Developing hierarchically porous MnO$_x$/NC hybrid nanorods for oxygen reduction and evolution catalysis

Pandey, J.; Hua, B.; Ng, W.; Yang, Y.; van der Veen, K.; Chen, J.; Geels, N.J.; Luo, J.-L.; Rothenberg, G.; Yan, N.

DOI
10.1039/C7GC00147A

Publication date
2017

Document Version
Final published version

Published in
Green Chemistry

License
Article 25fa Dutch Copyright Act (https://www.openaccess.nl/en/in-the-netherlands/you-share-we-take-care)

Citation for published version (APA):
Developing hierarchically porous MnOx/NC hybrid nanorods for oxygen reduction and evolution catalysis

Starting from abundant precursors and using a simple synthesis approach, we report a good bifunctional oxygen electro-catalyst. The composite nanorods, comprising manganese oxides and nitrogen-doped carbon, have hierarchical porosity. The nitrogen-doped carbon forms a contiguous 3D network, connecting all the isolated MnOx nanoparticles and ensuring superior electrical conductivity. This hybrid is among the best non-noble-metal ORR/OER catalysts in alkaline media, outperforming even Pt and RuO2 catalysts.

As featured in:
Green Chemistry
Phenolic acetals from lignins of varying compositions via iron(III) triflate catalysed depolymerisation

See Ning Yan et al., Green Chem., 2017, 19, 2793.
Developing hierarchically porous MnO$_x$/NC hybrid nanorods for oxygen reduction and evolution catalysis†

Jay Pandey,$^a$ Bin Hua,$^b$ Wesley Ng,$^a$ Ying Yang,$^c$ Koen van der Veen,$^a$ Jian Chen,$^d$ Norbert J. Geels,$^a$ Jing-Li Luo,$^b$ Gadi Rothenberg,$^d$ and Ning Yan$^d$$^a$$^a$

Electrochemical oxygen reduction and evolution reactions (ORR and OER) play a vital role in the field of energy conversion and storage. The problem is that both processes are sluggish, requiring precious-metal catalysts. Here, starting from abundant precursors and using a simple synthesis approach, we report the preparation of a good bifunctional oxygen electro-catalyst: a composite nanorod of manganese oxides and nitrogen-doped carbon. This material has hierarchical porosity, facilitating the mass transfer within the electrode. The nitrogen-doped carbon forms contiguous 3D network, connecting the isolated MnO$_x$ nanoparticles and ensuring superior electrical conductivity. Importantly, the MnO$_x$ particles contain manganese of mixed oxidation states; aligned with the nitrogen-doped carbon, this hybrid is among the best non-noble-metal ORR/OER catalysts in alkaline media, outperforming even Pt and RuO$_2$ catalysts.

Introduction

The ever-increasing quest for sustainability calls for green energy solutions.$^1$–$^5$ In this context, the oxygen reduction reaction (ORR) and the oxygen evolution reaction (OER) have become a focus of attention. They play pivotal roles in various emerging energy conversion and storage devices including reversible fuel cells and metal-air batteries.$^6$–$^9$ However, efficient ORR/OER is often catalyzed by noble-metal catalysts,$^9$–$^{12}$ posing a challenge for real-life applications. Besides, good ORR catalysts often lack OER activity and vice versa.$^{13}$–$^{15}$ This is problematic when switching the charging/discharging mode of the batteries or transforming fuel cells into electrolysis cells. Thus, inexpensive bifunctional ORR/OER catalysts feature high on industrial wish lists.

There are a number of promising alternatives indeed. For instance, transition metal oxides, including simple and perovskite oxides,$^{16}$–$^{18}$ demonstrate superior intrinsic bifunctional activity; several recent review papers have well summarized the research progress made to date.$^{5}$–$^{16}$ Among all the candidates, manganese oxide catalysts have attracted much attention.$^{13}$–$^{16}$, $^{18}$ Manganese is an earth-abundant metal, which forms a variety of oxide compounds, some of which have demonstrated great activities in ORR and/or OER.$^{18}$–$^{23}$ But their poor electrical conductivity may limit the electrochemical performance.$^{13}$–$^{21}$, $^{24}$–$^{25}$ Another typical example is the nitrogen-doped carbons that are good ORR catalysts with excellent activity towards both ORR and OER in alkaline media.$^{27}$–$^{29}$ But their susceptibility to passivation at high anodic potential may challenge their use in OER.$^{30}$–$^{31}$

In addition, optimizing the microstructure and morphology of the electrocatalysts is another promising approach for boosting their activities. It generates more active sites, forms composite structures, tunes the electronic structure, and creates hierarchical porosity (e.g., connected meso-/macropores).$^{32}$–$^{34}$ Nonetheless, making such structures is not an easy task. The synthetic routes are usually complicated and costly. In many cases, expensive and/or toxic reagents are also required.$^{13}$–$^{15}$, $^{28}$–$^{35}$, $^{37}$–$^{38}$ For these reasons, concerns arise regarding the scalable synthesis.

In this work, we report a green synthesis of a bifunctional oxygen catalyst comprising MnO$_x$ and N-doped carbon. Starting from two simple precursors (nitrilotriacetic acid and manganese acetate), this synthesis produces hybrid nanorods using a facile approach. The unique nano-structure grants excellent activity towards both ORR and OER in alkaline media, opening opportunities for practical application.
Results and discussion

Fig. 1 illustrates the synthesis procedure of the MnO/N-doped carbon hybrid nanorod. After the hydrothermal treatment of two precursors, nitritotriacetic acid (NTA) and manganese acetate, at 180 °C, a rod-like salt (MnNTA) was crystallized and precipitated. Pyrolysis of washed MnNTA nanorods at 900 °C in Ar yielded MnO/N-doped carbon hybrid rods (MnO/NC). A final mild-oxidation at 200 °C in air converted the surface of MnO to MnO2 (hereafter MnO/NC), yet without oxidizing the NC matrix. Fig. S1† compares the results of the thermogravimetric analysis and differential scanning calorimetry (TGA-DSC) for MnO/NC and MnO2/NC in air. The exothermic weight increase of MnO/NC starting from ca. 200 °C corresponded to the surface oxidation of MnO. The other exothermic peak at 300 °C was ascribed to the oxidation of carbon. In both materials, NC accounted for ~20 wt% of the hybrid, with 80 wt% MnOx (see Fig. S1†).

The powder X-ray diffraction (XRD) affirmed the surface oxidation in MnO/NC (see Fig. 2a). MnO is the main component as a result of the carbon reducing effect at 900 °C. The manganese oxide recrystallized after the oxidation, showing broadened peaks in MnO2/NC. New peaks also appeared, pertaining to the formation of Mn3O4 and MnO2 (the XRD data of MnNTA are shown in Fig. S2†). We also ran temperature programmed reduction measurements in hydrogen (H2-TPR, see Fig. S3†). MnO/NC showed little hydrogen consumption because MnO was irreducible in 5% H2 up to 800 °C.38 Conversely, MnO2/NC showed several distinct reduction peaks, suggesting that the surface of MnO2 was indeed oxidized during the 200 °C-oxidation step (vide infra).

N2-Adsorption measurements on the MnO2/NC hybrid revealed a type-IV isotherm with an adsorption–desorption hysteresis loop (see Fig. 2b). The specific surface area was 135 m2 g−1. The pore volume reached 0.48 cm3 g−1, comparable with those porous materials derived via either “soft-templating” or “hard-templating” methods.39–42 Calculating the micropore size distribution using the Horváth–Kawazoe (HK) model, we found that the median pore diameter was 0.84 nm (see Fig. 2c). Calculations using the Barrett–Joyner–Halenda (BJH) method revealed a mesopore size distribution ranging from 3.5 to 22 nm (see Fig. 2d). Alternatively, the pore size distribution obtained from the NLDFT method is shown in Fig. S4,† which revealed similar characteristics of the pore.

Note that the mesopore volume was up to 0.32 cm3 g−1, accounting for 67% of the total pore volume. Thus, this hierarchically porous structure meets the two requirements of a good electrocatalyst: efficient mass transfer and a large surface area with abundant active sites.28,35–37

We then used X-ray photoelectron spectroscopy (XPS) to examine the electronic structure of MnO2/NC (see the survey spectrum in Fig. S5†). The carbon matrix in the hybrid had a high content of nitrogen, demonstrating a N/C ratio close to 6.4 at%. The spectrum of the N 1s core-level displayed three distinct nitrogen moieties. Of these, 28 at% were pyridinic (Np) with a binding energy of 398.8 eV, which are believed to be the main active ORR sites.43 The peaks at 400.7 eV and 401.6 eV were assigned to graphitic (Ng) and oxidized (No) nitrogen, respectively. The Mn 2p3/2 core-level spectrum in Fig. 2f shows a broad shake-up satellite peak at 644.6 eV and reveals the mixed valence of Mn on the surface. The peaks at 640.8 eV, 641.8 eV and 643 eV were assignable to Mn2+ (37.0%), Mn3+ (38.3%) and Mn4+ (24.7%), respectively.24,44,45 Particularly, the latter two species were recognized to be critical in oxygen catalysis.46 This also in turn proved the successful surface oxidation of MnOx.

We also examined the morphologies of MnOx/NC. Fig. 3a and b contrast the scanning electron microscopy (SEM) images of the MnNTA salt and MnOx/NC. The rod-like shape remained stable after the pyrolysis and oxidation steps. The diameter of the rod was ca. 200 nm whereas the length was up to 3 μm. Interestingly, this structure was not observed when either Fe or
Mg was used instead of Mn (cf. Fig. S6†). Fig. 3c shows the transmission electron microscopy (TEM) image. The nanorod structure was covered with a “furry” surface which was then confirmed as the uniformly distributed manganese oxide nanoparticles, typically with a size of 20 nm (see Fig. S7†). The corresponding elemental mappings in Fig. 3d from the energy-dispersive X-ray spectroscopy (EDS) confirm the homogeneous dispersion of both MnO2 and NC species in the hybrid (note that the carbon grid was also shown). This is important for electrocatalysis: the contiguous NC matrix connects the MnO2 nanoparticles (which are poor conductors themselves), increasing the electrical conductivity of the catalyst.

The high resolution TEM micrograph (HRTEM) in Fig. 3e shows the advantageous microstructure of the hybrid catalyst. The black arrow indicates that the NC serves as an electronic connector, bridging two MnO2 nanoclusters. Worm-like structures were seen, typically representing micropores in carbon materials, which enhanced the surface area; the mesopores originated from the inter-particle cavities, which created channels for the mass transport to and from the active sites. Importantly, the micropores and mesopores are all connected, offering a classic hierarchically porous structure. The diffractogram inset from the selected area electron diffraction (SAED) implied the single crystal of MnO. Nevertheless, amorphous MnO2 species were observed on the surface (cf. the XRD, H2-TPR and XPS data). This amorphous structure might also be beneficial in the oxygen electrocatalysis.33

The electrocatalytic activities of MnO2/NC and the controls were all evaluated in 0.1 M KOH electrolyte using a standard rotating disc electrode (RDE; detailed experimental procedures are included in the ESI†). In the ORR experiments, the cyclic voltammetry of MnO2/NC in the N2-saturated electrolyte showed no redox peaks (Fig. S8†). In contrast, a remarkable reduction peak was seen when switching to the O2-saturated electrolyte, confirming that the oxygen reduction has occurred on the electrode. Subsequently, we performed linear sweep voltammetry (LSV) while varying the rotating speed of the RDE, a common approach for investigating the electrochemical reaction mechanism. In all cases from 400 rpm to 2400 rpm, high limiting-current densities were achieved and the onset potentials remained identical (Fig. 4a). The Koutecký–Levich plots in Fig. 4b were perfectly linear. Their slopes were all constant, from which the calculation indicated the preferable 4e−-transfer pathway of ORR catalyzed by the MnO2/NC hybrid. In addition, MnO2/NC demonstrated excellent robustness and resistance against methanol poisoning, superior to the commercial Pt/C catalyst. In the chronoamperometric measurements, no significant current loss was recorded after the methanol injection and 80% current was retained after the 10 h test (see Fig. S9†).

Fig. 4c compares the LSV of various catalysts. In the controls, the NC catalyst contained an essentially identical ratio of doped nitrogen with the counterpart in MnO2/NC, whereas the MnO2/C was formed by mixing MnO2 nanorods with carbon black (see the ESI† for details). The onset (Eorr-onset) and half-wave (Eorr-1/2) potentials are listed in Table 1, which are two important parameters for evaluating the activities of ORR catalysts. In general, both MnO2/NC and MnO2/C showed high Eorr-onset at 0.95 V (all values are given vs. RHE), ranking them among the best Mn-based ORR catalysts to date (see Tables S1 and S2†). This also implied the beneficial features of MnO2 with the unique rod-structure and the mixed Mn valence. However, MnO2/NC had a much higher Eorr-1/2 (0.80 V vs. 

\[ \text{ORR-1/2 (0.80 V vs. RHE)} \]

Fig. 4d shows the performance of various catalysts at different rotating speeds in the OER. The MnO2/NC catalyst demonstrated the highest onset potential (Eorr-onset) at 1.7 V, which is an exceptionally high value compared to other catalysts. This indicates the strong resistance against oxygen evolution reaction (OER) in the MnO2/NC catalyst. The LSVs in Fig. 4d also show the excellent stability of MnO2/NC in the OER, with no apparent degradation in the current density over the 1600 rpm test.

The results obtained from the electrochemical measurements confirmed the promising performance of MnO2/NC as a highly active and stable ORR catalyst, which could be beneficial in practical applications such as fuel cells and water splitting.
0.64 V), suggesting the synergistic effects between the constituents. Cooperative catalysis might be indeed contributing as the NC was also a good ORR catalyst ([E_{\text{orr-1/2}}] = 0.75 V). More remarkably, the 20 wt% NC was contiguous, connecting every individual MnO_{2} nanoparticle that was much less conductive. Hence, with a larger quantity of MnO_{2} (≥80 wt%) than those derived from conventional infiltration methods, this hybrid had more electrochemically “valid” active sites (see the discussions and the structural comparison in Fig. S10†). Moreover, the resulting hierarchical porosity aligned with the superior electrical conductivity also granted good activity (see above).

We then investigated the OER activity of MnO_{2}/NC. As the state-of-the-art, RuO_{2}/C possesses the lowest OER onset potential ([E_{\text{orr-onset}}]) and the “benchmark” potential at 10 mA cm^{-2} ([E_{\text{orr-10}}]). Pt/C performed badly in accordance with the literature, this state-of-the-art, RuO_{2}/C possesses the lowest OER onset potential. However, RuO_{2}/C worked nicely thanks to its unique microstructure. Its bifunctional oxygen catalysis. A small amount of a percolated NC network led to the formation of a porous structure, remarkably enhancing the conductivity and mass transfer in electrocatalysis. The mixed oxidation states of Mn cations along with the high ratio of nitrogen dopants contributed to the superior activity in both ORR and OER. The fact that this highly active bifunctional catalyst contains no noble metals aligned with a facile/green synthesis approach opens bona fide opportunities for the applications in new energy devices.

**Conclusions**

In summary, using two simple precursors and a scalable method, we developed a green and facile synthesis of hierarchically porous MnO_{2}/NC hybrid nanorods for efficient bifunctional oxygen catalysis. A small amount of a percolated NC network led to the formation of a porous structure, remarkably enhancing the conductivity and mass transfer in electrocatalysis. The mixed oxidation states of Mn cations along with the high ratio of nitrogen dopants contributed to the superior activity in both ORR and OER. The fact that this highly active bifunctional catalyst contains no noble metals aligned with a facile/green synthesis approach opens bona fide opportunities for the applications in new energy devices.

**Acknowledgements**

This work is part of the Research Priority Area Sustainable Chemistry of the UvA, http://suschem.uva.nl.

**Notes and references**
