Lifetime-Associated Two-Dimensional Infrared Spectroscopy Reveals the Hydrogen-Bond Structure of Supercooled Water in Soft Confinement

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DOI
10.1021/acs.jpclett.1c01595

Publication date
2021

Document Version
Final published version

Published in
The Journal of Physical Chemistry Letters

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Citation for published version (APA):
Lifetime-Associated Two-Dimensional Infrared Spectroscopy Reveals the Hydrogen-Bond Structure of Supercooled Water in Soft Confinement

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ABSTRACT: We demonstrate a method to address the problem of spectral overlap in multidimensional vibrational spectroscopy and use it to investigate supercooled aqueous sorbitol solutions. The absence of crystallization in these solutions has been attributed to “soft” confinement of water in subnanometer voids in the sorbitol matrix, but the details of the hydrogen-bond structure are still largely unknown. 2D-IR spectroscopy of the OH-stretch mode is an excellent tool to investigate hydrogen bonding, but in this case it seems difficult because of the overlapping water and sorbitol contributions to the 2D-IR spectrum. Using the difference in OH-stretch lifetimes of water and sorbitol we can cleanly separate these contributions. Surprisingly, the separated 2D-IR spectra show that the hydrogen-bond disorder of soft-confined water is independent of temperature and decoupled from its orientational order. We believe the approach we use to separate overlapping 2D-IR spectra will enhance the applicability of 2D-IR spectroscopy to study multicomponent systems.

Hydrogen bonding plays a crucial role in many structured liquids, the most well-known case being liquid water and aqueous solutions. Infrared spectroscopy is ideally suited to investigate hydrogen-bond structure and dynamics in such systems because of the strong dependence of the OH-stretch frequency on the length (strength) of the OH···O bond.1−4 In particular, two-dimensional infrared (2D-IR) spectroscopy on the OH-stretch mode is often used to investigate water and aqueous solutions and has provided many new insights.5−17 In contrast to conventional infrared spectroscopy, 2D-IR spectroscopy allows separate observation of the homogeneous and inhomogeneous contributions to the OH-stretch line shape: roughly speaking, the homogeneous broadening provides information on the fast dynamical fluctuations of the H-bond structure while the inhomogeneous broadening mirrors the (quasi)static distribution of hydrogen-bond strengths.18 However, 2D-IR is difficult to use in the case of aqueous solutions where both the solvent and the solute contain OH groups, and their OH-stretch bands overlap strongly. Alcohol—and polyalcohol—water mixtures are typical examples of such systems: they exhibit complex behavior, including subnanometer molecular phase segregation19 and liquid−liquid transitions,14,20 but overlapping IR-absorption bands render the application of 2D-IR spectroscopy difficult.

Here, we show that this problem can be overcome by exploiting the different vibrational $T_1$ lifetimes of the OH-stretching modes of the different species in a mixture, and we apply this approach to study the anomalous H-bond structure of supercooled aqueous sorbitol solutions. Recent neutron and X-ray scattering experiments have shown that the structure of water−sorbitol mixtures (70 wt % sorbitol, corresponding to a sorbitol molar fraction of $c \approx 0.19$) is extremely heterogeneous, with the water molecules forming small clusters (<2 nm) within the sorbitol amorphous matrix,21,22 which is characterized, in the supercooled phase, by the presence of nanometer/subnanometer voids that solidify at the glass-transition temperature ($T_g = 200$ K).21,22 Water is therefore in “soft” confinement in the sorbitol amorphous matrix and displays structural properties similar to those of water absorbed in porous silica substrates such as MCM41.22 Crystallization is inhibited by the small size of the pores within which water molecules are segregated, rather than by the increase in the viscosity accompanying the liquid-to-glass transition.22 Remarkably, already at room temperature the orientational order of the H-bond network within the pores is larger than that of neat water, and it increases upon cooling, suggesting that the suppression of ice formation is not due to a reduction of water tetrahedrality,22 but the details of the hydrogen-bond structure remain elusive. As we will see, by exploiting the different $T_1$ values of water and sorbitol OH-vibrations, we can cleanly separate the water and sorbitol 2D-IR spectra and so selectively probe the...
H-bond network of the nanoscopic water clusters across the supercooled phase and below \( T_g \).

We study isotope-diluted (OH:OD \( \approx 0.02 \)) aqueous sorbitol solutions with a sorbitol molar fraction \( c = 0.19 \). At this concentration crystallization is inhibited and the phenomenon of “soft” confinement is observed.\(^{22}\) The sample is prepared by mixing the appropriate amounts of sorbitol, \( \text{D}_2\text{O} \), and \( \text{H}_2\text{O} \) (after mixing, the OH- and OD-groups were statistically distributed between water and sorbitol molecules). We study dilute isotopic OD/OH mixtures rather than isotopically pure liquids to prevent resonant energy transfer.\(^{5,23}\) For the experiments we put 2 \( \mu \)L of liquid between two \( \text{CaF}_2 \) windows separated by a 25 \( \mu \)m Teflon spacer, resulting in an OH-stretch absorbance less than 1 OD at room temperature. A liquid-nitrogen cryostat was used to control the temperature of the sample with a stability of \( \pm 0.1 \) K. The cryogenic 2D-IR measurements were performed using a frequency-domain setup described previously.\(^{14}\) Further experimental details can be found in the Supporting Information.

In Figure 1 we show the linear IR absorption spectrum at several temperatures. At room temperature (295 K) the OH-stretching band is located at a frequency lower than that of pure water (Figure 1a). The IR absorption spectra of water and sorbitol are strongly overlapped, and this makes a spectral disentangling the water and sorbitol nonlinear 2D-IR responses possible.\(^{5,23,24}\) The broad-band pump \( \nu_1 \) component to the OH-stretch mode of water and the 1.6 ps component due to sorbitol, which we do not observe, probably because the OH-stretch analogue of this component decays too fast to be observable with our \( \sim 0.3 \) ps temporal resolution. The temperature-dependencies of the \( T_1 \) of water and sorbitol are plotted in panels c and d of Figure 2, respectively. The corresponding uncertainties, as extracted from the covariance matrix of the least-squares fit, are underestimated as systematic errors are not considered.

The large difference between \( T_1^{\text{HOD}} \) and \( T_1^{\text{sorb}} \) makes possible disentangling the water and sorbitol nonlinear 2D-IR responses in a fashion similar to that for the broad-band pump-probe measurements. Typical 2D-IR spectra, measured at pump-probe delays (\( \Delta t \)) corresponding to the \( T_1 \) values of water and
sorbitol, are shown in Figure 3a–d. The 2D-IR spectra show a strong negative absorption change associated with the $\nu_{OH} = 0 \rightarrow 1$ transition. The presence of two spectral components with different delay dependencies can be noticed by comparing panels a and c or panels b and d of Figure 3: for short pump−probe delays ($\Delta t = 0.7$ ps), the signal is dominated by the OH-stretching mode of HOD molecules, while for longer delays ($\Delta t = 1.6$ ps) the absorption bleach mostly reflects the sorbitol response. The investigated temperatures ($T \leq 270$ K) are sufficiently low that spectral diffusion is negligible on the time scale of the experiment, so that the delay dependence of the 2D-IR spectra directly reflects the vibrational relaxation of the excited states of the water and sorbitol molecules. It is then possible to extract the water (HOD) and sorbitol (â) 2D-IR spectra by fitting the time-dependent 2D-IR signal with a bimodal response function:

$$\Delta \alpha(\nu_1, \nu_2, \Delta t) = \Delta \alpha_{HOD}(\nu_1, \nu_2) \exp \left[ -\frac{\Delta t}{T_1^{HOD}(\nu_1)} \right] + \Delta \alpha_{sorb}(\nu_1, \nu_2) \exp \left[ -\frac{\Delta t}{T_1^{sorb}(\nu_1)} \right] + c(\nu_1, \nu_2)$$

(1)

where $c$ is a small contribution accounting for the temperature increase after the vibrational relaxation has occurred. The fitting procedure was performed using only the data in the 0.7–10 ps time-window to avoid coherent coupling effects, and $T_1^{HOD}$ and $T_1^{sorb}$ were fixed to the values obtained from the broad-band pump−probe measurements. The temperature and $\nu_{pump}$-dependence of the vibrational relaxation times was accounted for by interpolating the data plotted in Figure 2c,d. The resulting least-squares fits are shown in Figures S9 and S10, while the water (HOD) and sorbitol 2DIR spectra extracted from the 2D-IR spectra in Figure 3a,b are shown in Figure 3e–h.

For both water and sorbitol, the 2D-IR spectra are elongated along the diagonal. This indicates that the absorption band is inhomogeneously broadened, that is, there is a broad distribution of vibrational center frequencies, and therefore, the hydrogen-bond network is characterized by a strong structural disorder. Quantitative information can be obtained by analyzing the component-2DIR spectra of water and sorbitol with a Bloch line-shape model, in which the OH-groups are characterized by a Lorentzian homogeneous broadening and center frequencies ($\nu$) distributed according to a Gaussian. This model describes the experimental 2DIR spectra very well (see Figure 3i–l), and from a global least-
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Figure 3. (a and b) 2D-IR spectra measured at 220 and 140 K, respectively, for a pump–probe delay of $\Delta t = 0.7$ ps. (c and d) 2D-IR spectra at the same temperatures as in panels a and b but for $\Delta t = 1.6$ ps. (e and f) Extracted HOD components of the 2D-IR spectra at 220 and 140 K. (g and h) Sorbitol components of the 2D-IR spectra at 220 and 140 K, respectively. (i and j) Results of a least-squares fit of a Bloch line shape model to the data shown in panels e and f. (k and l) Results of a least-squares fit of a Bloch line shape model to the data shown in panels g and h. The black dashed lines show the main diagonal of the 2D-IR spectra.

squares fit to $\Delta \alpha_{\text{HOD}}$ and $\Delta \alpha_{\text{sorbitol}}$ we obtain the half-width at half-maximum of the Lorentzian line shape ($\gamma$), accounting for the homogeneous broadening, and the standard deviation of the central frequency distribution ($\sigma$), representing the inhomogeneous contribution to the line shape. The $T$-dependence of $\gamma$ and $\sigma$ for water (blue dots) and sorbitol (red squares) are shown in Figure 4. The homogeneous width $\gamma$, which represents the fast fluctuations in the H-bond structure, decreases upon cooling for both water and sorbitol, but interestingly, only the sorbitol response is sensitive to the $T_g$ of the mixture, while no change in the slope is observed for water. This observation supports the picture suggested in ref 22, that on approaching $T_g$ the mobility of water molecules decouples from the macroscopic viscosity. The inhomogeneous broadening, which reflects the hydrogen-bond strength distribution, is, in contrast, almost insensitive to $T$ (apart from a small gradual increase in the case of sorbitol). Thus, while the average H-bond strength of the water molecules increases upon cooling (Figure 1c), the width of the associated distribution does not change noticeably. This surprising result indicates that the H-bond-strength distribution and the orientational order within water clusters are independent, because the tetrahedrality of the H-bond network has been shown to increase significantly on approaching $T_g$.

Another intriguing aspect emerging from our transient absorption measurements is the increase of $T_1^{\text{HOD}}$ with decreasing $T$. In neat water, when the H-bond strength increases (i.e., $\nu_{\text{OH}}$ redshifts) the vibrational relaxation becomes faster as the energy gap with low-energy intra-molecular modes (such as the overtone of the HOH-bend) reduces. Here we observe the opposite trend (Figure 2). This trend might be explained if we consider another possible route for vibrational relaxation, e.g., the direct coupling of the OH-stretch with the H-bond network modes. The soft confinement might indeed prevent the formation of an extended H-bond network as in the case of pure water, thus reducing the coupling between the OH-stretching mode and low-frequency modes.

To summarize, we have used 2D-IR spectroscopy to investigate the hydrogen-bond structures of water in soft confinement. Exploiting the different vibrational lifetimes of the OH-stretching mode of water and sorbitol, we can cleanly separate the nonlinear responses of sorbitol and of the water clusters. We find that the average hydrogen bond length decreases upon cooling, and its dependence is sensitive to the liquid-to-glass transition of the solution. In contrast, the distribution of hydrogen bond strengths is almost temperature-insensitive and decoupled from the orientational order. The method here introduced can be applied to any mixture of molecules with overlapping 2D-IR spectra, provided that the $T_1$ values of the spectrally overlapping modes are sufficiently different. We therefore believe it can become a valuable tool to investigate mixtures or other multicomponent systems by means of 2D-IR spectroscopy.

**ASSOCIATED CONTENT**

**Supporting Information**

The Supporting Information is available free of charge at https://pubs.acs.org/10.1021/jacslett.1c01595.

5954  https://doi.org/10.1021/acs.jpclett.1c01595
Experimental details, absence of crystallization during FTIR measurements, T-dependence of the OH-stretching band, HOH bending mode of water, examples of fitting curves of transient absorption data, and absence of spectral diffusion (PDF)

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Notes
The authors declare no competing financial interest.

ACKNOWLEDGMENTS

This work was supported by The Netherlands Organization for Scientific Research (NWO) (Grant Number 680-91-13).

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(35) At 190 K, the 2D-IR measurements were performed at lower frequencies than the other temperatures. For this reason the sorbitol 2DIR spectrum, blue-shifted with respect to the water component, was not fully resolved. Hence, we could not properly estimate the related inhomogeneous broadening at this temperature.

