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Understanding the oxidative dehydrogenation of ethyl lactate to ethyl pyruvate over vanadia/titania†

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We studied the vapour-phase oxidative dehydrogenation (ODH) of ethyl lactate with air to give ethyl pyruvate over V2O5/TiO2 catalysts in a fixed-bed reactor. The nature of the vanadia species is changed by varying the vanadium surface density, and the corresponding structure of the VOx species was determined by XRD, UV-vis spectroscopy, XPS and H2-TPR. Monomeric and isolated vanadia species dominate at lower vanadium surface densities. As the surface density increases, two-dimensional polyvanadates and bulk-like vanadia crystallites become predominant. The activity per vanadium decreases with increasing vanadium surface density, indicating that the monomeric VOx species is better for pyruvate production and that the V–O–Ti bonds play an important role in the ODH of ethyl lactate. This is also confirmed by the superior catalytic performance of V2O5/TiO2 compared to vanadium supported on MgO, Al2O3, ZrO2 and CeO2.

In situ DRIFT spectroscopy coupled with mass analysis shows that the reaction can involve three possible adsorption modes of ethyl lactate on the V2O5/TiO2 surface. Under anaerobic conditions, 2-hydroxypropionate forms, giving ethyl acetate as the major product. Conversely, under aerobic conditions, oxygen that is chemisorbed on V2O5/TiO2 is active and easily replenished from the gas phase, converting the ethyl-propionate-2-oxide intermediate into ethyl pyruvate.

Introduction

Lignocellulosic biomass is one of the few sources of truly renewable carbon.1,2 But there’s a lot of it: even today, the world biomass production could meet the entire carbon demand of the chemical industry (excluding transportation fuels). Typically, biomass is first converted into simple derivatives, or ‘platform molecules’, that are then reacted further to industrial chemicals. Lactic acid (LA) and lactic esters are important platform molecules. They can be converted into a variety of bulk chemicals, including acrylic acid, propionic acid, pyruvic acid and acetaldehyde (Scheme 1).3,4

The oxidative dehydrogenation (ODH) of lactic acid to pyruvic acid is especially interesting due to the high demand for pyruvates in the pharmaceutical and agrochemical sectors.5 Today, pyruvates are still made via the classic dehydrative decarboxylation of tartaric acid in the presence of an excess of KHSO4.6,7 However, the extra reagent and the high temperature (300 °C) make this process unsustainable. Therefore, attention has focused on oxidative dehydrogenation of lactate as a ‘greener’ alternative.8 Several catalysts including Pd/Pt,9 iron phosphates,10 metal oxides (Mo, Ti, Zr, W and Sn),11 and binary oxides (TeO2–MoO3, SnO2–MoO3)12,13 were reported for this reaction in both the gas phase and liquid phase. Yet most of these also catalyse C–C scission, especially in the gas phase, giving acetaldehyde, CO and CO2.5 The catalytic challenge, therefore, is running a selective ODH reaction at a lower temperature, thus avoiding C–C bond scission. Supported/bulk vanadium oxides are used in a variety of ODH reactions at relatively low temperatures.14,15 The performance of supported vanadium catalysts depends on the type of support and the structure of the surface vanadium species.16 Compared with other oxide supports, titania interacts strongly with vanadia. The anatase phase in particular gives a stable VOx monolayer with a catalytic oxidation...
performance superior to that of rutile.\textsuperscript{17,18} We recently showed that TiO\textsubscript{2} itself can also catalyse the ODH of ethyl lactate to ethyl pyruvate in the liquid phase, due to its high affinity to ethyl lactate and oxygen.\textsuperscript{19} Meanwhile, Li et al. reported that MoVNB\textsubscript{x} supported on TiO\textsubscript{2} was more active than unsupported components for vapour phase ethyl lactate conversion.\textsuperscript{20} Cavani and co-workers elucidated the role of the components in MoV(Nb)TeO catalysts, showing that vanadium is crucial for achieving high ODH performance.\textsuperscript{21} Our preliminary tests showed that commercial V\textsubscript{2}O\textsubscript{5} gave higher yields of pyruvate compared with other oxides (MoO\textsubscript{3}, TeO\textsubscript{2} and MoVO\textsubscript{x}) at low temperatures (Fig. S1\textsuperscript{†}). Thus, we hypothesised that VO\textsubscript{x}/TiO\textsubscript{2} would be both active and selective in the ODH of lactates.

Here, we studied the structure–activity relationships in ethyl lactate ODH over V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} catalysts. We prepared a series of vanadium oxides supported on anatase titania, varying the vanadium loading via incipient wetness impregnation. Catalyst characterisation showed that different types of VO\textsubscript{x} species (monomeric, polymeric, and crystalline domains) formed on the TiO\textsubscript{2} surface. We then examined the correlation between the VO\textsubscript{x} structure and catalytic activity in the ODH of ethyl lactate with air in a fixed-bed reactor, comparing the catalytic properties of V\textsubscript{2}O\textsubscript{5} impregnated on TiO\textsubscript{2}, MgO, Al\textsubscript{2}O\textsubscript{3}, ZrO\textsubscript{2} and CeO\textsubscript{2}. Across this series, V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} showed a superior ODH activity, which we attribute to the cooperative effects between vanadia and titania.

In situ DRIFT spectroscopy showed that three adsorption modes of ethyl lactate are possible on the V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} surface. Under anaerobic conditions, 2-hydroxypropionate forms, giving ethyl acetate as the major product. Under aerobic conditions, oxygen chemisorbed on V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} is active and easily replenished from the gas phase, converting the ethyl-propionate-2-oxide intermediate into ethyl pyruvate.

Results and discussion

Synthesis and characterization of VO\textsubscript{x}/TiO\textsubscript{2} catalysts

First, we prepared five supported vanadium oxide catalysts by incipient wetness impregnation using aqueous solutions of ammonium metavanadate (NH\textsubscript{4}VO\textsubscript{3}) and oxalic acid, which were subsequently dried and calcined at 550 °C. The samples are denoted as n-V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2}, where n represents the wt% of V\textsubscript{2}O\textsubscript{5} on TiO\textsubscript{2} (n = 1%, 3%, 5%, 10%, 20%; see the Experimental section for details).

Fig. 1 shows the X-ray diffraction (XRD) patterns of the five catalysts. The crystalline structure of anatase TiO\textsubscript{2} remains stable after impregnation with vanadium oxide. We then examined the correlation between the VO\textsubscript{x} structure and catalytic activity in the ODH of ethyl lactate with air in a fixed-bed reactor, comparing the catalytic properties of V\textsubscript{2}O\textsubscript{5} impregnated on TiO\textsubscript{2}, MgO, Al\textsubscript{2}O\textsubscript{3}, ZrO\textsubscript{2} and CeO\textsubscript{2}. Across this series, V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} showed a superior ODH activity, which we attribute to the cooperative effects between vanadia and titania. In situ DRIFT spectroscopy showed that three adsorption modes of ethyl lactate are possible on the V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} surface. Under anaerobic conditions, 2-hydroxypropionate forms, giving ethyl acetate as the major product. Under aerobic conditions, oxygen chemisorbed on V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} is active and easily replenished from the gas phase, converting the ethyl-propionate-2-oxide intermediate into ethyl pyruvate.

Results and discussion

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Fig. 1 shows the X-ray diffraction (XRD) patterns of the five catalysts. The crystalline structure of anatase TiO\textsubscript{2} remains stable after impregnation with vanadium oxide. No vanadia peaks are detected as the V\textsubscript{2}O\textsubscript{5}/(V\textsubscript{2}O\textsubscript{5} + TiO\textsubscript{2}) ratio increases from 1% to 5%. This suggests that amorphous vanadia species are well dispersed on the TiO\textsubscript{2} surface. When the V loading is increased to 10% and 20%, the peaks of crystalline V\textsubscript{2}O\textsubscript{5} can be seen at 20.3° and 26.2°. This probably comes from polymeric vanadia species on the surface.\textsuperscript{22} Examining the surface morphology of V\textsubscript{2}O\textsubscript{5}/TiO\textsubscript{2} by scanning electron microscopy (SEM) upheld the XRD results (Fig. 2). All samples showed misshaped agglomerated granules and an increase in particle size with increasing vanadium content. The 20% loaded sample was hereafter discarded, because it contained essentially a separated V\textsubscript{2}O\textsubscript{5} phase.

We then studied the textural properties of the catalysts using nitrogen adsorption–desorption isotherms. All samples...
showed a type IV isotherm with an H1-type hysteresis loop characteristic of mesoporous materials (Fig. 3). As expected, increasing the V content lowered the BET surface area from 81 m² g⁻¹ to 51 m² g⁻¹ (Table 1). Actually, the structure of surface VOₓ species depends closely on the vanadium surface density (Vatoms nm⁻²). The surface densities of our samples are 0.8, 3.0, 4.8 and 12.8 Vatoms nm⁻², corresponding to a V₂O₅ content of 1, 3, 5 and 10 wt%, respectively. The theoretical surface density value for monovanadate is 2–3 Vatoms nm⁻². Thus, the vanadia species of 1-V₂O₅/TiO₂ and 3-V₂O₅/TiO₂ are probably isolated VO₄ species, while 5-V₂O₅/TiO₂ and 10-V₂O₅/TiO₂ are predominantly polymeric VO₄ species and V₂O₅ crystallites. To confirm the coordination environment of the vanadium species, we used UV-visible diffuse reflectance spectroscopy (UV-vis DRS, see Fig. 4A). The absorption intensity in the visible region is enhanced after depositing vanadium oxide on the titania support. 1-V₂O₅/TiO₂ showed charge-transfer bands centred around 280–350 nm, which are assigned to isolated tetrahedral V⁵⁺ species. As the V₂O₅ loading increased, the absorption bands shifted to higher wavelengths, reflecting the lowering of the charge transfer between oxygen and V atoms. This indicates the formation of highly coordinated polymeric vanadium species from isolated tetrahedral monomeric species (cf. similar structures observed for molybdenum). Increasing the vanadia content to 10 wt% showed crystalline V₂O₅ species (absorption at 490 nm). Defining the position of isolated and polymerized VO₄ units using UV-vis DRS is difficult, as their absorption bands can overlap with the strong absorption of the titania support. However, the absorption edge energy (Eg) can give quantitative information on the coordination number and the local structure of VOₓ species. The Eg value is determined by finding the intercept on the X-axis of the tangent line in the low-energy rise of the plot of [F(I/R)α]₁/₂ vs. hν, where F(I/R) is the Kubelka-Munk function and hν is the incident photon energy. The corresponding Eg values for V₂O₅/TiO₂ are shown in Fig. 4B, wherein the value of pure TiO₂ is about 3.27 eV. For the V₂O₅/TiO₂ catalysts, the Eg value gradually decreases with increasing surface vanadia content from 3.19 to 2.1 eV. A similar trend was observed by Danilevich et al. The high Eg value at low vanadium surface density corresponds to isolated surface VO₄ species. When the vanadium content is 1%, Eg ≈ 3.19 eV, which is close to the value of 3.21 eV for Na₃VO₄ (where vanadium exists as an isolated tetrahedral VO₄ species). Bulk NaVO₃ was reported to have a polymeric tetrahedral VO₄ structure with an Eg value of

Table 1  Textural properties of VOₓ/TiO₂ catalysts

<table>
<thead>
<tr>
<th>Sample</th>
<th>V content a (ICP, wt%)</th>
<th>V surface density (VO₄/nm²)</th>
<th>S_BET b (m² g⁻¹)</th>
<th>V_p b (cm³ g⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pure TiO₂</td>
<td>—</td>
<td>—</td>
<td>81.3</td>
<td>0.50</td>
</tr>
<tr>
<td>1-V₂O₅/TiO₂</td>
<td>0.58</td>
<td>0.8</td>
<td>81.2</td>
<td>0.52</td>
</tr>
<tr>
<td>3-V₂O₅/TiO₂</td>
<td>1.80</td>
<td>3.0</td>
<td>70.9</td>
<td>0.51</td>
</tr>
<tr>
<td>5-V₂O₅/TiO₂</td>
<td>2.70</td>
<td>4.8</td>
<td>66.3</td>
<td>0.49</td>
</tr>
<tr>
<td>10-V₂O₅/TiO₂</td>
<td>5.64</td>
<td>12.8</td>
<td>51.8</td>
<td>0.43</td>
</tr>
</tbody>
</table>

a Determined by ICP analysis. b Calculated based on N₂ sorption at 77 K.

Fig. 4  (A) UV-vis absorption spectra of TiO₂, V₂O₅ and V₂O₅/TiO₂ catalysts. (B) [F(I/R)α]¹/² plotted against the energy of the incident photon for the determination of edge energy for V₂O₅/TiO₂ catalysts; the inset shows the corresponding Eg values. The edge energies are determined from the intercept on the X-axis of a straight line fitted through the rise of the function [F(I/R)α]. (C) Schematic structures of V₂O₅/TiO₂ catalysts showing different types of vanadia species.
Thus, we assume that the 3% V₂O₅ on TiO₂ (with Eₓ = 2.84 eV) also contains polymerized VO₄ units in addition to isolated VO₄ species. Increasing the V₂O₅ content to 5% yielded an Eₓ value of 2.38 eV, indicating the presence of a higher amount of polyvanadate species. Note that the Eₓ value of 10% V₂O₅/TiO₂ (2.10 eV) is close to that of bulk V₂O₅ (2.05 eV), suggesting that the polymerized VO₄ species is aggregated to crystalline V₂O₅, in agreement with the XRD results.

We ran H₂-TPR experiments to further elucidate the oxidation state of surface vanadia species and the reducibility of V₂O₅/TiO₂ catalysts (see Fig. 5). The pristine TiO₂ support showed only a very weak peak around 540 °C. After depositing vanadia, the temperature of the maximum hydrogen consumption (Tmax) shifted to a lower value, compared with pure V₂O₅. This suggests that V₂O₅/TiO₂ is easier to reduce than pure TiO₂ or V₂O₅. Among our V₂O₅/TiO₂ catalysts, the Tmax shifted to a higher temperature with increasing V₂O₅ content. Generally, the reducibility of supported V₂O₅ catalysts is affected both by the type of surface vanadia species and vanadium coverage. Indeed, polymeric VO₂ species in the monolayer are reported to be more easily reduced than monomeric ones. In our case, the 3% V₂O₅/TiO₂ sample, which also contains polymerized VO₄ units, was more easily reduced compared to the 1% V₂O₅/TiO₂ sample, which has only monomeric VO₄ species, in agreement with published data. VOₓ species below or at monolayer coverage are more reducible than the VOₓ species in multilayers and in crystalline or bulk vanadia. Accordingly, reducing 5% V₂O₅/TiO₂ and 10% V₂O₅/TiO₂ was even more difficult, in agreement with this trend.

We then ran X-ray photoelectron spectroscopy (XPS) measurements to identify the chemical species on the surface (see Fig. 6). The V 2pₓ/₂ peak (Fig. 6a) was deconvoluted into two distinct peaks centered at 516.5 eV and 517.5 eV, corresponding to V⁴⁺ and V⁵⁺ species, respectively. The peaks at 464.7 eV and 458.9 eV are assigned to Ti 2pₓ/₂ and Ti 2pᵧ/₂ (Fig. 6b), suggesting that the Ti⁴⁺ state predominates. Similarly, the spectra of O 1s (Fig. 6c) were fitted with three peaks. The main peak was observed at a lower BE of 530.2 eV and assigned to lattice oxygen species (Oₗ). The peak at around 531.7 eV represented the surface chemisorbed oxygen (Oₛ), while that at 532.6 eV represented the chemisorbed water Oₙ. We also varied the liquid hourly space velocity (LHSV) of ethyl lactate from 0.5 h⁻¹ to 3.5 h⁻¹ (corresponding to a 1–8 ml h⁻¹ ethyl lactate feeding rate). The results are summarized in Table S1. As expected, with an increase in LHSV, the ethyl lactate conversion gradually decreased due to the decrease in contact time. The ethyl pyruvate selectivity
increased from 20% to 80% with the change in LHSV from 0.5 to 2.5 h\(^{-1}\). However, a further increase in LHSV did not favor pyruvate. This is probably due to the coverage of the active sites of the catalyst by the reactants, leading to coke formation.

Then, the activities of V\(_2\)O\(_5\)/TiO\(_2\) catalysts for the ODH reaction of ethyl lactate to ethyl pyruvate were investigated. The pure anatase TiO\(_2\) support gave only low yields. Similarly, control experiments with commercial bulk V\(_2\)O\(_5\) gave <30% pyruvate. Interestingly, even 1-V\(_2\)O\(_5\)/TiO\(_2\) gave a higher selectivity and yield than bulk V\(_2\)O\(_5\) (Fig. 7). Note that the specific surface area of 1-V\(_2\)O\(_5\)/TiO\(_2\) is 81 m\(^2\) g\(^{-1}\), which is much higher than that of V\(_2\)O\(_5\) (5 m\(^2\) g\(^{-1}\)). This suggests that highly dispersed vanadia is responsible for the catalytic activity. When the vanadium loading was increased to 3%, the selectivity to pyruvate reached a maximum of 78% at 62% conversion. Increasing the V loading further did not improve the catalytic performance. To better differentiate the influence of surface vanadia species on catalytic performance, control experiments were performed by increasing the LHSV and decreasing the air flow at low reaction temperatures (160 °C, 180 °C vs. 200 °C). Fig. 8 shows that all V\(_2\)O\(_5\)/TiO\(_2\) catalysts show higher ethyl lactate conversion than pure TiO\(_2\) and V\(_2\)O\(_5\). The sub-monolayer 1-V\(_2\)O\(_5\)/TiO\(_2\) catalyst gave 10% ethyl lactate conversion at 160 °C. Increasing the vanadia loading to 3% (near monolayer coverage) gave an ethyl lactate conversion of 28%. With a vanadia content of over 3%, the conversion almost levelled up. We therefore attribute the reactivity enhancement to non-polymeric vanadium species close to TiO\(_2\).

To further evaluate the catalytic performance of V\(_2\)O\(_5\)/TiO\(_2\), we calculated the turnover frequency (TOF) under a relatively low conversion of ethyl lactate, and the results are shown in Fig. 9. The TOF of V\(_2\)O\(_5\)/TiO\(_2\) was inversely proportional to the vanadium surface density, in close correlation with the absorption edge energy (\(E_g\), see the right axis in Fig. 9). A low V surface density with high \(E_g\), where vanadium is present as monomeric species, gave the highest TOF. As the VO\(_x\) surface density increases, isolated VO\(_4\) species agglomerate with their nearest neighbours, resulting in a lower ratio of V–O–Ti bonds and an increase in V–O–V and V=O bonds. Meanwhile, the V–O–Ti bonds within multilayers are less accessible than those in the surface monolayer, which may explain the lower selectivity at higher vanadium loadings. Ji et al. reported that the ratios of V\(^{V+}/(V^{V+} + V^{V3+})\) reflected the ratio [vanadium strongly interacting with the support]/(total vanadium)];\(^{24}\) in our case, the V loading increased from 1% to 10%, while the ratios of V\(^{V+}/(V^{V4+} + V^{V5+})\) decreased from 0.5 to 0.29 (Table S2†). This result confirms the decrease in V–O–Ti bonds. Earlier, Wachs et al. showed that the reactivity of VO\(_x\) species is unrelated to the terminal V=O bonds in the partial oxidation of methanol.\(^{47}\) Oxygen labelling experiments also demonstrated that the V=O bond is stable during butane oxidation.\(^{48,49}\) In our system, the terminal V=O bonds do not favour the TOF of ethyl lactate conversion, as we showed subsequently using DRIFTS study. The fact that the 1-V\(_2\)O\(_5\)/TiO\(_2\) catalyst, which has negligible V–O–V bonds, gave comparable catalytic activity, confirms that the bridging V–O–V bond...
Table 2  ODH of ethyl lactate to ethyl pyruvate over V2O5 on various supports

<table>
<thead>
<tr>
<th>Entry</th>
<th>Catalyst</th>
<th>BETa</th>
<th>Con.c</th>
<th>Sel.d</th>
<th>Yieldd</th>
<th>pHd</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>V2O5/Co3O4</td>
<td>59</td>
<td>1</td>
<td>32.2</td>
<td>Trace</td>
<td>6.75</td>
</tr>
<tr>
<td>2</td>
<td>V2O5/MgO</td>
<td>198</td>
<td>15.9</td>
<td>0.5</td>
<td>Trace</td>
<td>11</td>
</tr>
<tr>
<td>3</td>
<td>V2O5/ZrO2</td>
<td>89</td>
<td>28.5</td>
<td>43.9</td>
<td>12.5</td>
<td>5.9-6.1</td>
</tr>
<tr>
<td>4</td>
<td>V2O5/Al2O3</td>
<td>278</td>
<td>35.6</td>
<td>36.7</td>
<td>13.1</td>
<td>8.9</td>
</tr>
<tr>
<td>5</td>
<td>V2O5/TiO2</td>
<td>81</td>
<td>58.7</td>
<td>63.7</td>
<td>37.4</td>
<td>6.0-6.4</td>
</tr>
</tbody>
</table>

a Reaction conditions: 5 ml h⁻¹ ethyl lactate, 2.25 L h⁻¹ air flow (molar ratio of ethyl lactate/O2 = 2.3), 1.0 g catalyst (1 wt% of metal oxide on TiO2), 180 °C. b Calculated based on N2 sorption at 77 K. c Determined by GC using biphenyl as an external standard. d Based on ref. 50 and 51.

The apparent activation energy (Ea) was measured at a series of temperatures under 20% ethyl lactate conversion.

Fig. 10  Arrhenius plots for steady-state ethyl lactate conversion over supported V2O5 catalysts. Reaction conditions: LHSV = 4 h⁻¹, air flow = 1 L h⁻¹. Ethyl lactate consumption rate:

\[ \text{rate}_{\text{ethyl lactate}} = \frac{\text{mole of ethyl lactate per hour in the reactor (mol s⁻¹)}}{\text{mass of catalyst}} \]

does not play a critical role in our reaction.⁴⁷ For comparison, we prepared a mixture of V2O5 and TiO2 by mixing bulk V2O5 and TiO2 powder in a mortar. This physical mixture, which contained 3 wt% V2O5, gave only a yield of 5.6% ethyl pyruvate, compared to the 48.1% yield obtained over impregnated V2O5/TiO2 (with the same vanadium loading and under otherwise identical reaction conditions). Thus, we concluded that the catalytic performance is determined by the structure of surface vanadium species, and in particular, the V–O–Ti bonds play a critical role in the oxidative dehydrogenation of ethyl lactate.

Three types of bonds can be present in supported vanadia catalysts: V=O, V–O–V and V–O–M (where M is the support metal cation). To understand the importance of the V–O–M bonds, we studied vanadium oxide catalysts supported on four different supports: MgO, Al2O3, ZrO2 and CeO2, varying the support surface acidity.⁵⁰ For a fair comparison of V=O and V–O–V bonds, we prepared porous supports with a high surface area and subsequently impregnated 1 wt% of vanadium oxide on each support to form isolated VO2 species (see the Experimental section for details). The acidity of oxide supports is a key factor in many ODH systems.⁵² The relative acidity in this series is Al2O3 > TiO2 > ZrO2 > CeO2 > MgO.⁵³,⁵⁴ In each case, we first tested the support alone in our reaction at 180 °C. Only MgO showed some conversion (13%) but no selectivity (because ethyl lactate undergoes hydrolysis to ethanol and lactic acid on the basic MgO surface). After introducing vanadia, the catalytic activities of the supported vanadia catalysts are summarized in Table 2. After impregnation, the reactivity of V2O5/MgO was unchanged. Conversely, the catalyst supported on acidic Al2O3 showed high conversion but low selectivity, owing to a competing decarboxylation of ethyl pyruvate to acetaldehyde.¹³ TiO2-supported vanadium oxide showed the highest catalytic activity, while V2O5/ZrO2 and V2O5/CeO2 scarcely catalyzed the reaction under identical conditions. Fig. 10 shows the Arrhenius plots of ethyl lactate consumption rates. For V2O5/TiO2, the apparent activation energy (Ea) of ethyl lactate conversion (48 kJ mol⁻¹) is lower than that of V2O5/Al2O3 (57 kJ mol⁻¹), V2O5/ZrO2 (75 kJ mol⁻¹) and V2O5/CeO2 (127 kJ mol⁻¹). These results confirmed that the monomeric V–O–Ti species is more catalytically active, confirming further the importance of the vanadium–oxygen interactions.

In situ DRIFTS study. Further insight into the reaction can be gained by in situ diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS). Fig. 11 shows the in situ DRIFTS of temperature-programmed ethyl lactate desorption on 3-V2O5/TiO2 in the absence of oxygen. The characteristic bands of ethyl lactate adsorbed on V2O5/TiO2 were detected. The C=H vibrations were observed at 2990 cm⁻¹, 2940 cm⁻¹, 2883 cm⁻¹, 1454 cm⁻¹, 1387 cm⁻¹ and 1302 cm⁻¹, and ascribed to νas(CH3), νs(CH3), ν(C–H), δas(CH3), δs(CH3) and δ(C–H), respectively.⁵⁵,⁵⁶ These bands decreased in intensity with increasing temperature, indicating the degradation of ethyl lactate.⁵⁷ This process was also reflected by the changes of the carbonyl and carboxyl bands on the V2O5/TiO2 surface (four peaks at 1730 cm⁻¹, 1667 cm⁻¹, 1566 cm⁻¹ and 1420 cm⁻¹, assigned to ν(C=O), ν(C=O=O···M'), νs(COO) and νas(COO) vibrations, respectively).⁵⁵,⁵⁸,⁵⁹ Increasing the temperature broadened these carbonyl and carboxyl peaks, and the signals of νs(COO) and νas(COO) were blue-shifted to 1540 cm⁻¹ and 1445 cm⁻¹ over 250 °C. This is because of the formation of adsorbed acetate species.⁶⁰ Evidently, we detected ethyl acetate by monitoring the mass signals during the temperature-programmed ethyl lactate desorption (Fig. S6†). The C=O stretching of ethyl lactate weakened (1730 cm⁻¹ and 1667 cm⁻¹), accompanying two new shoulder peaks at 1780 cm⁻¹ and 1651 cm⁻¹. These bands were assigned to the carboxyl stretching of the α-keto group of the pyruvate,⁵⁷ supporting the experimental results (Fig. S5†). The surface chemisorbed oxygen of V2O5/TiO2 could participate in the oxidation of lactate without molecular oxygen. The bands at 1224 cm⁻¹ and 1141 cm⁻¹ may reflect the alcohol OH related C–O stretching vibrations of ethyl lactate.⁶¹

Analysis of the DRIFT spectra shows several possible adsorption modes of ethyl lactate on the catalyst surface...
Upon increasing the temperature to 150 °C, we observed the characteristic bands of α-keto in pyruvate (1795 cm\(^{-1}\) and 1651 cm\(^{-1}\))\(^{57}\), together with a red-shifted COO stretch at 1597 cm\(^{-1}\).\(^{55,62}\) The in situ DRIFTS also reflected the adsorption and dissociation of ethyl lactate on the V\(_2\)O\(_5\)/TiO\(_2\) surface in the presence of oxygen. Due to the formation of pyruvate species, the position of carbonyl ν(C=O) was shifted to 1746 cm\(^{-1}\),\(^{62}\) as compared to the ν(C=O) of ethyl lactate at 1730 cm\(^{-1}\); the peaks at 1141 cm\(^{-1}\) (C–O belongs to modes (a) and (b)) also started to weaken at 150 °C, accompanying the characteristic peaks of pyruvate, while the characteristic peaks of mode (c) were unchanged. This means that ethyl-propionate-2-oxide (in modes a and b) is the key intermediate to produce pyruvate (similar to the ethoxide species in ethanol oxidation).\(^{60,61}\) The decrease of the negative peak at 3662 cm\(^{-1}\) upheld this dehydration processes. Additionally, the intensity of pyruvate bands decreased as the temperature increased over 250 °C, indicating a decomposition process (Fig. S7†). The pyruvate peaks disappeared at 400 °C. Meanwhile, the new bands at 1520 cm\(^{-1}\) and 1445 cm\(^{-1}\) were attributed to surface carbonates ν\(_{as}\)(COO) and ν\(_{s}\)(COO),\(^{63}\) as pyruvate is easily over-oxidised to form the carbonates as side products.\(^{64}\)

The bands at 2040 cm\(^{-1}\) and 1380 cm\(^{-1}\) are attributed to the V=O bond overtone band and the V-O bond combination band (probably belonging to V-O-Ti), respectively (for a detailed analysis of the pristine V\(_2\)O\(_5\)/TiO\(_2\), see Fig. S8†).\(^{61}\) The presence of negative bands means that VO\(_x\) interacts with the adsorbed species. Under anaerobic conditions (Fig. 11), the V=O stretch remained constant with temperature, while V–O was blue-shifted (from 1360 cm\(^{-1}\) at 50 °C to 1380 cm\(^{-1}\) at 400 °C). In the presence of air, the V=O band is well preserved after the formation of the pyruvate band, and then vanishes at 350 °C along with pyruvate species; meanwhile, the V–O band, which diminished in intensity, is

### Table 3 Ethyl lactate adsorption modes and their characteristic vibration wavenumber

<table>
<thead>
<tr>
<th>Molecules</th>
<th>Vibrational mode</th>
<th>Wavenumber(^a) (cm(^{-1}))</th>
<th>Wavenumber(^b) (cm(^{-1}))</th>
</tr>
</thead>
<tbody>
<tr>
<td>(\text{O}_2\text{C}(-\text{C})\text{CH}_3)</td>
<td>(\nu(\text{C}=\text{O}))</td>
<td>1725</td>
<td>1730</td>
</tr>
<tr>
<td>(\text{H}_3\text{C}\text{C}=\text{O}\text{CH}_2\text{CH}_3)</td>
<td>(\nu(\text{C}=\text{O}))</td>
<td>1668</td>
<td>1667</td>
</tr>
<tr>
<td>(\text{OH})</td>
<td>(\nu_{as}(\text{COO}))</td>
<td>1569</td>
<td>1566</td>
</tr>
<tr>
<td>(\text{H}_3\text{C}\text{C}=\text{O}\text{CH}_2\text{CH}_3)</td>
<td>(\nu_{s}(\text{COO}))</td>
<td>1450</td>
<td>1420</td>
</tr>
<tr>
<td>(\text{H}_3\text{C}\text{C}=\text{O}\text{CH}_2\text{CH}_3)</td>
<td>(\nu(\text{C}=\text{O}))</td>
<td>1058</td>
<td>Not detected</td>
</tr>
<tr>
<td>(\text{H}_3\text{C}\text{C}=\text{O}\text{CH}_2\text{CH}_3)</td>
<td>(\nu(\text{C}=\text{O}))</td>
<td>1118</td>
<td>Not detected</td>
</tr>
<tr>
<td>(\text{H}_3\text{C}\text{C}=\text{O}\text{CH}_2\text{CH}_3)</td>
<td>(\nu(\text{C}=\text{O}))</td>
<td>1614</td>
<td>Not detected</td>
</tr>
</tbody>
</table>

AL = Alcoholic functionalities. \(^a\) Based on ref. 55 and 56. \(^b\) From our in situ DRIFTS.
still present at 400 °C in the presence of carbonates. These observations support the hypothesis that V=O bonds are not involved in the ODH of ethyl lactate while V–O ones play a key role in the reaction.

In Fig. 13, we propose a reaction network for the aerobic and anaerobic conversion of ethyl lactate. Ethyl lactate first absorbs and dissociates on the V₂O₅/TiO₂ surface, with adsorption modes (a), (b) and (c). Modes (a) and (b) dominate at low temperatures, while model (c) generates 2-hydroxypropionate on the surface through a hydrolysis process at higher temperatures. In the absence of molecular oxygen, increasing the temperature accelerates the formation of mode (c), subsequently giving ethyl acetate as the major product. Pyruvate species can also be produced but are limited by the surface chemisorbed oxygen. Under aerobic conditions, however, molecular oxygen reoxidizes the catalyst, replenishing the surface oxygen and promoting the oxidative dehydrogenation of ethyl lactate to ethyl pyruvate.

Fig. 12 In situ DRIFT spectra recorded during ethyl lactate temperature-programmed-desorption over 3-V₂O₅/TiO₂ with an air flow (with O₂).

Fig. 13 Proposed reaction pathway for the aerobic and anaerobic conversion of ethyl lactate over V₂O₅/TiO₂ catalysts.

Conclusions

We studied the catalytic oxidative dehydrogenation of ethyl lactate to ethyl lactate over V₂O₅/TiO₂ catalysts. To explore the structure–activity relationship for various vanadia species, a series of V₂O₅/TiO₂ catalysts with different surface densities were prepared via incipient wetness impregnation. As their surface density increases, the isolated VO₂ species agglomerate into polymeric and crystalline VO₅₂ species, leading to a decrease in TOF of ethyl lactate oxidation. The titania-supported vanadium oxide was superior to vanadia catalysts supported on MgO, Al₂O₃, ZrO₂ and CeO₂. This shows that V–O–Ti bonds play a key role in the oxidative dehydrogenation of ethyl lactate to ethyl pyruvate. In situ DRIFTS showed that the ethyl-propionate-2-oxide species on the catalyst surface are the key intermediates in this reaction via dehydrogenation. Molecular oxygen can replenish the surface oxygen, accelerating the oxidative dehydrogenation of ethyl lactate to ethyl lactate. This work allowed us to understand the chemical-physical features needed for a vanadium-oxide based catalyst to be active and selective in the ODH of ethyl lactate to ethyl pyruvate, thus opening new perspectives in the valorisation of bio-based platform molecules.

Experimental section

Materials and instrumentation

All chemicals were commercially available and used without further purification: anatase titania (Hombikat M311), (±)-ethyl L-lactate (Sigma–Aldrich, ≥98.0%, analytical standard), ammonium metavanadate (Acros Organics, 99.5%, analytical standard), oxalic acid (Sigma–Aldrich, ≥99.0%, analytical standard), vanadium(v) oxide (Alfa Aesar, 99.2%), magnesium oxide nanopowder (Strem Chemicals, S.A. ≥ 230 m² g⁻¹, >95%), aluminium oxide (Sasol, S.A. = 181 m² g⁻¹, 97%), zirconium(IV) oxide nanopowder (Sigma–Aldrich, S.A. ≥ 25 m² g⁻¹, 99.0%), cerium(IV) oxide (Alfa Aesar, S.A. = 30–50 m² g⁻¹, 99.5%). X-ray diffraction patterns were recorded on a Rigaku Mini Flex II diffractometer using Cu-Kα radiation (λ = 1.5406 Å) at 35 kV and 30 mA. Nitrogen adsorption-desorption isotherms were measured using a Quantachrome Autosorb-3B instrument after evacuating the samples at 523 K for 6 h. The specific surface areas were evaluated using the Brunauer–Emmett–Teller method. The vanadium loading was measured by inductively coupled plasma (ICP) atomic emission spectroscopy (AES) using a Thermo IRIS Intrepid II XSP. UV-visible diffuse reflectance spectra were collected using a Jasco V670 spectrophotometer with Spectralon as a standard in the range 200–1000 nm. Scanning electron micrographs were recorded using a Hitachi S-4800 microscope. Temperature programmed reduction (TPR) measurements were performed using a 1100 Series Thermo Electron TPDRO machine by using a stream of 5% H₂/N₂ and a heating rate of 5 °C min⁻¹. XPS spectra were collected using a Thermo Scientific K-ALPHA with Al-K radiation (1486.6 eV), monochromatized by a twin crystal monochromator, yielding a
focused X-ray spot with a diameter of 400 μm, at 3 mA × 12 kV when charge compensation was achieved with the system flood gun that provides low energy electrons and low energy argon ions from a single source. The alpha hemispherical analyzer was operated in the constant energy mode with survey scan pass energies of 200 eV to measure the whole energy band and 50 eV in a narrow scan to selectively measure the particular elements. Estimation of the intensities was carried out after the calculation of each peak integral, S-shaped background subtraction and fitting the experimental curve to a combination of Lorentzian (30%) and Gaussian (70%) lines. Binding energies (BE), referenced to the C 1s line at 284.6 eV, have an accuracy of ±0.1 eV. In situ diffusion reflectance infrared Fourier transform (DRIFT) spectra were recorded using a Bruker Vertex 70 spectrometer equipped with a Pike DiffusIR cell attachment. The cell window was made of ZnSe. Spectra were recorded using a MCT detector after 128 scans at a 4 cm⁻¹ resolution. The instrument is online with an EcoSys-P mass spectrometer from European Spectrometry Systems. In each experiment, the sample was pretreated at 450 °C in He for 30 min in order to obtain a clean catalyst surface. Afterwards, the carrier gas was switched to air in the test in the presence of oxygen. Then the IR backgrounds were collected every 50 degrees from 450 °C to 50 °C. Afterwards, 1-ethyl lactate (EL) pulsing was carried out at 50 °C. Then, the catalyst was kept under the carrier gas flow for 30 min in order to eliminate physisorbed molecules. The sample temperature was then increased at 5 °C min⁻¹ and spectra were collected every 50 °C. During the overall IR analysis, several mass signals (m/z) were monitored continuously: 4, 14, 15, 17, 18, 27, 28, 29, 31, 42, 43, 44, 45, 46, 58, 60, 61, 70, 74, 103, and 116. The conversion and selectivity were quantified using an Agilent 7820A GC equipped with a flame ionization detector (FID) and a dimethylpolysiloxane capillary column (VB-1, 30 m × 0.32 mm × 3.00 μm).

Preparation of V₂O₅/TiO₂

The supported vanadium oxide catalysts were prepared following the procedure reported by Srinivas et al. In a typical synthesis of V₂O₅/TiO₂ catalysts, TiO₂ was impregnated with aqueous solutions of ammonium metavanadate (NH₄VO₃) and oxalic acid, followed by drying and calcining for 4 h at 550 °C. The resulting solid was denoted as n-V₂O₅/TiO₂, where n represents the weight percent of V₂O₅ on TiO₂ (n = 1%, 3%, 5%, 10%, 20%).

Preparation of porous MgO, ZrO₂ and CeO₂

Porous MgO was prepared according to a previous report. 10 grams of magnesium hydroxide carbonate was calcined at 500 °C for 4 hours at a heating rate of 1 °C min⁻¹ and then cooled to room temperature.

Porous ZrO₂ was prepared following the procedure reported by Davshan et al. 7.8 grams of cetyltrimethylammonium bromide (CTMABr) was added to 50 ml water. Then the pH value was adjusted to 2 by adding 2.0 M HCl solution. Next, zirconium propoxide solution (70 wt% in 1-propanol) was added to the premixed solution under rigorous stirring for 1 h. After that, the mixture was transferred into an autoclave and heated at 180 °C for 24 h. After that, the suspension was filtered, washed with water and dried at 80 °C overnight.

Procedure for catalytic experiments

The oxidative dehydrogenation of ethyl lactate to ethyl pyruvate was carried out in a fixed-bed quartz reactor with an internal diameter of 4 mm and a length of 300 mm. The catalyst (1 g, 20–25 mesh) was placed in the middle of the reactor and the upper part was filled with quartz sands for preheating the ethyl lactate. Ethyl lactate was injected into the reactor at a rate of 5 ml h⁻¹ using a syringe pump and using air as the carrier gas and terminal oxidant. After each reaction period of 2 h, the products were collected in a cold trap and a calculated amount of biphenyl was added as an external standard for GC analysis.

Conflicts of interest

There are no conflicts to declare.

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