Distance dependence of singlet and triplet charge recombination pathways in a series of rigid bichromophoric systems

Roest, M.R.; Oliver, A.M.; Paddon-Row, M.N.; Verhoeven, J.W.

Publication date
1997

Published in
The Journal of Physical Chemistry. A

Citation for published version (APA):

General rights
It is not permitted to download or to forward/distribute the text or part of it without the consent of the author(s) and/or copyright holder(s), other than for strictly personal, individual use, unless the work is under an open content license (like Creative Commons).

Disclaimer/Complaints regulations
If you believe that digital publication of certain material infringes any of your rights or (privacy) interests, please let the Library know, stating your reasons. In case of a legitimate complaint, the Library will make the material inaccessible and/or remove it from the website. Please Ask the Library: https://uba.uva.nl/en/contact, or a letter to: Library of the University of Amsterdam, Secretariat, Singel 425, 1012 WP Amsterdam, The Netherlands. You will be contacted as soon as possible.
Distance Dependence of Singlet and Triplet Charge Recombination Pathways in a Series of Rigid Bichromophoric Systems

Martin R. Roest,† Anna M. Oliver,‡ Michael N. Paddon-Row,*† and Jan W. Verhoeven*†

Laboratory of Organic Chemistry, University of Amsterdam, Nieuwe Achtergracht 129, 1018 WS Amsterdam, The Netherlands, and School of Chemistry, University of New South Wales, Sydney, NSW 2052, Australia

Received: March 18, 1997

The recombination process of the charge-separated state in a series of rigid donor–bridge–acceptor compounds 1[n] follows two pathways, viz. recombination to the singlet ground state and recombination to the local triplet donor state. The rates of both these processes are shown to depend exponentially on distance, but with a significantly different damping factor. Thus at short distances singlet recombination is dominant, while at longer distances the triplet pathway eventually becomes the major process.

Introduction

The distance dependence of electron transfer reactions has been studied extensively during the last decade using systems of the type donor–bridge–acceptor, in which a nonconjugated bridge connects the electron donor and acceptor. A major concern in these studies is that the bridge should maintain a strictly defined relative orientation and distance between D and A and that the latter can be varied in well-defined increments. The series of bichromophoric compounds 1[n] (Scheme 1) satisfies these demands to a high extent since not only the distance is fully defined, but at the same time the relative orientation of the dimethoxynaphthalene (DMN) donor, D, and the dicyanonovinyl acceptor, A, is fixed in each member of the series and rather constant across the series. It has been shown that photoinduced charge separation (kcs) occurs, following local excitation of the DMN donor, for all members of the series (see Scheme 2).

This has allowed the study of the photoinduced charge separation (kcs) as well as the consecutive charge recombination (kcr) as a function of the bridge length. The latter may be expressed either in terms of the number of σ bonds, n, separating D and A or as the center–to–center distance between D and A (Rc) in angstroms. In studies on charge recombination within rigidly linked systems published until now, this process was described with a single, distance dependent rate. In the systems 1[n], however, recombination can occur either to the ground state (k1cr) or to a local DMN triplet state (k3cr), a situation that is typical for many other donor–acceptor systems as well. We now present a study in which the distance dependence of k1cr and k3cr are separately determined for 1[n].

\[ k_{et} = \frac{4\pi^2}{h} V^2 FCWD \]  (1)

\[ V = V_0^* \exp(-\alpha R) \]  (2)

\[ k_{cs} = k_{0,et}^* \exp(-\beta R) \]  (3)

For weak coupling, which appears a reasonable assumption in most members of series 1[n], the rate of electron transfer can be described by eq 1, in which V is the electronic coupling between the reactant and product states and FCWD is the Franck Condon weighted density of states. A major problem in interpreting the distance dependence of electron transfer rates is that both V and FCWD are in principle distance dependent, and it is not a simple task to separate these factors quantitatively. It has been argued, however, that the distance dependence of the electronic coupling plays a dominant role because this is expected to decrease exponentially with distance (eq 2). When FCWD is assumed to be virtually independent of distance, the...
SCHEME 2: Charge Separation and Recombination Processes in 1[n]a

![Diagram]

a Note that the energy of the charge separated state varies as a function of n and solvent from slightly above 1D*A (i.e. for n = 12 in saturated hydrocarbon solvents) to about 1 eV below (i.e. in highly polar solvents). Under all circumstances the local donor triplet (1D*A) is significantly below the CT state.

rate of electron transfer should thus also decrease exponentially with distance with a damping factor $\beta = 2\alpha$ (eq 3). An exponential distance dependence was indeed found for both $k_{cs}$ and $k_{cr}$ in 1[n]a-g,13-17

As already noted above, however, for $k_{cr}$ the situation is complicated. In fact charge recombination can occur via three distinctly different pathways. Thus, as indicated in Scheme 2, under certain conditions the CT state is high enough in energy to allow thermally activated repopulation of the locally excited singlet state. This “back electron transfer” ($k_{cr}$) occurs only for $n \geq 8$ in low-polarity solvents, as revealed by the delayed donor fluorescence emerging from it.15,16 While conditions can be chosen to avoid a contribution of $k_{cr}$ to $k_{cs}$, still two discrete recombination pathways are always accessible, leading either to the ground state ($k_{cs}$) or to the local triplet state of the DMN donor ($k_{cr}$), which is always significantly lower in energy than the CT state (see Scheme 2). This implies that the $k_{cr}$ values determined previously are in fact the sum of $k_{cr}^{1}$ and $k_{cr}^{2}$ and that the $\beta_{cr}$ values represent a weighted average of that for singlet and triplet recombination ($\beta_{cr}^{1}$ and $\beta_{cr}^{2}$).

The main purpose of the present study is to separate the singlet and triplet pathways, while no effort has yet been made to separate the influence of electronic coupling and FCWD for either of these pathways.

To separate the rates of the triplet and singlet recombination processes in 1[n], we now determined quantum yields for the formation of 1D*A in the compounds 1[n] in benzene and dioxane with transient absorption (TA) spectroscopy. As the charge separation in these compounds is complete in these solvents, except for 1[12], and $k_{cr}$ for 1[12] in benzene, the yield of all 1D*A must stem from recombination of the CT state. Benzene and 1,4-dioxane were chosen as solvents because the overall recombination rates in these solvents are known from previous studies.16

Results and Discussion

The overall rate of charge recombination in systems 1[n] has been studied quite extensively.1b-g,15-17 The main technique employed has been time-resolved microwave conductivity (TRMC), which allows direct detection of the decay of the dipolar CT state. In addition, the decay of the weak CT fluorescence emerging from this state has been measured in these cases where it was sufficiently strong. Consistent results for the decay rates were obtained with both techniques as tabulated in Table 1. From these data it is immediately clear that, in contrast to $k_{cr}$,1a,13 $k_{cr}$ displays a very strong solvent dependence and increases dramatically upon increasing solvent polarity.

Qualitatively this can be understood from the decrease of the energy gap separating the CT state from the ground state. Since recombination to the singlet ground state ($k_{cs}$) occurs under “inverted region” conditions where the driving force is much larger than the reorganization energy (-$\Delta G_{cr}^{1}$), this implies that upon decreasing the driving force, the FCWD for this process should increase (see eq 4). It should, however, be noted that under such conditions eq 4 can only give a qualitatively correct picture of this effect. In fact the FCWD is much larger than predicted by eq 4 because nuclear tunneling becomes dominant, as e.g. shown by the small temperature dependence of $k_{cs}$ in systems 1[n].15,17

The data compiled in Table 1 have been plotted in Figure 1. As noted earlier $k_{cs}$ appears to decrease exponentially with distance, and even though the rates are very strongly solvent dependent, the damping factor $\beta_{cs}$ appears rather solvent independent with an average value of 0.95 ± 0.10 per $\sigma$ bond ($\beta_{cs} = 0.83$ Å$^{-1}$).

\[
FCWD = \frac{1}{\sqrt{4\pi\kappa B T}} \exp\left[\frac{-(\Delta G + \lambda)^2}{4\kappa B T}\right]
\] (4)

In order to determine the contributions of $k_{cr}^{1}$ and $k_{cr}^{2}$ to $k_{cr}$, yields of formation of the 1D*A species were now determined from the TA spectrum of the compounds 1[n] in dioxane and benzene. To obtain the exact amounts of triplet, the molar extinction of the $T_1 \rightarrow T_n$ absorption had to be determined. Attempts to obtain this value by complete triplet energy transfer from benzophenone (BP) to donor model 2 did not succeed due

### Table 1: Rates of Charge Recombination $k_{cr}$ ($\times 10^8$ s$^{-1}$) Obtained Earlier for 1[n] in Various Solvents

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$k_{cr}$ (×10$^8$ s$^{-1}$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>trans-decalin$^{a}$</td>
<td>22.2 167 100 2000</td>
</tr>
<tr>
<td>benzene$^{a}$</td>
<td>125 1000 500 83.3</td>
</tr>
<tr>
<td>di-n-butyl ether$^{a}$</td>
<td>83.3 2.78 2.94 23.3</td>
</tr>
<tr>
<td>1,4-dioxane$^{a}$</td>
<td>90.9 (1.35) $e$ 3.75</td>
</tr>
<tr>
<td>1,4-dioxane$^{b}$</td>
<td>90.9 (1.35) $e$ 3.75</td>
</tr>
<tr>
<td>1,4-dioxane$^{c}$</td>
<td>90.9 (1.35) $e$ 3.75</td>
</tr>
<tr>
<td>1,4-dioxane$^{d}$</td>
<td>90.9 (1.35) $e$ 3.75</td>
</tr>
</tbody>
</table>

$^{a}$ Data in brackets refer to situations in which back electron transfer ($k_{cr}$) to the locally excited state has been shown to contribute significantly. $^{b}$ Data obtained by TRMC, ref 16. $^{c}$ Data obtained from CT fluorescence, ref 1c. $^{d}$ Recombination too fast to be measured. $^{e}$ Not measured. $^{f}$ No CT emission observed.
In the 12-bond compound the yield of charge recombination is 0.67, so in the 12-bond bichromophore the yield of direct formation of the local triplet from the local singlet excited state is negligible in the absence of a time window over which the decay of that triplet is negligible in the absence of benzophenone and an excess of 2. However, during such a time window the TA spectra (see Figure 2) show an isosbestic point at 483 nm. At this point the molar extinctions at 530 nm) 20,21 the molar extinction at 355 nm). Spectra max (n) of that triplet is negligible in the absence of benzophenone, even to incomplete quenching of the triplet of benzophenone, even at high 2 concentration, in a time window over which the decay of that triplet is negligible in the absence of 2. However, during such a time window the TA spectra (see Figure 2) show an isosbestic point at 483 nm. At this point the molar extinctions of both species taken separately.

Figure 2. Transient absorption spectra of the sample containing benzophenone and an excess of 2 in benzene (λ exc = 355 nm). Spectra shown were taken with an interval of 200 ns. Note that the full time window is short relative to the triplet lifetimes of both species taken separately.

The quantum yield for triplet formation in the donor model is 0.67, so in the 12-bond bichromophore the yield of direct formation of the local triplet from the local singlet excited state amounts to 0.13. After correction is made for this contribution, the rates of recombination from the charge-separated state to the singlet ground state (k1cr) and to the D*A state (k3cr) were calculated under the assumption that Φ1cr = Φ3cr = 1 (Table 3). For 1[12] in benzene this assumption is not valid, because repopulation of the local singlet state (k1be) occurs, as was discussed previously.15,16 Therefore the data for 1[12] in benzene will be omitted from the following analysis.

From the data in Table 3 it can be seen that both k1cr and k3cr increase sharply in dioxane as compared to benzene. More importantly, the data also show that while k1cr increases by several orders of magnitude at shorter charge separation distances, it becomes comparable or even smaller for larger separation distances. Thus the contribution of the triplet pathway (Φ3) increases strongly with the bridge length (see Table 2). Also Φ3 is always somewhat larger in the slightly more polar dioxane than in benzene. The latter appears in contradiction with polarity dependence results reported for a related rigid D—bridge—A system,12 but in that case the polarity was only changed within a series of saturated hydrocarbon solvents. Most importantly the results now obtained show that k1cr and k3cr strongly differ in distance dependence. Logarithmic plots versus the number of σ bonds (n) separating donor and acceptor are shown in Figure 4. Their near linearity22 implies that for both k1cr and k3cr the distance dependence is close to exponential but with a significantly smaller damping constant for k3cr.

Because for most systems k1cr ≫ k3cr, the overall damping factors (see below) are close to those now found for k1cr, i.e. β1cr = 1.04 per bond (0.91 Å⁻¹) in benzene and β3cr = 1.22 per bond (1.06 Å⁻¹) in dioxane. The damping factor for triplet recombination is dramatically smaller, viz. β3cr = 0.45 per bond (0.38 Å⁻¹) and 0.56 per bond (0.48 Å⁻¹) in benzene and dioxane, respectively.

### TABLE 2: Quantum Yields of Formation of the Local Donor Triplet (Φ1) Determined with Transient Absorption and Lifetime (τ1) of the Charge-Separated State (As Determined Earlier with TRMC) in the Compounds 1[n] in Benzene and 1,4-Dioxane

<table>
<thead>
<tr>
<th>Compound</th>
<th>Φ1 (Benzene)</th>
<th>τ1 (ns)</th>
<th>Φ1 (Dioxane)</th>
<th>τ1 (ns)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1[4]</td>
<td>b</td>
<td>0.8</td>
<td>b</td>
<td>0.8</td>
</tr>
<tr>
<td>1[6]</td>
<td>0.01</td>
<td>6</td>
<td>0.03</td>
<td>0.5</td>
</tr>
<tr>
<td>1[8]</td>
<td>0.03</td>
<td>40</td>
<td>0.09</td>
<td>2.5</td>
</tr>
<tr>
<td>1[10]</td>
<td>0.14</td>
<td>410d</td>
<td>0.37</td>
<td>43</td>
</tr>
<tr>
<td>1[12]</td>
<td>0.37</td>
<td>740d</td>
<td>0.64</td>
<td>297d</td>
</tr>
</tbody>
</table>

* Yield too small to be determined. + Value extrapolated from TRMC experiments on longer homologues.

### TABLE 3: Rates of Charge Recombination to the Singlet Ground State (k1cr) and to the D*A State (k3cr) for the Compounds 1[n] in Benzene and 1,4-Dioxane

<table>
<thead>
<tr>
<th>n</th>
<th>k1cr (×10⁵ s⁻¹)</th>
<th>k3cr (×10⁵ s⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>4</td>
<td>12000d</td>
<td></td>
</tr>
<tr>
<td>6</td>
<td>1650</td>
<td>20</td>
</tr>
<tr>
<td>8</td>
<td>242</td>
<td>8.5</td>
</tr>
<tr>
<td>10</td>
<td>21</td>
<td>3.4</td>
</tr>
<tr>
<td>12</td>
<td>b</td>
<td>b</td>
</tr>
</tbody>
</table>

* Value calculated from TRMC lifetimes under assumption of Φ1 = 0. Determination prevented by contribution of k1be.

### Figure 3. T1 → Tn absorption spectra of (A) benzophenone and (B) 2 in benzene.

Distance Dependence of Charge Recombination
to oscillate between singlet and triplet character (1 \( \rightarrow \) 3 state. Under such conditions it is mainly hyperfine interaction that may be quite slow in comparison to the singlet recombination. Furthermore it should be noted that for short bridges the latter process is significantly smaller (in view of the large exchange interaction is expected to be quite large, making the rate at which the CT state is going to acquire sufficient singlet character to produce \( ^3(D*A) \) upon charge recombination. For short bridges this is apparently not a very competitive process, which can be attributed to at least two factors. First of all the rate of singlet recombination \( k_{1cr} \) is high, but furthermore it should be noted that for short bridges the exchange interaction is expected to be quite large, making \( ^1(D^+A^-) \) and \( ^3(D^+A^-) \) distinct species, and the intersystem crossing between these \( (k_{isc}) \) requires spin–orbit coupling and may be quite slow in comparison to the singlet recombination.

For longer bridges, however, not only does the rate of the latter process decrease exponentially with increasing bridge length, but so does the exchange interaction, thereby leading to a strong attenuation of the singlet/triplet energy gap in the CT state. Under such conditions it is mainly hyperfine interaction that may actually exceed the rate of singlet recombination. For \( \omega_{isc} \gg k_{1cr} \), the ratio of the singlet and triplet recombination yields closely reflects, apart from spin statistics, the ratio of the FCWD-V\( ^2 \) factors for \( ^1(D^+A^-) \rightarrow DA \) and \( ^3(D^+A^-) \rightarrow DA \). Since the energy gap in the latter process is significantly smaller (in view of the large energy gap involved, charge recombination to the ground state in fact always remains deep in the inverted region), it is not unlikely that triplet recombination becomes the dominant pathway at larger bridge length just because it has a more favorable FCWD and as is actually observed here for 1[12] in dioxane.

Thus the fact that \( k_{3cr} \) exceeds \( k_{1cr} \) at longer distances can qualitatively be explained from the increased triplet character of the CT state at larger distances and the more favorable FCWD for the triplet recombination pathway.

Concluding Remarks

The electronic coupling (\( V \)) is often assumed to be the most important factor responsible for the distance dependence of electron transfer rates, and differences in “damping factor” between systems involving different media interposed between D and A are generally discussed in terms of structural effects of these media on \( V \). The main result of the present study is that charge recombination pathways to the ground state and the \( ^3D*A \) state in the bichromophoric compounds \( [1n] \) have very different damping factors. In both solvents studied, the rate of charge recombination to the \( ^3D*A \) state is much less distance dependent than that for charge recombination to the ground state, and as a result the contribution of the triplet pathway tends to become dominant for larger separation distances. This implies that the overall distance dependence of charge recombination cannot be described adequately by a single set of distance dependent \( V \) and FCWD factors but that in fact the distance dependence of two such sets should be taken into consideration as well as the distance dependence of the rate with which the CT state, prepared with singlet character, acquires triplet character.

Experimental Section

Transient absorption spectra were obtained using a Lumonics EX700 XeCl excimer laser (308 nm) as the excitation source and a 450 W high-pressure Xe-arc in right angle geometry as probe light, pulsed with a Müller Elektronik MSP05 pulser to enhance its brightness during the observation time gate of the detector. The probe light, after passing through the sample cell, was collected by an optical fiber and fed into a Jarrel-Ash monospec 27 model 1234 spectrograph in which the light was dispersed by a grating (150 grooves/mm) onto an MCP intensified diode array detector (EG&G 1421G, 25 mm, 1024 diodes). With this setup a spectral range of about 600 nm was covered with a bandwidth of 7 nm (250 \( \mu m \) slit). The detector was gated at 100 ns by an EG&G FPG05 fast pulse generator, and the start of the time window was delayed in 100 ns increments relative to the laser pulse to obtain subsequent spectra across the total decay time of the transients studied. The timing of the laser, the probe light and the optical multichannel analyzer (OMA) gate pulse were controlled by an EG&G OMA III Model 1460 console with a 1303 pulse delay generator and a digital delay generator (EG&G 9650). Spectra were averaged over 25 pulses for each delay to improve the signal to noise ratio.

To excite benzophenone in the triplet energy transfer experiment, the third harmonic of a Spectra Physics Quanta Ray GCR-3 Nd-YAG (355 nm) was used instead of the excimer laser. The sample was prepared by dissolving benzophenone in benzene to a concentration yielding \( A_{1000} \) = 1 and adding of an excess of 2. The sample was degassed by several freeze–pump–thaw cycles.

The samples for the transient spectra of the compounds \( [1n] \) were prepared by dissolving the respective compound in benzene (Merck Uvasol) or 1,4-dioxane (Merck Uvasol) to a concentra-
tion yielding $A_{1\text{cm}}^{308} = 1-1.5$, and degassed by several freeze–
pump–thaw cycles. The solutions were excited at 308 nm using
the excimer laser. The absorbance of the dimethoxynaphthalene
triplet was determined from the first spectrum in which the
dimethoxynaphthalene radical cation absorption at 408 nm was
absent (see Figure 5). At that delay time a dimethoxynaphtha-
lene triplet absorption remains, which is constant in the time
range used ($\tau(2) \approx 9 \mu s$). After each experiment the laser
power was determined by repeating the experiment with a
benzophenone sample and an absorbance at 308 nm which
equals the absorbance of the preceding sample. Quantum yields
were determined by eq 5:

$$
\Phi_{\text{sec}} = \frac{(A_{\text{DMN}})/\epsilon_{\text{DMN}}}{(A_{\text{BP}})/\epsilon_{\text{BP}}} \frac{A_{\text{BP}}^{308}}{A_{1[n]}^{308}} \tag{5}
$$

In this equation $A$ are the absorbances of the $T_1 \rightarrow T_n$
absorption of 1,4-dimethoxynaphthalene (DMN) at 450 nm and
benzophenone (BP) at 530 nm respectively, $\epsilon$ are the molar
extinction coefficients of these absorption bands ($\epsilon_{\text{max}}(\text{BP}) =
7630 \text{ M}^{-1} \text{ cm}^{-1}$ at 530 nm$^{20,21}$ and $\epsilon_{\text{max}}(\text{DMN}) = 7000 \text{ M}^{-1}\text{ cm}^{-1}$ at 450 nm), and $A_{\text{BP}}^{308}$ are the absorbances at the excitation
wavelength for both sample $I[n]$ and reference BP.

Acknowledgment. The present research was supported by
the Netherlands Foundation of Chemical Research (SON) with
financial aid from the Netherlands Organization for Scientific
Research (NWO) and the Australian Research Council.

References and Notes

(1) (a) Oevering, H.; Paddon-Row, M. N.; Heppener, M.; Oliver, A.
M.; Cotsaris, E.; Verhoeven J. W.; Hush, N. S. J. Am. Chem. Soc. 1987,
109, 3258. (b) Paddon-Row, M. N.; Oliver, A. M.; Warman, J. M.; Smit,
K. J.; de Haas, M. P.; Oevering, H.; Verhoeven, J. W. J. Phys. Chem. 1988,
92, 6958. (c) Oevering, H.; Verhoeven, J. W.; Paddon-Row, M. N.; Warman,
J. M. Tetrahedron 1989, 45, 4751. (d) Warman, J. M.; Smit, K. J.; Jonker,
S. A.; Verhoeven, J. W.; Oevering, H.; Kroon, J.; Paddon-Row, M. N.;
Oliver, A. M. Chem. Phys. 1993, 170, 379. (e) Hash, N. S.; Paddon-Row,
Phys. Lett. 1985, 117, 8. (f) Verhoeven, J. W.; Paddon-Row, M. N.; Hush,
N. S.; Oevering, H.; Heppener, M. Pure Appl. Chem. 1986, 58, 1285. (g)
Warman, J. M.; de Haas, M. P.; Paddon-Row, M. N.; Cotsaris, E.; Hush,
N. S.; Oevering, H.; Verhoeven, J. W. Nature 1986, 320, 615. (h) Pasman,
P.; Rob, F.; Verhoeven, J. W. J. Am. Chem. Soc. 1982, 92, 6958. (i) Mes,
G. F.; Jong, B. de; Ramesdonk, H. J. van; Verhoeven, J. W.; Warman,
Soc. 1984, 106, 6524. (j) Leland, B. A.; Jordan, A. D.; Felker, P. M.; Hopfield,
Niemczyk, M. P. J. Am. Chem. Soc. 1985, 107, 1080, 5562. (o) Heitele,


Misumi, S. J. Lumim. 1976, 12/13, 159.

46, 334.

(6) Roth, H. D.; Schilling, M. L. M. J. Am. Chem. Soc. 1980, 102,
4303.


1873.

1984, 78, 271.

1985, 118, 19.

(11) Van Haver, Ph.; Helsen, N.; Depaempelaere, S.; Van der Auweraer,

Chem. 1993, 97, 13138.

(13) Oevering, H. Ph.D. Thesis, University of Amsterdam, Amsterdam,

(14) Lawson, J. M.; Craig, D. C.; Paddon-Row, M. N.; Kroon, J.;

(15) Smit, K. J.; Warman, J. M.; Haas, M. P. de; Paddon-Row, M. N.;

(16) Warman, J. M.; Smit, K. J.; Haas, M. P. de; Jonker, S. A.; Paddon-
Row, M. N.; Oliver, A. M.; Kroon, J.; Oevering, H.; Verhoeven, J. W. J.

(17) Kroon, J.; Oevering, H.; Verhoeven, J. W.; Warman, J. M.; Oliver,

5353.

(19) Kroon, J. Ph.D. Thesis, University of Amsterdam, Amsterdam,


1987, 16, 239.

(22) In benzene as a solvent the compound [12] is known to give back
electron transfer to the local singlet state (refs 16). This results in a deviation
from the linear dependence for this compound in benzene.

1978, 58, 4860.