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Hopper Growth of Salt Crystals

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Supporting Information

ABSTRACT: The growth of hopper crystals is observed for many substances, but the mechanism of their formation remains ill understood. Here we investigate their growth by performing evaporation experiments on small volumes of salt solutions. We show that sodium chloride crystals that grow very fast from a highly supersaturated solution form a peculiar form of hopper crystal consisting of a series of connected miniature versions of the original cubic crystal. The transition between cubic and such hopper growth happens at a well-defined supersaturation where the growth rate of the cubic crystal reaches a maximum (~6.5 ± 1.8 μm/s). Above this threshold, the growth rate varies as the third power of supersaturation, showing that a new mechanism, controlled by the maximum speed of surface integration of new molecules, induces the hopper growth of cubic crystals in cascade.

The beauty of snowflakes is partly due to the amazing variety of shapes they come in. Yet, the underlying crystalline structure of snow is always the same; the thermodynamic equilibrium shape is hexagonal. The variety of shapes is due to the kinetics of ice growth that is clearly more important in determining the snowflake shape than the equilibrium shape dictated by the crystal structure and surface energy. In nature, crystals generally are rarely found only in their equilibrium state; instead, many minerals and salts appear as clusters of interconnected crystalline regions known as hopper crystals. The latter results from an anisotropic growth in which the edges of a crystal grow faster than the centers of its faces.

The morphology of such hopper crystals is common to many substances and minerals, such as bismuth, quartz (called skeletal or fenster crystals), gold, calcite, and halite (NaCl). Also here, it is suspected that the kinetics of growth is an important factor in determining the overall shape, but so far, little fundamental work has been done to understand hopper crystal shapes. In general, the conditions under which crystal growth occurs, i.e., parameters such as supersaturation, volume, and homogeneity of the solution, remain largely unknown. Understanding how crystal shapes come about is nonetheless of considerable importance in applications such as weathering of rocks and historical monuments,1−3 pharmacology,4,5 glaciology, and for the food industry.6 For the latter, as crystal morphology dictates the dissolution speed of crystals, the control over the crystal structure is an important variable to consider as a strategy to reduce sodium intake while maintaining salt intensity and saltiness perception. Here, we study hopper crystals formed from solution crystallization of the most common salt on earth: sodium chloride (NaCl). Sodium chloride is a key factor in many industrial and geological processes, e.g., in the food industry, oil well drilling, CO2 sequestration, erosion of rocks, sodification of soils, and weathering of historical monuments. Equilibrium NaCl crystals are well-known to have a cubic shape because the underlying crystal lattice has a cubic close-packed (CCP) structure; however, both in nature and in laboratory experiments, hoppered shapes are often encountered (Figure 1). Most of the reported hopper forms of sodium chloride in the literature are hollow cubes or pyramid shapes,7−9 and very few experiments on skeletal hopper growth have been reported10−12 (Figure 1).

Recently, it was reported that the metastability limit for the precipitation of NaCl from an aqueous salt solution is almost twice the saturation concentration.10,13 One of the striking consequences of this is that it appears to lead to the precipitation of a collection of interconnected cubes, a hopper
form called “skeletal growth”. Because of the singular importance of the salt, its atomistic pathways of crystal nucleation and the subsequent kinetics of cubic growth of NaCl have been extensively investigated. Nonetheless, although hopper growth has been reported for different salts, little is known experimentally about the transition from cubic to hoppered growth.

In this Letter, we report the primary nucleation and growth of NaCl crystals in small volumes at different supersaturations. We show that there is a specific supersaturation above which the transition from cubic to hoppered growth is observed. We demonstrate experimentally that this supersaturation corresponds to the maximum growth rate of the cubic crystal and explain why above this concentration hopper morphologies appear in the form of interconnected cubes. The experiments are performed under controlled evaporation (T = 23 ± 1 °C and RH = 50%) of small volumes (∼10⁻²–10⁻³ μL) of an initially undersaturated NaCl solution (Sigma-Aldrich purity > 99.9%, at initial concentration m₀ = 4.9 mol/kg) in microcapillaries (100–200 μm). The evaporation rate is measured by recording the displacement of the two menisci of the entrapped salt solution using a CCD camera connected to an optical microscope. The concentration at the moment of precipitation is determined from the volume difference (V₀ - Vₜ) between the initially introduced volume (V₀) and the solution’s volume at the onset of crystallization (Vₜ). For each experiment, the supersaturation at the onset of crystallization in the solution is defined as Sm = mₜ/m₀, where mₜ and m₀ are the molal concentrations when the crystal precipitates and at equilibrium, respectively. We use small volumes because the Peclet number (the ratio of advection due to the evaporative flux and diffusion) is small and the bulk solution is therefore homogeneous: in such small volumes, the Peclet number is on the order of 10⁻² close to the meniscus.

In all of our experiments, only one crystal was observed to precipitate; this is because the nucleation rate depends very steeply on the concentration, and in small volumes, the probability of the growth of other nuclei becomes very low. Due to the statistical nature of nucleation, the first formation of a crystal is observed for a range of supersaturations up to the metastability limit of 1.6 ± 0.2. The key observation is that for supersaturations of S < 1.45 we clearly see the growth of simple cubic crystals. These take a rectangular shape due to the confinement, i.e., walls of the microcapillary; the faces of the cube are crystallographically equivalent (1 0 0) faces. However, for supersaturations higher than S > 1.45 ± 0.05, hopper crystals rapidly appear in the form of a chain-like structure of cubic crystals: the higher the supersaturation at the onset of precipitation, the larger the number of interconnected cubic crystals and the number of branches of the hopper morphology (Figure 2).
To quantify the growth speed, the size of growing crystals is measured as a function of time. For the cubic crystals, the measurements are done along one side (Figure 2). In the case of the hopper morphology, the growth of the face of a single cubic unit is measured as well as the length of each branch \( (L_D) \) and total length \( (L_T) \) by summing up the number of branches.

As can be observed in Figure 3, the hopper crystals grow with a rate that can be more than 10 times higher than that of the cubic crystal. For both forms, cubic and hopper, the crystal length grows linearly with time near the onset of precipitation \((i.e., t < 20 \text{ s})\). The growth of the crystal subsequently slows down due to the consumption of ions and the concomitant decrease of the supersaturation in the surrounding salt solution. The kinetics of growth is then limited by the arrival of molecules from the bulk to the surface and is hence governed by diffusion. This is evident from Figure 3, where we compare the growth kinetics to the expected \( \sqrt{t} \) behavior when bulk diffusion is the rate-limiting step for crystal growth. As can be observed in Figure 3, the diffusive growth rate depends on the shape of the crystal. However, a full theoretical treatment of this difference is very difficult because it necessitates solving the Laplace equation (for the diffusion) with the appropriate boundary conditions. This is impossible if the boundary is of arbitrary shape and in addition is moving except in some special cases; this is beyond the scope of this paper. The agreement with simple diffusive dynamics is excellent for all but the shortest times \((t < 20 \text{ s})\). At very short times, the growth is more rapid than predicted by simple bulk diffusion, and hence, another mechanism must be at play.29 This is because at the onset of crystal precipitation the supersaturation is high and consequently the transport of the ions from the bulk to the crystal surface is not the limiting factor for growth. It is rather the incorporation of molecules into the surface of the critical nucleus that limits the growth at the early stage of growth.

Figure 3. Growth of NaCl crystals in time: (a) growth of the cubic crystal \((L_F, \text{open circles})\) and (b) hopper growth \((L_T, \text{red squares})\) as a function of the square root of time. The straight continuous lines indicate the diffusion-limited growth and describe the data well for times \(> 20 \text{ s}\). The inset shows the same data for short times in a linear plot. For times \(t < 20 \text{ s}\), the growth is linear in time with coefficients, (a) \(y = 6.2x, R^2 = 0.958\) and (b) \(y = 20x, R^2 = 0.814\).

Figure 4. (a) Growth rates during the first 20 s after crystallization for separate experiments (black squares) as a function of relative supersaturation \((S_m - 1)\) at the onset of precipitation. A linear fit (solid line) was made for relative supersaturations below 0.5, excluding the outlier data point at 7.3 \(\mu\text{m/s}\). The gray zone represents the region of the maximum growth rate: for each decile starting from \(S_m = 1 \geq 0.55\), the average value is given (red diamonds) with its spread in relative supersaturation \((\text{min–max range})\) and the standard deviation of the growth rate. The gray zone corresponds to the average ± standard deviation \((6.5 \pm 1.8 \mu\text{m/s})\) of the experiments at relative supersaturations above 0.5. (b) Arrhenius plot of the overall crystal growth coefficients. The error bars correspond to the spread in \(1/T\) (min–max range) and the standard error in \(K\).
For the cubic growth, looking at the initial growth behavior at \( t < 20 \text{ s} \), we find that it is characterized by a constant growth rate \( dL_c/dt \); this rate increases linearly with supersaturation up to \( S_m \approx 1.45 \), but surprisingly, above this threshold, it reaches a maximum value of about \( 6.5 \pm 1.8 \mu \text{m/s} \) (see Figure 4). This suggests that in the presence of a large excess of ions, the incorporation of ions into the growing crystal becomes the rate-limiting step. However, our results show that there is also a maximum speed with which this can happen; for the highest supersaturations, the growth rate in Figure 4a levels off. This coincides exactly with the onset of growth of the hopper morphology, suggesting that when the maximum growth speed of the cubic crystal is attained and the supersaturation is still very high the hopper crystals result from a change in growth mechanism.

The classical Burton–Cabreba–Frank (BCF) theory of crystal growth when the rate-limiting step for the growth of a crystal is the incorporation of molecules in the crystal lattice describes the experimental data in a satisfactory manner (see the SI). We therefore investigate the process for surface integration, which is given by the overall growth rate order \( m \) that characterizes how the growth rate depends on the supersaturation as\(^{30,31}\)

\[
\frac{dL}{dt} = K(S_m - 1)^m
\]  

(1)

with \( K \) being the overall crystal growth coefficient. The linear behavior \((g = 1)\) observed for \( S_m < 1.45 \) in Figure 4a confirms that indeed in this regime the growth of a cubic crystal is limited by the incorporation of ions into the crystal.

The overall growth rate coefficient \( K \) obtained from our experimental data is about \( 9.4 \mu \text{m/s} \). The latter together with some measured \( K \) values for cubic growth reported in the literature (see the SI) is plotted in Figure 4b as a function of the inverse temperature to establish an Arrhenius behavior

\[
\ln K = \ln C - \frac{\Delta G_m}{R} \frac{1}{T}
\]  

(2)

with \( C \) a constant and \( \Delta G_m \) the activation energy of the growth process. Our results obtained at much higher supersaturations are in line with those published for much lower concentrations \((10^{-4} \leq S_m < 1 \times 10^{-7})\). The combined data yield an activation energy \( \Delta G_m \approx 21.5 \pm 2.9 \text{ kJ/mol} \). Previously found activation energies are usually around \( 20 \text{ kJ/mol} \) very close to our finding.

We have also plotted the total growth rate of the hopper crystals \( dL_T/dt \) as a function of the supersaturation for experiments where \( S_m - 1 > 0.45 \) (Figure 5). The results can be fitted with a power law with an exponent of 3 showing that the overall growth rate order in eq 1 is \( g = 3 \). This result underlines that above \( S_m > 1.45 \) the limiting step for the growth of NaCl crystals is the total amount of (cubic) surface area. Because cubic growth is limited by the incorporation of ions into the surface, in these range of supersaturation, the only way to incorporate more ions from the supersaturated solution is to create new surfaces by forming a chain of several cubic crystals making the hopper morphology; the edges of the first cubic nucleus point in three directions into a highly supersaturated solution, and they can grow faster than the flat centers of its faces. In this way, they will serve as a point for the secondary nucleation of a new cubic crystal with its maximum growth rate, which will result in the “staircase of crystals” observed in Figure 2; the higher the supersaturation, the larger the number of interconnected cubic crystals. This creates a lot of surface area, which is thermodynamically unfavorable. The hopper growth is thus a kinetic effect, induced by the very high supersaturations that makes that the crystals find other ways to grow than the simple cubic single crystal.

Such sudden precipitation of interconnected cubes in small volumes leads to a sharp decrease of the local supersaturation, concomitant with the abrupt change of the growth rate measured for hopper branches (Figure 3). The kinetics of growth becomes then diffusion-limited, and if the volume of the surrounding salt solution is large enough to still supply the crystal with ions for a very long time, the hopper morphology will subsequently evolve toward the polyhedral growth of its extreme cubic units until complete evaporation (Figure 5).

In summary, we report on the mechanism of hopper formation of sodium chloride crystals as skeletal growth of interconnected miniature cubes in controlled evaporation experiments of small volumes of salt solution up to its metastability limit. Although such hopper growth has been reported for different salts, the mechanism behind hopper formation had not been elucidated experimentally before. We find a well-defined supersaturation at which the transition from cubic to hopper growth is observed: \( S_m \approx 1.45 \pm 0.05 \). This concentration corresponds to the point where the growth speed of the cubic crystal, which was increasing linearly with increasing supersaturation, reaches a maximum of \( 6.5 \pm 1.8 \mu \text{m/s} \). Up to this threshold, the mechanism of cubic growth is well described by BCF theory; by gathering the available

Figure 5. Top: Log–log plot of the growth rate of cubic and hopper crystals \( (dL_T/dt) \) as a function of the supersaturation. The hopper morphology is observed for relative supersaturations \( S_m - 1 > 0.45 \). (Bottom) Evolution of Hopper morphology with time: (a) \( t = 10 \text{ s} \) after precipitation at \( S_m \approx 1.68 \); (b) \( t = 9 \text{ min} \); (c) \( t = 39 \text{ min} \); (d) \( t = 43 \text{ min} \).
growth rate parameters reported in the literature and combining with our data, an activation energy of 21.5 kJ/mole is found. Above this threshold, because cubic growth is limited by the incorporation of ions into the surface, the only way to incorporate more ions from the supersaturated solution is then to create new surfaces. Because the edges of the first cubic nucleus point in three directions into a highly supersaturated solution, they can grow faster than the cubic nucleus point in three directions into a highly supersaturated solution. In this way, they will serve as a point for the range of applications.

**REFERENCES**

(9) Fontana, P. Die Vielfalt Der Salzkristalle; R + A Print, 2013.

