Redox-Mediated Alcohol Oxidation Coupled to Hydrogen Gas Formation in a Dye-Sensitized Photosynthesis Cell

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Solar energy is an attractive $\text{CO}_2$-neutral energy source, providing $\approx 4000$ times the energy consumption estimated for 2035.\[^{[1, 2]}\] Beyond widely deployed silicon-based photovoltaic (PV) technology, dye-sensitized solar cells (DSSCs)\[^{[3]}\] present a low-cost alternative with improved performance in low/diffuse light conditions. DSSCs use molecular dyes, wide-band gap semiconductors and redox mediators to absorb light and separate charges affording efficiencies that have surpassed 14%\[^{[4, 5]}\]. While the application of solar-to-electric energy conversion is increasing, long-term energy storage remains a challenge. Electricty generated from PV cells can be stored indirectly into Li-ion batteries for short-term storage, or PV cells can be attached to an electrolyzer to split water and obtain $\text{H}_2$ as a storable fuel for long-term storage.\[^{[6]}\] Substitution of redox mediators in a DSSC with catalysts to furnish a dye-sensitized photoelectrochemical cell (DSPEC) to realize oxidative and reductive transformations powered by light convert light energy directly into fuels providing an alternative strategy for a solar-to-fuel device.\[^{[7-10]}\] An exciting advantage for DSPECs is the opportunity to optimize the device by molecular-level design to improve light absorption, charge separation and attenuate catalytic events.\[^{[11, 12]}\] Most examples of these devices focus on the light-driven water splitting to form $\text{O}_2$ and $\text{H}_2$,\[^{[13]}\] however, the solar-to-hydrogen efficiencies are still relatively low (up to 1.5%).\[^{[14]}\] The efficiency is partly due to the high kinetic barrier of the oxygen evolution reaction (OER),\[^{[15]}\] potentially leading to charge recombination. Parallel to solar fuel related chemistry lies the field of organic photoredox catalysis, a rapidly advancing topic where light is used to synthesize valuable chemicals.\[^{[16]}\] These fields can be elegantly combined by replacing the difficult OER to improve the performance of the DSPEC.\[^{[17]}\] The OER can be substituted for an alternative oxidative reaction to generate bulk chemicals with higher market values than $\text{H}_2$, while photogenerating a storable fuel (i.e. $\text{H}_2$).\[^{[18]}\] Initial work along these lines show that the oxidation of 5-hydroxymethylfurfural and $\text{C}_2-\text{H}_4$ activation on BiVO$_4$ can be achieved in such devices.\[^{[19, 20]}\] Recently, Meyer et al., reported the oxidation of benzyl alcohol (PhCH$_2$OH) to benzaldehyde (PhCHO) using a light-driven, surface-bound molecular catalyst in a DSPEC device, with reported photocurrents of 12 and 35 $\mu$A$\,\text{cm}^{-2}$.\[^{[21]}\] In this communication, we report the development of a DSPEC with 2,2,6,6-tetramethylpiperidine 1-oxyl (TEMPO) as a redox mediator in solution for the light-driven oxidation of PhCH$_2$OH. We employed TEMPO$^{\bullet}$-couple in solution as a redox mediator as 1) it has been successfully utilized as a redox mediator in DSSCs, indicating TEMPO$^{\bullet}$ production in situ\[^{[22-26]}\] and 2) having the oxidation reaction in solution instead of at the surface bound species may lead to less charge recombination and therefore higher photocurrents. 3) TEMPO$^{\bullet}$ is used as a sacrificial oxidant for many relevant oxidation reactions in organic synthesis so the device could be used for different reactions.\[^{[27-31]}\] To properly study such DSPEC we have designed a new reactor that allows in situ monitoring of reaction progress.

We started with initial experiments to demonstrate the viability of light-driven oxidation of PhCH$_2$OH in a DSPEC using...
TEMPO as electron mediator to photogenerate TEMPO$^+$ as the oxidant (Figure 1). For the photooxidation of TEMPO, thienopyrroledione-based (AP11) was used as high-potential photo-sensitizer, with an oxidative power of 1.8 V upon illumination.$^{[32]}$ Although DSSC based on AP11 have not resulted in records in power conversion efficiency (PCE) it was selected because oxidation potential, providing room for future high potential redox mediators.

The initial DSPEC consisted of FTO|TiO$_2$|AP11 working electrode (WE), a ftted Pt-wire counter electrode (CE) in a glass three-electrode cell using a household lamp as illumination source (Figure S12) with the results summarized in Table 1. PhCHO generation was monitored by ATR-FT-IR (Figure S5) and photocurrent was recorded for 16 h (Figure S6). The photocurrent density ($J_{OC}$) of 0.1 mA cm$^{-2}$ and PhCHO conversion are consistent with the overall light-driven oxidation proceeding with a Faradaic efficiency of 74%. A bias-free environment is desirable for industrial applications and therefore, the light-driven reaction was repeated in a two electrode system by connecting the WE directly to the CE. The system still produced PhCHO albeit with slightly lower efficiency (35%). Control experiments in absence of one of the components, (TEMPO, light or dye) afforded negligible photocurrents (Figure S6) and also PhCHO was not produced under these conditions. These experiments demonstrate that only the full system produces photocurrent and PhCHO, indicating that the photooxidized dye is required to generate TEMPO$^+$, which acts as the in situ photogenerated chemical oxidant for oxidation of benzyl alcohol.

Next, we sought to optimize the photoreactor to yield maximum Faradic efficiencies and $J_{OC}$. We developed a gas-tight Teflon photoreactor with WE and CE compartments featuring inlets/outlets to enable on line monitoring of the chemical conversion by ATR-FT-IR and headspace analysis by gas chromatography (GC) (Figure 2A). We assembled different sized TEMPO$^+$-AP11- DSSCs to investigate possible losses in photocurrent ($J_{OC}$) due to the change in dimensions (surface area, and distance between the electrodes) of the system. The DSSC system applied in the photoreactor (0.64 cm$^2$, Figure 2A) and in a typical sandwich DSSCs (0.19 cm$^2$, Figure 2B) were measured under 1 sun (AM1.5G, 100 mW cm$^{-2}$) to allow literature comparison$^{[32, 33]}$ and with DSPEC irradiation sources: 50 mW cm$^{-2}$ LED lamp and a household lamp (details and spectra of lamps given in Figure S12). The results and $J$-V curves ($N=3$ for both typical sandwich and photoreactor sized DSSC) are shown in Figure S4 and Table S1. The sandwich DSSCs measured under 1 sun revealed open-circuit voltage ($V_{OC}$) of 0.74 V and $J_{OC}$ of 6.12 mA cm$^{-2}$. The AP11 sandwich cells with the TEMPO$^+$ electrolyte show enhancements to the $J_{OC}$ when compared to the reported AP11-DSSCs using an 1 /I$_3^-$ electrolyte.$^{[32]}$ The $J_{OC}$ and $V_{OC}$ are comparable to reported TEMPO$^+$-DSSCs with different dyes, indicating efficient dye regeneration by the TEMPO$^+$ redox mediator.$^{[30]}

To investigate the photo-performance of the system in the DSPEC set-up, the sandwich DSSC was upscaled to size of the photoreactor size (Figure 2A, 2B), and a LED lamp was used in these comparative photo-oxidation experiment (Figure 2C). Only a small reduction of 10% in $J_{OC}$ is observed in going from the typical sandwich cell to the larger reactor, indicating only small losses due to diffusion issues. From this we established that TEMPO$^+$ is efficiently generated by the dye in the larger set-up, which in principle could be used for the chemical oxidation reaction.

The substrate-to-product oxidation reaction in our DSPEC reactor was independently studied by using TEMPO(BF$_4$)$^{[34]}$ as a chemical oxidant for PhCHO in order to quantify the formation of PhCHO over time. TEMPO(BF$_4$) mimics the role of in situ

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**Table 1.** Control experiments to demonstrate light-driven benzyl alcohol oxidation in a DSPEC using TEMPO as electron mediator were performed in a three-electrode cell. WE: FTO|TiO$_2$|AP11 WE (0.79 cm$^2$) in 5 mL MeCN, 1.2 M LiTFSI, 1.0 M TEMPO, 0.1 M PhCH$_2$OH, CE: Pt-wire in 0.5 mL MeCN, 1.2 M LiTFSI and 1.0 M Ac$_2$O, separated with glass frit, 0 V vs. Ag/AgCl reference electrode (RE) and illuminated with a household LED lamp for 16 hours).

<table>
<thead>
<tr>
<th>Entry</th>
<th>Conditions</th>
<th>PhCHO produced ($\mu$mol)</th>
<th>Electrons ($\mu$mol/2)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>without TEMPO</td>
<td>0.4</td>
<td>1.7</td>
</tr>
<tr>
<td>2</td>
<td>without AP11</td>
<td>0</td>
<td>&lt;0</td>
</tr>
<tr>
<td>3</td>
<td>without cell</td>
<td>0</td>
<td>&lt;0</td>
</tr>
<tr>
<td>4</td>
<td>without light</td>
<td>0</td>
<td>&lt;0</td>
</tr>
<tr>
<td>5</td>
<td>without bias</td>
<td>13</td>
<td>–</td>
</tr>
<tr>
<td>6</td>
<td>total system</td>
<td>20</td>
<td>27.2</td>
</tr>
</tbody>
</table>

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**Figure 1.** Schematic representation of the TEMPO mediated DSPEC system for light-driven alcohol oxidation and reductive H$_2$ production. D = dye, D$^-$ = oxidized dye, D$^+$ = excited dye, CE = conductance band TiO$_2$ (anatase), with important redox potentials provided.$^{[32]}$

**Figure 2.** A) Schematic Figure of A) photo-reactor (area TiO$_2$, 0.64 cm$^2$, 6 mL) B) sandwich DSSC (area TiO$_2$, 0.19 cm$^2$, 4.71 $\mu$L) Chopped light of sandwich (blue) and reactor size (red) DSSCs (WE: FTO|TiO$_2$|AP11, CE: FTO|Pt, electrolyte: 1.2 M LiTFSI, 1.0 M TEMPO, 0.1 M TEMPO(BF$_4$) in MeCN) illuminated under 50 mW cm$^{-2}$ LED lamp.
photogenerated TEMPO$^+$ in the photosynthesis cell. Therefore, PhCH$_2$OH (1 equiv.) was added to the electrolyte solution containing TEMPO(BF$_4$) (2 equiv.) as the chemical oxidant. On line monitoring of chemically-driven PhCH$_2$OH oxidation by TEMPO(BF$_4$) in the photoanode compartment of the DSPEC was performed over 2 h enabled by the use of ATR-FT-IR (Figure 3A). At 16 hours $t = 0$ hours, the characteristic C-O stretch (1008 cm$^{-1}$) from PhCH$_2$OH is dominant and decreased over time, with concomitant growth of the C=O stretch from PhCHO (1703 cm$^{-1}$) over 2 hours. Further confirmation reaction progress was through GC analysis which was fully in line with the data obtained from ATR-FT-IR.

After all individual steps of the TEMPO mediated DSPEC system in the reactor had been confirmed we performed bulk photoelectrolysis experiments in the newly designed photoelectrode in order to demonstrate the simultaneous photo-generation of H$_2$ at the cathode and the oxidative organic transformation at the anode. The photocurrents and PhCHO production were monitored ($N = 3$) over 32 hours (Figure 3B). After an induction period of 2 hours, the photocurrent showed an average of 0.4 mA cm$^{-2}$ (Figure S10), which is a $\approx$ 4-fold increase compared to our initial DSPEC experiments carried out in the three-electrode cell (Table 1). In the first 2 hours the system needs to equilibrate as benzyl alcohol oxidation leads to formation protons, which, after diffusion through the membrane to the other part of the cell, are converted to hydrogen. The proton concentration has an effect on the catalytic rates and is also known to affect the potential of the conduction band of TiO$_2$ as seen in water splitting DSPECs.$^{[30]}$ Consistent with the preceding experiments, photocurrents and PhCHO production overlap and indicate an overall Faradaic efficiency approaching 100% when carried out in the designed photoelectrode (Figure 3B, Figure S9, Table S2). Sampling of the CE compartment headspace by on line GC confirmed H$_2$ as the only reduction product (Figure S7), in line with a fully operative DSPEC. Most importantly, the photocurrent of 0.4 mA cm$^{-2}$ does not degrade over the course of the experiment, demonstrating the stability of the system.

In conclusion, we have demonstrated the light-driven chemical oxidation of organic alcohols into aldehydes, using in situ photogenerated TEMPO$^+$ from the TEMPO that is present as redox mediator in a DSPEC, and coupled to reductive H$_2$ production at the cathode. Photocurrent is only produced when all components are present, that is, dye, light, and TEMPO, and photooxidation occurs with a near quantitative Faradaic efficiency. Effortless on line monitoring of the formed products with ATR-FT-IR and GC is possible in our newly designed DSPEC. The DSPEC generates a photocurrent density that is $\approx$ 10-fold higher than a system with a bound PhCH$_2$OH oxidation catalyst previously reported.$^{[21]}$ We hypothesize that using TEMPO as mediator-instead of a molecular catalyst attached to the semiconductor surface-result in less charge recombination as the oxidized redox mediator can diffuse away from the surface, which will be studied in more detail in the near future.$^{[36]}$

Next to the DSPEC we also report the DSSC properties of the AP11 dye in combination with the TEMPO$^{2+}$-redox couple which shows similar $J_{SC}$ and $V_{OC}$ as compared to the literature LEG4- TEMPO$^{2+}$.$^{[31]}$ The $J_{SC}$ is not as high as in record DSSC based on different dyes, however, we need the high oxidation power that is provided by the AP11 dye. Further optimization to improve efficiency of these devices should include optimization of dyes that provide both high efficiency and high potential. Next to higher efficiencies this may also broaden the scope for photo-oxidative organic synthesis. We are currently exploring this chemistry as well as the utility of our DSPEC for upgrading biobased feedstocks to industrially relevant chemicals.

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Conflict of interest

The authors declare no conflict of interest.

Keywords: catalysis • DSPEC • hydrogen formation • light-driven hydrogen production • photo oxidation • TEMPO
