Cathodic Disintegration as an Easily Scalable Method for the Production of Sn- and Pb-Based Catalysts for CO2 Reduction


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Cathodic Disintegration as an Easily Scalable Method for the Production of Sn- and Pb-Based Catalysts for CO₂ Reduction

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ABSTRACT: CO₂ electroreduction to formate powered by renewable energy is an attractive strategy to recycle air-based carbon. At the moment, the electrode materials showing high selectivity for formate at high current density are post transition metals such as In, Sn, Bi, and Pb. Scaling up the CO₂ electroreduction technology to industrial size requires, among other things, cheap and clean methods to produce cathode materials in the form of particles to fabricate the square meters of the electrode surface area needed for the industrial electrolyzers. We show here that it is possible to easily produce catalytic powders based on Sn and Pb via a process known as cathodic disintegration, driving the reaction with electric power and avoiding the use of organic solvents, stabilizers, and reducing agents. The catalysts produced with this method are highly selective for the reduction of CO₂ to formate and show promise for use in industrial electrolyzers. Moreover, the process of cathodic disintegration is quick and clean, it has a high atom efficiency, it uses dilute aqueous electrolytes as solvents, and it has the possibility to be driven by renewable energy.

KEYWORDS: cathodic disintegration, cathodic corrosion, nanoparticles, electrocatalysis, CO₂ reduction

■ INTRODUCTION
The electrocatalytic conversion of CO₂ is of considerable interest nowadays due to the possibility of using energy from renewable sources to convert this ubiquitous and rather inert waste product to value-added chemicals and fuels.¹ The reaction can yield several products such as CO, methane, ethylene, and formic acid. Among these, formic acid is of particular interest because of its possible use as a liquid fuel precursor and hydrogen carrier for fuel cell applications² and because of the existing market as a preservative in animal feed and in the production of leather. The formic acid salt (formate) can be used as a precursor to C₂ chemicals such as oxalic acid, glyoxylic acid, and glycolic acid, thus enabling further conversions.³,⁴ The small number of electrons (two) needed for the conversion of CO₂ to formic acid and formate, together with the high atom efficiency, make this conversion one of the economically most promising electrocatalytic conversions. Typically, catalysts based on metals such as Sn, In, Pb, Hg, and Bi are used, as they show high catalytic selectivity to convert CO₂ into formate.⁵ High productivities are obtained using gas diffusion electrode (GDE) configurations because GDEs allow us to overcome mass transport limitations arising from the low solubility of CO₂, which is a common drawback to achieve high reaction yields at high current densities.

In this work, we study a relatively unexplored method for the production of metal catalyst particles: cathodic disintegration of Sn and Pb.⁶,⁷ During the process, particles are produced directly upon cathodic polarization of metals in aqueous electrolytes, and it is possible to produce catalysts that can be readily coated on a GDE to catalyze the electrochemical conversion of CO₂ to formate. The method avoids the use of surfactants, stabilizers, organic solvents, reducing agents, and expensive chemicals, which can hamper the industrial use of catalysts prepared by other common reduction–deposition methods.⁸ In addition, the method is extremely quick.

Using cathodic disintegration, we produced catalysts based on unsupported particles of Sn, Pb, and a Sn50Pb50 alloy. These catalysts were tested by applying them as an ink to a 10 cm² GDE and the results indicate that they perform well in industrial conditions.

Moreover, we tested the possibility to directly make carbon-supported Sn particles using cathodic disintegration, inves-
tigating the effect of different applied currents during the disintegration process on the morphological characteristics of the catalyst and tested these carbon-supported catalysts in a GDE flow cell, benchmarking them against a carbon-supported Sn catalyst prepared by a simple chemical reduction method available in the literature.9 The best performing catalyst prepared by cathodic disintegration outperforms the benchmark in selectivity (average faradic yield of benchmark 73 vs 98% for our cathodic disintegration catalyst).

These findings open up the possibility to quickly produce catalysts in the form of a powder (with a high surface area and easily transferrable to GDEs) based on Pb, Sn, their alloys, and any metal that shows this behavior under cathodic polarization. Moreover, understanding cathodic disintegration in greater detail could allow for the tuning of important parameters such as particle size distribution and surface termination.

**EXPERIMENTAL SECTION**

**Materials and Chemicals.** NaOH (>97%), KHCO3 (99.5%), and H2SO4 (95% solution in water) were purchased from Acros Organics. Na2SO4 anhydrous (99%), graphite rods (3.05 mm diameter and 38.10 mm long 99.9995%), lead foil (0.76 mm thick 99.8%), tin foil (1 mm thick 99.8%), and high purity Sn and Pb wires (1 mm diameter, 99.998% trace metal basis) were obtained from Alfa Aesar. SnSO4·H2O (wt %) wire (3 mm diameter) was purchased from Merwan Inc. Nafion 5 wt% solution in lower aliphatic alcohols and water (15–20% water content) was purchased from Sigma-Aldrich. NaNO3 (>99%), NaH2PO4 anhydrous (>99%), SnCl2·2H2O, poly(vinylpyrrolidone) (PVP) (MW 55 000), ethylene glycol (EG), NaBH4 (99.99%), and trisodium citrate dihydrate (>99%) were obtained from Sigma-Aldrich. Diamond polishing paste (1 μm) was obtained from Kemet Europe BV. Carbon cloth 60% Te (1 mm thick 99.8%), and high purity Sn and Pb wires (1 mm thickness 50%, 99.8% for our cathodic disintegration catalyst). The best performing catalyst prepared by a simple chemical reduction method for the production of Sn/C nanoparticles that were used on GDEs in similar conditions.9 We dissolved 500 mg of SnCl2·H2O and 150 mg of poly(vinylpyrrolidone) (PVP) in 100 mL of ethylene glycol (EG). After the dissolution, we added 600 mg of Vulcan carbon (to obtain 30 wt % Sn/C) and used alternated stirring and sonication to obtain a homogeneous dispersion. At this point, 500 mg of NaBH4 was added directly to the solution, which was allowed to react for 30 min. At the end of the reaction, the catalyst was centrifuged and washed with acetone at least three times before drying overnight.

All of the inks based on the Sn/C catalysts were prepared by dissolving the catalyst powder in isopropanol with an amount of 5 wt % Nafion corresponding to a dry 20 wt % of the catalyst.

The catalyst-coated gas diffusion electrodes (GDEs) were prepared by airbrushing the catalytic ink on a strip of the bare GDE prepared by following a previously described method.1 The metal loading was calculated by the difference in the weight of the electrode before and after airbrushing and assuming that the ratio of a metal to an ionomer was retained during the airbrushing process. The results are shown in Table 1.

<table>
<thead>
<tr>
<th>catalyst</th>
<th>metal loading (mg/cm²)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Sn</td>
<td>0.20</td>
</tr>
<tr>
<td>Pb</td>
<td>0.15</td>
</tr>
<tr>
<td>Sn50Pb50</td>
<td>0.23</td>
</tr>
<tr>
<td>Sn/C (0.5 A)</td>
<td>0.24</td>
</tr>
<tr>
<td>Sn/C (0.7 A)</td>
<td>0.31</td>
</tr>
<tr>
<td>Sn/C (1 A)</td>
<td>0.30</td>
</tr>
<tr>
<td>Sn/C (1.2 A)</td>
<td>0.28</td>
</tr>
<tr>
<td>Sn/C chem. red.</td>
<td>0.27</td>
</tr>
</tbody>
</table>

**Particles and Surface Characterization.** X-ray diffraction (XRD) patterns of the unsupported particles were obtained by a Philips XPERT equipped with X’pert equipped in a 2θ range from 30 to 80°. Twenty microliters of the particles suspended in ethanol was dropcast on a zero-background quartz stage and analyzed in a diffractometer. The diffractograms of the same metal foils used for disintegration were also obtained for comparison.

Scanning electron microscopy (SEM) was performed on an Apreo SEM equipped with an energy-dispersive X-ray (EDX) analyzer. EDX was used to determine the atomic ratios of Sn and Pb in different areas of the Sn50Pb50 metal foil and the relative particles on the GDE; this way, we could obtain local particle compositional information (within the interaction volume limit of SEM-EDX of about 1 μm). EDX analysis was also used to investigate the distribution of Sn on the Sn/C powder and estimate the Sn loading on the carbon support.

Transmission electron microscopy (TEM) images were taken in bright-field mode on an FEI Tecnai 20 (type Sphera) operated with a LaB6 filament at 200 kV with a bottom-mounted CETA II 4k × 4k CMOS camera.

**Flow Cell Electrosynthesis.** The controlled potential electrolysis was carried out in a commercial two-compartment, 10 cm² GDE flow cell (ElectroCell Micro Flow Cell). The anode and the cathode were separated by a reinforced Nafion membrane N324. The anolyte was a 0.5 M H2SO4 solution, the anode catalyst was a Ti mesh coated with Ir/RuO2, and the anodic reaction was the oxygen evolution reaction 200 mg of Vulcan carbon and disintegrated. With these quantities, we aimed at producing 30 wt % Sn/C catalysts. We applied a current of 0.5, 0.7, 1, and 1.2 A while continuously stirring the solution until the wire was completely disintegrated, thus opening the circuit. When the disintegration was over, the suspensions were simply filtered with a pressure filter and a hydrophilic PTFE membrane with a pore size of 100 nm and rinsed with MilliQ ultrapure water.

For the benchmark catalyst, we used a previously described chemical reduction method for the production of Sn/C nanoparticles that were used on GDEs in similar conditions.9 We dissolved 500 mg of SnCl2·H2O and 150 mg of poly(vinylpyrrolidone) (PVP) in 100 mL of ethylene glycol (EG). After the dissolution, we added 600 mg of Vulcan carbon (to obtain 30 wt % Sn/C) and used alternated stirring and sonication to obtain a homogeneous dispersion. At this point, 500 mg of NaBH4 was added directly to the solution, which was allowed to react for 30 min. At the end of the reaction, the catalyst was centrifuged and washed with acetone at least three times before drying overnight.

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The catalyst-coated gas diffusion electrodes (GDEs) were prepared by airbrushing the catalytic ink on a strip of the bare GDE prepared by following a previously described method.1 The metal loading was calculated by the difference in the weight of the electrode before and after airbrushing and assuming that the ratio of a metal to an ionomer was retained during the airbrushing process. The results are shown in Table 1.
The catholyte was a 0.5 M KHCO₃ solution and the cathode was a Sn, Pb, Sn50Pb50, or Sn/C airbrushed GDE. A modified frame in the cathodic compartment allowed for the insertion of a leak-free Ag/AgCl reference electrode close to the surface of the GDE. The electrolyte solutions were circulated in the compartments at a flow rate of 50 mL/min with a peristaltic pump, and CO₂ was fed through the GDE in the cathodic compartment at a flow rate of 50 mL/min.

For the testing of the unsupported catalysts, the flow cell was connected to an Autolab PGSTAT100N potentiostat (with Nova software version 1.11) equipped with a 10 A current booster (Autolab BOOSTER10A). The cell was run controlling the potential of the cathode sequentially at −2.44, −1.94, and −2.94 V vs SHE for 1 h.

For the testing of the Sn/C catalysts (synthesized with cathodic disintegration as well as the benchmark), the same apparatus was used but the cell was simply connected to a power supply and operated at 200 mA/cm² for 4 h.

The samples of the catholyte were collected and analyzed for soluble products with a Metrohm 930 Compact IC Flex ion chromatograph equipped with a Metrohm Metrosep A Supp 7 4 × 250 mm column and a conductivity detector.

## RESULTS

Disintegration Screening for Catalyst Production. A voltage of up to 20.5 V was applied between a graphite rod and Sn or Pb electrodes in different electrolytes to screen whether or not the disintegration happened and to assess the stability of the particles in suspension. The results are shown in Table 2.

As can be seen from the table, the reaction happens in the majority of the electrolytes tested: it is possible to disintegrate Sn and Pb cathodes to a suspension of metallic particles in 1 and 0.1 M NaOH, 0.3 M trisodium citrate, 0.3 M Na₂SO₄ and 0.3 M NaH₂PO₄ but not in 0.5 M H₂SO₄ and 1 M NaNO₃. The resulting suspension, with the exception of 0.3 M Na₂SO₄, starts to flocculate and bigger particles appear and deposit on the bottom of the container.

Due to the stability of the particles in Na₂SO₄, this electrolyte was chosen for the subsequent synthesis of unsupported Sn, Pb, and Sn50Pb50 particles.

In the case of the carbon-supported Sn particles, a 0.1 M NaOH solution was used since the disintegration proceeds more quickly than in Na₂SO₄. Also, since Sn is expected to be in situ supported on the carbon present in the electrolyte, the requirement of keeping the suspension stable to prevent the agglomeration of the metallic particles is less stringent.

Characterization of Disintegration Products. Figures 1a–c shows a comparison between the XRD patterns of the metal foils used for the disintegration and the unsupported particles obtained from them. Figure 1d shows the results of the EDX analysis performed on the Sn50Pb50 foil and particles to compare the elemental compositions. The foils of pure Sn and Pb are characterized by a distribution of crystallite orientations, as shown in Figure 1a,b, indicating their polycrystalline nature. The particles are at least partly crystalline (we cannot exclude the presence of an amorphous phase) and the peaks do not show significant broadening, suggesting that the crystallite size is rather big. The Sn50Pb50 alloy, according to the phase diagram, should result in a hypoeutectic composition, leading to the formation of a lamellar structure of an α phase (Pb-rich solid solution) and a β phase (Sn-rich solid solution) with inclusions of α phase particles. Peaks from Sn and Pb are visible in the diffractogram of both foil and particles (Figure 1c). The peaks are shifted with respect to their position in the pure metals, indicating that Sn and Pb form solid solution phases, as expected.

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**Table 2. Summary of the Cathodic Disintegration Behavior in Different Electrolytes for Pb and Sn Foils**

<table>
<thead>
<tr>
<th>Electrolyte</th>
<th>Behavior of Cathode</th>
<th>Behavior of Suspension</th>
</tr>
</thead>
<tbody>
<tr>
<td>1 M NaOH</td>
<td>Disintegrates</td>
<td>Agglomerates</td>
</tr>
<tr>
<td>0.1 M NaOH</td>
<td>Disintegrates</td>
<td>Agglomerates</td>
</tr>
<tr>
<td>0.3 M Na₃citrate</td>
<td>Disintegrates</td>
<td>Agglomerates</td>
</tr>
<tr>
<td>0.3 M Na₂SO₄</td>
<td>Disintegrates</td>
<td>Stable</td>
</tr>
<tr>
<td>0.3 M NaH₂PO₄</td>
<td>Disintegrates</td>
<td>Agglomerates</td>
</tr>
<tr>
<td>1 M NaNO₃</td>
<td>No disintegration</td>
<td></td>
</tr>
<tr>
<td>0.5 M H₂SO₄</td>
<td>No disintegration</td>
<td></td>
</tr>
</tbody>
</table>

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![Figure 1. Comparison of the XRD foil (black) and particles (red) for (a) Sn, (b) Pb, and (c) Sn50Pb50. (d) Comparison of at. % of Sn and Pb from EDX in the Sn50Pb50 (Sn64Pb36 in at. %) foil (black) and particles (red).](https://dx.doi.org/10.1021/acssuschemeng.0c04875)
particles of Sn50Pb50 show a small shift compared to the parent foil, suggesting that some changes in the composition of the Sn–Pb solid solutions might occur during disintegration. The EDX analysis of the Sn50Pb50 metal foil and the corresponding particles shows that the values of atomic composition are mostly retained locally in the alloy particles. The morphology and size of the unsupported particles were investigated by SEM. As the particles were used for CO2 electrochemical reduction on a GDE-type electrode, the images were taken on the GDE itself and also used to analyze the catalyst distribution on the electrode surface. Some of the SEM pictures taken on the GDEs catalyzed with the unsupported Sn, Pb, and Sn50Pb50 particles are shown in Figure S1. The particle size and shape distribution are very broad, ranging from about 100 nm to micrometers (at least the observed ones). This is probably caused by partial agglomeration of the particles due to the absence of carbon supports or stabilizers. Even though the particles remain in suspension for a longer time in Na2SO4 (compared to the other electrolytes tested, in which the agglomeration is visible minutes after the end of the disintegration reaction), it does not mean that agglomeration can be entirely prevented. The bigger particles could result from a concomitant mechanical disintegration due to the partial dissolution of the electrode and intense gas evolution.

Interestingly, in the case of carbon-supported Sn particles produced with the disintegration method, we find a correlation between the current applied for the synthesis and the particle size, which in turn influences the efficiency of the support of the Sn particles on the Vulcan carbon. Figure 2 shows SEM-EDX mapping of the catalysts synthesized at the different currents, all at the same magnification and compared to the catalyst synthesized by chemical reduction. At the lowest current (0.5 A), we observe big pieces of Sn (order of 5 μm) on the carbon support, which are obviously too big to be supported. Moreover, there is a very inhomogeneous and sporadic distribution of Sn on the carbon. The situation becomes better at 0.7 A, but only on increasing the current to 1 A and especially to 1.2 A is the Sn properly distributed on the carbon support in the form of nanoparticles. As we will see later, this impacts the performance of the catalysts. Moreover, the EDX analysis (Figure S2) reveals that the Sn loading on
the Sn/C catalysts increases with increasing applied current, probably due to the inefficient support of the bigger particles produced at lower currents. These loose particles may be lost during the cleaning procedure. Sn/C 0.5 A, Sn/C 0.7 A, and Sn/C 1 A have Sn loadings lower than the expected 30 wt % (19.6, 19.4, and 26.5%, respectively), while Sn/C 1.2 A and Sn/C chem. red. have a loading slightly higher than the expected value (38.7 and 41.1%, respectively), which may be due to the loss of some of the carbon before particle deposition. Sn/C 1.2 A and Sn/C chem. red. have similar Sn loadings. In any case, since the loading is similar to the expected one, the results of the EDX analysis suggest that the Sn/C 1.2 A and Sn/C chem. red. have a loading slightly higher than the expected value (38.7 and 41.1%, respectively), which may be due to the loss of some of the carbon before particle deposition. Sn/C 1.2 A and Sn/C chem. red. have similar Sn loadings. In any case, since the loading is similar to the expected one, the results of the EDX analysis suggest that the major percentage of Sn is effectively transformed into particles and there is not a significant loss of Sn as soluble species. In the TEM images, it is difficult to find supported Sn particles in the catalysts synthesized at the lower currents, and in the case of the catalyst synthesized at 0.5 A, it is possible to see big pieces of Sn, thick enough to block the electron beam. In the case of the Sn/C 1.2 A, we can see clusters of nanoparticles on the carbon support (Figure 3). On the catalyst produced by chemical reduction, we observe the most homogeneous distribution of Sn on the surface, which is expected as the metal is distributed on and in the pores of the carbon support in the form of soluble Sn2+ before being reduced (Figure 2).

**Catalytic Activity for CO2 Reduction.** Figure 4 shows the geometrical current density of the three unsupported catalysts at three different potentials and the faradic yield (FY) toward formate. Since it is known from the previous literature5,11 that Sn, Pb, and Sn–Pb alloys produce formate as a main CO2 reduction product, we did not sample the gas outlet of our electrochemical cell. It is safe to assume that the rest of the current is employed for the evolution of hydrogen and possibly small amounts of CO.

The catalyst reaching the highest partial current toward formate is Sn at a potential of −2.44 V vs SHE with a partial current density of 144 mA/cm² and FY of about 61%. The most selective catalyst is the Sn50Pb50 alloy at −2.44 V vs SHE with a FY of 64%, however, the total current density is significantly lower for this catalyst, leading to a lower partial current density toward formate of 86 mA/cm².

For the carbon-supported Sn catalysts produced by cathodic disintegration (results in Figure 5), we can see that the average FY over 4 h is correlated with the quality of the catalyst distribution on the support. For the Sn/C 0.5 A and the Sn/C 0.7 A catalysts, the FY is, respectively, 72 and 63%, and this is probably due to the fact that the majority of the Sn is not properly supported but exists in the form of micrometric particles, loosely held in the carbon by the binder. We cannot exclude that a fraction of these particles sediments out of the ink during airbrushing. For Sn/C 1 A, we have a higher FY of 83%, while the Sn/C 1.2 A, which is the one that shows the best catalyst distribution in the EDX mappings, is the best one with a FY of 98%. The gray column shows the comparison with the benchmark Sn/C catalyst synthesized by chemical reduction. Sn/C 1 A and Sn/C 1.2 A outperform the benchmark prepared by chemical reduction regarding the selectivity toward formate.

**DISCUSSION**

**Particle Formation Mechanism.** As previously discussed in the Results section, it is possible to disintegrate Sn, Pb, and Sn50Pb50 wires in a variety of electrolytes. This phenomenon was first described as cathodic disintegration of Sn and Pb in the 1950s.6,7 In those papers, the authors explained the process as formation and immediate decomposition of volatile Sn and Pb hydrides upon cathodic polarization.

A more recent explanation for the phenomenon involves the generation of Zintl phases. These are anions based on negatively charged clusters of post transition metals and are known to be extremely unstable.12 The electrochemical formation of these species is documented for Sn and Pb in
an anhydrous aprotic organic solvent, in which they can be kept stable for some time by choosing the proper counter ion and keeping strictly anhydrous, anaerobic conditions. Subsequent addition of water to the solvent in which the Zintls are synthesized and stabilized causes the oxidation of the Zintl phase to particles of metallic Sn or Pb.13 In our case, the formation and decomposition of the Zintl phase (some form of sodium stannide or plumbide) could be happening in a rapid sequence due to the presence of water and lead to the observed disintegration.

A similar process is observed on transition metals and is referred to as cathodic corrosion,14 which has been suggested to have (ternary metal) hydrides as intermediates.15 It has been shown that, for this process, it is possible to have some control over particle size distribution and orientation.16,17 In our case, we know that below a certain threshold current, the particles produced are not supported properly on the carbon because of the formation of large particles in the micrometer range, but we still do not know how variations in the current density above that threshold can create properly supported particles of different sizes and how different electrolytes can influence the process. Most interestingly, there is abundant evidence of the possibility to reliably retain alloy composition of the parent wire in the nanoparticles in the case of cathodic corrosion with an alternating current.18,19 For now, we have evidence of the possibility to retain alloy composition in a single Sn–Pb alloy but more Sn–Pb alloys and Sn and Pb alloys with other elements can be screened, potentially yielding more selective or stable catalysts.

At this stage, it is still not entirely clear whether cathodic corrosion (transition metals) and cathodic disintegration (Sn and Pb), although similar, are the same phenomenon since they show a different rate dependence on electrolyte concentration and the disintegration of Sn and Pb happens easily in dilute electrolytes with direct currents, while the corrosion of transition metals requires either very concentrated electrolytes with direct current or less concentrated electrolytes with an alternating current.6,7,14–18

**Control over Particle Size for Supported Sn Particles.** Sn particles of decreasing size can be obtained by increasing the current density applied. This can be observed by comparing the EDX elemental mapping of the carbon-supported catalysts (Figure 2). The reasons for this behavior are not clear, but a possible explanation is that at the lower current densities, the corrosion process is slower and localized in areas of higher mechanical stress (for example, grain boundaries), causing a slow production of nanoparticles and a weakening of the structure, leading to the detachment of the observed larger particles for mechanical instabilities. At higher current densities, instead, the corrosion could happen more homogeneously on the surface, producing significant amounts of nanoparticles small enough to be supported on the carbon (the particle size of Vulcan carbon is between 30 and 60 nm). Some evidence for this behavior is provided by the fact that on disintegrating Sn wires in 0.1 M NaOH at different current densities but passing the same amount of charge, the weight loss increases with increasing current density (see Table S1). This suggests that the process becomes more efficient at higher current densities. In contrast to Sn, we note that for Pt, cathodic corrosion at higher currents leads to larger particles.16

**Catalytic Performance.** The values for the FY of the unsupported pure metal catalysts are in line with the previous literature using similar conditions.11 The significant agglomer-
CONCLUSIONS

We have shown that it is possible to generate catalytically active metallic powders based on Sn and Pb starting from a metallic plate or a wire by cathodic polarization in an appropriate electrolyte. It is also possible to produce a catalyst based on an alloy of Sn and Pb, retaining the composition of the parent wire. The process goes through the disintegration of the cathode and is extremely quick and clean, giving stable suspensions just by carefully choosing the nature and concentration of the electrolyte. This avoids the use of organic solvents, additives, stabilizers, and reducing agents.

The catalysts produced are selective for CO₂ reduction to formate at high current densities in a 10 cm² GDE-type flow cell, with the best performing catalyst being Sn, reaching maximum partial current densities toward formate of 144 mA/cm² with a faradic yield of 61%.

The catalysts can significantly improve their performance by being properly supported on carbon. The best performing Sn/C catalyst produced with disintegration shows a faradic yield of 98% at 200 mA/cm², showing promise to be used in industrial electrolyzers.

We are confident that careful optimization of our method of electrochemical particle production can yield catalysts with very good performances that can be produced easily, quickly, and in a clean way. Our method holds promise to be easily scalable to produce large quantities of the catalyst needed for the industrialization of the process of electrochemical conversion of CO₂ to formate as well as other (electro-)chemical conversions. We believe that the electrochemically driven production of nanoparticles, both from Sn and Pb as reported here as well as from transition metals as reported in the literature, is an environmentally friendly and promising alternative to conventional catalyst synthesis. The extension of the cathodic disintegration method to other p-block metals remains to be thoroughly investigated. Preliminary results suggest that a similar macroscopically observable process occurs for Bi but no visible disintegration is observed for In, at least in the range of conditions tested. In this case, it may be interesting to investigate the behavior of these metals under conditions similar to the ones that allow the cathodic corrosion of transition metals.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at https://pubs.acs.org/doi/10.1021/acssuschemeng.0c04875.

SEM pictures of unsupported catalysts on GDE, Sn loadings on C estimated with EDX, the weight loss of a Sn wire as a function of the applied current, the summary of synthesis conditions for Sn/C catalysts, and the calculation of atom efficiency for the synthesis of Sn/C catalysts (PDF)

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Notes

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