetriimide (BPTI) derivatives has been synthesized and characterized. These remarkably soluble BPTI derivatives show strong optical absorption in the range of λ = 300–500 nm and have a high triplet-state energy of 1.67 eV. A cyanophenyl substituent renders BPTI such a strong electron acceptor (E_ad = –0.11 V vs. the normal hydrogen electrode) that electron-trapping reactions with O_2 and H_2O do not occur. The BPTI radical anion on a fluorine-doped tin oxide (FTO) electrode is persistent up to tens of seconds (τ_{1/2} = 39 s) in air-saturated buffer solution. As a result of favorable packing, theoretical electron mobilities (10^{-2}–10^{-1} cm^2 V^{-1} s^{-1}) are high and similar to the experimental values observed for perylene diimide and C_60 derivatives. Our studies show the potential of the cyanophenyl-modified BPTI compounds as electron acceptors in devices for artificial photosynthesis in water splitting that are also very promising non-fullerene electron-transport materials for organic solar cells.

Introduction

Perylene-3,4,9,10-tetracarboxylic acid diimide derivatives (perylene diimides, PDIs) are well-known dyes and pigments with widespread application in industry.\[1\] As a result of their outstanding photochemical stability and possibilities for chemical functionalization,\[2\] they have been investigated intensively in academic research on luminescent solar concentrators,\[3\] biolabeling and -imaging,\[4\] building blocks for functional nanomaterials,\[5\] fluorophores for single-molecule spectroscopy,\[6\] etc. Self-assembled nanostructures of PDIs feature strong π–π stacking interactions between the perylene backbones and a good electron-accepting character because of the imide moieties.\[7\] Outstanding charge-transport properties were also found in functionalized PDIs.\[8\] For this reason they have been utilized intensively in various organic electronic applications, such as n-type semiconductors in organic field-effect transistors (OFETs),\[9\] and electron acceptors in organic photovoltaic cells (OPVs).\[10, 11\]

PDIs have been considered as potential acceptors in bulk heterojunction organic solar cells to replace fullerene.\[12\] However, the solar energy conversion efficiencies obtained in conjugated polymers\[13\] or small-molecule\[14\] photovoltaic blends with PDIs were significantly lower than those with fullerene-based acceptors. Recently, ultrafast spectroscopy studies suggested that the low internal quantum efficiency is the result of fast geminate recombination at the low-band-gap conjugated polymer–PDI interface.\[15\] We also showed that geminate and nongeminate charge recombination mechanisms result in PDI triplet-state formation in pyrene/PDI thin-film blends.\[16\]

One reason for the lower efficiency is charge loss because the energy of the triplet excited state of a PDI (E(T_1) ≈ 1.2 eV) is often below that of the charge-transfer state. Consequently, charge recombination from charged states to the triplet state of PDIs can be thermodynamically favorable and efficient. Theoretical work also indicates that fast charge recombination from the charge-transfer states to the lower-lying triplet state of a PDI can play a role in a phthalocyanine/PDI system.\[17\]

Recently, a high excited triplet-state energy (E(T_1) ≈ 1.68 eV) was reported for benzo[ghij]perylenetriimide (BPTI), which has a more expanded π-conjugated backbone than PDI.\[18\] This value is even higher than that of fullerene systems such as C_60 (E(T_1) ≈ 1.63 eV)\[19\] and [60]PCBM (E(T_1) ≈ 1.50 eV).\[20\] Consequently, in the design of long-lived charge-transfer states,\[21\] the BPTI chromophore could potentially provide the solution to avoid triplet formation by charge recombination. Furthermore, as a result of the small dielectric constant (ε_r ≈ 2–4) in organic solar cells, the large Coulombic attraction of photogener-
ated electron–hole pairs results in low quantum yields of separation into long-lived dissociated charges.[22] A recent theoretical study suggested that solar cells with over 20% power conversion efficiencies can be achieved by increasing the \( \varepsilon \), to 10 in the cell.[23,24] Although several factors, which include orientation, geometry, packing density, and molecular polarizability, can influence the dielectric response of molecule-based systems,[25] it was shown that the increase of the polarity of the building blocks is an effective approach to increase the dielectric constant of organic \( \pi \)-conjugated materials by amplifying the chemical polarity as well as the ground-state dipole moment. Molecular design strategies include the introduction of polar groups (e.g., fluorine atoms substituted on the chromophore backbone)[26] or side chains (e.g., ethylene glycol, nitrite, functionalized peptide)[27–29] into the molecular structure. Recently, Kulkarni et al. reported that a high dielectric constant (up to 16) of an organic \( \pi \)-conjugated material can be obtained from a parallel orientation of molecular dipoles ordered by a stacked supramolecular assembly of PDIs.[30] An impressively high dielectric constant of 15 was also obtained in a hyper-branched copper phthalocyanine polymer with a high molecular polarizability because of the presence of cyanophenyl groups.[31, 32] Consequently, to design new BPTI derivatives, we took into account the large local dipole of the cyanophenyl moiety, which was introduced in the five-membered ring imide position on the BPTI backbone.

To improve the solubility and processability of PDI derivatives, substituents can be added at the N atoms or at the bay positions.[33–34] Unfortunately, the prevention of \( \pi \)-\( \pi \) stacking with these strategies affects the electronic material properties negatively. Firstly, the low crystalline order as well as the low delocalization of the electronic charge density in the acceptor phase causes donor/PDI blends to exhibit high barriers to charge-transfer state dissociation, certainly compared to fullerene-based acceptors.[35] Secondly, it also results in a large free volume in the donor–acceptor (D–A) interfacial region and thereby decreases the average D/A electronic coupling.[36]

Electron D/A molecular systems that can undergo efficient photoinduced electron transfer are among the key requirements for solar-to-fuel artificial photosynthetic devices, a high electrochemical reduction potential of the electron donor (photosensitizer) radical cation, \( E(D/D^+) \) is required.[38–40] In other words, a rather poor electron donor should be employed. The reaction free energy criteria to achieve a high quantum yield of photoinduced charge separation in the assembled D/A molecular system then dictates the use of a strong electron acceptor.[41]

In the present work, we show that planar BPTI derivatives with a cyanophenyl substituent on the five-membered ring imide, and two alkyl groups on the six-membered imides, are excellent electron acceptors that combine a surprisingly good solubility with the possibility of favorable \( \pi \) stacking in the solid state that leads to good charge-transport properties (Scheme 1). The electrochemical and spectroscopic properties of these BPTI derivatives with different side chains were investigated. PDI-1, which is used extensively as an electron acceptor in solution-processed organic photovoltaics, was also studied for comparison.[44] Quantum-chemical calculations give an insight into the electronic structure, reorganization energy for electron transfer, and ground-state dipole moment of these BPTI derivatives and predict high electron mobilities. When bound to a fluorine-doped tin oxide (FTO)|TiO₂ electrode, BPTICNP-2 can act as an air- and water-stable negative-charge relay.

![Scheme 1. Chemical structures and synthesis of BPTI derivatives through (a) Diels–Alder reactions on the PDI scaffold followed by substitution reactions with (b) 3-aminopentane or (c) 4-aminobenzonitrile.](http://www.chemsuschem.org)

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Results and Discussion

Synthesis

To obtain our target BPTI compounds, we followed the conventional synthetic approach shown in Scheme 1 based on (a) Diels–Alder reactions on the PDI scaffold followed by substitution reactions with (b) 3-aminopentane or (c) 4-amino-benzonitrile. The reactions of the PDI derivatives PDI–1–PDI–4 and maleic anhydride in the presence of p-chloranil as an oxidizing agent result in the formation of diimido anhydrides BBDI–1–BBDI–4. Nitrobenzene was added to improve the solubility of the PDIs during the reaction with maleic anhydride. As a result of the very poor solubility of BBDI–1–BBDI–4 in common solvents at room temperature, column chromatography purification could not be used. Dry crude BBDI–1–BBDI–4 were used directly after washing with methanol and acetone for the subsequent condensation reaction. During the condensation reactions, the complete dissolution of BBDI–1–BBDI–4 in hot DMF (−140 °C) was observed. In contrast to BPTICNP–4, the compounds BPTI–1, BPTICNP–1, BPTICNP–2, and BPTICNP–3 are all highly soluble at room temperature in common laboratory solvents such as dichloromethane, chloroform, THF, and DMF. The high room-temperature solubility of BPTICNP–2 and BPTICNP–3, which have no α-branched alkyl chains or bulky groups on the imide positions, is quite unusual. Particularly, the solubility of BPTICNP–2 in chloroform is up to 15 mg mL⁻¹ (−15 °C) at room temperature; which is the required concentration for spin-coating in the fabrication of organic photovoltaic devices. The length of the alkyl chains influences the solubility of BPTICNP–2–BPTICNP–4 in chloroform dramatically. A solubility of 15 mg mL⁻¹ of BPTICNP–3 in chloroform can be achieved at −50 °C in a hot-water bath (Figure S1). Unfortunately, BPTICNP–4 did not show suitable solubility in chloroform or other organic solvents such as ortho-dichlorobenzene, DMF, or THF at room temperature.

Steady-state spectroscopic studies

Room-temperature absorption and fluorescence spectra were measured to evaluate the electronic structure of the BPTI derivatives (Figure 1 and Table 1). Absorption spectra of PDI–1, BPTI–1, and BPTICNP–1–BPTICNP–3 in chloroform are shown in Figure 1a. The benzo[ghi]perylene chromophore shows a blue-shifted absorption compared to that of the PDIs. The typical vibronic fine structure of the strongly allowed benzo[ghi]perylene core π−π⁺ transitions (λ ≈ 460 nm), 6.2 × 10⁷ cm⁻¹ m⁻¹) on all BPTIs is apparent, and three bands at λ = 467, 436, and 410 nm were observed in the longer-wavelength region. The fact that all BPTI compounds have the same S₂→S₁ transition energy shows that the different alkyl groups on the imide and the presence of the cyanophenyl group do not influence the optical band gap. The absorption bands in the range of λ ≈ 350–400 nm are assigned to the S₁→S₃ transition. There is a small redshift (635 cm⁻¹) on the S₂→S₃ transition of the BPTICNP analogues compared to BPTI–1. The transition dipole moment of S₂→S₃ is along the short-axis direction on BPTICNP, parallel to the direction of the cyanophenyl axis (blue arrow in Figure 1c).

The normalized fluorescence spectra are shown in Figure 1b. The substituents on BPTI have a negligible influence on the positions of the emission maxima.

The excited triplet-state energies of the BPTI analogues were determined from emission maxima in phosphorescence spectra in glassy matrices (2-methyltetrahydrofuran/iodoethane 1:1) at 77 K (Figure 2). The introduction of the polar cyanophenyl group in BPTICNP–1 or BPTICNP–2 does not result in any shift of the shortest-wavelength emission maxima. The phosphorescence band corresponds with a triplet excited-state energy of 1.67 eV for BPTI–1, BPTICNP–1, and BPTICNP–2. This value is very

![Figure 1](image-url)

**Figure 1.** (a) UV/Vis absorption spectra and (b) fluorescence spectra of PDI–1 (λ_ex = 480 nm) and BPTI (λ_ex = 440 nm) derivatives in CHCl₃. (c) Transition dipole moments of BPTICNP. The red and blue arrows correspond to the S₁→S₃ and S₂→S₅ optical transitions, respectively.

<table>
<thead>
<tr>
<th>Compound</th>
<th>λ_ex (nm)</th>
<th>ε (10⁶)</th>
<th>λ_em (nm)</th>
<th>S₂ (eV)</th>
<th>T₁ (eV)</th>
<th>T₁ (eV) theor.</th>
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</thead>
<tbody>
<tr>
<td>PDI–1</td>
<td>466</td>
<td>1.36</td>
<td>505</td>
<td>2.26</td>
<td>1.67</td>
<td>1.69</td>
</tr>
<tr>
<td>BPTI–1</td>
<td>466</td>
<td>1.36</td>
<td>505</td>
<td>2.26</td>
<td>1.67</td>
<td>1.69</td>
</tr>
<tr>
<td>BPTICNP–1</td>
<td>466</td>
<td>1.36</td>
<td>505</td>
<td>2.26</td>
<td>1.67</td>
<td>1.69</td>
</tr>
<tr>
<td>BPTICNP–2</td>
<td>466</td>
<td>1.36</td>
<td>505</td>
<td>2.26</td>
<td>1.67</td>
<td>1.69</td>
</tr>
<tr>
<td>BPTICNP–3</td>
<td>466</td>
<td>1.36</td>
<td>505</td>
<td>2.26</td>
<td>1.67</td>
<td>1.69</td>
</tr>
</tbody>
</table>

[a] The longest wavelength absorption maxima. [b] Excitation wavelength λ_ex = 480 nm in PDI–1 and λ_ex = 440 nm in all BPTI derivatives. [c] Excited at 460 nm. [d] Derived from the maxima of the phosphorescence band. [e] Theoretical triplet-state energy from TDDFT calculations with B3LYP with the 6-31G(d) basis set. [f] Ref. [20] [g] Ref. [19]. [h] Ref. [20].
similar to that of BPTI with branched alkyl side chains. All spectral data are listed in Table 1. The $T_1$ energies of the BPTI compounds are 0.57 eV higher than that reported for N-hexyl-heptyl-substituted PDI. Furthermore, the $T_1$ energies of BPTI compounds are higher by 40 meV than that of C$_60$ [E($T_1$) $\approx$ 1.63 eV] and 170 meV than that of [60]PCBM [E($T_1$) $\approx$ 1.50 eV]. The theoretical triplet excited-state energies obtained with time-dependent density functional theory (TDDFT) calculations using the B3LYP functional with the 6-31G(d) basis set are also included in Table 1. The calculated $T_1$ energies of all BPTI compounds are in close agreement with the experimental values. These results indicate that TDDFT with B3LYP/6-31(d) is a good approach to calculate the triplet excited-state energy of BPTI compounds.

Spectro-electrochemical characterization of reduced BPTI

The radical anion species of BPTIs were produced by electrochemical reduction in dichloromethane at room temperature. The absorption spectra of the radical anions of BPTI-1, BPTICNP-1, and BPTICNP-2, are shown in Figure 3. In all cases, new broad bands appear during the reduction process with isosbestic points at $\lambda$ = 480 nm for BPTI-1, and BPTICNP-1 and $\lambda$ = 486 nm for BPTICNP-2. The spectral features of the radical anions are throughout the visible and near-infrared (NIR) regions, which all possess low-energy transitions that correspond to D$_0$ $\rightarrow$ D$_n$. The radical anion BPTI-1$^-$ shows four absorption bands with maxima at $\lambda$ = 531 (e = 11 200), 644 (31 100), 720 (15 500), and 864 nm (13 000 m$^{-1}$ cm$^{-1}$) at an estimated conversion of $\approx$ 90%. In comparison with BPTI-1, the spectral features are not so clearly different for BPTICNP-1$^-$, which also displays four similar absorption bands with maxima at $\lambda$ = 531, 645, 721, and 864 nm. Clearly, the introduction of the cyanophenyl unit shifts the reduction potential to more positive values, but it does not induce spectral shifts in the radical anion spectra. The BPTICNP-2$^-$ radical anion also shows similar spectra with absorption bands with maxima at $\lambda$ = 529, 647, 717, and 856 nm and two enhanced shoulders at $\lambda$ = 595 and 762 nm.

Electrochemical studies

Cyclic voltammograms of PDI-1 and BPTI analogues in deaerated CH$_2$Cl$_2$ that contained 0.1 M NBu$_4$PF$_6$ are shown in Figure 4. For reasons of solubility, C$_60$ was measured in toluene/1,2-dichloroethane (1:1) solution. The measured half-wave potentials $E_{1/2}$ are collected in Table 2. Ferrocene was used as an internal standard, and the ferrocene/ferrocenium redox couple has a half-wave potential $E_{1/2}$ = 0.690 V vs. the normal hydrogen electrode (NHE) in CH$_2$Cl$_2$. All potentials are related to the NHE.

<table>
<thead>
<tr>
<th>Table 2. Redox properties and estimated LUMO energies of PDI-1, BPTI derivatives, and C$_60$ from cyclic voltammetry.</th>
<th>$E_{1/2}$ [V vs. NHE]</th>
<th>$E_{LUMO}$ [eV]</th>
</tr>
</thead>
<tbody>
<tr>
<td>Compound</td>
<td>X/X</td>
<td>X$^+$/X$^{2+}$</td>
</tr>
<tr>
<td>PDI-1</td>
<td>$-0.39$</td>
<td>$-0.58$</td>
</tr>
<tr>
<td>BPTI-1</td>
<td>$-0.41$</td>
<td>$-0.67$</td>
</tr>
<tr>
<td>BPTICNP-1$^{N}$</td>
<td>$-0.21$</td>
<td>$-0.43$</td>
</tr>
<tr>
<td>BPTICNP-2$^{N}$</td>
<td>$-0.11$</td>
<td>$-0.42$</td>
</tr>
<tr>
<td>BPTICNP-3$^{N}$</td>
<td>$-0.10$</td>
<td>$-0.43$</td>
</tr>
<tr>
<td>C$_60$</td>
<td>$-0.35$</td>
<td>$-0.73$</td>
</tr>
</tbody>
</table>

[a] The potential of 0.24 V vs. NHE corresponds to 0.0 V vs. SCE. $E_{LUMO}$ = $[E_{1/2}^{\text{ferrocene}} - 4.4]$ eV. Refs. [58, 59]. [b] The potential values are derived from DPV measurements.

Figure 3. UV/Vis spectral changes in the first electron reduction of (a) BPTI-1, (b) BPTICNP-1, and (c) BPTICNP-2 in deaerated CH$_2$Cl$_2$ that contained 0.1 M NBu$_4$PF$_6$ (0.1 M).
The first X/X⁻, the second X⁻/X²⁻, and the third X²⁻/X³⁻ redox couples of BPTI-1 were reversible and clearly detectable (Figure 4a). If we compare the first and second redox potentials of BPTI-1 and PDI-1, we can see that these values are only slightly negatively shifted in BPTI-1: the additional electron-withdrawing imide moiety on BPTI-1 does not improve the electron-accepting strength of BPTI. The results for C₈ and cyanophenyl-substituted BPTI (BPTICNP-1 and BPTICNP-2) are shown in Figure 4b. Although the first redox waves are not identified clearly in BPTICNP-1 and BPTICNP-2, a gradual shift of the onset potential to more positive values can be observed on going from C₈ to BPTICNP-2. This indicates that the cyanophenyl-substituted BPTIs have a stronger electron-accepting character than C₈. To obtain a clearer determination of the reduction potentials, differential pulse voltammetry (DPV) was applied to investigate BPTICNP-1, BPTICNP-2, and BPTICNP-3 (Figure 4c). The reduction potentials of these compounds are given in Table 2. Four clear reduction signals are observed in DPV.

To gain further insight into the electronic properties, the frontier molecular orbital (MO) energies of the HOMO and LUMO of the PDI and BPTI compounds were calculated at the B3LYP/6-31G(d) level. The energy levels and the electron distributions of the HOMO and LUMO of the PDI and BPTI compounds calculated at the B3LYP/6-31G(d) level are depicted in Figure 5. The effect of the electron-withdrawing imides is mainly a reduction of the HOMO energy in BPTI-1. The LUMO energy of BPTI-1 is slightly higher than that of PDI-1, which is consistent with the slightly negative shift of the first redox potential. The electron-withdrawing cyanophenyl group participates in the electron capture on BPTIs, as shown by the delocalization of electron density to the cyanophenyl moiety in the LUMO (Figure 5). This shifts the first reduction potential by +200 mV if we compare BPTICNP-1 to BPTI-1. The replacement of the 3-pentyl group with n-alkyl chains induces a further positive shift (~100 mV). The first reduction potentials of BPTICNP-2 and BPTICNP-3 with linear alkyl chains are more positive (~240–250 mV) than that of C₈. More importantly, BPTICNP-3 has its first reduction potential at ~0.10 V vs. NHE, which is very close to that of the primary plastoquinone electron acceptor Qₐ for which \( E(Q_a/ Q_{a}^-) = -0.08 \) V vs. NHE in the photosynthetic reaction center.\(^{[51]}\) The calculated energies (Figure 5) of the LUMO and HOMO shift by the same amount upon the introduction of the electron-withdrawing cyanophenyl group from BPTI-1 to BPTICNP-1. This leads to a constant HOMO–LUMO gap, in agreement with the optical absorption spectra.

For long-term stability, all the active materials in organic heterojunction solar cells or in photoanodes based on a D/A molecular assembly for water splitting are required to be chemically stable in water- and oxygen-containing environments. The effects of water on the degradation of the photoactive material in organic solar cells have been reported.\(^{[52–54]}\) A high degree of π-conjugation and alkyi side chains on photoactive materials can provide a hydrophobic interface that is water resistant. More importantly, the radical (di)anions of rylene diimide compounds are known to be remarkably chemically stable in a deoxygenated aqueous solution.\(^{[55–57]}\)

Unfortunately, these radical (di)anions are converted immediately to the neutral parents upon exposure to air.\(^{[56]}\) Likewise, serious electron mobility degradation was observed in PDI and fullerene derivatives for OFETs upon exposure to air.\(^{[59–61]}\) This is a result of electron trapping by absorbed O₂ and water. To make n-type organic semiconductors resistant to O₂/H₂O electron traps, the tuning of the LUMO energy to ~4.0 eV has been suggested as a design strategy.\(^{[59,60,62,63]}\) In air-stable n-type OFETs based on rylene diimide derivatives, strong electron-withdrawing groups, such as –CN, –F, –Cl, and –NO₂,
have been incorporated on bay positions or densely packed perfluoroalkyl chains were placed at the imide N,N positions to prevent H₂O and O₂ diffusion. In fullerene derivatives, only [84]PCBM, which has a low-lying LUMO level of −4.0 eV, shows air-stable electron-transport behavior. The widely used [60]PCBM, C₆₀, and even [70]PCBM as electron acceptors in organic solar cells are not suitable for the design of photoanodes for direct water oxidation in aqueous solution. For our BPTI compounds, the experimental LUMO level energies derived from the reduction potentials are shown in Table 2 and are 0.4–0.5 eV more negative than the computed values, but show the same trend. BPTI-1 has E LUMO = −3.75 eV, which is similar to the values of E LUMO = −3.77 eV for PDI-1 and E LUMO = −3.81 eV for C₆₀. Thus, it can be expected that the electron-transport performance of BPTI-1 will be sensitive to O₂ and H₂O. Functionalization of the BPTI core with the electron-withdrawing cyanophenyl group, however, results in lower LUMO energies: E LUMO = −3.95 eV for BPTICNP-1 and E LUMO = −4.06 eV for BPTICNP-2 and BPTICNP-3. Thus, the BPTICNP compounds are expected to be resistant to O₂/H₂O electron traps if they transport electrons through a BPTI film. Notably, although air-stable PDI-based OFETs have been obtained by functionalizing PDI cores with electron-withdrawing substituents and dense self-segregation of the perfluoroalkyl modification, the typically low triplet-state energy of PDI can still act as an energy sink in these PDI derivatives. Our BPTICNP compounds, however, show both a high triplet-state energy comparable to that of fullerene derivatives and E LUMO < −4.0 eV. Thus, they satisfy the requirement for resistance to O₂/H₂O electron traps formulated for OFETs and also preclude triplet charge recombination processes. Both these advantages imply that BPTICNP can be a suitable electron acceptor phase for an organic photovoltaic-based photoanode for direct water oxidation.

Furthermore, to investigate the stability of the radical anion of water-insoluble BPTI in air-saturated buffer solution, which is compatible with light-driven water oxidation conditions, a FTO | TiO₂, BPTICNP-2 electrode was prepared. The cyano group of the BPTICNP can act as an anchor by complexion to the Ti ions on the TiO₂ surface. The surface coverage Σ of the BPTICNP-2 dye on the FTO | TiO₂ | BPTICNP-2 electrode can be estimated from Beer’s law, and a value of 1.31 × 10⁻² mol cm⁻² was obtained. This value is one order smaller than the typical value (~10⁻¹ mol cm⁻²) in conventional dye-sensitized solar cells. Compared to the absorption spectra of BPTICNP-2 in chloroform, the loss of vibrational fine structure and a broader, redshifted spectrum of BPTICNP-2 on the electrode can be observed clearly (Figure 6a). This indicates that molecular aggregation of BPTICNP-2 occurs on the TiO₂ surface. In electrochemical studies of the FTO | TiO₂ | BPTICNP-2 electrode in aqueous solution, the initial reduction wave of BPTICNP-2 is more positive (~200 mV) than that of trapped states of TiO₂ with a potential of ~0.36 V vs. NHE (Figure S2). It can be concluded that electron transfer from the radical anion of BPTICNP-2 to TiO₂ will be thermodynamically unfavorable. Consequently, the BPTICNP-2 radical anion on the electrode can be prepared by the addition of the sacrificial reductant sodium dithionite (Na₂S₂O₄) in aqueous solution.

Upon immersion of the FTO | TiO₂ | BPTICNP-2 electrode in a Na₂S₂O₄ solution, the electrode presented a yellow color initially and then gradually turned greenish. The spectral features of the greenish electrolyte are similar to those of the electrochemically generated BPTICNP-2 radical anion in CH₃Cl₂, which evidenced the formation of the radical anion of BPTICNP-2 on the electrode (Figure 6b and c). Notably, the saturated and steady absorbance of the radical anion was only observed in a large excess of Na₂S₂O₄. This phenomenon probably results from the spin-pairing effect in molecular aggregation, which prevents further reduction. The aggregation stabilizes the radical anion species, which is also observed in PDI aggregation in solution and films. Upon the immersion of the greenish FTO | TiO₂ | BPTICNP-2 electrode in Na₂S₂O₄-free air-saturated phosphate buffer solution (pH 8.0, 0.1 M), the greenish electrode gradually turned yellow. The color change indicates that the BPTICNP-2 radical anion is oxidized slowly by the oxygen dissolved in water. The oxidation reaction was followed as a function of time by absorption spectroscopy, which shows a smooth conversion to the initial absorption spectrum of the neutral species (Figure S3). Importantly, the observed lifetime of the BPTICNP-2 radical anion is several tens of seconds (half-life t1/2 = 39 s). This long-lived BPTICNP-2 radical anion state present for seconds is the key requirement in the design of

![Figure 6](image-url)
long-lived charge transfer in a D/A assembly system in air-saturated buffer solution.

We found that the oxidation/reduction cycles can be repeated at least 10 times and we even repeated the reaction with a large excess of Na2S2O4 for days in air without any noticeable BPTICNP-2 decomposition or desorption from the electrode. These results indicate that BPTICNP-2 can undergo full and reversible chemical charge/discharge of excess electrons and is chemically resistant to superoxide or its byproducts (e.g., H2O2) in air-saturated phosphate buffer solution.

Theoretical electronic structure and electron-transport properties

Increasing the dielectric constant can make free charge carriers separate more easily from the bound electron–hole pairs on the D–A interface. The molecular structure concept in the design of the desired dielectric properties should consider the net polarity of the molecular structure, as argued in the Introduction. To improve the molecular dielectric properties, the first step is to increase the dipole moment in the whole molecule.[50] The introduction of strongly polar groups or side chains into the molecular structure can be an effective approach. The computed ground-state dipole moments of PDI-1, BPTI, and fullerene derivatives (C60 and [60]PCBM) are listed in Table 3. The electrostatic potential (ESP) surface maps are presented in Figure 7 for further insights into the ground-state charge distributions. In PDI-3, even though ESP shows opposite colors between the imides and the perylene core, the resulting local dipoles still cancel each other out because of symmetry, which yields a zero dipole for PDI-1. In BPTI compounds, the five-membered ring imide enhances the ground-state dipole moment μg to 1.30 Debye in BPTI-1. Furthermore, a negative ESP is clearly present on the cyanophenyl moiety on the cyanophenyl group. This indicates that this extra dipole can extend from the bayside imide on BPTI backbone. Hence, the strongly polar cyanophenyl group dramatically enhances the μg to 7.76 Debye in BPTICNP-1 and 7.77 Debye in BPTICNP-3. These values are two times higher than that of [60]PCBM.

At the microscopic level of n-type organic semiconductors, the charge-transport mechanism can be described as involving an electron transfer from a charged organic acceptor molecule A− to an adjacent neutral one A [Eq. (1)]:

\[ \text{A}^- + \text{A} \rightarrow \text{A} + \text{A}^- \]  

(1)

The basic assumption is that the charge is localized on a single molecule and charge migration takes place through an intersite hopping mechanism. In the context of semiclassical Marcus electron transfer theory,[74] the electron hopping rate k hop can be estimated by Equation (2)[75]:

\[ k_{\text{hop}}(H_{\text{sp}}, \lambda, T) = \frac{2\pi \hbar^2}{h^2} \frac{1}{\sqrt{4\pi \lambda k_B T}} \exp \left( -\frac{\lambda^2}{4k_B T} \right) \]  

(2)

in which T is the temperature, Hsp is the electronic coupling matrix element that reflects the strength of electronic interaction in orbital overlap between the adjacent molecules in the electron self-exchange process, and λ is the overall reorganization energy, which includes external (λA) and internal (λc) reorganization energies. If we assume that these hops can be described as electron transfer from a charged (and relaxed) molecule to a nearby neutral molecule, we can estimate the electron mobility μ [cm2/V·s] from the Einstein relationship [Eq. (3)]:

\[ \mu = \frac{qd^2}{k_B T k_{\text{hop}}} \]  

(3)

in which q is the elementary charge and d is the intermolecular distance between adjacent molecules.[74]

Thus, the theoretical electron mobility of organic semiconductors can be obtained from k hop and an estimate of the intermolecular distance at a given temperature. The k hop can be evaluated by quantum-chemical calculations of electronic coupling and reorganization energy. Herein, we set T to room temperature (298 K).

For the reorganization energy, λA corresponds to the polarization of the surrounding medium upon electron transfer in an organic solid-state system and this value is often assumed to be negligible.[77] Therefore, λ here is only the internal reorganization energy of the isolated active molecules that accounts for the changes in the geometry of the two acceptor molecules upon electron transfer. In the present model, the reorganization energy for electron transfer in Equation (2) can be expressed as follows [Eq. (4)]:

\[ \lambda_{\text{int}} = E^A(Q_A) - E^A(Q_B) = E^A(Q_A) + E^A(Q_B) - E^A(Q_A) - E^A(Q_B) \]  

(4)
in which $R$ is the reactant state and $P$ is the final state, both of which are with an anionic and a neutral molecule $A$. $Q_0$ indicates the optimal nuclear coordinates for state $X$, which can be $A$ or $A^-$. $\mathbb{E}(Q_0)$ indicates the energy of state $X$ calculated at $Q_0$.

The internal reorganization energies of PDI-1, BPTI, and fullerene derivatives ($C_{60}$ and [60]PCBM) were also computed at the B3LYP/6-31G(d) level, and the results are collected in Table 3. To obtain optimal charge transport in organic semiconductor systems, the reduction in the reorganization energy is important. The calculated $\lambda_{in}$ values of electron transfer in PDI-1 and PDI-3 are very close to those obtained in benchmark studies of PDI with alkyl side chains. BPTI-1 possesses a slightly lower reorganization energy ($\sim 12$ meV) than PDI-1. BPTICNP-1 and BPTICNP-3 have slightly higher reorganization energies ($\sim 17$ meV) than BPTI-1. The increase of the reorganization energy in BPTICNP-1 or BPTICNP-3 is induced by the increase of the degree of the freedom upon the introduction of the cyanophenyl moiety. Particularly, the dihedral angle between the five-membered ring imide and the cyanophenyl moiety increases from $143^\circ$ in the neutral state to $148^\circ$ in the anion. As expected, all reorganization energies of the BPTI compounds are higher ($\approx 98–124$ meV) than the values for $C_{60}$ and [60]PCBM.

Another key requirement to tune the charge hopping rate and subsequently, the charge mobility in organic semiconductors is to identify the correlations between effective electronic coupling and the molecular stacking structure. We further studied the relative configurations and electronic properties of possible $\pi$ stacked dimers of BPTICNP-3, which we constructed from insights derived from the crystal structure of PDI-3. In the PDI-3 solid state, a slipped parallel structure with a 1.12 Å shift of the short axis, a 3.05 Å shift of the long axis, and an interplanar distance of 3.28 Å was observed. In this geometry, there is a favorable electrostatic interaction between the negative surface potentials on four carbonyl groups and positive potentials on the bay area (Figure 7). Moreover, there are strong $\pi-\pi$ and van der Waals interactions from the aromatic perylene cores and linear alkyl chains on adjacent PDI molecules. We constructed the anti and syn conformations of slipped parallel stacking geometries of the BPTICNP-3 dimer and calculated the binding energy and electronic coupling with various displacement distances as shown in Figure S4 and Figure 8 (anti- or syn-BPTICNP-3 with respect to the arrangements of the cyanophenyl groups on two monomers). We fixed the $\pi-\pi$ displacement between the two BPTICNP-3 planes at 3.5 Å, a common $\pi-\pi$ stacking distance, and shifted the other two dimensions. The calculated binding energies, electronic couplings, and deduced electron mobilities of syn-BPTICNP-3 dimers at different relative positions are shown in Figure 8. For comparison, the energy and electronic coupling of a $C_{60}$ dimer was also calculated. The slipped and translated structures of syn-BPTICNP-3 and $C_{60}$ dimers are also shown in Figure 8. The binding energies of the ground-state stacking geometries are computed with DFT using the M06-2X functional with the 6-31G(d) basis set and further with abasis set superposition error (BSSE) correction. Electronic couplings were calculated by the direct coupling method with the Hartree–Fock (HF) scheme and the 6-31G(d) basis set.

Figure 8. Calculated properties of the syn dimer of BPTICNP-3 and of the $C_{60}$ dimer. Binding energies (kcal mol$^{-1}$; –■–), electronic couplings (cm$^{-1}$; –○–) and electron mobilities (cm$^2$V$^{-1}$s$^{-1}$; – ● –) are shown as a function of $Y$-axis displacement for the BPTICNP-3 dimer (a, b) and as a function of intermolecular distance for $C_{60}$ (c). To the left of (a) the structure of the dimer and the $X$, $Y$, and $Z$ coordinates used are presented. The translated BPTICNP-3 is highlighted in black. For the syn-BPTICNP-3 dimer the interplanar distance (Z-axis) was fixed at 3.5 Å. A fixed shift along the X-axis was taken as 1 Å (a) or as 2.5 Å (b). The marked gray area represents the range of local minimum of binding energy.
The electrostatic interaction between the bayside imide and benzoperylene core and the strongly polar cyanophenyl group potentially promotes molecular interactions in the syn conformations of slipped parallel geometry. Therefore, we built two syn conformations with $\Delta X$ fixed at 1 and 2.5 Å (Figure 8a and b, respectively). By varying $Y$, binding energies $E_b = -59.8$ and $-56.4$ kcal mol$^{-1}$ are obtained in the most stable dimer conformations (Figure 8a and b, respectively).

These $E_b$ values are both more negative than those in anti conformations (see Figure S4). A fully optimized syn-BPTI dimer structure could be obtained from the geometry in Figure 8b. This optimal BPTI dimer structure is similar to that indicated in Figure 8a and has $E_b = -69.9$ kcal mol$^{-1}$. This indicates that a barrierless pathway exists from the local minimal geometry shown in Figure 8b to that in Figure 8a. This favorable syn conformation of BPTICNP-3 indicates that unidirectional dipole moments can be obtained in the entire stacked BPTICNP-3 structure, which then lead potentially to a high dielectric constant. Notably, the calculated electron mobilities in the possible stacked BPTICNP-3 dimer conformations shown in Figure 8 are of the same order as the experimental ones in PDI derivatives with linear alkyl chains. The $C_{40}$ dimer with various displacement distances was also calculated, and the results are shown in Figure 8c. The local minima of the binding energy $E_b = -34.5$ kcal mol$^{-1}$ can be observed at a translated distance with $\Delta X = 3.25$ Å, which is very close to the adjacent distance in the $C_{40}$ crystal. Our calculated electron mobility $\mu = 0.12$ cm$^2$V$^{-1}$s$^{-1}$ ($H_{lo} = 128$ cm$^{-1}$) is in the range of the experimental values for $C_{40}$-based field-effect transistors.

Conclusions

We have established the efficient gram-scale synthesis of benz(g)ilperylenemetrimerides (BPTIs). Among these, BPTICNP-2 and BPTICNP-3, which have simple linear alkyl chains, are remarkably soluble in organic solvents (>15 mg mL$^{-1}$). The electrochemical and spectroscopic properties of these compounds were evaluated by cyclic voltammetry, differential pulse voltammetry, steady-state spectroscopy, and DFT calculations. The spectroscopic and electrochemical results were consistent with those of DFT calculations. High triplet-state energy levels of ~1.67 eV for BPTI compounds were found. These properties make them interesting components for bulk heterojunction solar cells or electron donor/acceptor molecular systems as compared to frequently used perylene diimides (PDIs), because triplet formation through charge recombination can be avoided. As expected, the introduction of the polar cyanophenyl group on the BPTI core unit not only improves the electron-accepting strength but also increases the ground-state dipole moment of the BPTI compounds. The addition of the cyanophenyl group on BPTI lowers the LUMO energy such that O$_2$/H$_2$O electron traps can be avoided in devices: the BPTICNP-2 radical anion on the fluorine-doped tin oxide (FTO)/TiO$_2$ BPTICNP-2 electrode showed a half-life $t_{1/2}$ of 39 s in air-saturated phosphate buffer solution. Importantly, the fully reversible chemical charge/discharge properties and chemical resistance to superoxide or its byproducts was found for the FTO/TiO$_2$/BPTICNP-2 electrode. From the theoretical investigation of the charge-transfer properties in BPTICNP-3 and $C_{40}$, we predict that BPTICNP-3 will reach similar electron mobilities to PDIs with linear alkyl chains and will also exhibit a comparable electron mobility to $C_{40}$. These findings will be explored in following studies as we believe that it will enable the design of new nonfullerene acceptors in solution-processed organic photovoltaics and assembled donor/acceptor molecular systems for artificial photosynthesis devices. Experimental confirmation of the dielectric constants and electron mobility in BPTICNP-2 and BPTICNP-3 will be established in future work.

Experimental Section

Photophysical measurements

UV/Vis absorption spectra were measured by using a HP/Agilent 8453 UV/Vis spectrophotometer using 10 mm path-length quartz cuvettes at RT. Emission spectra were measured by using a Spex Fluorolog 3 spectrometer equipped with double-grating monochromators in the excitation and emission channels. The excitation light source was a 450 W Xe lamp, and the detector was a Peltier cooled R636-10 (Hamamatsu) photomultiplier tube. The RT emission spectra were corrected for the wavelength response of the detection system. Phosphorescence measurements were performed at 77 K by using an Oxford Instruments liquid nitrogen nitrogen DN 1740 cryostat equipped with an ITC4 control unit. Samples were degassed with three freeze–pump–thaw cycles at $\approx 10^{-4}$ mbar.

Electrochemical measurements

The electrochemical measurements were performed by using a potentiotstat/galvanostat MacLab model ML160 controlled by NOVA software (1.8 version for Windows) using a conventional single-compartment three-electrode cell with a Pt working electrode, an Ag wire as the reference electrode, and a Pt wire as the counter electrode. All measurements in deaerated dichloromethane were performed with freshly distilled solvent with a solute concentration of $\approx 1.0$ mM in the presence of NaBu$_4$PF$_6$ (0.1 M) as the supporting electrolyte and a scan rate of 100 mVs$^{-1}$ in cyclic voltammetry and 5 mVs$^{-1}$ in DPV. Ferrocene (Fc) was added at the end of the experiment as internal standard to calibrate the redox potentials. The Fc/Fc$^+$ redox couple with a half-wave potential $E_{1/2}$ was set as 0.690 V vs. NHE in CH$_2$Cl$_2$.

Spectro-electrochemistry

UV/Vis spectro-electrochemistry was performed by using an optically transparent thin-layer (200 μm) electrochemical (OTTLE) cell equipped with CaF$_2$ optical windows and a Pt minigrid working electrode. The steady-state absorption spectra were recorded by using an HP1100 spectrometer in the course of the thin-layer cyclic voltammetry scanning process ($\nu = 2$ mVs$^{-1}$) controlled by a MacLab potentiostat with NOVA software. Solutions of BPTIs (1 mM) in dea-
Preparation of the FTO | TiO₂ | BPTICNP-2 electrode

The FTO | nanostructured TiO₂ electrode was purchased from solaronix (ref. 73101). The FTO | TiO₂ electrode with a 3 cm x 0.7 cm area was sensitized by soaking in a 1 M solution of BPTICNP-2 dye in THF solution overnight at RT. Following sensitization, the electrode was rinsed with dichloromethane and dried at RT. The solution was cooled, and the reaction mixture was stirred to the desired relative positions. An optimal 8 and E_sol solution of BPTICNP-2 dye in mol cm⁻³ was prepared by immersing the FTO | TiO₂ electrode in a 0.2 M Na₂CO₃ solution for 10 min in air, and the radical anion of BPTICNP-2 was obtained as 1.31 x 10⁻⁸ mol cm⁻².

Spectro-electrochemistry of FTO | TiO₂ | BPTICNP-2 electrode in aqueous solution

For the cyclic voltammetry measurements in aqueous solution, a three-electrode cell setup was used with the FTO | TiO₂ | BPTICNP-2 electrode as the working electrode, a Pt wire counter electrode, and a Ag/AgCl 3 M KCl reference electrode. Voltages were converted to NHE potentials by using E(NHE) = E(Ag/AgCl)+0.197 V. The calibration of the stability of the radical anion of BPTICNP-2 in air-saturated phosphate buffer solution was investigated by time-resolved UV/Vis absorption spectroscopy. The absorption spectra of FTO | TiO₂ in 0.1 M pH 8.0 potassium phosphate (KPi) buffer solution was set as the blank.

Computational details

All computations were performed using a development version of Q-CHEM. The water molecules studied in the present work were fully optimized at the B3LYP/6-31G(d) level. The MOs and electrostatic potentials were calculated at the same level. Graphical outputs of the computational results were plotted using the Gauss View software program (ver. 3.09). For the excited state calculation, we employed TD-DFT with the B3LYP functional and 6-31G(d) basis set.

Synthesis

The products were identified by ¹H and ¹³C NMR spectroscopy (Bruker AMX 400) and field desorption mass spectrometry (FD-MS) performed by using an AccuTOF GC v 4g, JMS-T100GCv mass spectrometer equipped with an FD Emitter with an emitter voltage of 10 kV. All PDI derivatives (PDI-1, PDI-2, PDI-3, and PDI-4) were obtained according to the literature: N⁺,N⁺-bis(3-pentyl)-perylene-3,4,9,10-tetracarboxylic diimide (PDI-1), N⁺,N⁺-bis(1-dodecyl)perylene-3,4,9,10-tetracarboxylic diimide (PDI-2), N⁺,N⁺-bis(1-octyl)perylene-3,4,9,10-tetracarboxylic diimide (PDI-3), and N⁺,N⁺-bis(1-buty l)perylene-3,4,9,10-tetracarboxylic diimide (PDI-4). All BPDI derivatives (BPDI-1, BPDI-2, BPDI-3, and BPDI-4) were prepared using the same reaction conditions: N⁺,N⁺-bis(3-pentyl) benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic-2,3,8,9-bisimide-11,12-anhydride (BPDI-1), N⁺,N⁺-bis(1-dodecyl)benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic-2,3,8,9-bisimide-11,12-anhydride (BPDI-2), N⁺,N⁺-bis(1-octyl)benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic-2,3,8,9-bisimide-11,12-anhydride (BPDI-3), and N⁺,N⁺-bis(1-butyl)benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic-2,3,8,9-bisimide-11,12-anhydride (BPDI-4). A mixture of PDI (1.5 g), maleic anhydride (40 g), p-chloranil (5.0 g), and nitrobenzene (5 mL) was stirred at 220 °C for 24 h under a N₂ atmosphere. The progress of the reaction was monitored using the complete disappearance of the typical 525 nm absorption band of PDIs in chloroform or DMF. After cooling to RT, the deep-brown product solution was poured into methanol/ethyl acetate (1:1 v/v, 400 mL) and then sonicated for 30 min. The yellowish BPDI solid was collected by vacuum filtration and washed with methanol, ethyl acetate, and acetone. The crude BPDI was used directly in the next step.

N⁺,N⁺,N⁻Tris(3-pentyl)benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic triimide (BPTI-1)

A mixture of the crude BPDI-1 (1.0 g, 1.6 mmol) and excess 3-amipropentane (1.5 mL, 13 mmol) in DMF (150 mL) was heated to 140 °C for 6 h under a N₂ atmosphere. The solution was cooled, and the reaction mixture was concentrated under reduced pressure. The yellow product was precipitated by the addition of methanol (100 mL). The precipitate was filtered and washed with methanol. The resulting residue was dissolved in chloroform and purified by column chromatography (silica gel, chloroform) to afford BPTI-1 (952 mg, 56% yield calculated from initial material PDI-1) as yellow powder. ¹H NMR (400 MHz, CDCl₃): δ = 10.44 (s, 2H), 9.32 (d, 2H, J = 8.0 Hz), 9.12 (d, 2H, J = 8.0 Hz), 5.21–5.14 (m, 2H), 4.33–4.28 (m, 1H), 2.40–2.22 (m, 6H), 2.07–1.93 (m, 6H), 1.03–0.96 ppm (m, 168.91, 133.42, 130.13, 128.02, 127.82, 127.35, 125.47 (br), 125.22, 125.05, 124.12, 123.62, 123.38 (m, 1H), 240–222 (m, 6H), 2.07–1.93 (m, 6H), 1.03–0.96 ppm (m, 168.91, 133.42, 130.13, 128.02, 127.82, 127.35, 125.47 (br), 125.22, 125.05, 124.12, 123.62, 123.38 (br), 123.31 (br), 123.25 (br) 58.35, 55.69, 52.70, 25.26, 11.62, 11.58 ppm; FD-MS (10 kV): m/z: calcd for 693.2839: C₃₀H₂₅N₉O₆; found 693.2875.

N⁺,N⁺-Bis(3-pentyl)-N⁻(4-cyanophenyl)benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic triimide (BPTCNP-1)

A mixture of crude BPDI-1 (1.0 g, 1.6 mmol) and excess 4-aminobenzonitrite (760 mg, 6.4 mmol) in DMF (150 mL) was heated to 140 °C for 18 h under N₂. The solution was cooled, and the reaction mixture was concentrated under reduced pressure. The yellow product was precipitated by the addition of methanol (100 mL). The precipitate was filtered and washed with methanol. The resulting residue was dissolved in chloroform and purified by column chromatography (silica gel, chloroform) to afford BPTCNP-1 (364 mg, 56% yield calculated from initial material PDI-1) as yellow powder. ¹H NMR (400 MHz, CDCl₃): δ = 10.44 (s, 2H), 9.32 (d, 2H, J = 8.0 Hz), 9.12 (d, 2H, J = 8.0 Hz), 25.70, 25.26, 11.62, 11.58 ppm; FD-MS (10 kV): m/z: calcd for 739.2373: C₃₀H₂₅N₉O₆; found 739.2364.

N⁺,N⁺-Bis(3-pentyl)-N⁻(4-cyanophenyl)benzog[ghi]perylene-2,3,8,9,11,12-hexacarboxylic triimide (BPTCNP-1)
The reaction was continued for 6 h. After cooling to RT, the precipitate was filtered and washed with methanol. The resultant residue was dissolved in chloroform and purified by column chromatography (silica gel, chloroform) to afford BPTICNP-1 (921 mg, 54% yield calculated from initial material PDI-2) as an orange powder. \( ^{1}H \) NMR (400 MHz, CDCl\(_{3}\)): \( \delta = 9.68 \) (s, 2H), 8.68–8.61 (m, 4H), 8.01 (s, 4H), 4.25 (t, 4H, \( J = 8.00 \) Hz), 1.87–1.77 (m, 4H), 1.4–1.20 (m, 36H), 0.87–0.84 ppm (m, 6H); \( ^{13}C \) NMR (100 MHz, CDCl\(_{3}\)): \( \delta = 163.79, 162.55, 162.04, 135.28, 133.49, 131.93, 129.91, 128.13, 128.08, 128.46, 127.97, 124.84, 123.71, 123.1, 122.74, 121.77, 118.43, 112.43, 41.43, 32.18, 31.17, 29.98, 29.95, 29.90, 29.69, 29.63, 28.31, 27.46, 22.94, 14.37 ppm; FD-MS (10 kV): \( m/z \) = calcd for 920.4512: \( C_{39}H_{34}N_{6}O_{2} \); found 920.4508.

A mixture of crude BPTICNP-3 (1.0 g, 1.4 mmol) and excess 4-aminobenzonitrile (670 mg, 5.6 mmol) in DMF (150 mL) was heated to 140 °C for 18 h under N\(_{2}\) atmosphere. The solution was cooled, and the reaction mixture was concentrated under reduced pressure. The yellow product was precipitated by the addition of methanol (150 mL). The precipitate was filtered and washed with methanol. The resultant residue was dissolved in chloroform and purified by column chromatography (silica gel, chloroform) to afford BPTICNP-2 (921 mg, 54% yield calculated from initial material PDI-2) as an orange powder. \( ^{1}H \) NMR (400 MHz, CDCl\(_{3}\)): \( \delta = 9.68 \) (s, 2H), 8.68–8.61 (m, 4H), 8.01 (s, 4H), 4.25 (t, 4H, \( J = 8.00 \) Hz), 1.87–1.77 (m, 4H), 1.4–1.20 (m, 36H), 0.87–0.84 ppm (m, 6H); \( ^{13}C \) NMR (100 MHz, CDCl\(_{3}\)): \( \delta = 163.79, 162.55, 162.04, 135.28, 133.49, 131.93, 129.91, 128.13, 128.08, 128.46, 127.97, 124.84, 123.71, 123.1, 122.74, 121.77, 118.43, 112.43, 41.43, 32.18, 31.17, 29.98, 29.95, 29.90, 29.69, 29.63, 28.31, 27.46, 22.94, 14.37 ppm; FD-MS (10 kV): \( m/z \) = calcd for 920.4512: \( C_{39}H_{34}N_{6}O_{2} \); found 920.4508.

**Keywords**: electrochemistry • electron transfer • photosynthesis • pi interactions • water splitting

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**Acknowledgements**

This work is part of the research programme of the Foundation for Fundamental Research on Matter (FOM), which is part of the Netherlands Organisation for Scientific Research (NWO). This research is financed in part by the BioSolar Cells open innovation consortium, supported by the Dutch Ministry of Economic Affairs, Agriculture and Innovation. We thank Dr. Bo-Chao Lin, Dr. Zhi-Qiang You, and Dr. Chou-Hsun Yang for helpful discussions in computational studies.