Hydrogenase Mimics in M12L24 Nanospheres to Control Overpotential and Activity in Proton-Reduction Catalysis

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Hydrogenases are fascinating metalloenzymes that can reversibly convert protons into molecular hydrogen at high rates with virtually no overpotential.[1] This reversible interconversion is of great interest in view of the transition from our current fossil fuel based society to one that is powered by renewable energy sources. As such, hydrogenase enzymes provide a powerful blueprint for the development of catalysts inspired by nature.[2] Intensive studies on the iron-iron hydrogenases showed their detailed operational mechanism and the key features that render these enzymes superb catalysts,[3] revealing an important function for the internal proton relay,[4] that is, the amine moiety in the azadithiolate bridge, and for the Fe₄S₄ cluster ligated to the proximal iron of the H-cluster, which functions as electron reservoir. In parallel, many groups around the world made synthetic analogues of the active site at which the actual proton reduction takes place.[5] Installment of proton relay moieties has been successfully achieved and demonstrated to improve the catalytic function of synthetic models.[6] Less attention has been given to the redox-active Fe₄S₄ cluster, nevertheless recent work on synthetic models with appended electron reservoirs demonstrated that such function also improves the catalytic properties.[7] Interestingly, despite all efforts, up to now, there are no synthetic mimics that can perform the proton reduction reaction at low overpotential. This suggests that the protein environment, that is, the second coordination sphere around the active site, may play a more important role than initially anticipated.[8] Recent experiments, in which synthetic mimics of the active site are installed in the inactive apo-hydrogenase enzyme, show full competence enzymatic activity, hinting at the importance of the protein matrix or second coordination sphere around the H-cluster.[9] Introduction of a synthetic second coordination sphere around the hydrogenase mimics has been attempted using diverse strategies, chief among which liposomes,[10] micelles,[11] cyclodextrins,[12] peptidic scaffolds[13] and polymeric matrices[14] yet information on catalysts activity and their overpotential remains rare.

Supramolecular cages represent an alternative strategy to control the second coordination sphere, and have proven successful in inducing enhanced activity and selectivity to the encapsulated catalysts.[15] We recently showed that encapsulation of a single hydrogenase mimic into a tight supramolecular cage effectively provides a second coordination environment[16] and this strategy resulted in lower catalytic overpotentials.[17] In this work we report a self-assembly strategy to install multiple mimics of the hydrogenase active site into very spacious M₁₂L₂₄ nano-spheres that contain ammonium salts as functional groups. We show that encapsulation of ammonium salts, effectively providing substrate preorganization and intermediates stabilization. We show that catalyst encapsulation in these nanocages reduces the catalytic overpotential for proton reduction by 250 mV as compared to the uncaged catalyst, while the proton-rich nanosphere environment created around the catalyst ensures that high catalytic rates are maintained.

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previously reported caged catalyst yet maintaining very similar catalytic rates.\textsuperscript{[17]}

In order to create a nano-environment able to effectively preorganize protons around di-iron hydrogenase models, modified M\textsubscript{x}L\textsubscript{1}\textsubscript{y}L\textsubscript{z} \textsubscript{34−} Fujita-type cages are employed. Such cages provide sufficiently large space within their cavity that can be easily decorated with various customized functional groups,\textsuperscript{[18a]} that is, catalyst and acidic functions. The di-iron catalyst functionalized building block Fe\textsubscript{2}BB features a short aliphatic linker that connects the ditopic bis(pyridyl) structure to a monocarboxylic acid benzenedithiolate di-iron complex through an amide bond as shown in Figure 1. Ammonium groups are installed on different building blocks, BBNH\textsubscript{+}, in order to provide a proton rich local environment. This sython features a short aliphatic chain terminated with a tertiary amine moiety. This base is about five orders of magnitude stronger than the pyridyl moieties present on the building block itself,\textsuperscript{[19]} allowing its selective protonation by just the addition of sub-stoichiometric amounts of pyridinium hexafluorophosphate as confirmed by single crystal X-ray diffraction (Figure S3). A third building block BB that doesn’t contain functional groups at the endohedral position of the ditopic bis(pyridyl) building block is also prepared as shown in Figure 1.

Cages were prepared using self-assembly strategies as reported previously.\textsuperscript{[18b,120]} Stirring a 5:19 ratio of Fe\textsubscript{2}BB and BBNH\textsubscript{+} (or HNEt\textsubscript{3}PF\textsubscript{6}) in the presence of a palladium source in MeCN at 60°C overnight provided the two respective cages [Pd\textsubscript{x}(Fe\textsubscript{2}BB)\textsubscript{y}(BBNH\textsubscript{+})\textsubscript{z}]\textsubscript{43−} and [Pd\textsubscript{x}(Fe\textsubscript{2}BB)\textsubscript{y}(BB)]\textsubscript{30−}.\textsuperscript{[21]}

The formation of the large nano-cages is confirmed by \textsuperscript{1}H-NMR, DOSY and HR-CSI-MS analysis. The \textsuperscript{1}H-NMR spectrum (Figure 2) shows a typical downfield shift of the pyridyl protons upon coordination of the caged catalyst. The di-iron catalyst is in a proton-rich nano-environment where the positively charged acid and the negatively charged \zeta (details are found in SI). These experiments show that we can prepare systems in which the hydrogenase model is effectively encapsulated into a nano-confined space with multiple mimics in one cage. For cage [Pd\textsubscript{x}(Fe\textsubscript{2}BB)\textsubscript{y}(BBNH\textsubscript{+})]\textsubscript{43−} containing acidic protons, the di-iron catalyst is in a proton-rich nano-environment where substrates are effectively preorganized around the catalyst within the cavity defined by the cage structure.

Next the electrochemical proton reduction catalysis displayed by the caged catalysts was evaluated.\textsuperscript{[20]} For the cage [Pd\textsubscript{x}(Fe\textsubscript{2}BB)\textsubscript{y}(BB)]\textsubscript{30−}, where BB is the standard ditopic bis(pyridine) cage building block featuring a non-acidic aromatic proton at the endo position, the voltammograms show a reduction event around −1.3 V vs. Fe\textsubscript{3+}/Fe\textsubscript{2+} consistent with the reduction of the di-iron moiety (Figure S24). Formation of the reduced catalyst is supported by IR-spectroelectrochemical measurements showing its clear signature in the carbonyl region, indication that the caged di-iron catalyst is stable under electrochemical conditions (Figure S35). Sequential additions of external weak acid, HNEt\textsubscript{3}PF\textsubscript{6}, not sufficiently strong to protonate the pyridyl groups or the non-reduced iron-iron bond, causes the appearance of a new peak at −1.7 V vs. Fe\textsubscript{3+}/Fe\textsubscript{2+} (Figure S25).

This peak increases in intensity with the amount of acid added, in line with proton reduction catalysis at this potential. The modest increase in peak current intensity suggests that this catalytic process is rather slow. The external acid is able to diffuse into the cage cavity, but this may be relatively slow due to electrostatic repulsion between the positively charged acid and the positively charged cage shell. Interestingly,
comparison of this catalytic peak potential to that obtained for the uncaged free $\text{Fe}_{2}\text{BB}$ reveals an anodic potential shift of about 230 mV towards more favorable potentials (Figure S26). The local environment around the catalyst lowers the overpotential for the catalytic proton reduction reaction probably due to stabilization of reduced reaction intermediates by the positive cage framework. So the positively charged cage results in more favorable overpotential for proton reduction catalysis and at the same time reduces the catalytic activity; $k_{\text{cat}}$ is estimated by foot of the wave analysis\(^{23}\) to be $7.42 \times 10^7 \text{ mol}^{-1}\text{s}^{-1}$ and TOF$_{\text{max}}$ calculated to be $44 \text{ s}^{-1}$ a decrease of two order of magnitude compared to the uncaged catalyst ($k_{\text{cat}} \text{Fe}_{2}\text{BB} 1.51 \times 10^5 \text{ mol}^{-1}\text{s}^{-1}$).

We hypothesized that creation of a local acidic environment as designed for cage $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3}\text{]^{33+}}$ featuring acidic quaternary ammonium groups, would not show slow diffusion rate limitation because of the pre-organization. When this cage is subjected to electrochemical analysis, the voltammogram displays a reduction event $-1.7$ vs. $\text{Fc}^{0/0}$ as shown in Figure 3 and Figure S30. Addition of increasing equivalents of external acid to this same cage solution reveals a current increase of the peak at $-1.7$ vs. $\text{Fc}^{0/0}$, in line with a proton reduction event. The catalytic rate constant observed for the proton preorganized-encapsulated di-iron catalyst is estimated to be in the order of $1.03 \times 1.0 \text{5 mol}^{-1}\text{s}^{-1}$ for the proton reduction event. The catalytic rate constant of magnitude faster than the nano-confined catalyst in $\text{MeCN}$ with 30 equivalents of external acid (red) and free catalyst $[\text{Fe}_4(\text{ZnL})_6][\text{Fe}_2(\text{F}_4\text{bdt})(\text{PPy}_3)(\text{CO}_5)]^{8}\text{]^{8}}$. The voltammograms obtained for the un-caged $\text{Fe}_{2}\text{BB}$ catalyst in the presence of acid and those obtained for the cage sample $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ reveals that the catalytic half wave potential ($E_{1/2\text{cat}}$) is shifted anodically by $250 \text{ mV}$ as shown in Figure 3. The cage effects are clear when plotting the properties for proton reduction catalysis in Tafel plots as shown in Figure 4. The encapsulation of the di-iron catalyst leads to a reduction of the overpotential, for both cages investigated to only about $350 \text{ mV}$, which is among the lowest overpotential reported for this class of hydrogenase mimics. This suggests that the effect is unrelated to proton preorganization of the acidic moieties within the cage cavity but rather a cage effect possibly due to stabilization of negatively charged reaction intermediates (Figure S36). Because of the smaller size, the active site is closer to the bulk solution allowing a more rapid reaction with substrates. As such, substrate preorganization for this system is not needed for fast rates as substrate diffusion is not limiting catalysis. In contrast, the current $\text{M}_{13}\text{L}_{13}$ nano-cage is much larger with a diameter of $5 \text{ nm}$ and a volume over 30 times bigger. It can accommodate several catalysts as shown by CSI-MS data and those are on average further away from the cage windows (Figure S36). As such, the two systems are rather different and so are the rates of substrate diffusion towards the caged catalyst. More detailed experiments are required to confirm these hypotheses.

Importantly, the voltammograms obtained for the uncaged $\text{Fe}_{2}\text{BB}$ catalyst in the presence of acid and those obtained for the cage sample $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ reveals that the catalytic half wave potential ($E_{1/2\text{cat}}$) is shifted anodically by $250 \text{ mV}$ as shown in Figure 3. The cage effects are clear when plotting the properties for proton reduction catalysis in Tafel plots as shown in Figure 4. The encapsulation of the di-iron catalyst leads to a reduction of the overpotential, for both cages investigated to only about $350 \text{ mV}$, which is among the lowest overpotential reported for this class of hydrogenase mimics. This suggests that the effect is unrelated to proton preorganization of the acidic moieties within the cage cavity but rather a cage effect possibly due to stabilization of negatively charged reaction intermediates by the positively charged cage framework. Instead, preorganiza-

![Figure 3. Cyclic voltammograms for cage $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ in $\text{MeCN}$ with 30 equivalents of external acid (red) and free $\text{Fe}_{2}\text{BB}$ in $\text{MeCN}$ with 10 equivalents of external acid (black). For proton preorganized cage system, proton reduction takes place at about $250 \text{ mV}$ milder potential. Scan speed $0.1 \text{ Vs}^{-1}$; $\text{Fe}_{2}\text{BB}$ 1 mm; cage $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ 0.02 mm thus $\text{Fe}_{2}\text{BB}$ 0.1 mm due to solubility reasons (see Figure S28).](Angew. Chem. Int. Ed. 2020, 59, 18485–18489 © 2020 The Authors. Published by Wiley-VCH GmbH www.angewandte.de 18487)

![Figure 4. Tafel plot for free $\text{Fe}_{2}\text{BB}$ (black), cage $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ (blue), cage $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ (red) and cage $[\text{Fe}_4(\text{ZnL})_6][\text{Fe}_2(\text{F}_4\text{bdt})(\text{PPy}_3)(\text{CO}_5)]^{8}\text{]^{8}}$ extrapolated at $30 \text{ mV}$ external acid concentration, showing that catalyst encapsulation in proton-rich environment leads to a drop in catalytic overpotential of $250 \text{ mV}$ with respect to free diffusion $\text{Fe}_{2}\text{BB}$ while increasing the turnover frequency by two orders of magnitude compared to catalyst encapsulation in proton-poor environment lacking substrate preorganization. Cage $[\text{Pd}_{12}(\text{Fe}_{2}\text{BB})(\text{BBNH}_4)^{3\text{]}^{33+}}$ catalyzed proton reduction at $290 \text{ mV}$ milder overpotential as compared to previously reported cage $[\text{Fe}_4(\text{ZnL})_6][\text{Fe}_2(\text{F}_4\text{bdt})(\text{PPy}_3)(\text{CO}_5)]^{8}\text{]^{8}}$ yet at similar rates.](Angew. Chem. Int. Ed. 2020, 59, 18485–18489 © 2020 The Authors. Published by Wiley-VCH GmbH www.angewandte.de 18487)
tion of proton substrates within the cavity of the supramolecular assembly has a beneficial effect as it allows for higher catalytic rates, stressing the importance of proton relays around the di-iron moiety.

The previously reported $[[\text{Fe}_2(\text{ZnL})_6][\text{Fe}_2(\text{F,bdt})(\text{PPy}_3)(\text{CO}_3)]^{8+}]$ system was based on a ligand template approach to encapsulation, which requires the catalyst to have a coordinated phosphine ligand, whereas the current system has an hexacarboxyl di-iron derivative. Such a coordinated phosphine ligand increases the electron density at the di-iron core and this typically results in faster catalytic rates but at the expense of higher overpotentials. As shown in Figure 4, $[[\text{Fe}_2(\text{ZnL})_6][\text{Fe}_2(\text{F,bdt})(\text{PPy}_3)(\text{CO}_3)]^{8+}]$ features the highest rate but also the largest overpotential while the combination of the more electron deficient hexacarboxyl catalyst and cage effect enables catalysis at 350 mV overpotential which represents a 290 mV reduction of catalytic overpotential as compared to the previously reported system. At the same time, substrate preorganization provided by the modified M$_{12}$L$_{24}$ cage allows for maintaining high proton reduction rates and in fact very comparable to those obtained by the electron richer monophosphine catalyst encapsulated in the smaller cage.

The strategy presented in this work allows to create a special environment around synthetic hydrogenase mimics, leading to improved performance in electrocatalytic proton reduction catalysis. The M$_{12}$L$_{24}$ cages provided a flexible platform to achieve a better understanding of second coordination sphere effects in catalysis and clear insights for future developments. Whereas we here demonstrate the effect of the cage and local concentration of protons (substrate), further modification to closely mimic the essential amino acid residues found around the structure of the natural H-cluster may be possible. Introduction of synthetic mimics into nano-environments such as the cavity of preferably precious-metal-free supramolecular cages decorated with such residues could further lower the overpotential of synthetic models, finally approaching enzymatic rates and efficiencies, a strategy that is currently explored in our laboratories.

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**Conflict of interest**

The authors declare no conflict of interest.

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