[Ni(cod)2][Al(ORF)4], a source for naked nickel(I) chemistry


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Abstract: The straightforward synthesis of the cationic, purely organometallic Ni\(^{II}\) salt \([\text{Ni(cod)}]^{+}\), \([\text{Al(OR}\text{\textsc{f}}^{\text{II}}]^{+}\) was realized through a reaction between \([\text{Ni(cod)}]_{2}\) and \(\text{Ag[Al(OR}\text{\textsc{f}}^{\text{II}}]_{2}\) (\text{cod} = 1,5-cyclooctadiene). Crystal-structure analysis and EPR, XANES, and cyclic voltammetry studies confirmed the presence of a homooleptic Ni\(^{II}\) olefin complex. Weak interactions between the metal center, the ligands, and the anion provide a good starting material for further cationic Ni\(^{II}\) complexes.

Nickel is traditionally used in many homogeneous catalytic processes,[3] for example, the Reppe carbonylation, the cyclo-tetramerization of acetylene, di- or trimerization reactions of ethylene, as well as the Shell higher olefin process (SHOP), where the catalytically active species are Ni\(^{II}\) and Ni\(^{III}\) compounds.[3] However, other oxidation states of nickel have been reported for intermediates or isolated compounds.[3] With a d\(^{1}\) electron configuration, mononuclear Ni\(^{II}\) is a rather uncommon oxidation state. Previously isolated Ni\(^{II}\) complexes were typically stabilized by electron-rich ligands, such as phosphanes,[4] amines,[5] carbenes,[6] \(\beta\)-diketiminates,[6c,7] or \(\text{Cp}^{*}\) (cyclpentadienyl),[6c,8] or were incorporated in aluminophosphates.[9] All mononuclear complexes include strongly \(\pi\)-donating C, N, P, S, O, or halogen atoms in their ligands.[4–8,10] An open question is whether Ni\(^{II}\) leads to a better performance in catalysis than Ni\(^{III}\) or Ni\(^{IV}\). Towards this goal, Stephan’s dinuclear Ni\(^{IV}\) \(\beta\)-diketiminate complexes were used by the groups of Driess and Limberg for small-molecule activation.[7k,10–12] Furthermore, Ni\(^{III}\) catalysts were used in Kumada cross-couplings[9g,10d] and olefin oligomerization or polymerization reactions.[10d–f] Nevertheless, knowledge regarding Ni\(^{II}\) olefin complexes is very scarce. The first report of a Ni\(^{II}\) olefin complex was the marginally stable \((\text{cod})\text{Ni}[\text{X}]\) (\text{cod} = 1,5-cyclooctadiene; \(\text{X} = Br, I\)), which was published in 1967 without any characterization.[10g] Later, Saraev et al. described a poorly characterized Ni\(^{II}\) olefin species as an intermediate in EPR studies starting from \([\text{Ni(cod)}]^{+}\).[12]

To study a previously unknown homooleptic olefin coordination sphere of Ni\(^{II}\), by analogy with the “naked” Ni\(^{III}\) complexes of Wilke et al.,[13] our aim was to produce a stable, cationic Ni\(^{II}\) olefin complex in combination with a weakly coordinating anion (WCA). The closest known approximation to this goal thus far is Grützmacher’s \([\text{Ni}(\text{trop-NH})(\text{OCOCF})_{2}]^{+}\) complex (trop-NH = bis(SH-dibenzo[ad]cyclo-hepten-5-yl)amine)[10d] and the intermediate \([\text{Ni(cod)}]^{+}\) species of Saraev and co-workers.[12] A straightforward access to the \([\text{Ni(cod)}]^{+}\) salt of very weakly coordinating perfluoro-tort-butoxy aluminate, \([\text{Al(OR}\text{\textsc{f}}^{\text{II}}]^{+}\) (\text{OR} = OC(CF\text{\textsc{f}}_{3}\text{$_{3}$})), is provided by the oxidation of \([\text{Ni(cod)}]_{2}\) with Ag[Al(OR\text{\textsc{f}}^{\text{II}})] in CH\(_{2}\)Cl\(_{2}\) at room temperature (Scheme 1; orange crystals in 61% yield after recrystallization; the reaction was calculated to be exergonic at the COSMO/PBE0/def2-TZVP level of theory; for details, see the Supporting Information, Figure S3).

Crystal-structure analysis as well as IR, EPR, XANES, and cyclic voltammetry (CV) measurements confirmed the existence of a homooleptic Ni\(^{II}\) cod complex. The powdered

\[
[\text{Ni(cod)}]^{+} + [\text{Ag(CH}_{2}\text{Cl}_{2}]^{+}
\]

in CH\(_{2}\)Cl\(_{2}\):

\[
\Delta^{0}_{v}(\text{CH}_{2}\text{Cl}_{2}) = -247 \text{ kJ mol}^{-1}
\]

in the gas phase:

\[
\Delta^{0}_{v}(g) = -40 \text{ kJ mol}^{-1}
\]

\[
\Delta^{0}_{f}(g) = +63 \text{ kJ mol}^{-1}
\]

\[
[\text{Ni(cod)}]^{+} + \text{Ag}^{0} + 3 \text{CH}_{2}\text{Cl}_{2}
\]

Scheme 1. Oxidation of \([\text{Ni(cod)}]_{2}\).

[4] Dipl.-Chem. M. M. Schwab, Dr. D. Himmel, Dr. D. Kratzert, Dr. V. Radtke, P. Weis, Prof. Dr. I. Krossing
Institut für Anorganische und Analytische Chemie
Freiburger Materialforrschungszentrum (FMF)
Universität Freiburg
Stefan-Meier-Strasse 21, 79104 Freiburg (Germany)
E-mail: krossing@uni-freiburg.de
Dr. S. Kacprzak, Prof. Dr. S. Weber
Institut für Physikalische Chemie
Universität Freiburg
Albertstrasse 21, 79104 Freiburg (Germany)
Dr. K. Ray
Institut für Chemie, Humboldt-Universität zu Berlin
Brook-Taylor-Strasse 2, 12489 Berlin (Germany)
Dr. E.-W. Scheidt, Prof. Dr. W. Scherer
Institut für Physik, Universität Augsburg
Universitätsstrasse 1, 86159 Augsburg (Germany)

Prof. Dr. B. de Bruin
Universiteit van Amsterdam
Vanderhoff Institute for Molecular Sciences
Science Park 904, 1098 XH Amsterdam (The Netherlands)
Prof. Dr. S. Weber
Freiburg Institute for Advanced Studies (FRIAS)
Albertstrasse 19, 79104 Freiburg (Germany)

Supporting Information, including an extensive introduction to Ni\(^{II}\) chemistry, details on the synthesis, crystallographic data, cyclic voltammetry values, EPR, NMR, UV/Vis, and IR spectra with a table of calculated and experimental vibrational frequencies, the XANES and SQUID measurements, as well as calculation methods, and ORCID(s) from the author(s) for this article are available on the WWW under http://dx.doi.org/10.1002/anie.201506475.
product is stable at room temperature and, astonishingly, did not show oxygen or air sensitivity over weeks. By contrast, in solution, the salt 1 is highly sensitive towards dioxygen. Weakly coordinating solvents such as CH2Cl2 or ortho-difluorobenzene (o-DFB) neither replace the cod rings nor coordinate to the nickel center. CV measurements of 1 in o-DFB showed an electrochemically irreversible oxidation for the redox pair NiII/Ni0 at E1/2 = +0.4962 V versus Fc/Fc+ (κ = 2.4 × 10−4 cm s−1; Table S8). The reduction of NiI at E1/2 = −0.7 V versus Fc/Fc+ includes a more complicated two-electron transfer, which will be analyzed later. In THF solution, 1 disproportionated into black solid Ni0 and yellow dissolved [NiIV(THF)][Al(OR)4]3. The molecular structure of 1 was determined by single-crystal X-ray crystallography and is shown in Figure 1a.

The overall structure of 1 is between tetrahedral and square-planar. More precisely, the torsion angle θ of the planes, which is defined by the C=C bond centroids of adjacent cod rings, is 53.1° (NiI, d6; Figure 1b). For comparison, we also prepared and crystallized tetrahedral [Ag(cod)]+[Al(OR)4]2− (2, d6+, θ = 88.5°; Figure 1b) as well as square-planar [Rh(cod)]+[Al(OR)4]− (3, d6, θ = 10.0°; Figure 1b). Neutral [Ni(cod)]0 has a torsion angle of 84.5° (d10)14. Compared to [Ni(cod)]+, the Ni-C bonds in 1 are elongated by 6–15 pm, and the C=C bonds are actually shorter by 1 pm than 3–4 pm. They are within 135 to 136 pm, similar to the undistorted C=C bonds in free cod (134 pm).15 This is probably induced from the positively charged Ni atom, which allows only minimal π-back bonding. NMR spectroscopy and quantum-chemical calculations support this hypothesis: The resonance of the olefinic proton in the 1H spectrum (Figure S6) is not paramagnetically shifted, which would be the case if there was a Fermi contact interaction with the SOMO of nickel. This finding suggests that the unpaired electron spin density is mainly centered on the metal. Calculations at the PBE0/def2-TZVPP level of theory reproduce the molecular structure well within 0.6° (torsion angle) and 4 pm (Ni–C distances; Figure S1, Table S1), and a Mulliken population analysis (Table S3, PBE0/def2-TZVPP, but also B3LYP/def2-TZVPP) localizes over 90% of the cation spin density on the nickel center (Figure 2a, inset). EPR measurements of a concentrated solution of 1 in CH2Cl2 with a non-reactive ionic liquid ([MeN(octyl)]+ [Al(OR)4]−, 0.1m) as a glass-forming additive that prohibits aggregation and an ordered orientation of the ions in frozen solution showed the typical signal of a nickel atom with d6 configuration (Figure 2a).12 The experimental spectrum was simulated with g tensor principal values of gxx = 2.390, gyy = 2.061, and gzz = 2.047 (dashed line). After magnification of parts of the experimental spectrum, an additional small signal from a second component became visible (5% signal intensity).16 Similarly to the spectrum of a frozen solution (Figure 2a), a powdered sample of 1 shows contributions from two components (Figure S5). Importantly, the g tensor components of the main species in the solid state (90% of the signal) are very similar to those of the main component in the frozen solution (a comparison of all experimental g tensors is given in Table S7)17.

X-ray absorption near-edge spectroscopy (XANES; Figure 2b) studies were performed at the Ni K-edge (on a powdered sample of 1 diluted in boron nitride) to directly probe the metal oxidation states in 1, and provide support for the EPR-derived Ni2+ assignment. Complex 1 exhibits an edge inflection energy of approximately 8341 eV, which is typical for nickel in the +1 oxidation state18,19. A shoulder along the rising edge is observed at 8334.5 eV and corresponds to a 1s−→4p shake-down transition,19 which is in accordance with the distorted structure in between tetrahedral and square-planar. This transition is strongest in four-coordinate square-planar Ni complexes, but also observed in five-coordinate square-pyramidal geometries (it is not present in either Td or Oh geometries).

It is important to note that spectra were collected at both 19 K and 298 K and found to be identical at both temperatures; this excluded the possibility of any temperature-dependent spin- or oxidation-state isomerism in 1. The temperature dependence of the inverse magnetic susceptibility χ−1(T) of 1 in a magnetic field (Figure S11) follows Curie’s law, indicating independent spins on the Ni site. From the slope of the fit, an effective magnetic moment of 1.86 ±
0.05 μμ was calculated. Taking into account the average g-factor of 2.166 ± 0.112 (EPR) and assuming a spin $\frac{1}{2}$ system with $J = S = 0.5$, led to the theoretical value of $\mu_{eff} = 1.876 \pm 0.097$ μμ. The good agreement of the experimental result with theory is in line with a Ni$^{3+}$ center in 1.

Salt 1 is a good starting material for the synthesis of other Ni$^3^+$ salts: In preliminary studies, we substituted the cod ligands of 1 by o-donors such as PPh$_3$ and 1,3-bis(diphenylphosphino)propane (dppp). The resulting products, [Ni(PPh$_3$)$_2$]$_2$-$\text{Al(OR)}_3$)$_3$] and [Ni(dppp)$_2$]$^2^+$-$\text{Al(OR)}_3$)$_3$], respectively, indicate the high potential of 1 as a precursor for further Ni$^3^+$ complexes as we thus obtained a propitious Ni$^{3+}$ salt in a direct oxidation route from commercially available chemicals.[21]

The stabilization by the [Al(OR)$_3$)$_3$]$_2$ WCA allows handling of 1 at room temperature and storage as an air-stable powder over months, as well as its use in highly oxygen-sensitive solutions in CH$_2$Cl$_2$ and 1,2-difluorobenzene. Preliminary experiments show that a simple exchange of the olefin ligands provides access to a variety of new (possibly catalytically active) Ni$^3^+$ complexes, which will be investigated in an upcoming full paper.

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tensor components of the $[\text{Ni(cod)}]$− cation are clearly sensitive to the used geometry (Table S4). Whereas the computed $g_x$ and $g_y$ values are in reasonable fair agreement with the experimental values, the computed $g_z$ components are underestimated. The DFT-calculated $g$ tensor components of the $[\text{Ni(cod)}]$− cation are available from http://www.iolitec.de.

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