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Reversible Redox Chemistry and Catalytic C(sp³)–H Amination Reactivity of a Paramagnetic Pd Complex Bearing a Redox-Active o-Aminophenol-Derived NNO Pincer Ligand

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Supporting Information

ABSTRACT: The synthesis, spectroelectrochemical characterization (ultraviolet–visible and nuclear magnetic resonance), solid state structures, and computational metric parameters of three isostructural PdCl(NNO) complexes 1 [PdCl(NNOISO)], 2 ([PdCl(NNO−ISO)])−, and 5 ([PdCl(NNOIBQ)])+ (NNO = o-aminophenol-derived redox-active ligand with a pendant pyridine) with different NNO oxidation states are described. The reduced diamagnetic complex 2 readily reacts with halogenated solvents, including lattice solvent from crystalline pure material, as supported by spectroscopic data and density functional theory calculations. Thorough removal of chlorinated impurities allows for modest catalytic turnover in the conversion of 4-phenylbutyl azide into N-protected 2-phenylpyrrolidine, which is the first example of a palladium-catalyzed radical-type transformation facilitated by a redox-active ligand as well as the first C–H amination mediated by ligand-to-substrate single-electron transfer.

INTRODUCTION

Redox-active ligands are highly relevant for many metalloenzymatic transformations.1 In synthetic inorganic chemistry, these systems can act as a (multi)electron reservoir, as a reactive ligand, or as a site for radical reactivity when combined with transition metals.2,3 Various classes of redox-active ligands have been developed and intensely investigated in the past several decades, finding applications in catalysis1 and materials science.5 A distinct class consists of the redox-active o-aminophenol-derived ligands, which have been studied in great detail by various groups.6–10 These ligands can coordinate to transition metals in a bisanionic amidophenolato (NOAP), a monoanionic iminosemiquinonato (NOISO), and a neutral iminobenzoquinone oxidation state (NOIBQ) (Scheme 1).

Analysis of the metric parameters in the NO ring obtained from single-crystal X-ray crystallography can provide key information about the oxidation state of the redox-active ligand, as significant dearomatization takes place upon one- and two-electron oxidation of an NOAP ligand. On the basis of a comprehensive analysis of structural data, Brown has developed a method that allows for the quantification of the oxidation state for o-aminophenol- and o-catechol-derived ligands.11 Using this method, a numerical “metrical oxidation state” (MOS) can be assigned to determine the oxidation state for aminophenol-derived systems with, e.g., d⁸ metals.

Single-electron transfer from a limited number of mononuclear metalloradical centers has been exploited to induce selective reduction of substrates, allowing radical-type transformations.12–16 However, single-electron transfer from a redox-active ligand to a substrate without changing the oxidation state of the accompanying metal is much rarer.17 This concept may hold promise for imparting radical-type reactivity onto redox-inert main group or “noble” metal centers and thus could potentially expand the scope for substrate coordination and reaction pathways significantly. As a first entry in this emerging field of research, we recently reported intramolecular redox-active ligand-to-substrate single-electron transfer from a low-coordinate palladium center..

Scheme 1. Three Possible Oxidation States of o-Aminophenol-Derived Ligands

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transfer to generate a reactive substrate-centered radical, using complex 1 \([	ext{PdCl(NNO}^{\text{IBQ}}\text{]}\) as a well-defined starting species.\(^{18}\) As proof-of-principle reactivity, the intramolecular sp\(^3\) C–H amination of 4-phenylbutyl azide (3) using complex 2 (generated by \textit{in situ} reduction of 1 with CoCp\(_2\)) was demonstrated (Scheme 2), but catalytic turnover to the desired pyrrolidine 4 by intramolecular ligand-to-substrate single-electron transfer on Pd\(\text{II}\) (bottom) was not achieved. To identify and circumvent potential bottlenecks for catalytic applications of the combination of this redox-active pincer ligand and Pd\(\text{II}\), we herein detail the coordination chemistry and reactivity of the Pd(NNO) complexes in all three ligand oxidation states (Scheme 1). Furthermore, we describe halogen atom abstraction reactivity as a potential deactivation pathway and a strategy for invoking modest catalytic turnover in the conversion of azide 3 to pyrrolidine 4. This results in the first catalytic example of a radical-type transformation in the coordination sphere of Pd\(\text{II}\) facilitated by a redox-active ligand.

\section*{RESULTS AND DISCUSSION}

\subsection*{Structural Comparison of Pd(NNO) Species 1, 2, and 5 with Different NNO Oxidation States.}

We previously reported the straightforward synthesis of brown paramagnetic square planar complex 1, Pd\(^{\text{II}}\)Cl(NNO\(^{\text{IBQ}}\)), which crystallized from chloroform with 1 equiv of lattice solvent per Pd complex.\(^{18}\) A combination of cyclic and differential pulse voltammetry showed a fully reversible one-electron oxidation \((E^{1/2}_{\text{ox}})\) at 0.10 V versus Fe/Fe\(^+\) and a one-electron reduction event \((E^{1/2}_{\text{red}})\) at –0.96 V versus Fe/Fe\(^+\) of 1 in a CH\(_2\)Cl\(_2\) solution.\(^{19}\) Chemical reduction using cobaltocene was shown to be facile, resulting in purple crystalline diamagnetic complex 2, [CoCp\(_2\)]\([\text{PdCl(NNO}^{\text{AP}}\text{]})].\(^{18}\)

Chemical oxidation of 1 to 5 (Scheme 3) proceeded equally smoothly with several oxidants, including Ag\(^+\) salts and thionylchloride tetrafluoroborate. An equimolar reaction of 1 with acetylferricenium tetrafluoroborate allowed for the isolation of [Pd\(^{\text{II}}\)Cl(NNO\(^{\text{IBQ}}\)]BF\(_4\) (5) as a red solid in good yield. Notably, the resonances in the \(^1\)H NMR spectrum of isolated 5 were significantly broadened in several solvents (CD\(_2\)Cl\(_2\), CD\(_3\)CN, and CDCl\(_3\)), especially for the two protons on the redox-active NO ring. In contrast, when the oxidation was performed in CD\(_3\)CN with a small excess of SelectFluor, which is a strong oxidant, only sharp signals were observed (see the Supporting Information). The protons on the redox-active NO ring appear as two well-separated singlets at 7.43 and 6.74 ppm in CD\(_3\)CN and are clearly separated from all pyridine protons. Vapor diffusion of diethyl ether into this mixture resulted in red single crystals of 5 and colorless crystals of the one-electron-reduced organic byproduct, which could be separated by hand. However, the \(^1\)H NMR spectrum of the redissolved red crystals again showed broadened resonances.\(^{20}\)

The structure for 5 was confirmed by X-ray structure determination. This allowed a direct comparison of the metric parameters of the redox series of 1, 2, and 5. The overall geometries of the anionic Pd-containing fragment of complex 2 (Figure 1, left) and the cationic Pd fragment of 5 (right) are almost isostructural to that of the distorted square planar 1 (center). However, the metric parameters for the NO ring in 2, which were reproduced by density functional theory (DFT) calculations (b3-lyp, def2-TZVP), confirm the amidophenolate (AP) oxidation state (MOS = −1.76 ± 0.17).\(^{11}\) The unit cell displays short contacts between hydrogen atoms of cationic cobaltocene ions and C and O atoms of the amidophenolate, suggestive of partial negative charges on the NO ring. Likewise, the metric parameters for the NO ring in 5 are characteristic of the iminobenzoquinone (IBQ) ligand oxidation state (MOS = −0.17 ± 0.05).\(^{11}\) The unit cell shows a disordered tetrafluoroborate anion with short contacts (<2.7 Å) to both C1 and C6 of two separate NNO\(^{\text{IBQ}}\) ligands, indicative of partial positive charges in the NO ring. Wieghardt et al. observed similar close contacts between BF\(_4\) anions and (di)cationic Pd(NNO\(^{\text{IBQ}}\)) species.\(^{21}\)

With the series in hand, some clear trends in metric parameters can be observed. Going from the NNO\(^{\text{AP}}\) to the NNO\(^{\text{IBQ}}\) ligand oxidation states, we observed a clear elongation of the C1−C2, C3−C4, and C5−C6 bonds of the NO ring, whereas bonds C2−C3, C4−C5, C1−C0, and C6−N1 are significantly shortened. These trends nicely fit with the resonance structures in Scheme 1.\(^{22}\) Another clear trend is the shortening of the Pd1−Cl and Pd1−N2 bonds going from the NNO\(^{\text{AP}}\) to the NNO\(^{\text{IBQ}}\) ligand oxidation state, reflecting the more electron-deficient nature of the metal center due to a decreased level of electron donation from the NO ring. The increased extent of pyramidalization of N1 upon reduction from the NNO\(^{\text{IBQ}}\) to the NNO\(^{\text{AP}}\) ligand oxidation state is ascribed to an increase in the sp\(^3\) character of N1 (Figure 1). The Pd1−O1 and Pd1−N1 bond lengths do not vary strongly in the redox series, which may reflect a balance between geometric factors (planar orientation of N1 in 5 vs slightly...
pyramidalized N1 in 2) and electronic factors (more covalent bonding character in 2). Very recently, Mukherjee et al. also reported isostructural PdII complexes with an aminophenol-derived redox-active tetradentate ligand in all three ligand oxidation states, showing similar trends.23

Ultraviolet–Visible (UV–vis) Spectra of Pd(NNO) Complexes 1, 2, and 5. To characterize the spectral features of the NNO ligand in its three oxidation states, we performed UV–vis titrations of the one-electron oxidation and one-electron reduction of complex 1 in MeCN to generate 5 and 2, respectively (Figure 2). For the ligand-centered one-electron oxidation of 1, aliquots containing 0.2 equiv of SelectFluor were added, resulting in clear conversion to 5 (Figure 2, left) after the addition of 1.0 equiv. Four isosbestic points are present (at \( \lambda = 371, 428, 513, \) and 636 nm), and no additional spectral changes were visible (in UV–vis or \(^{19}\text{F} \text{NMR}\) upon addition of more SelectFluor, confirming that the latter functions as a one-electron oxidant and that no secondary reactions occur. The
intraligand charge-transfer bands at $\lambda = 730, 815, \text{ and } 918 \text{ nm}$, characteristic for the NOISQ fragment, completely disappear upon oxidation, which is in agreement with generation of the NOIBQ derivative. Additionally, new bands at $\lambda = 401 \text{ and } 550 \text{ nm}$ appear that may be attributed to an NOIBQ fragment.

For the one-electron reduction of 1, aliquots containing 0.2 equiv of CoCp$_2$ were added until clear conversion to 2 (Figure 2, right) was observed after 1.2 equiv. The small excess required is ascribed to partial oxidation by traces of O$_2$ as a result of the highly sensitive nature of complex 2. Isosbestic points were found at $\lambda = 348, 389, 505, \text{ and } 698 \text{ nm}$, and no spectral changes beyond 400 nm were observed after the addition of more CoCp$_2$ (Figure 2, right). Similar to the one-electron oxidation to 5 and the recently reported PNO analogue, bleaching of the NOISQ intraligand charge-transfer bands is in agreement with ligand-centered reduction to the NNOAP ligand oxidation state. Two additional disappearing bands ($\lambda = 430 \text{ and } 469 \text{ nm}$) are attributed to metal-to-ligand charge transfer (MLCT) to the electron-deficient NOISQ ring, similar to that seen for the system reported by Mukherjee et al. The red-shifted, high-intensity band at $\lambda = 305 \text{ nm}$ is ascribed to the presence of [CoCp$_2$]$^+$.

NMR Spectroscopic Investigations of the Stability of Complex 2 in Chlorinated Solvents. Paramagnetic complex 1 is bench-stable as a solid and soluble in dichloromethane and related halogenated solvents, with no sign of decomposition after prolonged standing. Addition of a slight excess (1.2 equiv) of CoCp$_2$ to a solution of 1 in CD$_2$CN allowed for assignment and direct comparison of all $^1$H NMR resonances of the reduced species 2 with the data obtained for 5. The two protons on the amidophenolate ring appear at 6.48 and 6.24 ppm for 2. The smaller separation between the two signals compared to that in the data obtained for 5 reflects the aromatic character of the NO ring in 2. Reaction of 1 with CoCp$_2$ in CD$_2$CN results in clean formation of 2, with identical initial NMR signatures as observed in CD$_3$CN, but the spectral features dramatically changed over time in dichloromethane. The cobaltocenium protons initially appeared as a strongly broadened and upfield-shifted signal that sharpened and moved downfield over time (Figure 3). Correspondingly, the initially observed signals for 2 broadened and eventually disappeared. Re-addition of CoCp$_2$ regenerated the spectral features of 2 and shifted the [CoCp$_2$]$^+$ upfield. These observations suggest that 2 undergoes a follow-up reaction in CD$_2$Cl$_2$ to (re)generate 1. Starting from an equimolar mixture of the two compounds at 298 K results in a 1.4 x 10$^3$:1 2:1 ratio [$K_{eq}$ of 1.8 x 10$^6$ (see the Supporting Information)], based on the relative redox potentials of CoCp$_2$ and 1. Thus, reduction of 1 with CoCp$_2$ provides an almost quantitative conversion to 2. The slight excess of CoCp$_2$ is in electrochemical self-exchange with [CoCp$_2$]$^+$ cations in solution, causing an upfield shift and broadening of the [CoCp$_2$]$^+$ resonance. The [CoCp$_2$]$^+$ resonance slowly moving downfield upon reduction of 1 to 2 shows that the CoCp$_2$ concentration decreases over time. Abstraction of a chloride atom from the solvent upon chloride dissociation from 2 to regenerate 1 and [CoCp$_2$][Cl] would account for gradual disappearance of the [Pd(NNOAP)$^-$$]$ resonances and reappearance upon addition of additional CoCp$_2$ (Scheme 4). Performing the reduction in CD$_3$CN with CHCl$_3$, 1 equiv present as lattice solvent in crystalline 1 per Pd complex, and an excess of CoCp$_2$ led to the gradual disappearance of the $^1$H NMR resonance of chloroform. Addition of PPh$_3$ to 2 in CH$_2$Cl$_2$ generated the neutral complex.
Pd(NNO\textsuperscript{AP}) (PPh\textsubscript{3}) and halted the shifting of the [CoCp\textsubscript{2}]\textsuperscript{+} resonance, as reaction of 2 with solvent is prevented. Partial decomposition into an unidentified species can also be observed in the NMR spectra depicted in Figure 3. Cold-spray ESI-MS analysis of the resulting mixture showed m/z values corresponding to NNO-containing species (with and without Pd) with added CH\textsubscript{2}Cl fragments, suggestive of the formation of chloromethyl radicals and subsequent reaction with the organic ligand.\textsuperscript{27}

**DFT Calculations.** To understand the reactivity of 2 with CH\textsubscript{2}Cl\textsubscript{2} we performed DFT calculations (b3-lyp, def2-TZVP) on potential intermediates generated by homolytic C–Cl cleavage, based on the NMR spectroscopic observation that 1 is regenerated after reduction with CoCp\textsubscript{2} in the presence of CH\textsubscript{2}Cl\textsubscript{2} (or CHCl\textsubscript{3}).\textsuperscript{28} We have included the COSMO solvation model with the dielectric constant of CH\textsubscript{2}Cl\textsubscript{2} (\(\varepsilon = 8.9\)). The energy profile in Figure 4 involves the following intermediates: (A) [PdCl(NNO\textsuperscript{AP})\textsuperscript{−}] + CH\textsubscript{2}Cl\textsubscript{2}, (B) Pd(NNO\textsuperscript{AP})(CH\textsubscript{2}Cl\textsubscript{2}) + Cl\textsuperscript{−}, (C) Pd(NNO\textsuperscript{AP})(CH\textsubscript{2}Cl\textsubscript{2}) + Cl\textsuperscript{−}, (TS) Pd(NNO\textsuperscript{AP})(Cl\textsuperscript{−}···CH\textsubscript{2}Cl) + Cl\textsuperscript{−}, (D) PdCl(NNO\textsuperscript{SO})\textsuperscript{−} + *CH\textsubscript{2}Cl + Cl\textsuperscript{−}, and (E) PdCl(NNO\textsuperscript{SO}) + ClCH\textsubscript{2}CH\textsubscript{2}Cl + Cl\textsuperscript{−}. Dissociation of chloride from [PdCl(NNO\textsuperscript{AP})\textsuperscript{−}] is endergonic by 21.4 kcal mol\textsuperscript{−1}. Subsequent coordination of CH\textsubscript{2}Cl\textsubscript{2} to the “naked” neutral Pd(NNO\textsuperscript{AP}) fragment (A) is exergonic by −8.9 kcal mol\textsuperscript{−1}.\textsuperscript{29} Both a closed-shell singlet (CSS) and triplet solution were found for Pd(CH\textsubscript{2}Cl\textsubscript{2})(NNO\textsuperscript{AP}) (C), with the CSS being lower in energy by 5.7 kcal mol\textsuperscript{−1}. The barrier (\(\Delta G_{298}^\ddagger\)) on the open-shell singlet (OSS) surface for homolytic cleavage of the Cl–CH\textsubscript{2}Cl bond relative to species C was found to be only 13.5 kcal mol\textsuperscript{−1} (TS).\textsuperscript{30} The reaction to form PdCl(NNO\textsuperscript{SO})\textsuperscript{−} and a short-lived chloromethyl radical (^{3}CH\textsubscript{2}Cl)\textsuperscript{31} is exergonic (\(\Delta G_{298}^\ddagger = −18.2\) kcal mol\textsuperscript{−1}) (D). Although we have not detected dichloroethane from chloromethyl radical dimerization (vide supra) under the current experimental conditions, this reaction to generate state E makes the overall reaction exergonic by −25.0 kcal mol\textsuperscript{−1}, with the rate-determining step (highest barrier) being initial chloride dissociation. It can be suggested that other coupling reactions with this reactive ^{3}CH\textsubscript{2}Cl intermediate may result in similar energy profiles. The initial chloride dissociation was poorly described (highly endergonic) by calculations in the gas phase (see the Supporting Information), increasing the relative energies of B–E. As expected, the COSMO solvation model significantly decreased the barrier for this initial step. However, these calculations do not include the cobaltocenium ion. Ion pairing of the chloride anion will undoubtedly result in further stabilization of this fragment. As a result, it is reasonable to assume that this in turn would lower the energies for B–E even further.

As the DFT calculations indicate facile Cl–CH\textsubscript{2}Cl bond cleavage upon coordination of CH\textsubscript{2}Cl\textsubscript{2} to the reduced Pd(NNO\textsuperscript{AP}) fragment, we decided to look more closely into this homolytic bond cleavage. Notably, solely a transition state was found on the open-shell singlet (OSS) surface. No transition state could be found on the CSS (involves localized charges) or triplet surface (spin-forbidden reaction). The HOMO of Pd(CH\textsubscript{2}Cl\textsubscript{2})(NNO\textsuperscript{AP}) consists of two paired electrons located on the redox-active o-amidophenolate fragment. Interestingly, upon homolytic bond cleavage, an intramolecular single-electron transfer from the NNO\textsuperscript{SO} ligand to the coordinated Cl–CH\textsubscript{2}Cl fragment can be observed computationally. The spin density plot of the transition state for homolytic bond cleavage nicely illustrates the generation of an NNO\textsuperscript{SO} ligand radical and a chloromethyl radical (Figure 5).

![](image)

**Figure 4.** DFT-calculated free energy profile (\(\Delta G_{298}^\ddagger\)) using the COSMO solvation model for the homolytic cleavage of the Cl–CH\textsubscript{2}Cl bond (b3-lyp, def2-TZVP).

**Figure 5.** DFT (b3-lyp, def2-TZVP)-calculated HOMO plot for Pd(CH\textsubscript{2}Cl\textsubscript{2})(NNO\textsuperscript{AP}) (left) and spin density plot for the transition state of homolytic cleavage of the Cl–CH\textsubscript{2}Cl bond on the OSS surface (right).

Having established that complex 2 reacts with chlorinated reagents and given the fact that complex 1 contains CHCl\textsubscript{3} in the crystal lattice, we decided to reinvestigate the previously stoichiometric intramolecular radical-type sp\textsuperscript{3} C–H amination of azide 3 toward pyrrolidine 4 (Scheme 5). As reported for other systems, di(tert-butyl) carbonate (Boc\textsubscript{2}O) is added as an in situ N-protecting group to avoid catalyst deactivation by under the current experimental conditions, this reaction to generate state E makes the overall reaction exergonic by −25.0 kcal mol\textsuperscript{−1}, with the rate-determining step (highest barrier) being initial chloride dissociation. It can be suggested that other coupling reactions with this reactive ^{3}CH\textsubscript{2}Cl intermediate may result in similar energy profiles. The initial chloride dissociation was poorly described (highly endergonic) by calculations in the gas phase (see the Supporting Information), increasing the relative energies of B–E. As expected, the COSMO solvation model significantly decreased the barrier for this initial step. However, these calculations do not include the cobaltocenium ion. Ion pairing of the chloride anion will undoubtedly result in further stabilization of this fragment. As a result, it is reasonable to assume that this in turn would lower the energies for B–E even further.

As the DFT calculations indicate facile Cl–CH\textsubscript{2}Cl bond cleavage upon coordination of CH\textsubscript{2}Cl\textsubscript{2} to the reduced Pd(NNO\textsuperscript{AP}) fragment, we decided to look more closely into this homolytic bond cleavage. Notably, solely a transition state was found on the open-shell singlet (OSS) surface. No transition state could be found on the CSS (involves localized charges) or triplet surface (spin-forbidden reaction). The HOMO of Pd(CH\textsubscript{2}Cl\textsubscript{2})(NNO\textsuperscript{AP}) consists of two paired electrons located on the redox-active o-amidophenolate fragment. Interestingly, upon homolytic bond cleavage, an intramolecular single-electron transfer from the NNO\textsuperscript{SO} ligand to the coordinated Cl–CH\textsubscript{2}Cl fragment can be observed computationally. The spin density plot of the transition state for homolytic bond cleavage nicely illustrates the generation of an NNO\textsuperscript{SO} ligand radical and a chloromethyl radical (Figure 5).

![](image)

**Scheme 5.** Conversion of Azide 3 to Boc-Protected Pyrrolidine 4, Linear Boc-Protected Amine 6, and Nitrile 7*.

*Reagents and conditions: (i) 1 (0.1 equiv), CoCp\textsubscript{2} (0.1 equiv), Boc\textsubscript{2}O (1 equiv), CH\textsubscript{2}Cl\textsubscript{2} 100 °C.
product inhibition. Gratifyingly, thorough removal of all CHCl₃ lattice solvent in crystalline 1 by freeze-drying from benzene resulted in 2.8 turnovers for the conversion of 3 to 4 after 70 h at 100 °C in benzene in the presence of 1 molar equivalent of cobaltocene (CoCp₂) as a reductant to generate 2 in situ. Analysis of the product distribution over time showed that pyrrrolidine 4 (major product) and amine 6 (minor product) are predominantly formed in the first hours of the reaction (Figure 6). As the formation of these products slows, some (previously unidentified) nitrile 7 is produced. The use of additional equivalents of CoCp₂ did not increase the rate of turnover. Attempts to increase the rate of catalytic turnover by changing the reaction temperature, solvent, reagent stoichiometry, or the addition of additives were unsuccessful. Higher temperatures likely induce faster catalyst decomposition, while organoazide activation is inhibited at lower temperatures. Other amine protecting groups (FMocCl, CBzCl, and Ac₂O) gave no conversion. For example, addition of FMocCl generated an undefined orange species, supposedly because of fast reaction with complex 2. This suggests that, although the Pd(NNO) system is catalytically competent in the absence of chlorinated solvent, other deactivation pathways are still operative, limiting the overall turnover for this system. We are currently investigating this in more detail.

**CONCLUSIONS**

In summary, we have successfully prepared and characterized isostructural Pd complexes containing a redox-active NNO ligand in all three ligand oxidation states. The reduced complex 2 reacts with chlorinated reagents, which leads to deactivation in the intramolecular radical-type sp³ C–H amination of 4-phenylbutyl azide 3 toward pyrrrolidine 4. A possible mechanism studied by DFT calculations is chloride atom abstraction via intramolecular single-electron transfer from the redox-active ligand to a coordinated solvent molecule. The presence of only 1 molar equivalent of CHCl₃ (as lattice solvent) in the reaction mixture suffices to limit 2 to stoichiometric reactivity in C–H amination. In the absence of chlorinated reagents, (limited) catalytic turnover to 4 was achieved. This is the first catalytic example of a radical-type reaction in the coordination sphere of palladium(II) facilitated by single-electron transfer from a redox-active ligand. We are currently investigating other catalytic reactions with this platform as well as with other transition metals. We are also exploring productive (electrochemical) pathways involving C–X bond activation.

**EXPERIMENTAL SECTION**

**General Methods.** All reactions were conducted under an atmosphere of dry dinitrogen or argon using standard Schlenk techniques unless noted otherwise. Compounds 1–3 were prepared as previously reported. Acetylferroncium was prepared according to the literature procedure. All other reagents were purchased from commercial suppliers and used without further purification. THF, benzene, pentane, hexane, and diethyl ether were distilled from sodium benzophenone ketyl. CHCl₃ and methanol were distilled from CaH₂ and toluene was distilled from sodium under nitrogen. NMR spectra were recorded on a Bruker DRX 500, Bruker AMX 400, Bruker DRX 300, or Varian Mercury 300 spectrometer at 298 K unless noted otherwise. Cyclic voltammetry measurements were performed in CH₂Cl₂ (1 × 10⁻³ M) containing N(α-Bu)PF₆ (0.1 M) at room temperature under a N₂ atmosphere using a Pt electrode. All redox potentials are referenced to Fe/Ce⁺².

**Complex 5 ([PdCl(NNOIBQ)]BF₄).** A solution of 1 (24 mg, 0.05 mmol) in benzene (4 mL) was added to a suspension of acetylferroncium tetrafluoroborate (15 mg, 0.05 mmol) in benzene (1 mL). The mixture was stirred for 2 h after which the precipitate was collected by filtration. The dark red precipitate was washed with benzene (1 mL) and dried in vacuo, affording complex 5 as a red solid (21 mg, 78%). H NMR (300 MHz, CD₃CN) (obtained from in situ oxidation with Selectfluor): δ 8.70 (d, J = 5.2 Hz, 1H, Py-H), 8.21 (dd, J = 8.6, 7.9, 1.6 Hz, 1H, Py-H), 7.56 (m, 2H, Py-H), 7.43 (s, 1H, ap-H), 6.74 (s, 1H, ap-H), 2.09 (s, 6H, gem-Me), 1.29 (s, 9H, tBu), 1.27 (s, 9H, tBu). ¹⁹F NMR (282 MHz, CD₃CN): δ −151.80 (s). MS-ESI⁺ (m/z) calc for C₂₂H₃₀BClF₄N₂OPd: 872.2840, found 872.2893. C₂₂H₃₀BClF₄N₂OPd, fw = 567.14, brown plate, 0.61 mm.

**Isolation of Side Product 7.** In a glovebox, 1 (2.9 mg, 0.006 mmol, 0.1 equiv), (4-azidobutyl)benzene (10.5 mg, 0.060 mmol, 1.0 equiv), and Boc₂O (13.1 mg, 0.060 mmol, 1.0 equiv) were dissolved in benzene (3.0 mL) and placed in a 25 mL pressure tube. A solution of CoCp₂ (1.2 mg, 0.006 mmol, 0.1 equiv) in benzene (2.0 mL) was added CoCp₂ (1.2 mg, 0.006 mmol, 0.1 equiv) in benzene (2.0 mL) was added. The pressure tube was tightly sealed and heated to 100 °C. The mixture was stirred for 24 h after which the precipitate was collected by filtration and dried in vacuo. The dark red precipitate was washed with benzene (1 mL) and dried in vacuo, affording complex 5 as a red solid (21 mg, 78%). H NMR (300 MHz, CD₃CN) (obtained from in situ oxidation with Selectfluor): δ 8.70 (d, J = 5.2 Hz, 1H, Py-H), 8.21 (dd, J = 8.6, 7.9, 1.6 Hz, 1H, Py-H), 7.56 (m, 2H, Py-H), 7.43 (s, 1H, ap-H), 6.74 (s, 1H, ap-H), 2.09 (s, 6H, gem-Me), 1.29 (s, 9H, tBu), 1.27 (s, 9H, tBu). ¹⁹F NMR (282 MHz, CD₃CN): δ −151.80 (s). MS-ESI⁺ (m/z) calc for C₂₂H₃₀BClF₄N₂OPd: 479.108, found 479.865. Analytical data for this compound are in agreement with literature values.

**General Procedure for an Intramolecular sp³ C–H Amination Experiment.** In a glovebox, 2 (2.9 mg, 0.006 mmol, 0.1 equiv), (4-azidobutyl)benzene (10.5 mg, 0.060 mmol, 1.0 equiv), and Boc₂O (13.1 mg, 0.060 mmol, 1.0 equiv) were dissolved in benzene (3.0 mL) and placed in a 25 mL pressure tube. A solution of CoCp₂ (1.2 mg, 0.006 mmol, 0.1 equiv) in benzene (2.0 mL) was added. The pressure tube was tightly sealed and heated to 100 °C. After 24 h, the mixture was cooled to ambient temperature and filtered. 1,5,5-Trimethyloxazene was added as an internal standard. All volatiles were evaporated using rotary evaporation, and the residue was analyzed by ¹H NMR spectroscopy. When multiple reactions were performed to monitor the reaction at various time intervals, all reaction vessels were prepared under identical conditions in the glovebox and chemicals were taken from stock solutions.

**Isolation of Side Product 7.** In a glovebox, 1 (9.6 mg, 0.02 mmol, 1.0 equiv) was dissolved in CD₂Cl₂ (0.3 mL). To this brown solution was added CoCp₂ (11.4 mg, 0.06 mmol, 3.0 equiv) in THF, and Boc₂O (13.1 mg, 0.060 mmol, 1.0 equiv) was dissolved in benzene (3.0 mL) and placed in a 25 mL pressure tube. A solution of CoCp₂ (1.2 mg, 0.006 mmol, 0.1 equiv) in benzene (2.0 mL) was added. The pressure tube was tightly sealed and heated to 100 °C. After 24 h, the mixture was cooled to ambient temperature and filtered. 1,5,5-Trimethyloxazene was added as an internal standard. All volatiles were evaporated using rotary evaporation, and the residue was analyzed by ¹H NMR spectroscopy. When multiple reactions were performed to monitor the reaction at various time intervals, all reaction vessels were prepared under identical conditions in the glovebox and chemicals were taken from stock solutions.

**Isolation of Side Product 7.** In a glovebox, 1 (9.6 mg, 0.02 mmol, 1.0 equiv) was dissolved in CD₂Cl₂ (0.3 mL). To this brown solution was added CoCp₂ (11.4 mg, 0.06 mmol, 3.0 equiv) in CD₂Cl₂ (0.3 mL). The bright purple solution was transferred to a J Young NMR tube. ¹H NMR spectra were recorded at regular time intervals. After a few hours, a black precipitate was observed and the solution turned more brownish. To add new CoCp₂, the NMR tube was returned to the glovebox and a small amount CoCp₂ (approximately 3 mg) was added.
reflections were measured on a Bruker D8 Quest Eco diffractometer equipped with a Triumph monochromator ($\lambda = 0.7073$ Å) and a CMOS Photon S0 detector at a temperature of 150(2) K up to a resolution of ($\sin \theta/\lambda$)$_{max} = 0.60$ Å$^{-1}$. Intensity data were integrated with the Bruker APEX2 software. Multiscan absorption correction and scaling were performed with SADABS. All reflections were unique ($R_{int} = 0.0874$), of which 3779 were observed ($I > 2\sigma(I)$). The structures were determined using intrinsic phasing with SHELXT. Least-squares refinement was performed with SHELXL-2013 against $F^2$ of all reflections. Non-hydrogen atoms were refined freely with anisotropic displacement parameters. The H atoms were placed at calculated positions using the instructions of AFIX 13, AFIX 43, or AFIX 137 with isotropic displacement parameters having values 1.2 or 1.5 times the $U_{eq}$ of the attached C atoms; 470 parameters were refined with 866 restraints. R1/wR2 ($I > 2\sigma(I)$): 0.0339/0.0679. R1 was made for the condensed phase (1 L mol$^{-1}$) and scaling were performed with SADABS; 4607 reflections were integrated ($I > 0.0874$), of which 3779 were observed ($I > 2\sigma(I)$). Estimated condensed phase (1 L mol$^{-1}$) calculations on the b3-lyp, def2-TZVP optimized geometries.

Computational Details. Geometry optimizations were conducted using Turbomole coupled with the PQS Baker Optimizer via the BOPt package at the DFT level using the BP86 or b3-lyp functional and the def2-TZVP basis set. The corrected broken-symmetry energy $e_{BS}$ of the open-shell singlets ($S=1$) was obtained from the total energy $e_{S=1}$ of the optimized single-determinant broken symmetry solution and energy $e_{S=0}$ from a separate unrestricted triplet calculation at the same level, using the approximate correction formula:

$$e_{BS} \approx \frac{S_{S=1}^2 e_{S=1} - S_{S=0}^2 e_{S=0}}{S_{S=1}^2 - S_{S=0}^2}$$

The tert-butyl groups on the NNO ligand were replaced with Me groups to decrease the computational time. Energies including the COSMO solvation model were obtained by performing single-point calculations on the b3-lyp, def2-TZVP optimized geometries. Estimated condensed phase (1 L mol$^{-1}$) free energies and enthalpies were obtained from these data by neglecting the enthalpy RT term and subsequent correction for the condensed phase reference volume [$S_{CV} = S_{S=1} + R \times \ln(1/24.5)$]. For the reactions involving reaction with the solvent, a correction of $-8$ kcal mol$^{-1}$ was made for the condensed phase energies as the there is no entropic penalty for bringing the molecules together.

ASSOCIATED CONTENT

2 Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.inorg-chem.6b01192.

Details for the redox reaction between 1 and CoC$_5$NMR spectra, and DFT calculations (PDF)

Crystallographic data (CIF)

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Notes

The authors declare no competing financial interest.

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(20) The reason behind the observed spectral broadening remains unclear, but it may be related to fluxional coordination behavior of the NNOISQ ligand. Alternatively, workup and spectroscopy under ambient conditions in the absence of excess oxidant may result in partial regeneration of I, leading to broadened spectral features (due to a self-exchange process or the presence of a paramagnetic species).
For the corresponding gas phase calculations, see the Supporting Information.

Using the pure GGA BP86 functional provided similar results. The applied broken symmetry approach is an approximation of the multideterminant problem of bond-splitting reactions, hence introducing some (unavoidable) additional error into the computed barrier.


Most likely resulting from catalyst decomposition. This compound is also the main product at higher reaction temperatures. Nitrile 6 is thus an unlikely source of H atoms to generate linear amine S.

Reaction with benzyl bromide appears to be very facile. For a related study, see: Broere, D. L. J.; Modder, D. K.; Blokker, E.; Siegler, M. A.; van der Vlugt, J. I. Angew. Chem., Int. Ed. 2016, 55, 2406−2410.


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